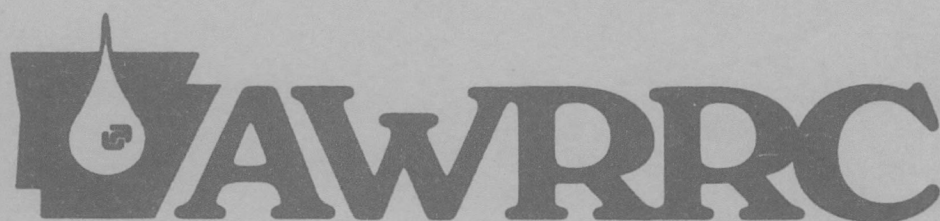


**CHEMISTRY OF THE SPRING WATERS
OF THE OUACHITA MOUNTAINS
EXCLUDING HOT SPRINGS, ARKANSAS**

by
George H. Wagner and Kenneth F. Steele
Department of Geology



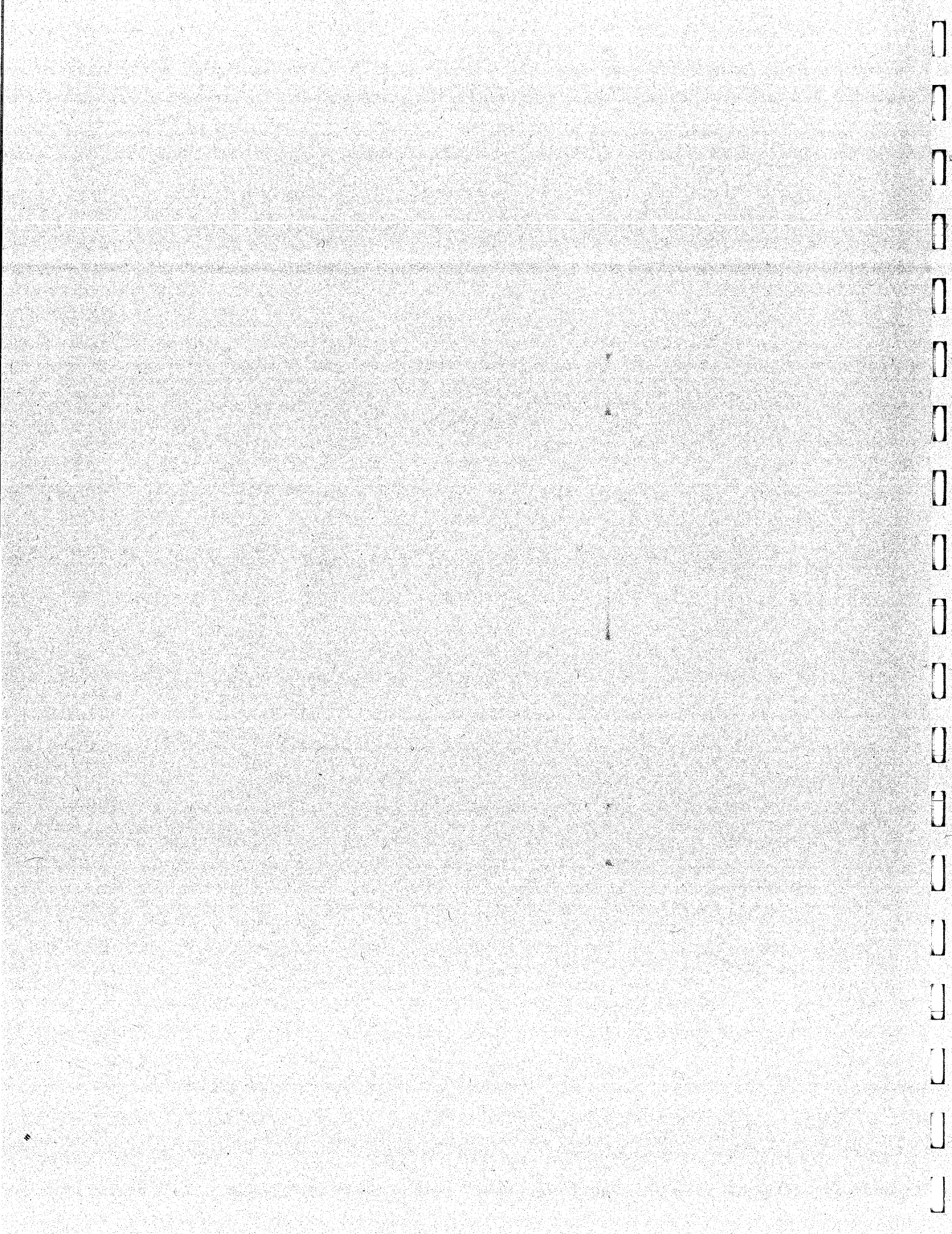
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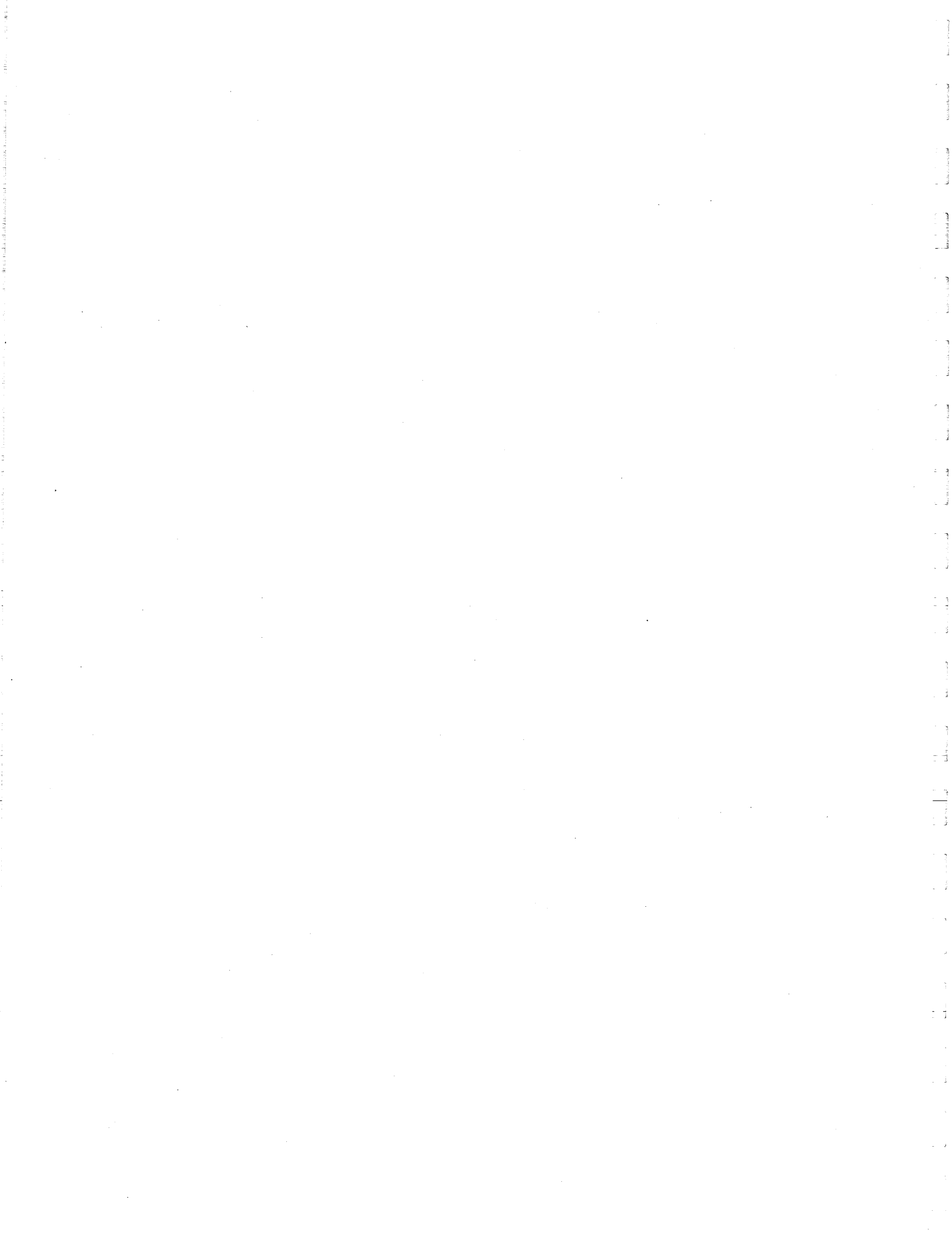
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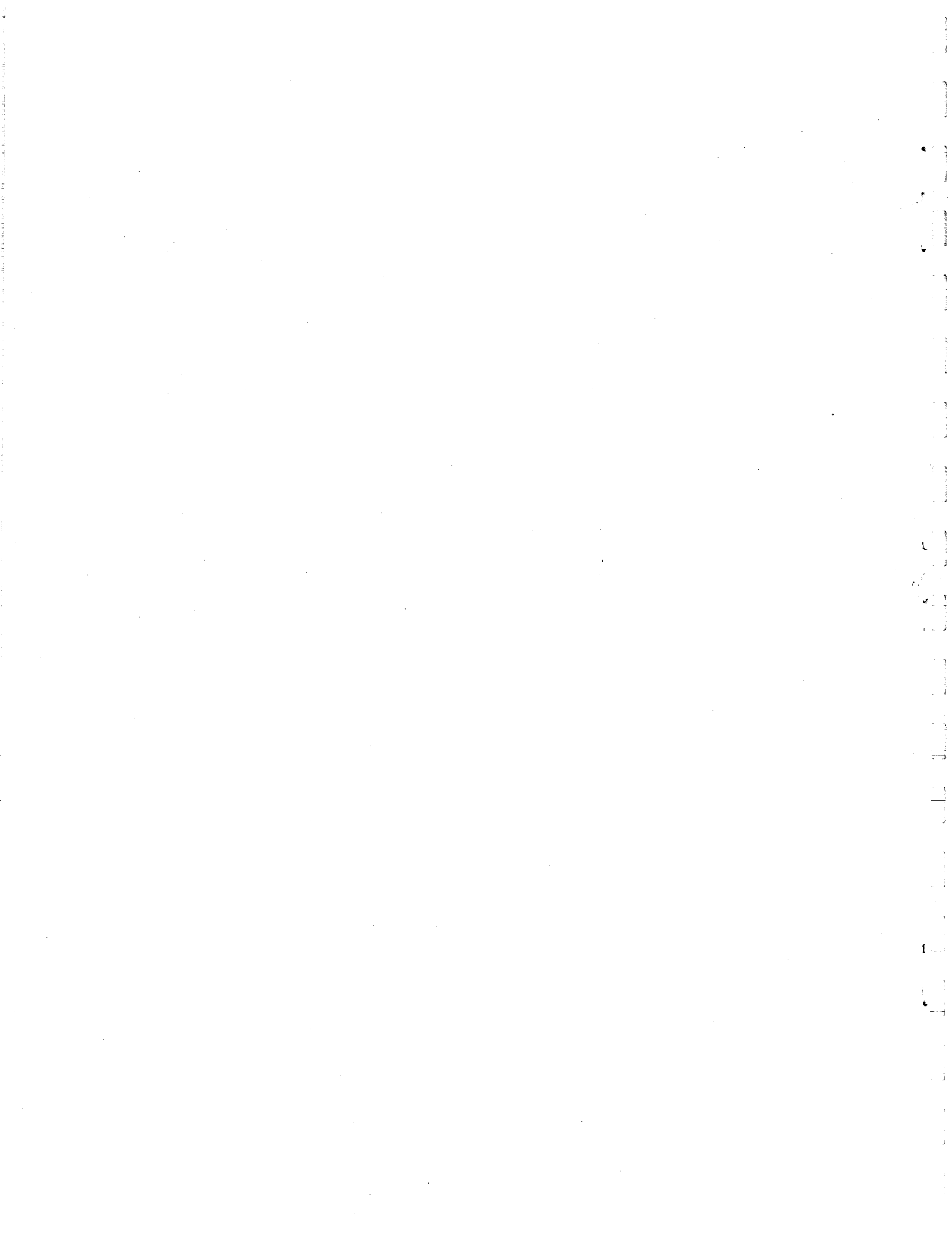


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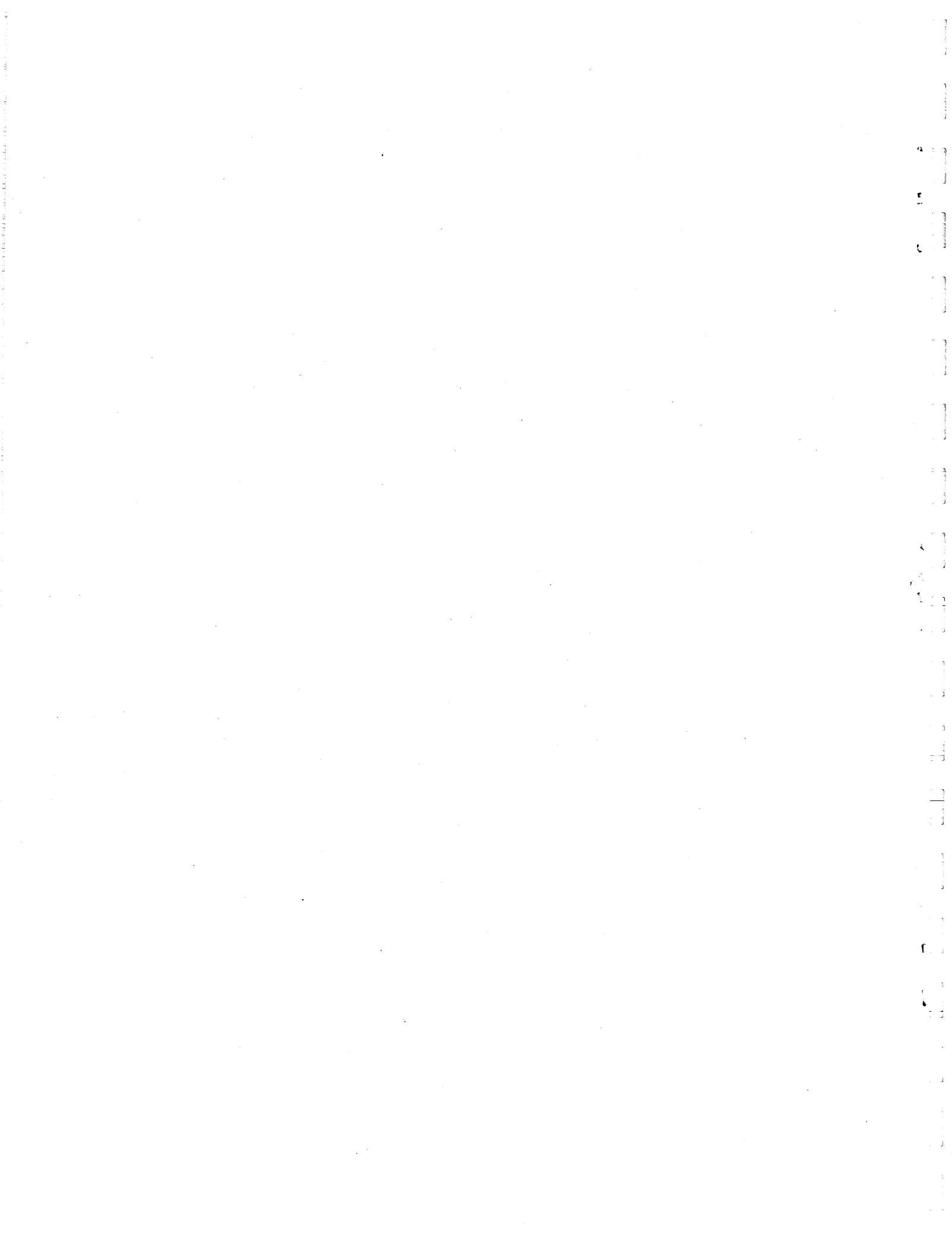
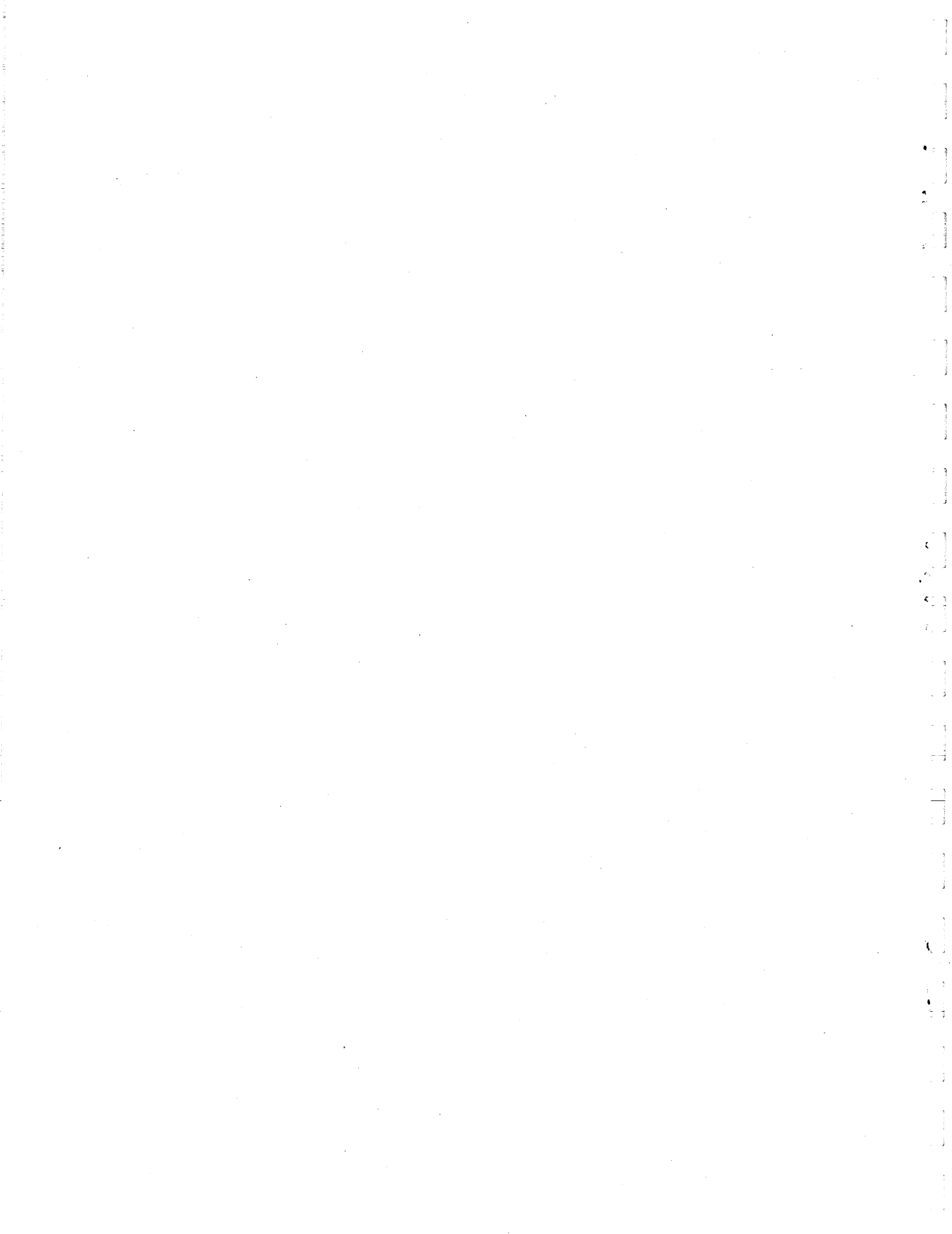


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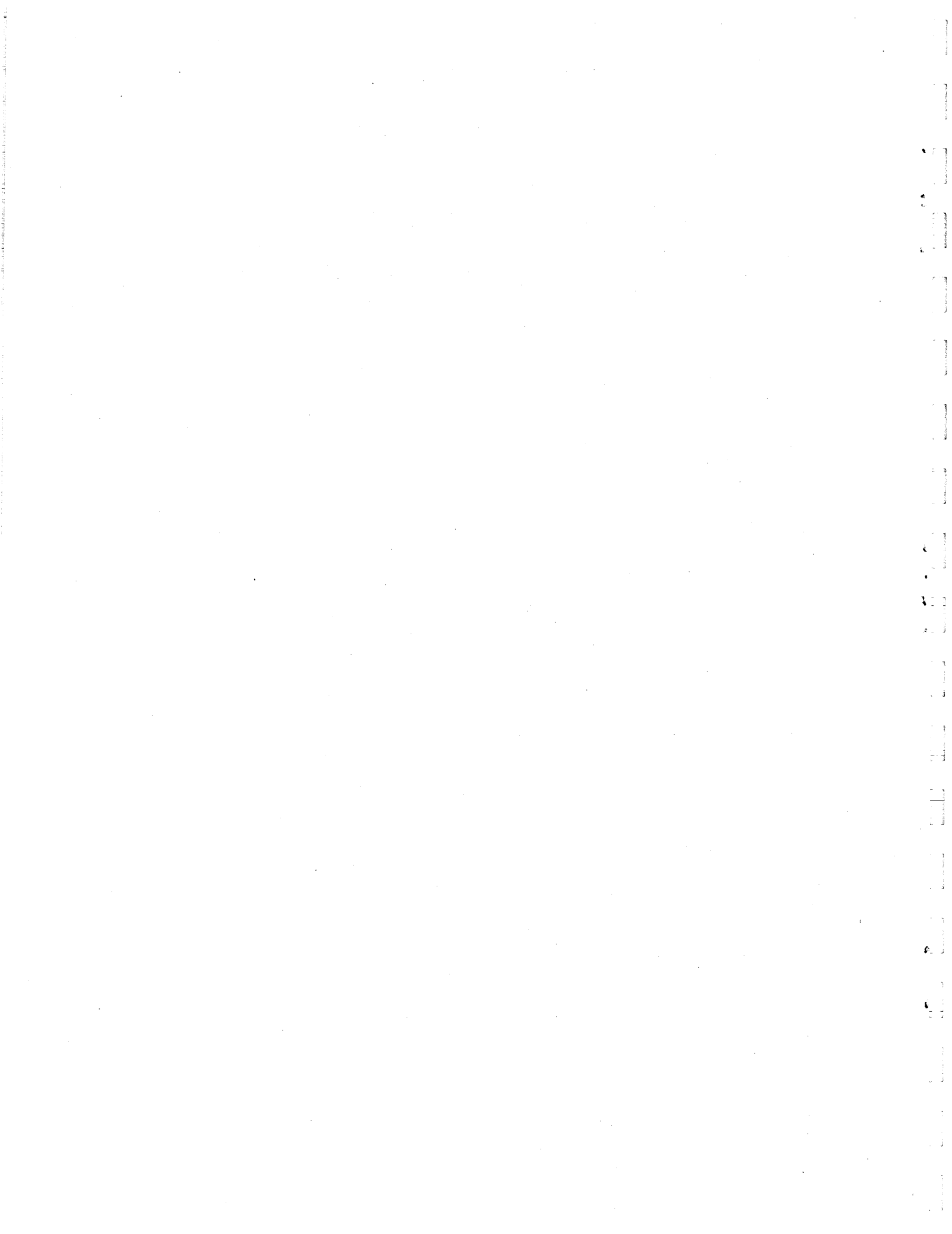
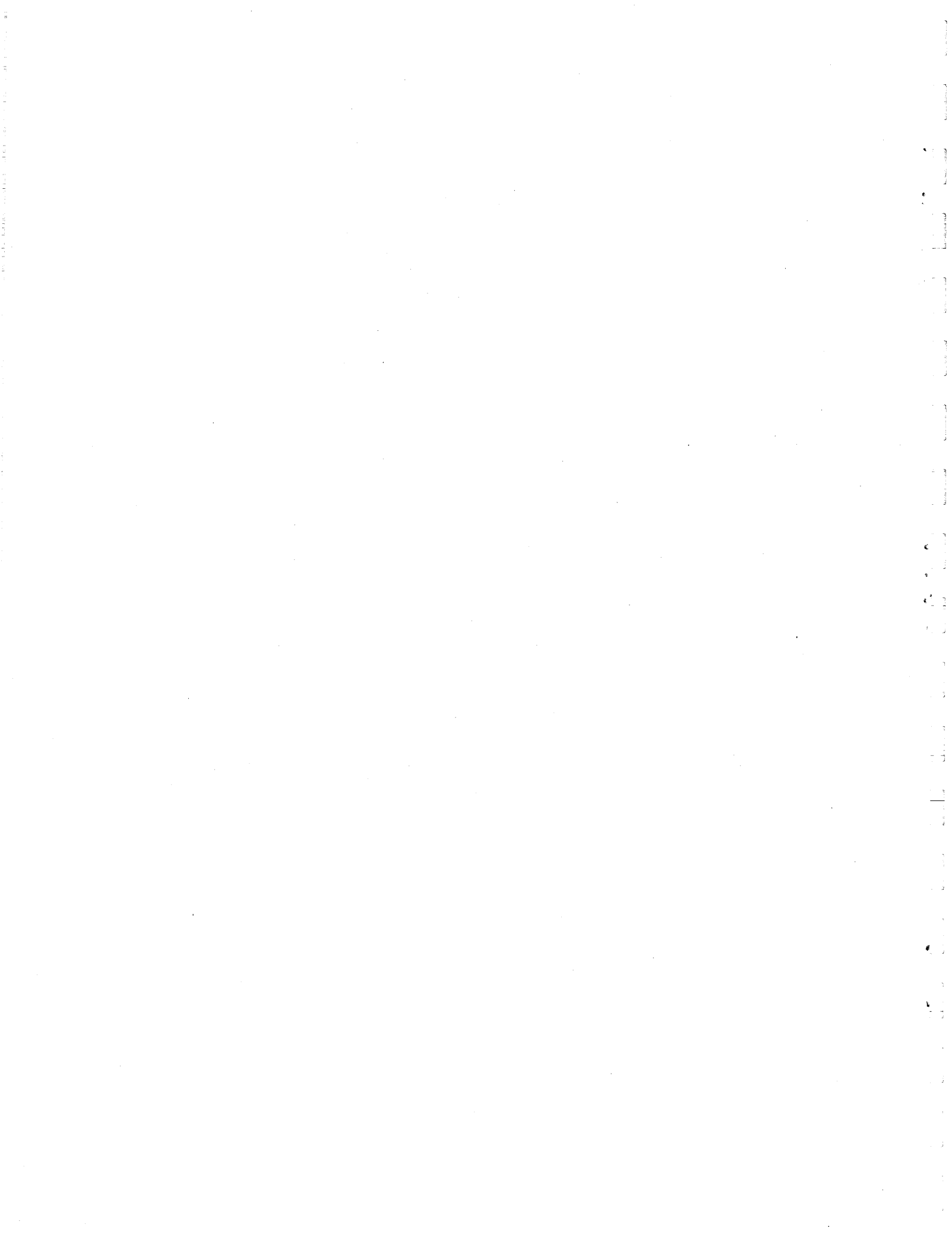


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1. Introduction

This report is based on the chemical analysis of the waters from 93 springs and 9 wells. Springs, when free from metal plumbing, provide an uncontaminated source of the ground water and it was desired to obtain water uncontaminated with metals. A few wells were added to the list, usually because of their unique location in the sampling grid.

An uncontaminated ground water sample is of interest for its quality (purity), its subsurface temperature and its aid in mineral exploration. These three interests are covered in the main sections of this report which follow the experimental and geological sections.

Study Area I is a 135 x 80 km area which extends west from Hot Springs, Arkansas, to the Oklahoma line and it is located north of the $34^{\circ} 42' 30''$ N. latitude line. It encompasses the core area of the Ouachita Mountains. Area II is 135 x 23km and abuts Area I on the south. It is known as the Athens Plateau and includes some of the Gulf Coastal Plains. Figure 1-1 shows the location of the study area and the springs. Springs were located from notations on USGS topographic maps, from the literature on warm springs in Arkansas (Billingsley and Hubble; Miser and Purdue, 1929; Bryan, 1922) and from discussions with local inhabitants.

For convenience tables and figures are collected together in the rear of the report and titles with page location are summarized in the Table of Contents.

2. Experimental

In the collection, storage, and analysis of the water samples, the recommendations and procedures of the Environmental Protection Agency were followed. These procedures are outlined in EPA (1974) and APHA (1971). All samples were collected and stored in polyethylene containers. Some polyethylene caps for the 1 gallon containers are compounded with a filler. These types of caps were avoided and only clear polyethylene caps were used. All polyethylene containers were rinsed with concentrated nitric acid and allowed to stand capped from one to several hours, then rinsed several times with dionized distilled water before use.

At each site the routine was to collect 5 liters of raw water, 4 liters of which were immediately filtered through a 0.4 μ m (142mm) Nucleopore filter and stored in a 1 gallon polyethylene container to which 12 ml of 1:1 HNO₃ were added. An additional 500 ml of 0.4 μ m filtered water was stored in a polyethylene container at 4^oC for SiO₂, SO₄, Cl, PO₄, NO₃, and NH₃ analyses. Anion analyses were made within 3 days after collection. Alkalinity, pH, and conductivity were measured at the site on unfiltered samples.

A special filter devised by the Savannah River Laboratory was used for the filtrations. This filter was pressurizable to 40 psi with freon, used a 142 mm filter, and filtered one liter of water per filling. New filters were prewashed with at least 100 ml of raw water, via filtration to remove any soluble contaminants.

All cation analyses were by either atomic absorption (AA), spectrophotometry or flame emission. Both a Perkin Elmer Model 303 AA unit and a Jarrel Ash modernized Solid State Model 82-500 unit were used. The particular flames and other adaptations of these units are summarized in Table 1-1 along with the anion analytical schemes. Heavy metals (Fe, Co, Ni, Co, Zn, Pb) were first chelated and extracted by methyl isobutyl ketone after Nix and Goodwin (1970), before being determined by AA.

Temperature was measured by a mercury thermometer which was calibrated against an A.S.T.M. certified thermometer. The pH was measured with a Markson Model 80 pH meter. Conductivity was measured with a KCl (soln.) calibrated YSI Model 33 S-C-T conductivity meter. Total alkalinity was measured by titration to the methyl red end point using a methyl red- bromocrescol green mixed indicator and 0.02 N H_2SO_4 . All measurements were made on the acidified raw water except for barium. For this element the acidified raw water was first concentrated 10-25 times by evaporation from a teflon beaker on a hot plate in order to improve the sensitivity. No loss in this evaporation technique was found for Na, K, and Mg. Therefore, no loss of Ba is expected.* However, Sb could not be retained by an evaporative concentration process and was always determined on the raw water.

Mercury was determined by atomic absorption using the flameless method (EPA, 1974, pp. 118). The samples of raw water were first treated with strong oxidizing agents to convert any organic mercury compounds to inorganic forms. The method thus measures organic and

*On a few samples with detectable Ba in the raw water no loss was found.

inorganic mercury. The oxidizing procedure was as follows: 1) add 5 ml of 5.6 N HNO_3 to 100 ml of the raw water sample (from 1 gallon polyethylene storage container which has been acidified with 1:1 HNO_3); 2) wait 15 seconds and add 5 ml of 18N H_2SO_4 ; 3) wait 45 seconds and add 15 ml of 5% potassium permanganate and 8 ml of a 5% aqueous solution of potassium persulfate; 4) heat 2 hours at 95°C ; 5) cool to room temperature, add 5 ml of hydroxylamine hydrochloride-NaCl solution (15 g/100 ml of each); 6) add 5 ml of stannous chloride solution (10%) and analyze immediately. A Perkin Elmer aerator and analyzer cell fitted to a Perkin Elmer Model 303 spectrophotometer were used.

Antimony was analyzed by the volatile hydride method after Fernandez (1973). A Perkin Elmer (PE) Hydride Generation Sampling System (303-0849) was adapted to a Jarrel Ash Solid State modernized Model 82-500 atomic absorption spectrophotometer. A 20 ml sample of spring water (from acidified storage sample) was added to the PE generator along with 8 ml of conc. HCl. A pellet (10/32 inch) of solid sodium borohydride (from Alfa Inorganic Division of Ventron Corp., Beverly, Mass. 01915) was added to the generator to form hydrogen and antimony hydrides which were fed to a hydrogen-nitrogen flame. Peak heights on a recorder were calibrated against antimony standards.

The following table gives a comparison of SiO_2 analysis by atomic absorption (DeVine and Suhr, 1977) and by the colorimetric method adopted for this work.

Spring No.	Colorimetric Method (ppm)		Atomic Absorption Method (ppm)
	Analysis 1	Analysis 2	
501	7.7	8.7	7.7
502	3.2	4.4	7.2
503	18.3	19.1	23.0
504	12.1	13.5	15.6
505	7.7	9.1	12.2
506	12.1	13.9	15.6
507	13.7	9.5	11.6
508	7.7	10.0	11.6
509	10.3	8.8	11.2
510	6.6	3.7	6.1
511	10.2	8.1	8.9
512	8.9	9.3	8.9

Analytical data are summarized in Tables 1-2 to 1-4. Precision for most cation analyses in Table 1-2 to 1-4 is estimated at $\pm 10\%$. This is generally true when the concentration being analyzed is several times the detection limit -- the "less than" values in the various columns. These detection limits may vary in a given column because of change in sample size, instrument noise at the time, or size of the blanks. Where the concentrations are low the per cent error is much larger and is given by $(\text{detection limit})/(\text{concentration}) \times 100$. The reproducibility of the various anion analyses in the hands of various investigators has been studied in part by EPA and APHA. Reproducibility for our median concentration ranges are given in the concluding lines of Table 1-2 using the precision data of EPA (1974) and APHA (1971) for this range when available.

3. Geology

The following two study areas were selected based on contrasting lithologies and mineralizations.

a. Area I: (Mn, Fe, Ba mineralization with minor Cu, Zn, Pb mineralization) As shown in Figure 1-1 this area includes parts of Montgomery, Polk, Pike, Howard, and Garland counties in west-central Arkansas. This is the core area of the Ouachita Mountains with relief of 152 to 304 m. and ridges 450 to 701 m above sea level. The region has some small farms, but is mostly forested and in the Ouachita National Forest.

A stratigraphic column for Area I is shown in Figure 1-2. Formations range in age from Cambrian to Carboniferous. Shales, Arkansas Novaculite (chert), and sandstone predominate with only minor limestone (in the Crystal Mountain and Womble Shale Formation). The upper member of the Arkansas Novaculite can be highly calcareous. Folding is intense and the east-to-west ridges have a corrugated surface due to the steeply dipping strata.

Manganese mineralization is widespread. Manganese oxide minerals lithiophorite ($\text{Li}_2\text{Al}_8\text{Mn}_{12}\text{O}_{35}\cdot 14\text{H}_2\text{O}$), cryptomelane ($\text{K}_x\text{Mn}_8\text{O}_{16}$), psilomelane, $(\text{Ba}, \text{H}_2\text{O})_4\text{Mn}_{10}\text{O}_{20}$, and pyrolusite (MnO_2) are scattered among over 100 manganese ore prospects and mines. These minerals particularly lithiophorite, may contain of the order of 1 wt. % each of the base metals Co, Ni, Cu, and Zn (Wagner, et. al., 1979). These minerals and iron oxides are found in the bedding planes and cracks of the lower and upper division of the Arkansas Novaculite. Native copper has also been found in the novaculite.

Two barite districts are known. In each case the barite is in the first few meters of the lower Stanley Shale Formation. The barite

is formed as a result of replacement or fracture filling. Quartz veins in shale are associated with silver-containing sphalerite and galena in the eastern extremity of Area I.

b. Area II: (Hg, Sb, Sr, Ba mineralization with minor Zn, Pb mineralization) This area consists mainly of the Athens Plateau in the Gulf Coastal Region of Arkansas. Ages of the rocks vary from Mississippian (Stanley Shale), Pennsylvanian (Jackfork sandstone and Atoka sandstone, shale, and siltstone), Cretaceous (Trinity limestone, gravel, siltstone, sandstone and some barite), and Quarternary (Terrace deposits and alluvium of gravel, sand, and silt). Along the northern edge there is a gradation from Area I to Area II type rocks. In the central part of the area there are kimberlite necks at Murfreesboro.

Three mineral districts exist. A mercury district with several mines covers essentially the central and eastern half of Area II (Clardy and Bush, 1976). An antimony district covers the western half of Area II with the antimony and mercury districts overlapping in the central part of Area I (Stroud et. al., 1969). Strontium ores have been mined in Howard county from the DeQueen Limestone member of the Trinity Formation. Similar deposits occur intermittantly across the area. Ore minerals for the three districts are mainly cinnabar (HgS), stibnite (Sb_2S_3), celestite ($SrSO_4$) and strontionite ($SrCO_3$). Some barite ($BaSO_4$) is found with the strontium area. In Howard and Pike counties barite occurs as a cementing material in gravels and sandstone.

Gypsum deposits occur in the central part of Area II.

Quartz veins in sandstone adjacent to Stanley Shale were mined for lead, zinc, and silver in two deposits in Sevier county, in the western part of Area II.

than the wells. The springs also have a higher median Ca^{++} concentration which may be due to the more alkaline nature of the well water. The higher pH values for well water may be related to chemical and physical reactions in storage tanks, e.g. the loss of carbon dioxide could raise the pH of the water causing precipitation. The springs median and upper range values for SiO_2 exceed those for the wells probably due to the high density of samples in the Caddo Gap area where some warm springs are located, and also the few wells samples were deep ones which could intersect thermal waters.

The higher values exhibited by the wells may be the result of more intimate contact of water with rocks or other factors. However, because the well water was not filtered, it is likely that the presence of suspended material has increased concentration values for many parameters. Thus, most of the differences between the well and spring waters in Table I-1 may simply reflect the differences between unfiltered and filtered groundwater.

2. Water Chemistry

The large number of analyses for a variety of elements and anions provides an opportunity to: 1) seek correlations among the various elements and 2) examine the saturation of the waters with respect to some of the more common minerals. These two facets are examined in the following subsections.

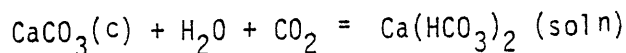
a. Base Metals: The manganese concentrations of the waters were found to correlate with their concentrations of Fe, Co, Ni and Zn. Correlation coefficients calculated after Lepeltier (1969) from Figures I-1 to I-11

are summarized below. The correlations are for Areas I and II as a whole.

<u>Metal</u>	<u>Correlation Coefficient</u>	
	<u>With Mn</u>	<u>With Fe</u>
Fe	0.652	-
Co	0.454	0.228
Ni	0.778	0.500
Cu	-0.050	0.034
Zn	0.430	0.391

Correlation of the base metals with Fe is somewhat poorer than with Mn as noted by the smaller correlation coefficients in the above table. Copper did not correlate with either Mn or Fe. These correlations with Mn are particularly interesting because base metal contents of stream sediments in this area were found in previous studies (Wagner, et al., 1978) to correlate with their Mn contents. Similar chemistries and perhaps similar source rocks account for the correlation between elements.

b. Alkaline Earth Metals (Ca, Mg, Sr and Ba): Calcium is by far the major element, in general, in the various spring waters. This is due to the widespread distribution of limestone and calcareous cements, both of which are mainly CaCO_3 , in sedimentary rocks and the ready solubility of CaCO_3 in waters containing CO_2 via the equation:



The carbon dioxide (CO_2) is available from the air (0.03%) and soil (respiration and decay of plants).

The concentrations of other alkali metals correlate with the concentrations of calcium. The following table shows correlation coefficients

calculated after Lepeltier (1969) from the plots in Figures I-12, III-12, III-13 and III-16.

<u>Element</u>	<u>Correlation Coefficient with Ca</u>
Mg	0.797 (Areas I and II)
Sr	0.925 (Areas I and II)
Ba	0.123 (Area I), 0.329 (Area II)

The chemistries of the four elements are similar and they have common sources, except for Ba. The barium ion does not fit into the CaCO_3 lattice, being too large. Mg^{++} and Sr^{++} are common impurities in limestone. Ba is more likely to come from barite, igneous rocks, or feldspar fragments in sedimentary rocks. These differences of Ba are reflected in the lower correlation coefficient.

Barium occurs widely in our study Area I as the manganese mineral psilomelane, $(\text{Ba}, \text{H}_2\text{O})_4\text{Mn}_{10}\text{O}_{20}$. In stream sediments correlation coefficients for Mn-Ba were about 0.9 (Wagner, et al., 1978). No significant correlation has been found for Mn-Ba based on their concentrations in the waters of Areas I and II ($p = 0.164$). This contrasts to the Mn-base metals correlation which persists in sediments and groundwater, as noted in the previous subsection.

Skougstad and Horr (1963) have summarized a large number of Sr analyses for surface and groundwaters of the United States. The sulfates and carbonates of Ca, Sr, and Ba are relatively insoluble. It is of interest to determine the relative saturation of the spring waters with

respect to the six minerals corresponding to these compositions. Table I-2 lists the minerals, their solubility products, the activity ratios for ions with minerals in equilibrium, and the median values for these quantities for Areas I and II. Equilibrium values of activity ratios when drawn on correlation log - log plots for the two ions show which springs are saturated with which mineral.

General statements can be made by comparing median values and equilibrium values of activity ratios. For the celestite (SrSO_4) = gypsum ($\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$) equilibrium the mean value of $a_{\text{Sr}^{++}}/a_{\text{Ca}^{++}}$ is greater than the equilibrium value in both Area I and II and indicates that celestite is generally the stable mineral, not gypsum. Barite (BaSO_4) is even more favored than gypsum and is thus the favored sulfate mineral. Calcite is the most favored carbonate mineral.

In Figure I-13 the solubility products of strontianite (SrCO_3) and calcite (CaCO_3) are plotted versus temperature. Figure III-11 has similar data for barite (BaSO_4) and celestite (SrSO_4). We have calculated for each spring using its surface temperature, the percent saturation with respect to the four minerals: barite, celestite, calcite, and strontianite. Median values and the ranges are shown in the following table. Activity coefficients were determined from ionic strength and the graph of Hem (1970), p. 21. Carbonate ion concentrations were determined from the $C_{\text{HCO}_3^-}$ (from alkalinity), pH, and the equilibrium constant for $K_{\text{HCO}_3^-}$ which is plotted versus temperature in Figure I-13A.

% Saturation of the Spring Waters with BaSO₄, SrSO₄, CaCO₃ and SrCO₃
At Surface Water Temperature

<u>Mineral</u>	<u>Area</u>	<u>Median</u>	<u>Range</u>	<u>% of Springs</u>	
				<u>Above 10% Saturation</u>	<u>Above 50% Saturation</u>
barite, BaSO ₄	I	1.3*	0.16-96	15	7
	II	8.5*	0.40-97	45	16
celestite, SrSO ₄	I	0.0075	<0.0001-0.50	0	0
	II	0.0047	0.0011-1.97	0	0
calcite, CaCO ₃	I	0.065	0.000013-154	25	7
	II	0.022	0.00019-79	19	4
strontianite, SrCO ₃	I	0.033	0.04 x 10 ⁻⁴ - 8	0	0
	II	0.002	0.002 x 10 ⁻⁴ - 31	8	0

*Using all values and counting < values at ½ their < value.
Area I had a median of 2% using finite values only.
Area II had a median of 8% using finite values only.

Only barite and calcite show any appreciable saturation in the spring waters. In Area II there were 8% of the springs with over 10% saturation with strontianite but none were over 50%. Both barite and calcite have some springs over 50% saturated.

c. Alkali Metals (Li, Na, K): Figures I-14 to 18 show plots of Na versus Mg, K, Sr, Ca, and Li. Only Li shows a reasonably good correlation coefficient ($p = 0.509$) with Na. However, the Mg-Na, Sr-Na, and Ca-Na plots have a minimum envelope to the data. The reason for this is unknown at the moment but it is interesting that the minimum parallels more or less the concentration ratios for the ions in rain water. Curves A, B, and C depict in Figure I-14 the course of the ionic ratio of Mg/Na based respectively on A (absorption of Mg from rain water solution and replacement with Na), B (either addition of Mg or absorption of Na with no

replacement by Mg) and C (addition of Na to the rain water with no change in Mg). Mere loss of water proceeds along the rain-ratio line. Since most of the points of Figure I-14 are above the rain-ratio line the combined process of water loss (by evaporation and transpiration) and Mg addition predominate. In Figure I-15 for K-Na, the process of water loss and K absorption predominate. Figure I-16 for Sr-Na is predominantly water loss and Sr addition. Figure I-17 indicates both Ca addition and Ca removal processes at work along with water loss.

A manganese mineral containing lithium, lithiophorite, with the formula $\text{Li}_2\text{Al}_8\text{Mn}_{12}\text{O}_{35}\cdot 14\text{H}_2\text{O}$ is widespread in our study area (Wagner et al, 1979). This mineral might provide both Li and Mn to groundwater. Figure I-5 shows a correlation plot for Mn-Li with $p = 0.289$, a rather poor correlation.

Vatin-Pérignon et al.(1979) has listed the Li/Na ratio of a large number of different kinds of waters .

II. GEOTHERMOMETRY

1. Introduction

A survey of spring water temperatures in Areas I and II has been carried out utilizing two methods: 1) the quartz geothermometer and 2) the surface temperature of the water. By this means areas of geothermal activity should be located. A third method, the Na-Ca-K geothermometer, was evaluated for the study area but it yielded erroneous results because it is designed for areas where feldspars are abundant. Feldspars are rare to absent in the rocks of the study area.

The approach taken was to construct histograms for the total study area (Areas I & II) of the temperature obtained from the SiO_2 and surface methods, calculation of the mean and standard deviation, assuming a normal distribution, and defining values greater than one standard deviation unit above the mean as being anomalous (Swanberg and Morgan, 1978). The details of the method are discussed below.

a. Quartz Geothermometer: Many papers have been published regarding the solubility of silica as an empirical function of temperature in an aqueous environment (Fournier, et al., 1974); (Fournier and Truesdell, 1973); (Mahon, 1966) and (Mariner and Willey, 1976). The relationship is expressed quantitatively according to the equation of Truesdell (Swanberg and Morgan, 1978):

$$T_{\text{SiO}_2} = \left[1315 / (5.205 - \text{Log}_{10} \text{SiO}_2) \right] - 273.15$$

where T_{SiO_2} is the silica geotemperature in degrees centigrade and SiO_2 is expressed in parts per million. The method is useful in estimating reservoir base temperatures of geothermal systems. However, three principle

assumptions must be made in order for the method to be utilized quantitatively (Swanberg and Morgan, 1978): 1) there is no mixing of non-thermal waters with the waters migrating from the geothermal reservoir to the surface, 2) equilibrium in the geothermal system must be established between the waters and the surrounding rocks, and 3) there is no silica precipitation. The general reluctance of quartz to precipitate from supersaturated solutions (Truesdell, 1976) and slow water-rock reactions at low temperatures (Rimstidt, 1977) aid in supporting these assumptions. It must be kept in mind that temperatures calculated from the SiO_2 concentrations are a minimum temperature due to the possibility of dilution taking place by waters encountered near the surface.

b. Surface Temperature: The surface temperature method involved establishing a mean surface temperature of the waters sampled in the study area. We assume after Swanberg and Alexander (1979) that waters at the surface having a temperature of one standard deviation above the mean may be of geothermal origin. Figure II-3 is a histogram of the surface temperatures of the study area. The modal value is 17°C and compares with a mean value of 16.8°C and standard deviation of 4.3°C .

c. Heat Flow: Heat flow is defined as the product of the thermal conductivity of a substance and the thermal gradient in the direction of the flow of the heat (AGI Glossary, 1972, p. 326). Swanberg and Morgan (1978) showed that regional heat flow can be predicted from the silica concentration of groundwater by using the following empirical equation:

$$T_{\text{SiO}_2} = m'q + b, \quad (1)$$

where T_{SiO_2} is the silica geotemperature in degrees centigrade ($^{\circ}C$), q is heat flow in milliwatts (w) per square meter (m) and m' and b are constants of $0.67^{\circ}Cm^2/w$ and $13.2^{\circ}C$, respectively. The physical significance of the intercept b is that this value should represent the national mean annual air temperature because the absence of heat flow implies that all thermal energy be solar in nature (Swanberg and Morgan, 1978).

d. Geothermal Gradient: The geothermal gradient represents the increase in the rock temperature per unit of distance as one proceeds from the surface of the earth downwards. It is expressed in the following equation:

$$\text{Geothermal Gradient} = \text{Heat Flow/Thermal Conductivity of the Rocks} \quad (2)$$

where the geothermal gradient is in $^{\circ}C$ per kilometer, heat flow is in milliwatts per square meter and the thermal conductivity is in milliwatts per kilometer per $^{\circ}C$.

The stratigraphic column of the study area (Figure 1-2) includes numerous strata of variable thicknesses. The dominant rock types in the area are sandstones, shale, and chert. The thermal conductivities of sandstone and shale are quite similar (Handbook of Chemistry and Physics, 1972-1973). The thermal conductivity value of chert could not be located and was estimated at $1.2w/km^{\circ}C$. Due to the relatively small difference in the thermal conductivities of the three rock units and also due to the actual thicknesses of the rock units in the Ouachitas being unknown, an approximate thermal conductivity value of $1.3w/km^{\circ}C$ was assigned to the Ouachita sediments.

e. Circulation Depth: It is not known to what depth the waters from the springs are circulating, or to which formations they are confined. However, based on the mean SiO_2 temperature and the mean geothermal gradient, a mean depth of circulation can be calculated by the following equation:

$$\text{Depth} = (T_{\text{SiO}_2} - A) / (\text{Geothermal Gradient}) \quad (3)$$

where depth is in kilometers, T_{SiO_2} is in $^{\circ}\text{C}$, the geothermal gradient is in $^{\circ}\text{C}/\text{km}$, and A is the mean surface temperature. The surface temperature of the rock is assumed to be the same as the surface temperature of the spring water.

2. Results and Discussion

The frequency distribution of the silica concentrations found in the Ouachita spring samples is shown in Figure II-1. It can be seen that the bulk of the samples can be accounted for in the 5-15 ppm range. The calculated mean was 11.5 ppm and the modal value was 10 ppm. Geotemperatures calculated from silica concentrations are shown in Table II-1 and represented in a frequency plot in Figure II-2. The mean silica geotemperatures in the study are was 44.3°C with a modal value of 37.5°C . Because the mean and the modal values are approximately equal and the 14.1°C standard deviation less than the 25°C suggested by Swanberg and Morgan (1978), the value of 44.3°C is taken to be representative of the entire Ouachita Province (Areas I & II). The provinces that most closely resemble the Ouachita Province in Swanberg and Morgan's (1978) paper are the Eastern United States (North), Eastern United States (Central), and Middle Rocky Mountains (Wyoming). In the North and Central Eastern United

States, the silica concentrations are slightly less than the Ouachita Province, which indicates a slightly lower calculated subsurface temperature. In contrast, the Rocky Mountain region yielded slightly higher silica values than the study area corresponding to higher subsurface temperatures.

According to Swanberg and Morgan (1978), geotemperatures of greater than one standard deviation unit above the mean (i.e. greater than 58.4°C in this case) is an effective cut-off temperature in searching for geothermal areas. Using this criteria, there are approximately 14 springs out of the 102 water sites sampled that are possibly of "thermal origin or have a component of thermal water". Table II-2 lists these springs and their associated silica geotemperatures.

Using the surface temperature method, 7 springs (Table II-2) have temperatures of one standard deviation above the mean (i.e. greater than 21.1°C in this study). Only two springs, 923W and 924 in the Caddo Gap area, have both anomalous surface and geotemperatures. The two hottest springs, based on surface temperatures, 915(35°C) and 921(30°C) also in the Caddo Gap area do not have anomalous geotemperatures.

The data show that low temperatures at the surface need not be indicative of low temperatures calculated from the SiO₂ concentrations. This observation can be explained by considering the path taken by the waters on their way to the surface. Once silica is in solution it is reluctant to precipitate out even though the temperature of the water drops as would be the case for a long distance of travel. If only a short distance of travel is involved, the silica concentration should remain constant;

however, the water temperature should be higher. Such a case then should be indicative of a heat source near the surface. However, it must be kept in mind that the silica geotemperature values represent a minimum temperature due to the possibility of the waters being diluted on their way to the surface.

Due to faulting being so intense in the study area, it is believed that the anomalous silica values are associated with faults which penetrate, stratigraphically, deeper than their surrounding counterparts.

Minimum heat flow was calculated using the method of Swanberg and Morgan (1978), as stated above. Table II-1 shows the results of these calculations. Sample 502 gave an unrealistic negative value. This sample is extremely low in silica and is interpreted as containing primary meteoric water.

Figure II-4 shows a contour map of the heat flow in the study area. Areas of highest heat flow are located in the vicinity of Caddo Gap, the Cossatot Mountains, Lake Greeson, Mena and east of Dierks Lake in Sevier County. In the Lake Greeson area, the high heat flow is believed to be associated with the epithermal Hg deposits.

Table II-1 shows the geothermal gradient values calculated for the study area. In general, the geothermal gradient in the study area ranges from 20-50 °C/km and has a mean value of 35.7. Due to the lack of wells drilled in the study area, the geothermal gradients given by the AAPG are only rough extrapolations, yet based on the AAPG values through the middle of the study area, their average of 30.08 °C/km corresponds quite well with our mean of 35.7 °C/km.

Equation (1) was developed for regional use where T_{SiO_2} and q have mean values for the region being studied. In Table II-1 we have applied the equation to the calculation of heat flow, q , for individual springs. Using the mean value for our study area (Area I & II) for T_{SiO_2} we now calculate the mean value of q using equation (1), and from equation (2) and (3) the mean heat flow, geothermal gradient and depth of circulation for our study area. These values are summarized below:

<u>Parameter</u>	<u>Symbol</u>	<u>Mean for Area I & II</u>	<u>Units</u>
Subsurface Spring Water T	T_{SiO_2}	44.3	$^{\circ}C$
Heat Flow	q	46.4	milliwatts/m ²
Geothermal Gradient	-	35.7	$^{\circ}C/km$
Surface Spring Water T	A	16.8	$^{\circ}C$
Depth of Circulation of Spring Water	-	0.77	km

The mean circulation depth of 0.77 km seems reasonable for our study area.

3. Comparison of Historical Data and Present Work

Miser and Purdue (1929) and Bryan (1922) have previously recorded data on warm springs in Area I. Although Miser's data were published in 1929, the temperature measurements were made in 1910 and 1916. As will be noted in the comparison on the next page of these previous measurements with ours, there has been no significant temperature changes in the springs. Our spring 921 is possibly not the same as Barton's spring and may be $\frac{1}{4}$ - $\frac{1}{2}$ mile to the south of it. We locate 921 at SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 26, T4S, R25W. The references list it as "along sec. 23-26 line.

<u>Present No. Spring</u>	<u>Past Name</u>	(1) <u>T^oC(year)</u>	(2) <u>T^oC(year)</u>	<u>Present Work, T^oC</u>
945	(Little Missouri River Bank)	23.3(1916)	-	23.0
915	(Bed of Caddo River, North Spring)	35.0(1910)	-	35.0
		34.4(1916)		
	(Bed of Caddo River, South Spring)	35.4(1910)		
		36.0(1916)		
921	W. B. Barton	listed, no T	31.7(?)	30.0
924	Redland Mountain	25.0(1916)	-	25.0

(1) Miser and Purdue, 1929; Bryan, 1922

(2) Billingsley, et al.

Miser and Purdue (1929) list some chemical data on two of the springs which are compared below to our data. All data are in ppm.

<u>Analysis</u>	<u>Caddo River Springs</u>			<u>Barton's Spring</u>	
	<u>N. Spring</u>	<u>S. Spring</u>	<u>Our Data</u>	<u>Miser and P.</u>	<u>Our Data</u>
SiO ₂	18.7	15.0	17.2	15.4	10.5
Ca	41.7	38.9	42.0	39.8	50.4
Mg	4.2	2.2	2.45	2.3	2.38
Na	3.4	7.5	5.4	9.1	3.83
K	0.000	0.000	1.00	1.3	1.10
SO ₄	2.32	1.42	0.70	8.3	3.2
Cl	7.1	4.8	3.5	8.0	2.5
Li	-	-	-	0.6	0.0046

The data on the Caddo River Spring are in close agreement, except for K and SO_4 . Only Mg and K are in close agreement for Barton's Spring. The literature data acknowledged some contamination with other waters during the Barton Spring sampling. Also, as pointed out earlier, our spring 921 while in the same area as Barton's Spring may be $\frac{1}{4}$ - $\frac{1}{2}$ mile south of it.

It is of interest to compare our data in a very general way with that of the hot and warm springs near Hot Springs, Arkansas. The hot springs have a uniform SiO_2 concentration of about 42 ppm which corresponds to a subsurface temperature of 94°C . (Bedinger, et al., 1974). The average concentration of SiO_2 for the warm springs (i.e. greater than 20°C) near Hot Springs was 23.6 ppm, corresponding to an average source temperature of 70.6°C . These warm springs have a fairly uniform elemental concentration which indicates a common source. The warm springs in our study areas varied substantially in chemical composition which indicate a variety of sources.

4. Summary

The area around Caddo Gap has the springs with the higher surface temperatures (30 and 35°C). Other warm springs, Redland Mt. (25°C) and a "warm" spring (23°C) are the next hottest. Miser first measured the temperatures of these springs in 1910 and 1916. The temperatures are the same today (1979).

Subsurface temperatures of the various spring waters based on the silica geotemperature were highest in the Fancy Hill barite district (911W with $91.2^\circ\text{C } T_{\text{SiO}_2}$), next highest in the Lake Greeson mercury district (907W with $89.2^\circ\text{C } T_{\text{SiO}_2}$) next highest in the Caddo Gap area (923W with

82.9°C T_{SiO_2}) and next highest in the antimony area of Sevier County (951W with T_{SiO_2} of 78.3°C). The preponderance of wells in this list is to be noted.

These data are of geologic interest and are discussed further in the Exploration Section III. In economic terms the hottest springs might be of balneological or direct heating use. They cannot be used for power generation which requires about 180°C for turbine power generation (Muffler, 1973).

III. EXPLORATION

1. Introduction

Stream sediments and soils have received much attention in geochemical exploration for mineral deposits. In general, mineral deposits have indicators in nearby sediments and soils either through clasts or chemical solutioning and reprecipitation. Chemical analyses of sediments and soils detect the element of interest and the vector of the amount of the element points, in theory, toward the deposit.

Analyses of surface and ground waters have likewise been used for exploration (see p. 271 of Hawkes and Webb, 1962). Water samples are generally very dilute in mineral content compared to solid samples and analytical schemes are strained. Surface waters have an advantage of traveling far and with strategic sampling can be used to broadly survey a large drainage basin. Surface waters have a disadvantage of being in contact with air and converting some metallic ions, Fe^{2+} and Mn^{2+} for example, into the less soluble "ic" forms. It is not unusual to see this process in operation via the precipitation of iron and manganese oxides where a seep or spring flows into air or a stream. Precipitation of Fe_2O_3 or MnO_2 coprecipitates many trace metals and removes them from solution.

For geochemical exploration, ground water has certain advantages over surface waters: 1) lower Eh and greater solubility for certain metallic ions in the "us" form; 2) longer contact with rocks and soils with greater chance of equilibrium and higher metallic content;

3) higher CO₂ content due to organic decay processes which give it a lower pH and more dissolving potential; 4) more concentrated in mineral content due to no dilution by direct rainfall or surface runoff and, in forested areas, a 2X to 3X concentration due to evaporation and transportation; 5) in a faulted or steeply dipping strata, spring waters may have a deeply buried conduit and in the process "sample" buried rock structures and minerals which have no surface exposure. This last advantage is also a disadvantage for springs in that all the rock units to which spring waters have been exposed are unknown.

In the present work about 100 springs were sampled in west-central Arkansas (see Figure 1-1). This is a highly mineralized area as described above in the "Geology" section and consists of two areas of distinct mineralization. Thus, an opportunity is provided for testing the potential of spring waters in these areas as an exploration tool. Chemical analyses of the waters are summarized in Tables 1-2 through 1-4.

Analyses were made for the following elements:

- 1) Base metals (Fe, Mn, Zn, Cu, Co, Ni, Pb, Hg, and Sb)
- 2) Alkaline earth metals (Ca, Mg, Sr, and Ba)
- 3) Alkali metals (Na, K, and Li)
- 4) Anions (NO₃, NH₃, PO₄, SO₄, Cl, and SiO₂)

In addition, the following measurements were made: water temperature, pH, specific conductivity, and alkalinity.

The scheme for testing spring water analyses as an exploration tool is to compare the location of springs of anomalously high concentrations of metal content with the location of known mineral deposits of the same metal. This will be done on a district (Areas I and II), as well as on a fractional section basis using Stroud et. al. (1969) for the location of mineral deposits.

We are using section, township, and range for locating springs and mineral deposits. Spring locations are based on the latest U.S.G.S. topographic maps and field observations. Stroud, et. al. (1969) also uses section, township, and range for locating mineral deposits. Any other system would be awkward as would any distance unit other than mile, due to sections being 1 mile square. All distances given in the discussion of individual springs are in miles.

To determine anomalous values, the concentration of each metal for the various springs was plotted graphically as cumulative frequency curves after Lepeltier (1969). Anomalous values were those exceeding threshold, the 95% value on the high side. Figures III-1 through Figure III-5 compare several of these cumulative frequency curves for Areas I and II. Figures III-6 through Figure III-10 show several additional cumulative frequency curves for the areas. These curves shown are illustrative of the types of curves exhibited by Areas I and II and are not meant to be complete.

Most of the cumulative frequency curves for Areas I and II show two segments of similar slopes which are slightly offset. This indicates that there are two types of aquifers in the areas with slightly different

chemistries. Several reasons such as different lithologies, depths, and temperature could explain the different chemistries. Although the slopes and intercepts are different for Areas I and II, the general shape of the curves for a given element are remarkably similar for the two areas. An exception is the PO_4 curve for the Area II which is quite linear with a single branch. Area I curves are based on 71 data sets whereas Area II curves are based on only 31. It is of interest to note that in both areas the data follow a lognormal distribution which is a requirement for linearity in this type of plot. To our knowledge, the lognormal distribution of chemical data on a natural water system has not been previously reported.

Threshold values and anomalous values determined from the above plots are summarized by analysis or measurement in Table III-1. Median, threshold, and anomalous values are compared for Area I and Area II, element by element in this table.

Before examining Table III-1, the different mineralizations of the two areas are worth repeating. Area I has Mn, Fe, Ba mineralization with minor Cu, Pb, and Zn mineralization. Area II has Hg, Sb, Sr, and Ba mineralization with minor Zn and Pb mineralization.

2. Background Values of Non-Mineralizing Elements

Median values which reflect the backgrounds for Areas I and II are compared below for temperature, pH, specific conductivity, and alkalinity with the ranges of the values given in parenthesis.

	Water Temperature (°C)		pH Units	Sp. Cond. μ mhos/cm(25°C)	Total Alkalinity mg/L CaCO ₃
	Surface	Subsurface*			
Area I	17(5-35)	40.0(7.1-91.2)	6.4(3.8-7.9)	107(14-385)	60(0-291)
Area II	16(8-25)	45.9(11.2-89.2)	5.8(4.1-8.7)	61(<5-563)	15(0-265)
II/I (%)	94	115	91	57	25

*Calculated from SiO₂ content.

The median subsurface water temperature is slightly higher for Area II, but the hotter individual springs were in the eastern part of Area I. Other parameters in the above table are comparable or lower for Area II. The lower median pH for Area II should favor higher concentrations of most elements. This is countered by the higher electrolyte content (and thus lower activity coefficients) of Area I as indicated by its higher median specific conductivity. Higher electrolyte content would favor the solubility of barite, which was found in several cases to be near its saturation value.

Anion concentrations for the two areas are compared in the following table which gives median values and ranges in parenthesis.

Anion Median (Range) Concentrations (ppm)			
	<u>NO₃ as N</u>	<u>NH₃ as N</u>	<u>PO₄ as P</u>
Area I	0.66(<0.01-5.5)	0.1(<0.01-1.52)	0.07(<0.01-0.54)
Area II	1.0(0.06-14.3)	0.2(0.02-1.52)	0.10(0.02-0.41)
II/I (%)	151	200	143
	<u>SO₄</u>	<u>Cl</u>	<u>SiO₂</u>
Area I	2.3(<0.3-32)	2.3(0.25-87)	10.3(3.2-34.9)
Area II	3.2(0.3-55)	3.5(0.5-9.5)	12.1(5.4-37.1) *
II/I (%)	139	152	117

In all cases, Area II has the greater median anion concentration but the ranges are comparable, except for chloride. For NO_3 , NH_3 , and PO_4 this is most likely due to the greater agricultural and silvi-culture activities of Area II.

Other ions to be considered for background purposes for the present study are shown below. Median values, with ranges in parenthesis, are shown for Areas I and II.

Alkali and Alkaline Earth Ions, Median (Range) Concentrations

	<u>Na</u>	<u>K</u>	<u>Ca</u>	<u>Mg</u>
Area I	1.8(0.95-152)	0.56(0.04-2.9)	13.5(0.03-70)	1.9(0.15-14.9)
Area II	3.9(0.85-162)	1.10(0.21-8.5)	2.7(0.6-78)	1.4(0.19-23.7)
II/I (%)	217	196	20	74

Ranges for these four ions are comparable for the two areas but median values for Na and K for Area II are about twice those for Area I and for Mg and Ca much less for Area II.

In summary, for the non-mineralizing ions in the spring waters of the two areas, no order of magnitude differences are found. The greatest difference is in the median value of Ca ion which for Area I is five times the value in Area II. It is generally the most abundant ion in Area I, controls alkalinity (through HCO_3), specific conductance, and ionic strength. In Area II, Na competes more strongly with Ca and Mg ions for control of these parameters.

3. Mineralizing Elements

a. Base Metals: Median values, with ranges in parenthesis, are given below for nine base metals found in the spring waters for Areas I and II. Seven of these, Fe, Mn, Zn, Cu, Hg, Sb, and Pb, have primary

mineral deposits in at least one of the areas. All, except Pb, Hg, and Sb, may be associated with Mn oxide minerals in Area I. Median values are given with ranges in parenthesis.

Base Metal Median (Range) Concentrations (ppb)					
	<u>Fe</u>	<u>Mn</u>	<u>Zn</u>	<u>Cu</u>	<u>Co</u>
Area I	56(2-5220)	19(1-1220)	18(2-144)	4(1-85)	5(<2-31)
Area II	49(7-10082)	14(4-370)	9.5(2-48)	3(4-14)	11(2-39)
II/I (%)	87	74	53	75	220
	<u>Ni</u>	<u>Hg</u>	<u>Sb*</u>	<u>Pb</u>	
Area I	9(2-39)	<0.1(<0.1-2.1)	<0.2(<0.2-0.3)	10(1-37)	
Area II	20(2-54)	0.2(<0.1-1.2)	<0.5(<0.2-0.5)	30(<10-59)	
II/I (%)	222	>200	--	300	

*Based on a limited number of only 8 samples for Area II

For Fe, Mn, Zn and Cu, Area II Has a lower median and a lower range (except for Fe). For the other five metals the ranges are comparable for the two areas but with Area II having a higher median. Considering the differences in sample densities, 71 samples for Area I and 31 for Area II, and the analytical difficulty in these low ppb ranges, these differences are judged to be small. Manganese which is decidedly more abundant as manganese oxide minerals in Area I does have a much higher range in Area I, but only about a 25% greater median. Hg which has a definite mineralized area in Area II, but none in Area I has a higher median but lower range in Area II. Threshold values are decidedly greater for Fe, Mn, Zn, and Cu in Area I as shown in the following table.

Base Metals, Threshold Concentrations*(ppb)

	<u>Fe</u>	<u>Mn</u>	<u>Zn</u>	<u>Cu</u>	<u>Co</u>	<u>Ni</u>	<u>Pb</u>
Area I	5000	900	100	11	20	28	31
Area II	580	220	33	7	32	37	42
II/I (%)	12	24	33	64	160	132	135

*95% values on the high side of cumulative frequency curve, except Hg=1 ppb.

Hg and Sb are not included in the above table. They have too few finite analytical values for each area for a cumulative frequency treatment. Arbitrarily, we define the threshold value for Hg as 1 ppb. Sb will not be treated further in this discussion except to say that in spite of good definite Sb mineralization in Area II we have been unable by our chemical methods to detect any significant finite amounts of Sb in the spring waters of Area II.

The number of anomalous springs and their average anomalous concentrations are summarized in the following table for Area I and II.

Base Metals Average Anomalous Concentration (ppb)
(Number of Anomalous Springs in Parenthesis)

	<u>Fe</u>	<u>Mn</u>	<u>Zn</u>	<u>Cu</u>	<u>Co</u>	<u>Ni</u>	<u>Hg</u>	<u>Pb</u>
Area I	5222(2)	1220(1)	134(4)	34(4)	27(4)	36(4)	2.0(2)	35(2)
Area II	6399(4)	353(3)	46(2)	12(3)	38(2)	54(1)	1.1(3)	59(1)
II/I (%)	122	29	34	35	141	150	55	169

Area I has the higher anomalous values for Mn, Zn, and Cu. In summary, a higher metal content for Area I is indicated for Fe, Mn, Zn, and Cu when using median, threshold and average anomalous concentrations

(except Fe) as criteria. Manganese which is well documented to have more mineralization in Area I gives the best indication of this in the threshold and average anomalous values, not the median value. Area II has a higher median for Hg but lower anomalous average.

We now examine specific springs for the correlation of their water chemistry with the local mineralization. Table III-1 lists by analysis the springs in Area I and II which are anomalous for that analysis. Table III-2 for Area I and III-3 for Area II lists the springs in order of sample number which have anomalous analyses. Those springs having anomalous base metal values are summarized in the following table along with any other anomalous measurements or analyses for that spring. The anomalous elements and their concentrations are listed.

Springs with Anomalous Base Metal Concentrations

Area I

Spring No. or Well (W)	Related Mineralization in Local Area of Spring*	<u>Measurement (Value) for Which Spring Is Anomalous</u>
511	A	Cu (85 ppb)
513	D	Cu (15 ppb), SiO ₂ and subsurface temperature (63.2 ⁰ C)
517	B	Mn (1220 ppb), Cu (21 ppb), Ni (29 ppb), PO ₄ (0.54 ppm), SO ₄ (12 ppm)
520W	A(Ba)	Pb (33 ppb), NH ₃ (0.64 ppm), Li (23 ppb), Ba (420 ppb)
523	B	Cu (14 ppb), surface temperature (22 ⁰ C)
528	D	Hg (21 ppb)
536	A	Zn (143 ppb)

*Using Stroud, et al. (1969) and Miser (1917)

A = within 1 mile of mineralization related to the anomalous metal concentration
 B = " 2 " " " " " " " " " "
 C = " 4 " " " " " " " " " "
 D = over 4 miles from " " " " " " " "

<u>Spring No. or Well No.</u>	<u>Related Mineralization in Local Area of Spring*</u>	<u>Measurement (Value) for Which Spring Is Anomalous</u>
537	B	Zn (110 ppb)
544	D	Hg (1.9 ppb)
900	A	Zn (137 ppb)
911	C	Zn (144 ppb), PO ₄ (0.28 ppm), Sn (500 ppb)
916W	A	Fe (522 ppb), Co (23 ppb), Ni (39 ppb), Pb (37 ppb), SO ₄ (13.1 ppm), subsurface temperature (91.2 ⁰ C)
938	A	Co (22 ppb), Ni (36 ppb)
939	A	Co (30 ppb), NO ₃ (4.8 ppm)
943	B	Co (31 ppb), Ni (39 ppb)
Area II		
904	D	Cu (14 ppb), Zn (48 ppb), pH (4.1 units), NO ₃ (14.3 ppm), Cl (9.5 ppm)
908	A	Fe (10082 ppb), Co (39 ppb)
909	A	Hg (1.1 ppb)
910	B	Hg (1.2 ppb), SiO ₂ and subsurface temperature (62.9 ⁰ C)
919	D	Cu (13 ppb), surface temperature (21 ⁰ C)
920	A	Cu (9 ppb), subsurface temperature (68.8 ⁰ C)
924	A	Mn (345 ppb), surface temperature (25 ⁰ C), SiO ₂ and subsurface temperature (60.4 ⁰ C)
928	A	Fe (4218 ppb), Mn (345 ppb), SiO ₂ and subsurface temperature (75.8 ⁰ C)
932	D	Fe (1375 ppb), Pb (59 ppb), SiO ₂ and subsurface temperature (73.4 ⁰ C)
936	D	Mn (370 ppb), Co (38 ppb), Ni (54 ppb), SiO ₂ and subsurface temperature (70.0 ⁰ C)
937	D	Fe (9930 ppb), Li (86 ppb), pH (8.7 units), sp. cond. (563 μ mhos), alkalinity (265 mg/L CaCO ₃)
948	D	Zn (45 ppb), pH (4.1 units), Cl (8.0 ppm)

Stroud, et al. (1969) has a comprehensive listing of the known mineral prospects and mines for Areas I and II and their locations by 1/16 sections. Table III-4 lists the 1/16 section locations for the springs sampled in this study. Using Table III-4 and Stroud, et al. (1969) the anomalous springs in the above table will be checked for known mineralization in their local area.

For spring 511 located NE 1/4 SW 1/4 Sec. 14, T3S, R29W we quote from Stroud p. 329, "Fractures in Arkansas Novaculite in T3SRs28-30W, and T4S, R31W, contain stains and thin seams of malachite, azurite, chrysocolla, chalcopyrite, and native copper".

Although there is no known mineralization in spring 517's section, there is a phosphatic manganese deposit 3 1/2 miles due south.

520W is in a barite district. Pb mineralization is unknown locally, but is represented by several small mines 10 miles due north at Silver, Arkansas.

Spring 523 is not in an area of known Cu mineralization, However, a small galena mine, now abandoned, is 2 1/2 miles due west. Spring 528 is in a Mn district and within 1 mile of a hematite deposit (Pointed Rock Tunnel) and at least one Mn prospect (E. S. White) containing psilomelane. No Hg minerals are known in this area.

Spring 536 is in the same 1/16 section as an abandoned Mn mine (manganite and psilomelane minerals). No Zn mineralization is known in the local area, but manganese oxide minerals, particularly cryptomelane, can have 1 wt. % of Zn (Wagner, et al., 1979).

Spring 537 is not in an area of known Zn mineralization, but is only 1 1/2 miles northwest of spring 536. The comments on 536 would apply here as to a source of Zn in cryptomelane.

Spring 544 is within 1 1/2 miles of two known Mn oxide mineral deposits. One 1 1/2 miles to the west has produced several tons of ore. The deposit to the southwest is associated with limonite. No Hg minerals are known in the area.

Spring 900 is within a mile of an abandoned Pb, Zn, and copper mine and mill.

Spring 911 is not in an area with local Zn mineralization. However, manganese deposits with possible Zn are located 4 miles to the south.

Artisian well 916W is in a barite district. Manganese minerals occur with the barite at the McKnight Claim, 1 mile away. The Mn minerals contain appreciable Fe, Co, and Ni.

Spring 938 is located 1 mile southeast of the abandoned Featherstone Mn Mine (psilomelane and limonite). Co and Ni are known trace metals; up to 1 wt. % in many Mn minerals of the area (Wagner, et al., 1979).

Spring 939 is 1 mile northeast of a Mn prospect (U.S. prospect) which might contain Co. A similar Mn prospect is 1 1/2 miles to the southeast.

Spring 943 is in a barite district with no locally known Co or Ni mineralization. Mn ores which might contain Co and Ni are known 3 miles northeast and 3 miles northwest.

Spring 904 has no known minerals adjacent, but it is 2 1/2 miles north of the kimberlite at Murfreesboro. This deposit of peridotite

has some barite veins associated with it, but no known Cu or Zn minerals.

Spring 908 is in a mercury district with ore minerals being cinnabar and associated minerals of Fe (pyrite and marcasite), Cu (sulfides) and As (arsenides). Co due to a similar chemistry to Fe would probably come from the Fe minerals.

Spring 909 is within 1/2 mile of a Hg mine (section 32 mine) which has produced over 600 flasks of Hg from cinnabar.

Spring 910 is 4 miles south and down dip of a Hg district.

Spring 920 is in a barite and mercury district. Copper sulfides are found with cinnabar ores.

Spring 924 is less than 1/4 mile from two Mn prospects (psilomelane, manganite and iron oxides).

Spring 928 is on the southeast edge of a titanium sands and clay district with ilmenite (FeTiO_3) the main ore minerals. The sands have a ferrogenous cement.

Springs 936 and 937 on the northern edge of a Hg district with cinnabar the main ore mineral and small amounts of Cu (sulfides), Fe (pyrite and marcasite) and As (arsenides). Mn minerals which could account for the Mn, Co, and Ni are 10 miles north. Sb minerals (stibnite) are included in some of the Hg deposits.

Spring 948 has no known mineral deposits nearby. Pb and Zn minerals (sphalerite) are known in a prospect 9 miles to the northwest. Sb minerals are known 7 miles to the northwest in an antimony district. Barite minerals are known 3 miles to the south as a cement in sand and gravel.

In summary, 41% of the springs with anomalous base metal concentrations were within 1 mile, 18% were within 2 miles, and 7% within 4 miles of known mineralization containing the anomalous elements. The remainder of the springs, 31%, had no known related mineralization within 10 miles.

b. Barium: Barite ($BaSO_4$) is prevalent in both Area I and Area II. The deposits differ in type. In Area I the barite ore is either replacement or open space fillings in Stanley Shales. Area II has this type plus two other types: 1) barite veinlets penetrate peridotite and intrusive breccia of Cretaceous igneous rocks, 2) barite veins, concretions, and cement in gravels and sandstone as a principal mineral. Area I contains several Mn oxide mineral deposits with psilomelane $(Ba, H_2O)_4Mn_{10}O_{20}$.

In the following table, medians, ranges, thresholds, and the average of the anomalous concentrations are summarized for Areas I and II.

Ba Concentrations (ppb) in Spring Waters of Areas I and II

	<u>Median (Range)</u>	<u>Threshold</u>	<u>No. of Anomalous Values</u>	<u>Average Anomalous Concentrations</u>
Area I	<15(1-420)	320	1	420
Area II	40(10-920)	320	2	915
II/I (%)	>267	100	-	218

Area II has a higher median, range, and average anomalous value than Area I. Both areas have few springs with anomalous concentrations considering the widespread mineralization of barite in each area. Both areas have similarly shaped cumulative frequency curves (Figure III-4) and the same threshold (95% value on the high side).

Barite is only very slightly soluble in pure water (2.6 ppm of BaSO_4 which is equivalent to 1.5 ppm of Ba^{++} and 1.1 ppm of SO_4^- at 25°C). The solubility product, defined as follows, is a constant.

$$M_{\text{Ba}^{++}} \times M_{\text{SO}_4^-} = \text{constant} = 1.24 \times 10^{-10} \text{ (at saturation, } 25^\circ\text{C)}$$

where $M_{\text{Ba}^{++}}$ and $M_{\text{SO}_4^-}$ are gram moles/kg

A consequence of this relationship is that when SO_4^- concentrations increase the saturation concentration of Ba^{++} decreases. For example, in a solution containing 11 ppm of SO_4^- at 25°C , the maximum or saturation concentration of Ba^{++} becomes 0.15 ppm.

In water containing other ions the solubility product becomes an activity product, defined as follows:

$$a_{\text{Ba}^{++}} \times a_{\text{SO}_4^-} = \text{constant}$$

$$a_{\text{Ba}^{++}} = \gamma_{\text{Ba}^{++}} \times M_{\text{Ba}^{++}}$$

$$a_{\text{SO}_4^-} = \gamma_{\text{SO}_4^-} \times M_{\text{SO}_4^-}$$

(1)

$\gamma_{Ba^{++}}$ and $\gamma_{SO_4^{=}}$ are the activity coefficients respectively of barium ion and sulfate ion. The solubility of barite increases in the presence of other ions due to a decrease in the activity coefficient. The amount of other ions is measured by ionic strength. See Hem (1970) for a more complete discussion of the relationship.

In view of the dependence of Ba^{++} concentration on sulfate concentration, presence of other ions (ionic strength) and, of course temperature, high concentrations of Ba^{++} alone may not be a good relative indicator for barite among various localities. This would be particularly true if the solutions are near saturation and equilibrium is a limiting factor on solubility.

For the above reasons the % saturation of the various spring waters with $BaSO_4$ were calculated using the following relationship which considers the factors of sulfate concentration, ionic strength, and temperature.

$$\% \text{ saturation} = 100 (a_{Ba^{++}} \times a_{SO_4^{=}})_T / (\text{solubility product of } BaSO_4)_T$$

where the subscript T indicates that activities and solubility products are for the same temperature. Activity coefficients were read graphically from page 21 of Hem (1970) using ionic strengths calculated from our analyses for the various ions. Substituting activity coefficients and the measured concentrations of Ba^{++} and $SO_4^{=}$ into equation (1) above gave the corresponding activities. The solubility products used as a function of temperature were taken from the literature and are shown geographically in Figure III-11.

The % BaSO₄ saturation at the water temperature at the surface and at the subsurface were calculated. Subsurface temperatures of the spring waters were calculated from silica solubilities as described in section II of this report. In the following table, the medians, ranges, thresholds, and average anomalous values for % BaSO₄ saturation for Areas I and II are summarized.

% BaSO₄ Saturation of Spring Waters for Area I and II

	<u>Median (Range)</u>		<u>Threshold</u>	<u>No. Springs Anomalous</u>	<u>Average Anomalous %</u>
	<u>Surface T</u>	<u>Subsurface T</u>	<u>Surface T</u>	<u>Surface T</u>	<u>Surface T</u>
Area I	2(0.16-96)	1.1(0.09-39)	70	2	91
Area II	8(0.40-97)	6.5(0.24-38)	84	2	94
II/I (%)	400	591	120	100	103

Only finite values ("less than" values not used) were used in deriving the above table. This gave 31 data sets (out of 71 possible) for Area I and 27 sets (out of 31 possible) for Area II. The above analysis may be considered a cumulative frequency curve analysis of the highest population for Area I.

Area II has the greater % saturation median and threshold. This is possibly a consequence of the highly disseminated nature (BaSO₄ cement) of some of the Area II barite deposits whereas those in Area I are veinlets and less exposed to groundwater. The ranges, number of anomalous values and average anomalous values are about the same for the two areas.

It is interesting to note that many of the spring waters have X% or more of BaSO₄ saturation. Note in Table III-5 that springs are saturated with SrSO₄ to median saturation of only 0.00X%.

Figures III-12 and III-13 show a broadly defined correlation between the concentration of Ba and Ca in Areas I and II. Due to the similar chemistry of Ba and Ca and perhaps similar source rocks, this might be expected. Those few points which lie outside the envelopes of these data are considered to represent anomalous springs. The same goes for springs whose data plot outside the envelope of the Ba⁺⁺-Sr⁺⁺ concentration data (Figure III-14) and Ba⁺⁺ concentration vs specific conductivity data in Figure III-15.

Five definitions of anomalous springs have been given above. Using these five definitions, the anomalous springs are summarized below:

<u>Sample No. of Springs with Anomalous Ba</u>		
<u>Definition of Anomalous</u>	<u>Area I</u>	<u>Area II</u>
1. Ba ⁺⁺ concentration exceeds threshold	520W	929, 944
2. % BaSO ₄ saturation exceeds threshold	916W, 943	927W, 932
3. Ba ⁺⁺ -Ca ⁺⁺ concentration outside area norm	930	908, 929, 944
4. Ba ⁺⁺ -Sr ⁺⁺ concentration outside area norm	902	908, 929, 944
5. Ba ⁺⁺ -specific conductivity outside area norm	930	929, 944

520W is an artesian well in a barite district (Pigeon Roast Mountain). It is interesting to note that spring 519 in the same 1/16 section is not anomalous in Ba but has 110 ppb which is well above the mean of <15 for Area I.

916W is in the barite district of Fancy Hill. Although the % BaSO_4 saturation is anomalous the concentration is not. However, the 200 ppb concentration is well above the <15 ppb mean for the area.

Spring 943 is in the Pigeon Roost Mountain barite district and in the same 1/16 section as 520W. The Ba^{++} concentration of spring 943 while not anomalous is high, 190 ppb.

Spring 930 is in an area of no local mineralization. However, the Pike Gravels are present which 5 miles to the southwest have a barite cement in appreciable quantities. Three miles to the southeast of 930 there are barite veins in periodotite, the Murfreesboro kimberlite. Spring 930 has 60 ppb of Ba.

Spring 902 is in the same 1/16 section as an abandoned mine for Pb, Zn, and Cu. Galena, sphalerite, pyrite, chalcopyrite, silver and gold are the minerals in quartz veins along a fault. The nearest barite deposits are 7 miles east in the Pigeon Roost District.

Springs 929 and 944 are the same spring, sampled at different times, 3-4-79 (929) and 5-16-79 (944). This spring abuts an abandoned mine based on barite cement in Pike Gravels. Interestingly, the spring is anomalous based on 3 of the 4 definitions. Only the % saturations of 45.1 and 63.1 are not anomalous. Perhaps the threshold value (84%) is too high a criterion.

Artesian well 927W is not in an area of known Ba mineralization. The nearest barite deposits are 15 miles to the north. Titaniferous

sands (ilmenite) are located 4 miles to the west of 927W. The Ba⁺⁺ concentration of 927W is 80 ppb.

Spring 932 has no known barite deposits nearby. The nearest deposit is 8 miles east. The Ba⁺⁺ concentration of this spring is high, 260 ppb, and almost to threshold (320 ppb).

Spring 908 is in a mercury district with no known barite deposits nearby. The nearest known deposits of barite are 10 miles northeast. Spring 907 which is 2 miles southwest of 908 (60 ppb Ba) had a reasonably high barium concentration of 270 ppb.

In summary, only definition (1) identified solely springs near or in barite deposits. Each of the other definitions selected springs in barite areas which were missed by definition (1). However, definitions (2), (3), and (4) also selected some springs not in known barite mineralization. All springs selected by each criteria had rather high concentrations of Ba and would have been selected by definition (1) with a smaller threshold value.

c. Strontium: Strontium mines and prospects are in the central part of Area II. Celestite (SrSO₄) is the main ore mineral with minor amounts of strontionite (SrCO₃). No strontium ores are known in Area I.

Median values, ranges, thresholds, and anomalous concentrations for springs of the two areas are summarized below.

Sr Concentrations (ppb) in Spring Waters of Areas I and II

	<u>Median (Range)</u>	<u>Threshold</u>	<u>No. of Anomalous Values</u>	<u>Average Anomalous Concentration</u>
Area I	55(2-2200)	380	5	987
Area II	30(<10-1400)	150	2	857
II/I (%)	55	39	40	87

All values are greater in Area I. This is not believed to be due to the small number of spring water samples from the strontium district in Area II. Four of the springs (933, 934, 935, 948) are within 2 miles of known Sr mineral deposits but four of the springs (904, 920, 929, 930) are only a few miles from the Sr district.

Strontium concentration in the spring waters increases as the calcium concentration increases as shown in Figure III-15. These two alkaline earths have similar chemistries and usually common source rocks. Thus, those concentrations outside the envelope of the data on the high side in Figure III-16 may be defined as anomalously high in Sr. Using this definition and the usual one of above threshold concentration gives the following springs which are anomalously high in Sr. We have not used % saturation of the waters with SrSO_4 as a criterion of anomalous springs because the waters are very unsaturated with SrSO_4 (see Table III-5), so much so that equilibrium saturation is not a barrier to dissolving. There is essentially no CO_3 in the spring waters.

Springs with Anomalously High Concentrations (ppb) of Sr

<u>Anomalous Definition</u>	<u>Sample No. of Springs</u>	
	<u>Area I</u>	<u>Area II</u>
1. Sr^{++} concentration exceeds threshold	501W, 526, 527, 911, 913	927W, 951W
2. Sr^{++} - Ca^{++} concentration outside area norm	526, 527, 913	920, 927W, 929, 944

Definition (1) includes two springs in Area I not selected by definition (2). The opposite is true in Area II.

Well 501W is not in an area of known Sr mineralization. Quartz crystal mines are within 1 mile.

Springs 526 and 527 are not in an area of Sr mineralization. A barite prospect in which Sr may be in secondary minerals, is about 1 mile away.

Spring 911 is not in an area of known Sr mineralization. About 4 miles south of spring 911 is a barite district.

Spring 913 is not in an area of known Sr mineralization. A barite district lies 6 miles south.

Spring 920 is not in an area of known Sr mineralization. A known barite deposit is less than 1 mile away. Known Sr deposits (celestite) are 6 miles to the southwest and gypsum deposits, in which Sr minerals occur to a minor extent, are 6 miles to the southeast.

Well 927 is not in an area of known Sr mineralization.

Spring 929 and 944 are the same spring sampled at different times. This spring is adjacent to a barite deposit.

The following are springs which did not have anomalous Sr concentration but are within 2 miles of known Sr mineralization: 933 (6 ppb), 934 (23 ppb), 935 (13 ppb), and 948 (92 ppb).

In summary, no spring with anomalously high Sr concentration was near known Sr mineralization. Rather, Sr acted as an indicator for barite in which Sr minerals occur. Figure III-14 notes a correlation of Ba with Sr.

d. Lithium and Potassium: Each of these metals exists in Area I as essential elements of two frequently encountered manganese oxide minerals -- lithiophorite ($\text{Li}_2\text{Al}_8\text{Mn}_{12}\text{O}_{35}\cdot 14\text{H}_2\text{O}$) and cryptomelane ($\text{K}_x\text{Mn}_8\text{O}_{16}$) (see Wagner, et al., 1979). Other sources for these two elements are clays and igneous rocks.

Listed below are the median values, ranges, thresholds, and anomalous concentrations for Li in springs of Areas I and II.

Li Concentrations (ppb) in Spring Waters of Areas I and II

	<u>Median (Range)</u>	<u>Threshold</u>	<u>No. of Anomalous Values</u>	<u>Average Anomalous Concentration</u>
Area I	2.5(<1-70)	18	4	46
Area II	2.3(<1-86)	54	2	72
II/I (%)	92	300	50	156

The median values and ranges are essentially the same for Areas I and II. Threshold and average anomalous concentrations are higher for Area II. There is no indication of higher Li concentrations in Area I resulting from or favoring the formation of lithiophorite.

Springs which have anomalously high Li concentrations are listed below.

Springs With Anomalously High Li Content

Spring No.	501W	520W	526	527	927W	937
Li Concentration (ppb)	21	23	70	70	58	86
Na Concentration (ppm)	48.4	6.5	131	152	80	162
Li/Na Weight Ratio ($\times 10^3$)	0.43	3.5	0.53	0.46	0.72	0.53
Li/Na Weight Ratio ($\times 10^3$) (Median for Area)	1.4				0.59	

520W, 526, and 527 of the above waters are near (1-2 miles) Mn mineralization and barite. The others are not near Mn mineralization. A common denominator of the above waters is a high Na content. When judged as a Li/Na ratio these high Li waters are not so abnormal.

Listed below are the medians, ranges, thresholds, and average anomalous values for the K content of springs in Areas I and II.

K Concentration (ppm) in Spring Waters of Areas I and II

	<u>Median (Range)</u>	<u>Threshold</u>	<u>No. of Anomalous Values</u>	<u>Average Anomalous Concentration</u>
Area I	0.56(0.04-2.9)	1.35	3	2.0
Area II	1.10(0.21-8.5)	6.0	1	8.5
II/I (%)	196	444	33	425

Area II has greater values in all categories, except number of anomalous values. There are no indications of higher Area I values due to the occurrences of cryptomelane in Area I.

The individual springs with anomalous K content are given below.

Springs With Anomalously High K Concentration

	<u>Area I</u>			<u>Area II</u>	
Spring No.	506	520W	922	948	904
K Concentration (ppm)	1.70	1.40	2.90	8.5	3.8

Spring 506 is within 1 mile of abandoned Mn mines. The Mn mineralization of this area contains cryptomelane. Slate deposits are also nearby. None of the other springs are in areas of Mn mineralization or have anomalous manganese concentrations.

4. Springs with Anomalous Temperatures

In Section II of this report the springs with anomalously high surface and subsurface temperatures were determined. Subsurface temperatures were determined based on the concentration of SiO_2 . These springs are summarized below along with their anomalous concentrations of metal ions, if any, and local mineralization, if any.

<u>Spring No.</u>	<u>Anomalous Temperature</u>	<u>Metals With Anomalous Concentrations</u>	<u>Known Local Mineralization</u>
503	Subsurface (60.4°C)	none	none
513	Subsurface (63.2°C)	Cu	none
523	Surface (22°C)	Cu	Pb, Zn mineralization in 2 miles
548	Surface (22°C)	none	Mn mineralization in 1-2 miles
907W	Subsurface (89.2°C)	none	Hg and Sb mineralization in 1-2 miles
910	Subsurface (62.9°C)	Hg	Hg mineralization in 1-2 miles
914	Subsurface (60.7°C)	none	none
915	Surface (35°C)	none	Mn mineralization in 1 mile
916W	Subsurface (91.2°C)	Fe,Co,Ni,Pb	Mn, Fe, Co, Ni, Bi mineralization in 1 mile
920	Subsurface (68.8°C)	Cu	Ba, Hg with minor Cu in 1 mile
921	Surface (30°C)	none	none
923W	Surface (22°C)	none	none
	Subsurface (82.9°C)	none	none
924	Surface (25°C)	Mn	Mn mineralization in 2 miles
	Subsurface (60.4°C)	Mn	Mn mineralization in 2 miles
927W	Subsurface (65.7°C)	Sr,Li	none
928	Subsurface (75.8°C)	Fe,Mn	ilminite sands in 1-2 miles

<u>Spring No.</u>	<u>Anomalous Temperature</u>	<u>Metals With Anomalous Concentrations</u>	<u>Known Local Mineralization</u>
932	Subsurface (73.4 ⁰ C)	Fe,Pb	none, Hg district 5 miles away
936	Subsurface (70.0 ⁰ C)	Mn,Co,Ni	Hg mineralization in 1-2 miles
945	Surface (23 ⁰ C)	none	Mn, Cu mineralization in 1-2 miles
951W	Subsurface (78.3 ⁰ C)	Sr	Sb mineralization in 1 mile
952	Surface (24 ⁰ C)	none	Pb, Zn, Sb mineralization in 2 miles

From the above table we find the following:

	<u>% With Anomalous Concentration Of Some Metals</u>	<u>% With Local Mineralization</u>	<u>% With Anomalous Concentration With Local Mineralization*</u>
Surface T Anomalous	25	75	100
Subsurface T Anomalous	71	57	60
All Other Springs, Areas I & II	31	--	54

*relevant minerals within 2 miles

Springs with anomalously high subsurface temperatures are:

1) more likely (71% frequency) to have an anomalous concentration of some metal than springs generally (31%) or springs with an anomalously high surface temperature (25%), 2) as likely (60% vs 54%) as other springs to indicate known, local mineralization by these anomalies. Thus overall springs with anomalously high silica content are more likely than other springs to indicate local mineralization.

5. Summary:

- a. A number of parameters and elemental concentrations for spring waters gave a lognormal distribution which were amenable to cumulative frequency curve analysis for median (50%), threshold (95% and anomalous (>95%) concentrations.

- b. Threshold and anomalous concentrations of Mn, Zn, and Cu in the spring waters of Area I were greater than in Area II and reflected better than median or concentration ranges the more frequent mineralization of these metals in Area I. Hg concentrations in spring waters of the Hg-mineralized Area II had a higher median value and springs with anomalously high concentrations were near Hg mineralizations. However, in non-Hg-mineralized Area I the two springs with anomalously high Hg concentrations had no known Hg mineralization nearby.
- c. Individual springs with an anomalous concentration of a base metal were in 41% of the cases within 1 mile, and 18% of the cases within 2 miles, of known mineralization of the anomalous element.
- d. Spring waters of both Area I and Area II were commonly saturated to X% and above with BaSO_4 (barite) whereas the median SrSO_4 saturation was only 0.00X%.
- e. Ba and Sr concentrations in the spring waters correlated with Ca concentration.
- f. High Ba concentrations in spring waters were more reliable for detecting nearby barite mineralization than criteria based on:
 - 1) % BaSO_4 saturation,
 - 2) Ba/Ca concentration ratio above area norm,
 - 3) Ba/Sr concentration ratio above area norm.

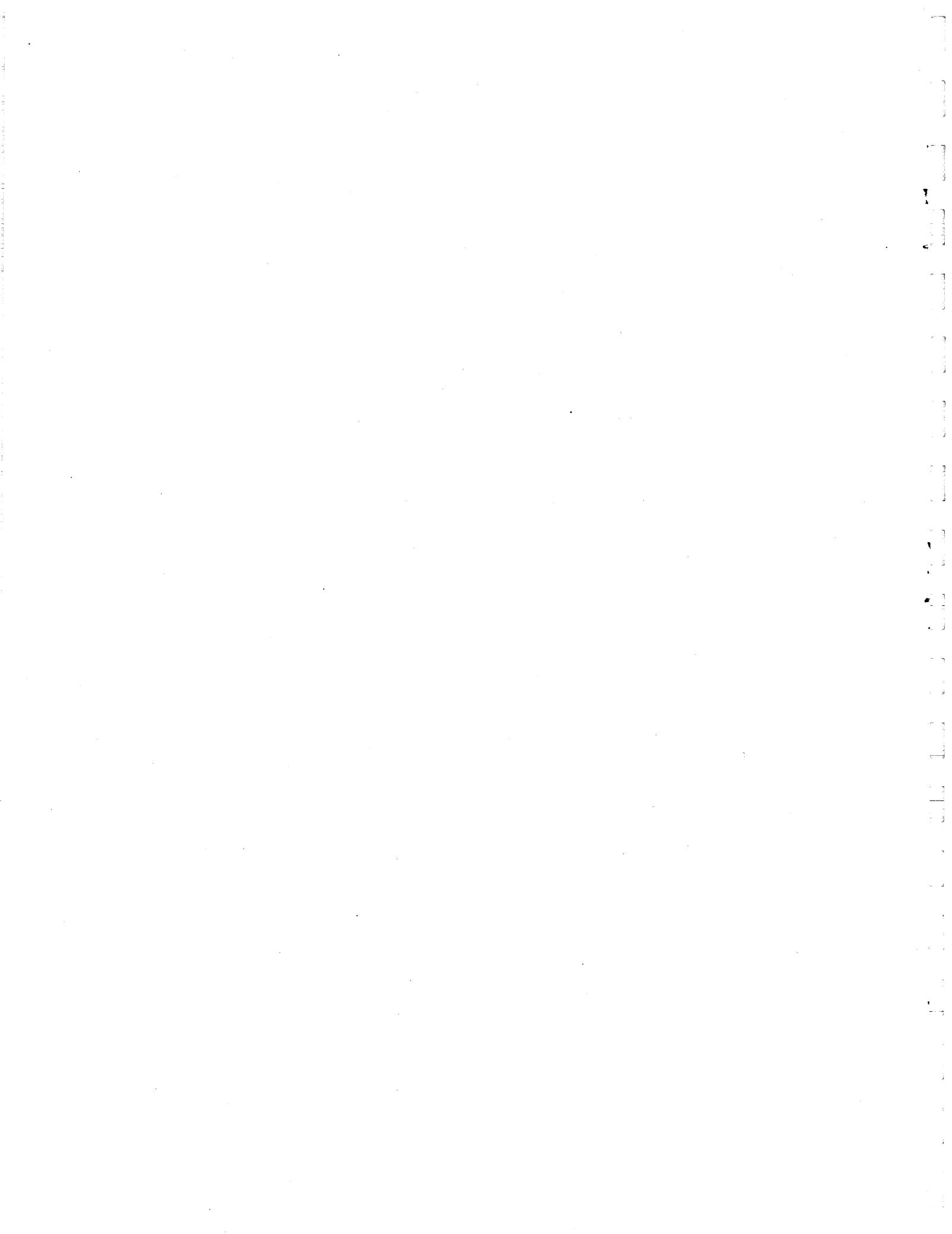
- g. More Sr mineralization in Area II was not reflected in greater median or threshold concentrations of the area's spring waters. Springs within 2 miles of known Sr mineralization (celestite) did not have anomalously high concentrations of Sr. Anomalous Sr concentrations were an index of nearby barite mineralization, in which Sr minerals occur as minor constituents.
- h. Concentrations of Li and K in Area I spring waters were not high and gave no indications of the frequently uncountered lithiophorite and cryptomelane of this area. Li tended to increase with Na in both Areas I and II.
- i. Springs with anomalously high silica content have a high incidence of anomalous high metal concentrations and are more likely than other springs to indicate local mineralization.

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TABLES

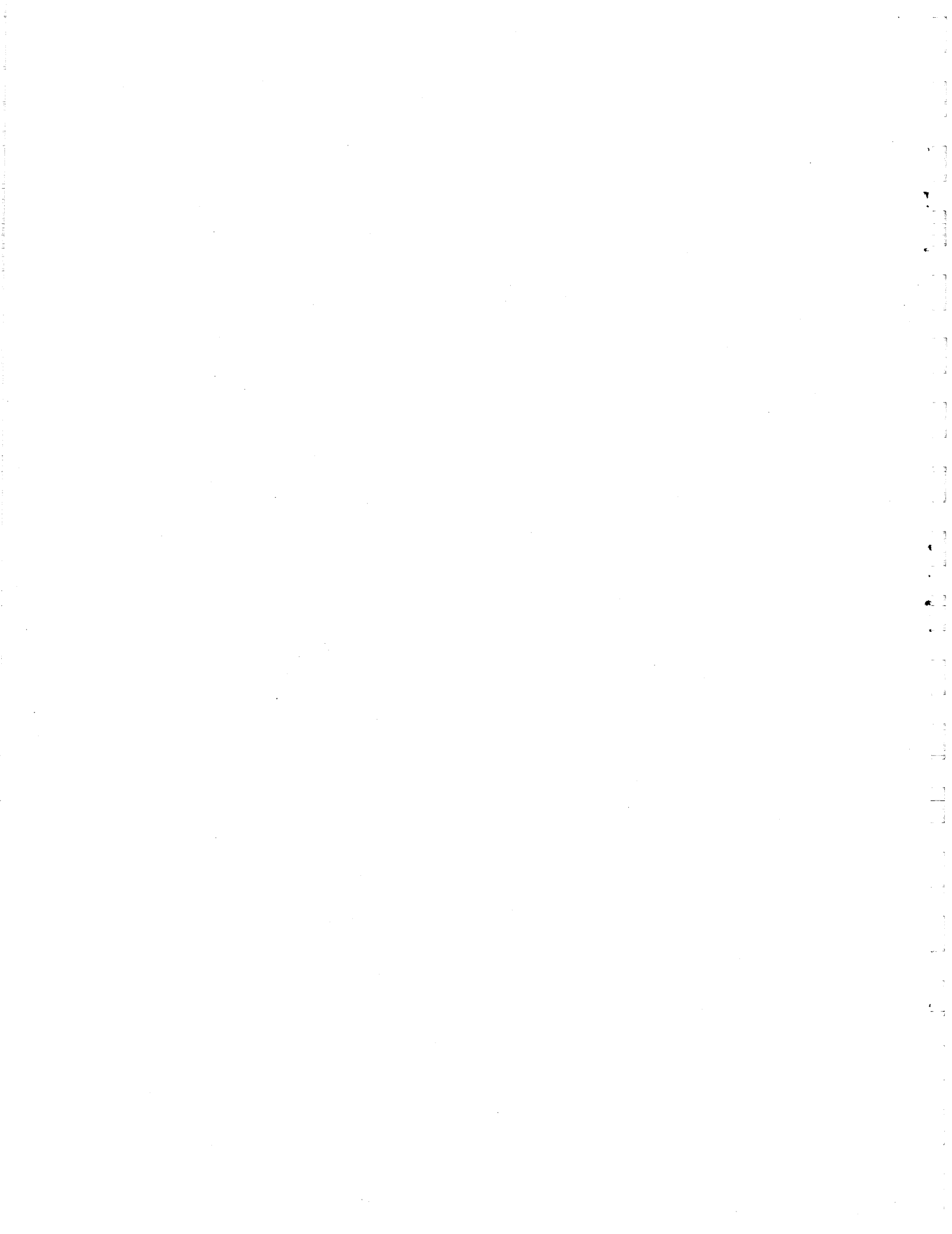


Table 1-1
Summary of Analytical Methods

<u>Measurement</u>	<u>Anions</u>	
	<u>Method</u>	<u>APHA (1971) Page</u>
NO ₃	Cd reduction	--
NH ₃	Nessler	226
PO ₄ (ortho, dissolved)	Ascorbic acid	532
SO ₄	Turbidimetric	334
SiO ₂	Colorimetric	303,336
Cl	Mercuric nitrate	97
Alkalinity(total)	Titration to methy red end pt.	52

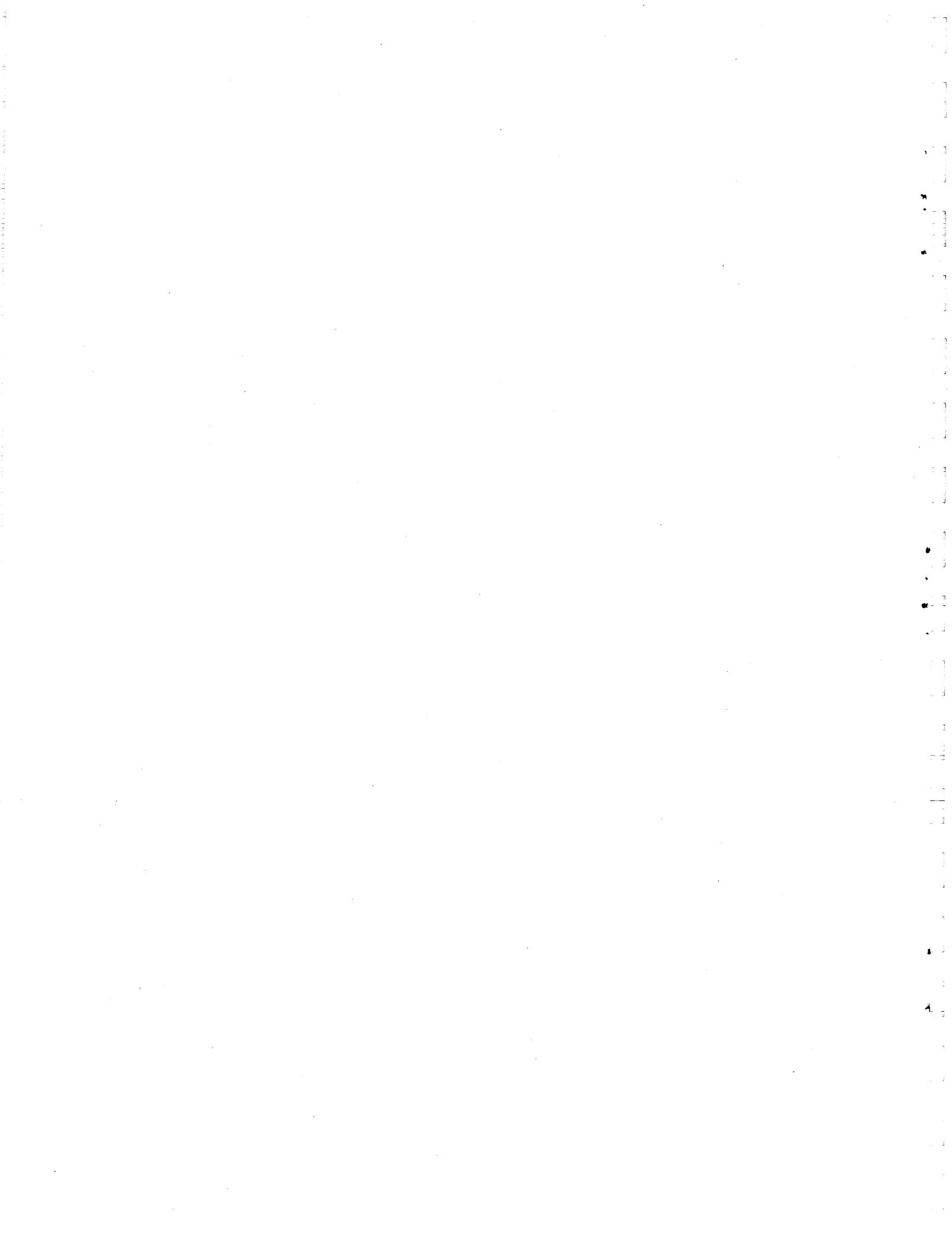


Table 1-1
Summary of Analytical Methods
(continued)

<u>Element</u>	<u>Instrument*</u>	<u>Cations</u>		
		<u>Mode</u>	<u>Flame</u>	<u>Additive</u>
Na	JA	emission	H ₂ - N ₂	none
K	"	"	"	"
Li	"	"	"	"
Ca	"	absorption	C ₂ H ₂ - N ₂ O	NaCl
Mg	"	"	"	"
Sr	"	emission	"	"
Ba	"	concentration	"	KCl
Fe	"	absorption	C ₂ H ₂ - Air	via chelation
Co	"	"	"	" "
Ni	"	"	"	" "
Cu	"	"	"	" "
Zn	PE	"	"	" "
Pb	PE	"	"	" "
Hg	PE	"	flameless	**
Sb	JA	hydride generation absorption	H ₂ - N ₂	***

PE = Perkin Elmer Model 303 Atomic Absorption Spectrophotometer

* JA = Jarrel Ash modernized Model 82-500 Atomic Absorption Spectrophotometer

** method of EPA(1974), p. 118

*** method of Fernandez (1973)

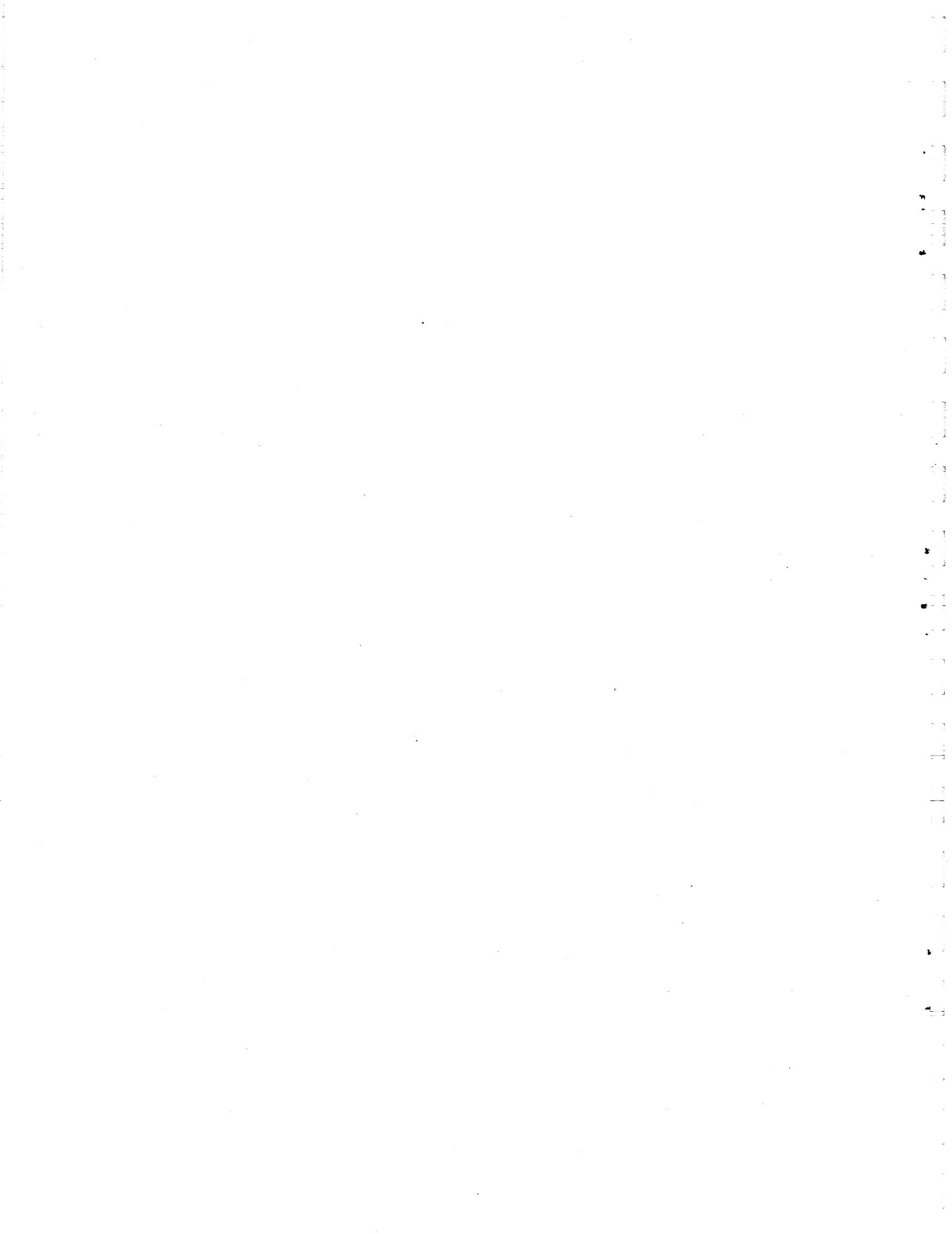


Table 1-2
 Chemical Analyses of Ouachita Spring Waters
 Anion Analyses and Miscellaneous Measurements

Name of Spring or Well (W)	Lab No.	Water T °C	pH units	Sp. Cond. μmhos per cm 25°C	Total Alkalinity mg/l as CaCO ₃	NO ₃ as N ppm	NH ₃ as N ppm	PO ₄ as P ppm	SO ₄ ppm	Cl ppm	SiO ₂ ppm
Iron (W)	501	15.0	7.4	252	291	2.28	0.11	0.03	0.3	0.3	7.7
Iron	502	5.0	6.5	14	2	<0.01	0.02	0.04	0.3	1.8	3.2
Black	503	15.0	7.5	324	10	<0.01	0.05	0.04	32.0	0.3	18.3
Collier	504	16.5	6.7	61	44	<0.01	<0.02	0.06	1.1	2.3	12.1
Caddo Valley	505	15.0	7.3	168	225	<0.01	<0.02	0.13	1.4	0.3	7.7
Crystal A	506	18.0	7.6	197	162	<0.01	0.42	0.08	1.1	0.3	12.1
Goat Brown	507	13.0	6.7	89	60	0.06	0.04	0.04	1.1	3.5	13.7
Bogg	508	13.0	7.2	372	130	0.06	0.02	0.05	2.3	0.5	7.7
Queen Wilhelmina S. Park	509	10.0	5.8	29	5	2.02	0.09	0.02	2.6	1.8	10.3
Queen Wilhelmina S. Park	510	9.0	6.4	66	2	2.68	0.13	0.05	0.6	3.8	6.6
Silver World	511	16.0	6.9	109	90	<0.03	0.06	0.03	4.7	1.5	10.2
Abernathy	512	16.0	7.0	144	70	0.37	0.09	0.12	2.0	1.0	8.9
Three Sisters	513	16.0	6.9	244	150	0.76	0.05	0.12	6.9	3.5	19.4
(No Name)	514	18.0	5.6	21	5	1.07	0.04	0.03	0.9	1.5	10.2
(No Name)	515	17.0	6.0	21	5	<0.02	<0.02	0.09	1.3	2.0	9.6
(No Name)	516	21.0	6.3	103	66	0.02	0.11	0.13	4.0	2.0	12.0
Wyatt	517	17.0	5.7	64	12	<0.02	0.25	0.54	12.0	2.5	12.0
(No Name)	518	15.0	5.8	18	5	0.65	0.12	<0.01	0.6	2.0	10.6
(No Name)	519	18.0	6.6	135	84	0.08	0.33	0.01	2.9	2.0	9.4
Pigeon Roost*(W)	520	17.0	7.2	208	160	0.02	0.64	0.03	2.6	1.8	13.5
Dripping	521	17.0	5.6	26	10	0.54	0.29	0.03	1.9	2.0	9.6
Strawn	522	18.0	7.3	228	125	0.05	0.38	0.03	4.0	2.3	9.5
(No Name)	523	22.0	3.8	37	0	-	0.09	0.04	<0.3	2.0	12.0
(No Name)	524	15.5	5.1	30	30	<0.02	0.44	<0.01	<0.3	3.0	15.5
(No Name)	525	15.0	6.1	48	25	0.15	0.05	0.05	3.0	2.0	9.7
Gillham A	526	17.0	7.5	383	170	0.37	0.54	0.06	0.3	86	16.8
Gillham B	527	17.5	7.6	385	230	0.04	0.47	0.11	<0.3	87	16.8
Bard	528	15.5	6.8	184	140	0.03	0.22	0.07	7.6	1.3	16.2
(No Name)	529	17.5	6.4	35	50	2.30	0.08	0.08	<0.3	1.5	9.2
(No Name)	530	15.0	6.8	42	12	--	0.08	0.05	3.0	1.5	9.2

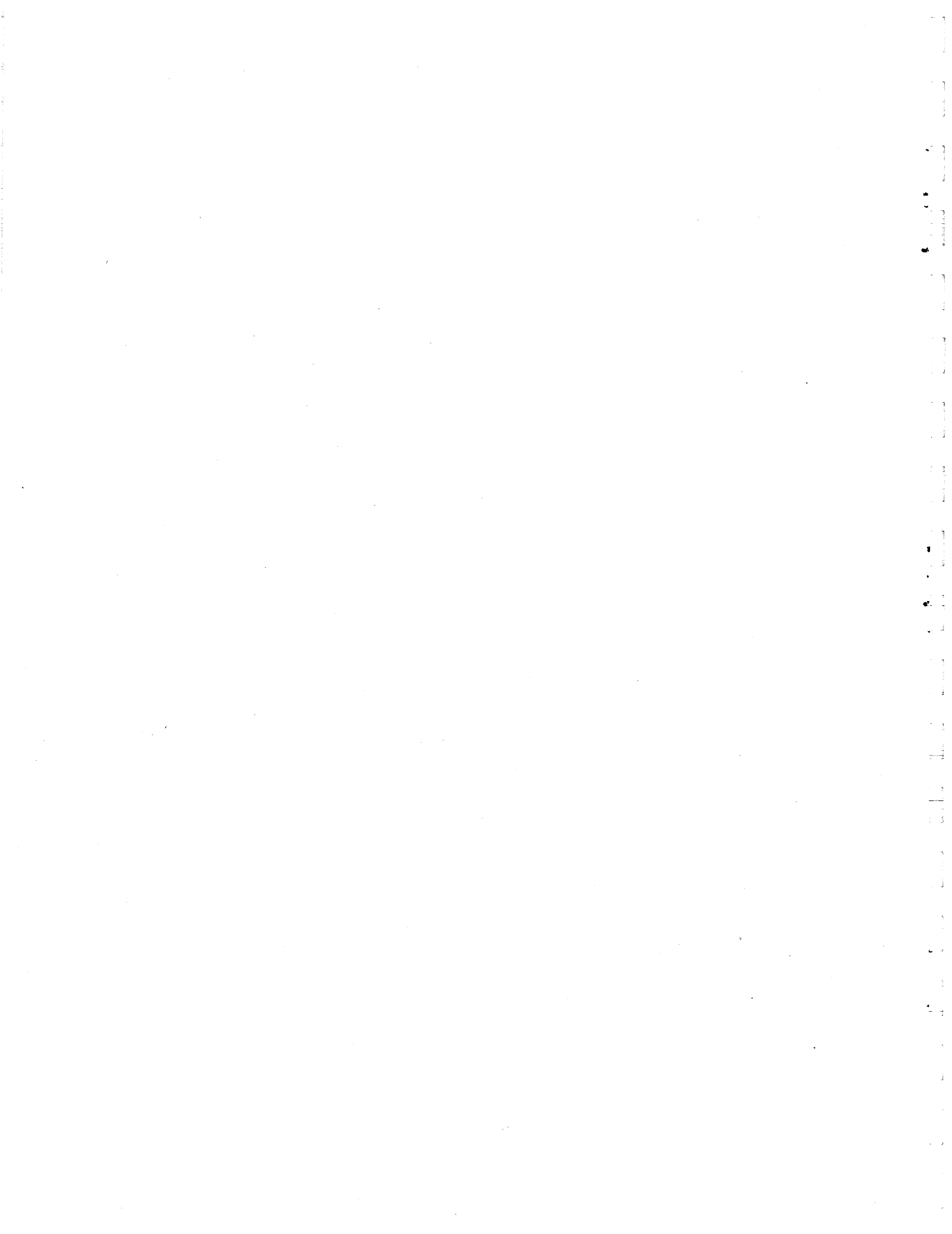


Table 1-2 Con't

No.	TOC	pH	Sp. Cond.	Alkalinity	NO ₃	NH ₃	PO ₄	SO ₄	Cl	SiO ₂
531	17.5	6.9	175	90	--	0.09	0.14	4.8	2.0	10.1
532	18.0	6.4	80	50	2.30	0.08	0.08	3.4	2.3	9.7
533	17.5	7.3	230	90	0.05	0.09	0.05	4.8	2.5	10.1
534	15.5	5.5	27	10	--	0.15	0.05	2.0	1.8	7.0
535	16.0	6.4	41	20	--	0.11	0.03	0.3	1.3	7.6
536	18.0	7.0	274	90	2.50	0.04	0.12	5.6	1.8	9.2
537	15.0	4.4	24	40	2.30	0.08	0.07	3.0	1.5	8.3
538	19.0	6.5	118	80	2.50	0.11	0.18	6.0	2.0	8.6
539	17.0	4.4	23	0	1.10	0.07	0.06	3.4	2.0	7.6
540	18.0	4.6	23	20	0.92	0.07	0.03	0.3	2.3	7.2
541	16.0	4.7	24	20	0.75	0.05	0.07	1.3	2.5	8.8
542	16.0	5.1	17	10	0.79	0.04	0.05	1.0	2.3	8.5
543	20.0	7.1	259	180	0.55	0.11	0.10	4.5	4.0	9.7
544	18.0	7.3	274	160	--	0.06	0.07	2.3	3.0	10.1
545	20.0	7.6	279	230	0.79	0.07	0.10	1.3	1.3	9.7
546	21.0	7.3	234	130	0.92	0.04	0.13	5.6	2.5	10.1
547	20.0	7.5	266	180	--	0.09	0.08	3.4	2.5	10.4
548	22.0	7.0	263	60	1.70	0.17	0.02	4.1	2.0	10.4
899	17.0	7.9	104	150	0.63	0.10	0.06	<0.3	2.7	10.8
900	18.0	6.2	34	100	0.87	0.13	0.06	0.3	1.8	9.3
901	18.0	6.9	131	112	0.19	0.21	0.04	1.4	1.8	9.1
902	20.0	6.4	28	50	0.67	0.16	0.08	0.7	3.8	10.5
903	18.0	6.3	108	70	0.76	0.14	0.10	1.0	2.5	11.6
904	18.0	4.1	108	0	14.3	0.10	0.13	3.6	9.5	8.7
905	20.0	6.2	303	125	1.35	0.13	0.20	2.8	4.5	9.9
906	20.0	4.3	55	10	1.62	0.05	0.30	1.7	4.5	5.4
907	19.0	6.2	196	67	0.31	0.19	0.11	5.9	2.8	37.1
908	18.0	7.1	57	25	1.20	1.39	0.04	2.1	2.3	11.9
909	19.0	7.2	17	5	0.68	0.21	0.08	2.8	3.0	13.2
910	18.0	5.8	57	2	0.83	0.47	0.20	12.6	5.0	19.3
911	16.0	5.4	183	175	1.19	0.47	0.28	5.1	2.3	11.6
912	16.5	5.8	135	70	1.07	0.38	0.21	4.0	1.5	9.3
913	17.0	4.1	267	135	0.68	1.08	0.25	1.4	3.5	14.0
914	20.0	5.5	46	22	0.42	0.10	0.09	0.3	3.0	18.2
915	35.0	5.7	200	122	0.17	0.16	0.08	0.7	3.5	17.2
916	18.0	5.8	97	40	0.11	0.20	0.24	7.6	3.0	34.9
917	18.0	7.2	57	17	0.67	0.10	0.14	10.2	3.0	11.6
918	20.0	5.7	55	37	0.97	0.02	0.08	2.5	2.5	8.7
919	21.0	5.9	81	12	3.77	0.20	0.08	0.7	6.5	17.2

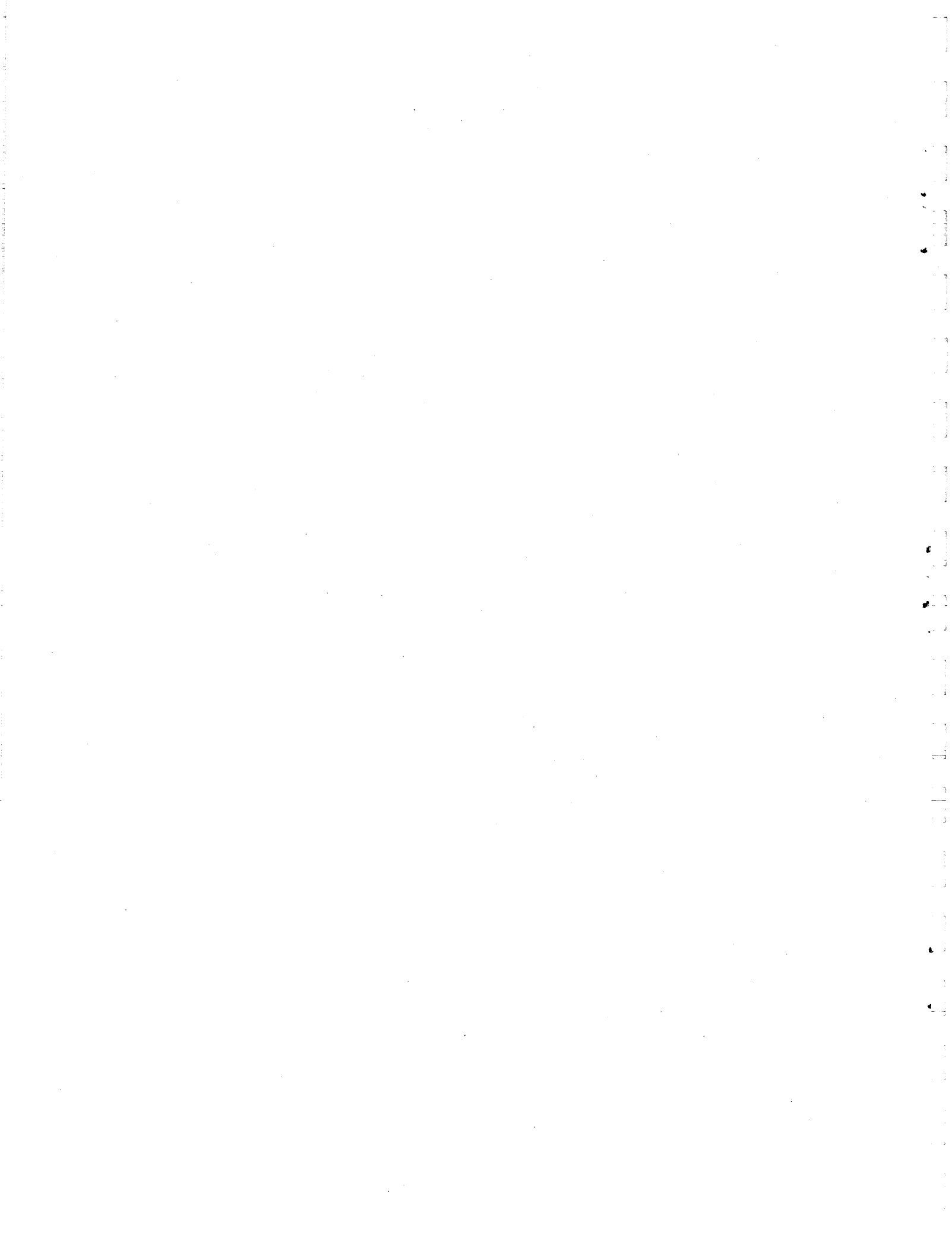


Table 1-2 Con't

No.	TOC	pH	Sp. Cond.	Alkalinity	NO ₃	NH ₃	PO ₄	SO ₄	Cl	SiO ₂
(No Name)	20.0	4.8	<5	5	0.77	0.21	0.03	0.3	5.0	22.9
(No Name)	30.0	6.4	227	130	3.26	0.10	0.12	3.2	2.5	10.5
Chalybeate	10.0	6.2	107	6	12.0	0.28	0.05	4.6	13.0	7.7
Caddo Gap (W)	22.0	7.4	138	100	0.98	0.23	0.15	5.0	2.5	31.9
Redland Mt.	25.0	7.2	210	144	3.86	0.22	0.08	9.7	1.5	18.5
Lolla Bell A	12.0	7.2	19	11	0.53	1.52	0.04	5.5	3.0	10.5
Lolla Bell B	12.0	7.0	19	9	0.53	1.52	0.04	6.8	1.5	10.0
Artesian Well (W)	18.5	7.6	373	186	10.50	1.02	0.34	29.5	3.5	21.3
Mineral	16.0	6.9	65	21	0.13	0.52	0.08	5.9	3.0	27.3
Barite Pit	10.0	5.0	20	1	0.53	1.31	0.02	0.51	2.5	10.5
Murfreesboro Area	11.0	5.1	19	1	1.02	1.52	0.03	4.1	3.0	7.2
Murfreesboro Area	11.0	5.0	15	2	0.83	0.36	0.09	8.7	0.8	4.8
Lake Greeson Area	13.0	6.8	112	58	0.38	0.23	0.14	5.5	4.8	25.6
Dierks Area	10.0	4.8	39	0	1.06	0.30	0.08	5.9	5.0	7.6
Cox Residence	12.0	4.7	63	0	1.36	0.29	0.06	5.5	4.0	7.4
Dierks Area	10.0	5.1	13	0	5.23	0.15	0.05	2.8	1.5	6.6
L. Greeson, Sulfur	11.0	6.8	109	56	--	0.50	0.18	4.1	3.5	23.5
L. Greeson, Possum	8.0	8.7	563	265	0.55	0.42	0.10	3.7	1.5	12.1
Salem Area	12.0	6.9	107	50	0.80	0.25	0.05	8.2	2.5	9.9
Bethesda	19.0	7.1	224	121	4.80	0.10	0.09	2.4	2.0	13.2
Mine Creek Area	16.0	5.9	248	170	0.20	0.00	0.13	2.9	2.5	9.2
Rock	17.0	6.5	299	176	1.24	0.05	0.09	4.3	3.0	13.2
West	15.0	5.3	36	24	1.24	0.10	0.06	1.3	1.5	7.9
Pigeon Roost	15.0	5.5	218	42	0.30	0.02	0.12	7.1	2.0	11.4
Barite Pit	18.5	4.6	23	12	0.08	0.06	0.10	1.0	3.0	10.8
"Warm"	23.0	7.1	116	73	5.50	0.00	0.15	2.6	2.5	11.1
Athens Area	16.0	5.8	18	24	1.91	0.04	0.23	2.6	2.8	13.6
Umpire Area	18.0	5.6	23	18	2.70	0.09	0.13	0.5	1.8	12.4
S. Cox	16.0	4.1	177	0	19.7	0.07	0.05	1.0	8.0	8.0
Cossatot	13.0	4.7	37	24	0.89	0.20	0.12	1.5	2.5	16.0
Church	18.0	4.7	205	18	4.70	0.14	0.10	2.6	5.0	12.4
Defore (W)	16.0	7.4	484	200	0.09	0.02	0.41	55.0	3.5	29.0
(No Name)	14.0	4.7	61	12	5.00	0.05	0.13	1.0	5.5	12.9
Mena Park	16.0	4.5	71	12	2.41	0.00	0.12	8.5	6.0	12.9
Median	16.5	6.2	85	37	0.9	0.15	0.08	3	2.9	11
EPA or APHA precision data	-	0.15	4	3.2	-	0.06	<0.01	0.3	-	≈2

* Cored, plastic piping

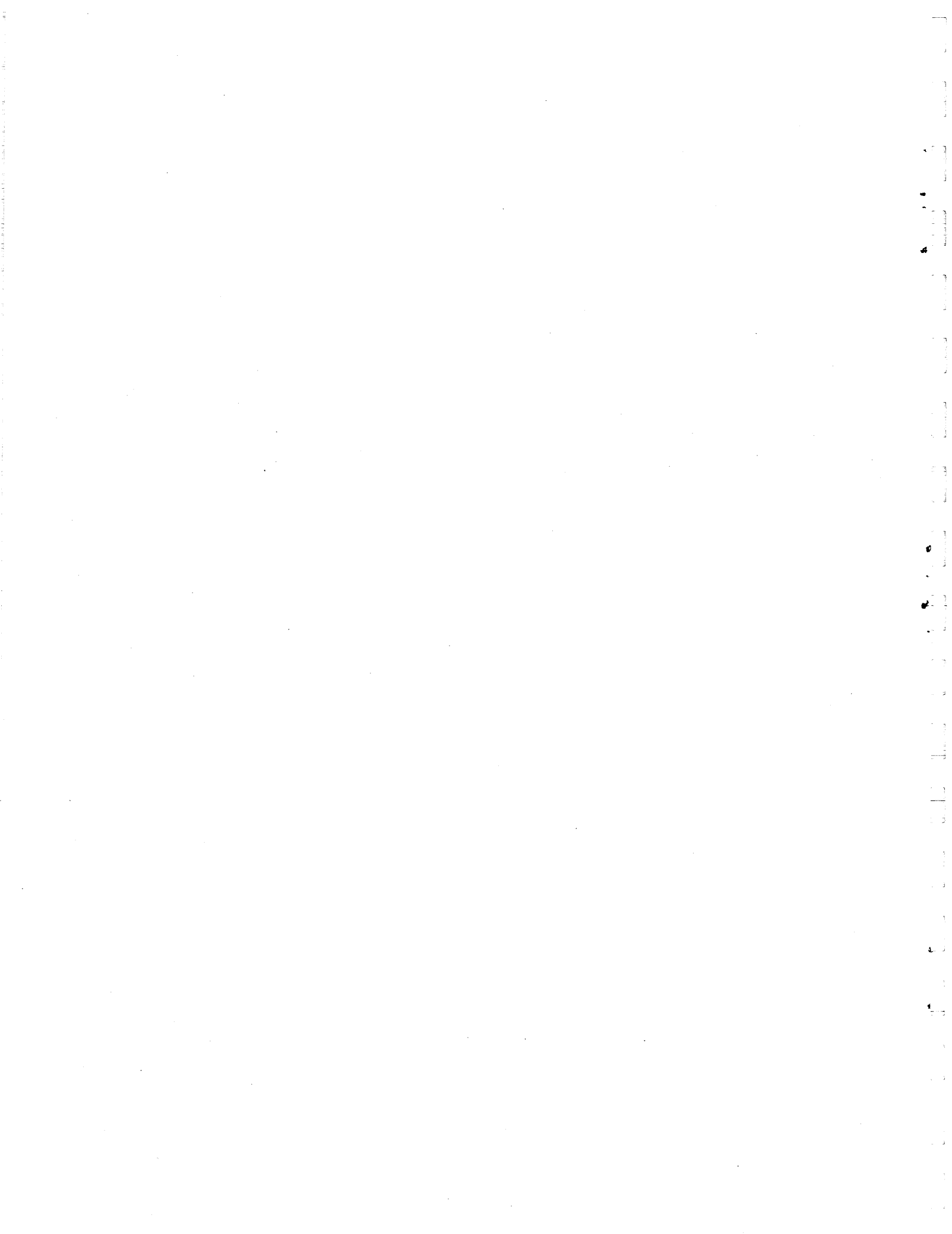


Table 1-3
Chemical Analyses of Ouachita Spring Waters
Alkali and Alkaline Earth Metals

Name of Spring or Well(W)	Lab No.	Na ppm	K ppm	Li ppb	Ca ppm	Mg ppm	Sr ppm	Ba ppm
Iron (W)	501	48.40	0.50	21.0	43.20	12.60	0.597	0.04
Iron	502	0.95	0.12	2.0	0.60	0.81	0.003	<0.01
Black	503	10.00	0.30	8.0	60.80	14.30	0.356	<0.01
Collier	504	2.00	1.25	3.0	8.30	1.95	0.020	0.02
Caddo Valley	505	1.10	0.21	2.0	40.00	1.31	0.140	<0.01
Crystal A	506	1.50	1.70	8.0	40.20	6.21	0.062	<0.01
Goat Brown	507	2.20	0.30	4.0	17.70	1.47	0.031	0.02
Bogg	508	2.80	0.34	3.0	41.50	1.89	0.063	<0.01
Queen Wilhelmina Park	509	2.00	0.42	1.0	1.52	0.93	0.014	<0.01
Queen Wilhelmina Park	510	2.40	0.48	1.0	1.47	1.11	0.011	0.02
Silver World	511	1.30	0.12	1.0	24.50	0.97	0.117	<0.01
Abernathy	512	1.10	0.31	2.0	32.70	1.12	0.138	<0.01
Three Sisters	513	3.10	0.80	7.0	41.20	6.33	0.089	<0.01
(No Name)	514	1.10	0.42	1.9	1.47	0.84	<0.002	<0.01
(No Name)	515	1.10	0.38	2.2	1.47	0.79	<0.002	<0.01
(No Name)	516	1.60	1.05	4.0	14.70	3.02	0.027	<0.01
Wyatt	517	1.20	0.78	3.7	4.25	1.08	<0.002	<0.01
(No Name)	518	1.50	0.23	1.0	0.73	0.62	0.004	<0.01
(No Name)	519	6.00	1.30	3.7	25.00	2.92	0.105	0.11
Pigeon Roost* (W)	520	6.50	1.40	23.0	37.70	3.18	0.317	0.42
Dripping	521	0.95	0.30	1.9	1.91	0.95	0.008	<0.01
Strawn	522	1.80	0.49	4.2	54.40	2.86	0.234	<0.01
(No Name)	523	2.60	0.45	1.9	1.06	0.83	0.012	0.02
(No Name)	524	4.30	1.05	1.9	4.64	1.50	0.015	0.01
(No Name)	525	1.50	0.37	2.2	9.53	1.04	0.033	<0.01
Gillham A	526	131	1.20	70.0	9.99	3.06	0.838	0.22
Gillham B	527	152	1.10	70.0	9.77	3.18	0.802	0.23
Bard	528	6.20	0.72	3.0	31.00	12.40	0.140	<0.01
(No Name)	529	1.50	0.44	1.6	4.05	1.22	0.012	<0.01
(No Name)	530	1.50	0.46	1.6	4.52	1.96	0.016	<0.01
McClaine A	531	2.10	0.80	4.6	44.20	2.73	0.131	0.01
McClaine B	532	1.80	0.58	2.6	17.30	1.56	0.072	0.01
McClaine C	533	2.10	0.75	4.6	50.80	2.95	0.160	<0.01
McClaine D	534	1.20	0.56	<1.0	3.13	0.76	0.018	0.02
W. Jones	535	1.10	0.37	<1.0	6.24	0.68	0.035	<0.01
Jones Valley	536	1.40	0.48	2.4	39.70	2.16	0.197	<0.01
Buttermilk A	537	1.30	0.29	<1.0	1.91	0.35	0.013	<0.01
Buttermilk B	538	1.50	0.40	1.9	35.60	1.22	0.168	<0.01
(No Name)	539	1.30	0.29	<1.0	1.46	0.32	0.010	<0.01
Mcellhanon	540	1.80	0.57	<1.0	2.37	0.46	0.014	<0.01
(No Name)	541	1.60	0.31	<1.0	1.09	0.46	0.003	<0.01
(No Name)	542	1.20	0.04	<1.0	0.033	0.15	0.002	<0.01
Crystal B	543	1.70	0.48	2.6	66.5	2.04	0.161	<0.02
Brogdams A	544	1.90	0.36	2.6	60.1	2.54	0.150	<0.01
Brogdams B	545	1.80	0.46	3.1	61.3	2.14	0.167	<0.02
Brogdams C	546	2.10	0.51	3.3	56.4	2.16	0.143	<0.02

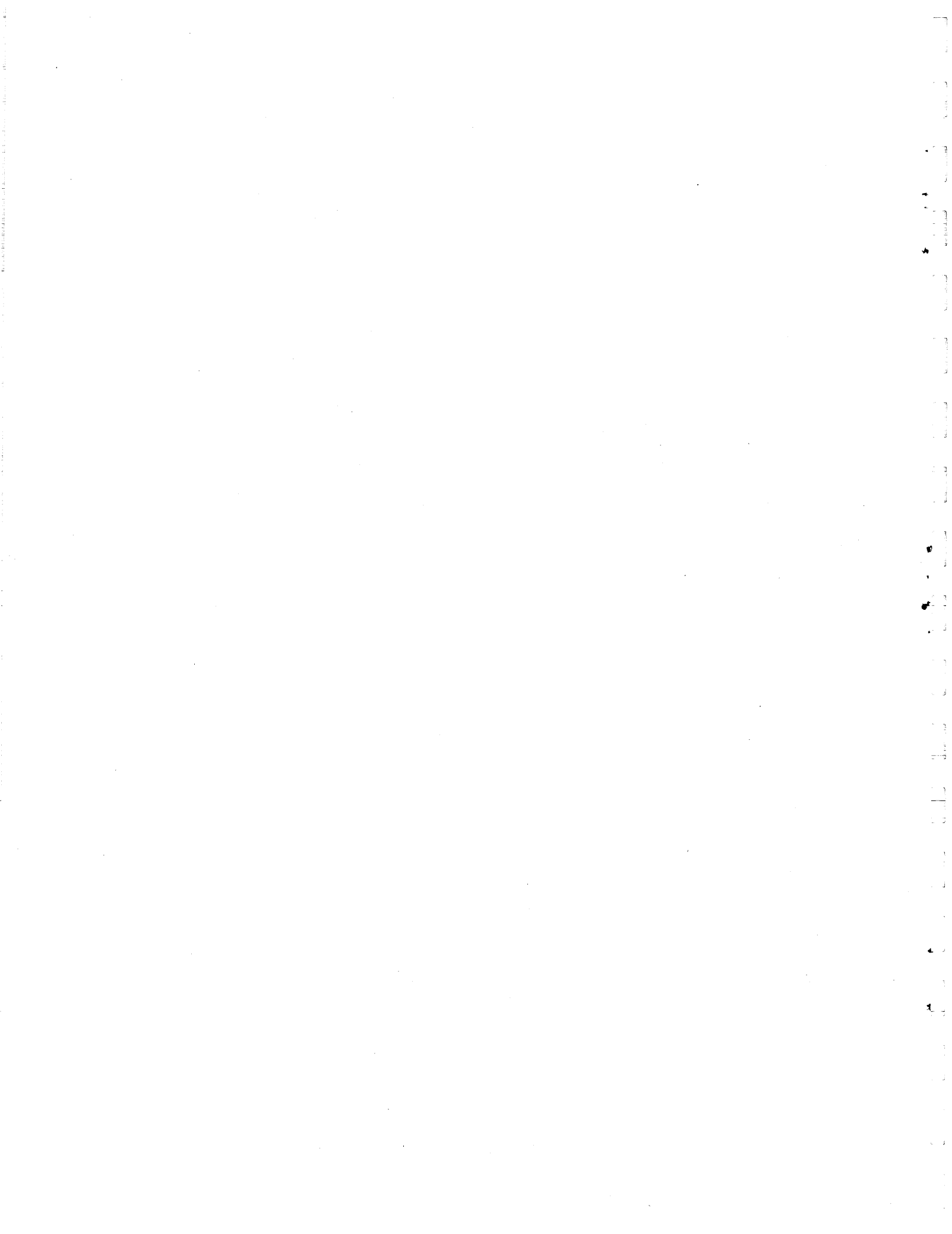


Table 1-3 Con't

Name of Spring or Well (W)	Lab. No.	Na ppm	K ppm	Li ppb	Ca ppm	Mg ppm	Sr ppm	Ba ppm
Brogdams D	547	2.20	0.48	3.1	58.9	2.33	0.141	<0.02
Brogdams E	548	2.30	0.55	2.9	61.3	2.41	0.149	<0.02
Burrows	899	3.25	0.50	3.8	53.3	2.17	0.119	0.03
Lead Mine Road	900	1.31	0.24	<1.0	5.60	0.89	0.010	0.06
Lead Mine Road	901	1.75	0.37	1.9	35.90	6.16	0.089	0.01
(No Name)	902	1.80	0.24	2.2	2.36	1.08	0.003	0.04
(No Name)	903	1.70	0.24	2.2	25.40	1.21	0.045	0.01
(No Name)	904	10.00	3.80	1.9	4.97	3.05	0.057	0.18
Womack Well (W)	905	82.5	1.15	28.5	8.56	2.67	0.148	0.02
Womack	906	5.15	4.30	<1.0	2.36	1.15	0.018	0.07
Lolla Bell (W)	907	20.00	1.00	16.5	13.10	19.60	0.101	0.27
(No Name)	908	2.10	0.62	7.5	1.30	3.48	0.004	0.06
(No Name)	909	1.00	1.40	1.8	0.75	0.88	0.002	0.02
(No Name)	910	0.85	0.67	14.0	4.47	1.00	0.011	0.08
(No Name)	911	1.50	0.20	2.5	70.10	2.83	0.500	0.06
(No Name)	912	1.23	0.27	1.0	25.20	1.13	0.160	0.02
Sulfur	913	16.50	0.90	17.3	27.54	14.92	2.200	0.07
Elliott	914	1.55	0.24	4.0	1.37	1.91	0.015	0.02
Caddo River	915	5.40	1.00	11.5	42.05	2.45	0.051	0.03
Artesian*Well (W)	916	2.00	0.95	10.0	7.30	2.52	0.044	0.20
Artesian* Well (W)	917	1.13	1.05	8.2	4.18	1.91	0.035	0.03
(No Name)	918	1.65	0.27	2.1	6.83	1.61	0.032	0.03
(No Name)	919	7.60	2.80	4.3	4.18	2.90	0.048	0.03
(No Name)	920	1.23	0.46	2.1	0.60	0.74	0.070	0.02
(No Name)	921	3.83	1.10	4.6	50.41	2.38	0.100	<0.01
Chalybeate	922	10.00	2.90	2.1	6.68	7.75	0.080	0.15
Caddo Gap (W)	923	13.00	0.65	12.5	18.30	2.80	0.088	0.10
Redland Mt.	924	2.50	0.95	7.0	53.35	23.70	0.071	0.02
Lolla Bell A	925	2.00	0.70	2.0	1.54	0.96	0.016	0.02
Lolla Bell B	926	2.50	0.38	2.0	0.91	1.15	0.011	0.03
Artesian Well (W)	927	80.00	5.20	58.0	26.68	4.31	1.44	0.08
Mineral	928	4.60	3.90	20.0	2.66	1.40	0.049	0.01
Barite Pit	929	2.40	0.21	2.1	0.91	0.30	0.032	0.91
Murfreesboro Area	930	1.60	0.38	1.0	1.22	0.61	0.013	0.06
Murfreesboro Area	931	0.90	0.49	1.6	0.91	0.64	0.007	0.02
Lake Greeson Area	932	10.50	1.41	26.0	8.60	4.55	0.042	0.26
Dierks Area	933	3.20	2.11	1.0	1.07	0.77	0.006	0.03
Cox Residence	934	2.60	1.75	1.0	4.51	1.86	0.023	0.03
Dierks Area	935	1.25	0.45	1.6	0.91	0.58	0.013	0.03
Lake Greeson, Sulfur	936	11.00	2.60	17.5	5.57	3.92	0.041	0.15
Lake Greeson Possum	937	162.5	0.48	86.0	1.22	0.19	0.030	0.03
Salem Area	938	1.90	0.80	2.1	20.74	1.40	0.018	0.04
Bathesda	939	2.85	0.80	4.0	46.36	6.44	0.030	<0.02
Mine Creek Area	940	1.24	0.46	2.6	53.90	1.70	0.320	<0.02
Rock	941	3.15	0.42	3.3	65.66	5.68	0.105	<0.02
West	942	0.93	0.49	<1.0	5.66	0.76	0.020	0.02
Pigeon Roost	943	2.00	0.70	8.2	12.64	2.35	0.068	0.19
Barite Pit	944	1.40	0.30	2.0	0.39	0.34	0.020	0.93
"Warm"	945	1.70	0.73	2.2	19.52	2.66	0.033	<0.02
Athens Area	946	2.85	0.78	2.0	0.78	0.47	0.010	0.02

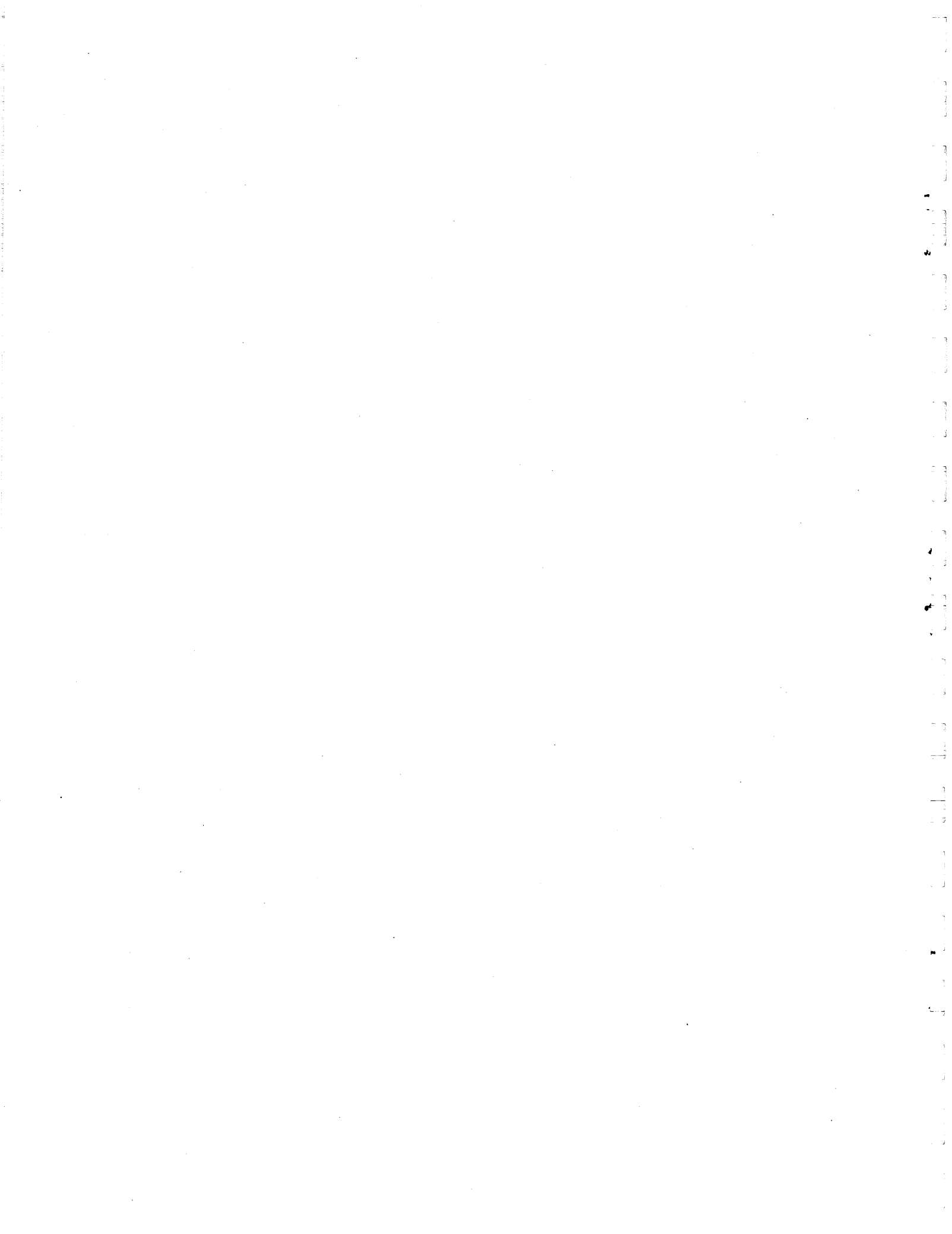


Table 1-3 Con't

Name of Spring or Well (W)	Lab No.	Na ppm	K ppm	Li ppb	Ca ppm	Mg ppm	Sr ppm	Ba ppm
Umpire Area	947	2.70	0.62	1.6	1.37	0.64	0.010	<0.02
S. Cox	948	6.80	8.50	4.6	9.45	3.45	0.092	0.26
Cossatot	949	5.00	0.95	2.0	2.73	0.64	0.033	<0.02
Church	950	6.25	1.20	2.0	2.82	1.47	0.026	<0.02
Defore Well (W)	951	21.50	1.25	45.0	78.39	9.37	0.275	0.02
(No Name	952	6.50	1.30	2.6	2.62	1.93	0.020	0.03
Mena Park	953	6.70	0.78	2.6	3.68	1.88	0.020	<0.02

* Cored, plastic piping

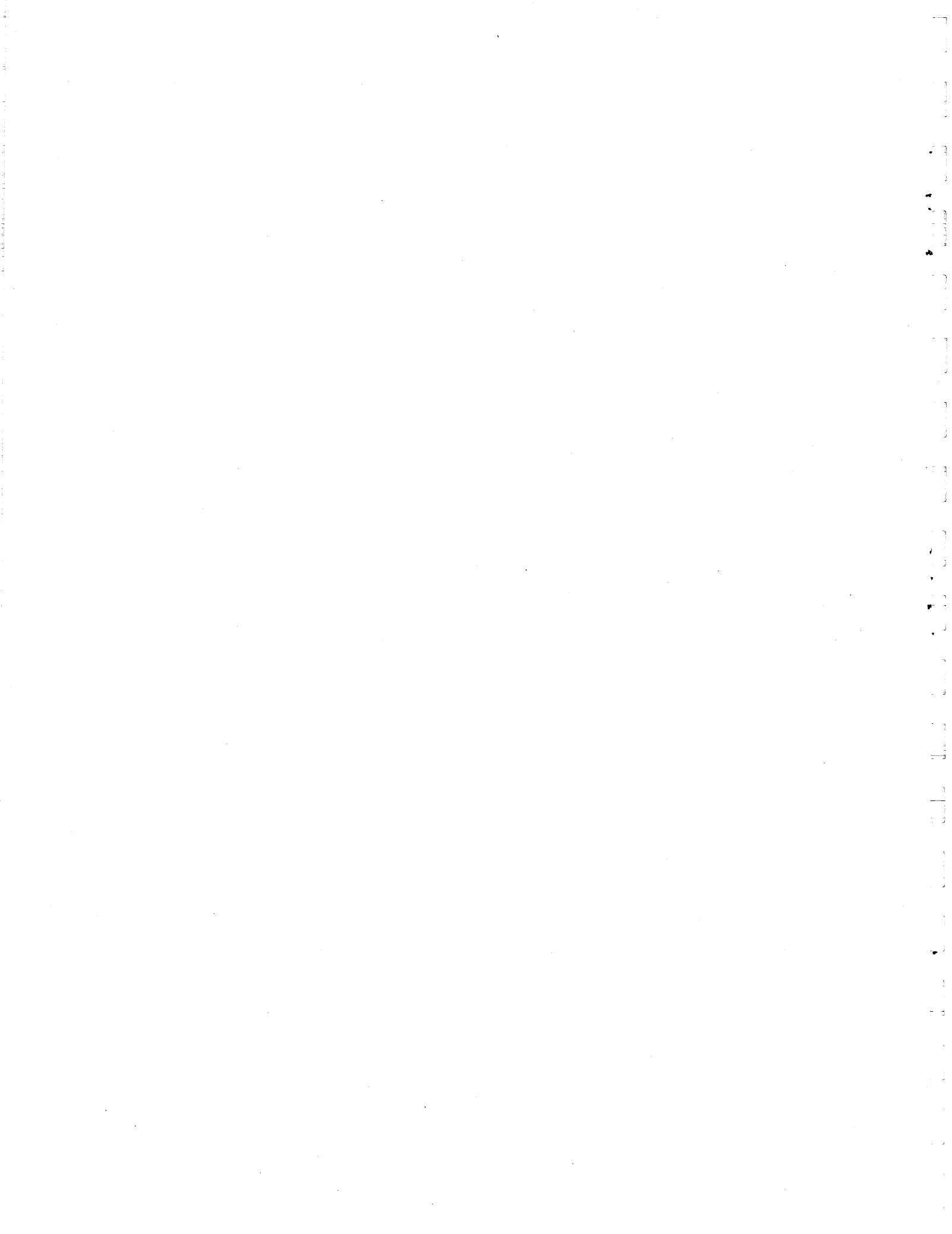


Table 1-4

Chemical Analyses of Ouachita Spring Waters for
Heavy Metals (all data in parts per billion)

Name of Spring or Well(W)	Lab									
	No.	Fe	Mn	Zn	Cu	Co	Ni	Pb	Hg	Sb
Iron (W)	501	4	20	1,480	14	<3	<3	<10	<0.1	<0.2
Iron	502	275	12	17	5	<3	<3	<10	<0.1	<0.2
Black	503	36	195	18	1	8	<3	<10	<0.1	<0.2
Collier	504	<2	<5	6	1	<3	<3	<10	<0.1	0.2
Caddo Valley	505	97	25	31	1	<3	<3	<10	<0.1	0.2
Crystal A	506	<2	<5	<2	1	<3	<3	<10	<0.1	0.2
Goat Brown	507	7	<5	25	7	<3	<3	<10	<0.1	0.3
Bogg	508	22	10	2	3	<3	<3	<10	<0.1	0.25
Queen Wilhelmina S. Park	509	29	6	9	7	<3	<3	<10	<0.1	0.12
Queen Wilhelmina S. Park	510	85	8	17	6	<3	<3	<10	<0.1	0.12
Silver World	511	472	35	26	85	8	8	18	<0.1	<0.2
Abernathy	512	165	30	26	5	8	<3	<10	<0.1	0.2
Three Sisters	513	210	18	24	15	<3	3	<10	0.1	<0.2
(No Name)	514	6	17	17	8	<3	<3	<10	<0.1	<0.2
(No Name)	515	3	29	17	8	<3	<3	18	<0.1	<0.2
(No Name)	516	663	316	32	5	13	12	18	<0.1	<0.2
Wyatt	517	3620	1220	69	21	20	29	29	<0.1	0.2
(No Name)	518	43	25	16	9	<3	<3	13	<0.1	<0.2
(No Name)	519	3	5	5	5	<3	<3	10	<0.1	<0.2
Pigeon Roost*(W)	520	323	272	10	3	13	8	33	0.72	<0.2
Dripping	521	7	14	20	7	<3	<3	13	<0.1	<0.2
Strawn	522	880	206	9	7	11	8	18	<0.1	<0.2
(No Name)	523	15	4	10	14	<3	2	<5	<0.1	<0.2
(No Name)	524	4632	664	15	4	19	17	8	0.2	<0.2
(No Name)	525	47	24	14	3	3	7	<5	<0.1	<0.2
Gillham A	526	11	15	2	1	2	4	<5	0.47	<0.2
Gillham B	527	6	15	3	3	2	4	<5	0.30	<0.2
Bard	528	3151	469	5	2	12	9	<5	2.10	<0.2
(No Name)	529	27	20	11	6	7	4	<5	<0.1	<0.2
(No Name)	530	11	2	20	2	2	4	<5	<0.1	<0.2
McClaine A	531	1974	448	28	7	7	18	15	<0.1	<0.2
McClaine B	532	135	9	26	8	2	9	<5	<0.1	<0.2
McClaine C	533	1495	405	22	9	7	18	18	<0.1	<0.2
McClaine D	534	198	62	23	5	8	10	<5	<0.1	<0.2
W. Jones	535	51	6	10	6	4	10	<5	<0.1	<0.2
Jones Valley	536	739	278	143	2	7	21	<5	<0.1	<0.2
Buttermilk A	537	496	9	110	4	6	21	<5	<0.1	<0.2
Buttermilk B	538	804	65	43	1	4	18	<5	<0.1	<0.2
(No Name)	539	68	7	66	8	<2	12	<5	<0.1	<0.2
Mcellhanon	540	17	3	46	4	<2	9	11	0.16	<0.2
(No Name)	541	19	9	49	6	<2	11	<5	<0.1	<0.2
(No Name)	542	19	6	37	6	<2	9	<5	0.18	<0.2
Crystal B	543	699	210	39	6	2	17	<5	0.1	<0.2
Brogdams A	544	503	189	45	4	<2	15	5	1.86	<0.2
Brogdams B	545	643	205	28	4	<2	15	10	<0.1	<0.2
Brogdams C	546	760	210	27	4	5	17	7	0.29	<0.2

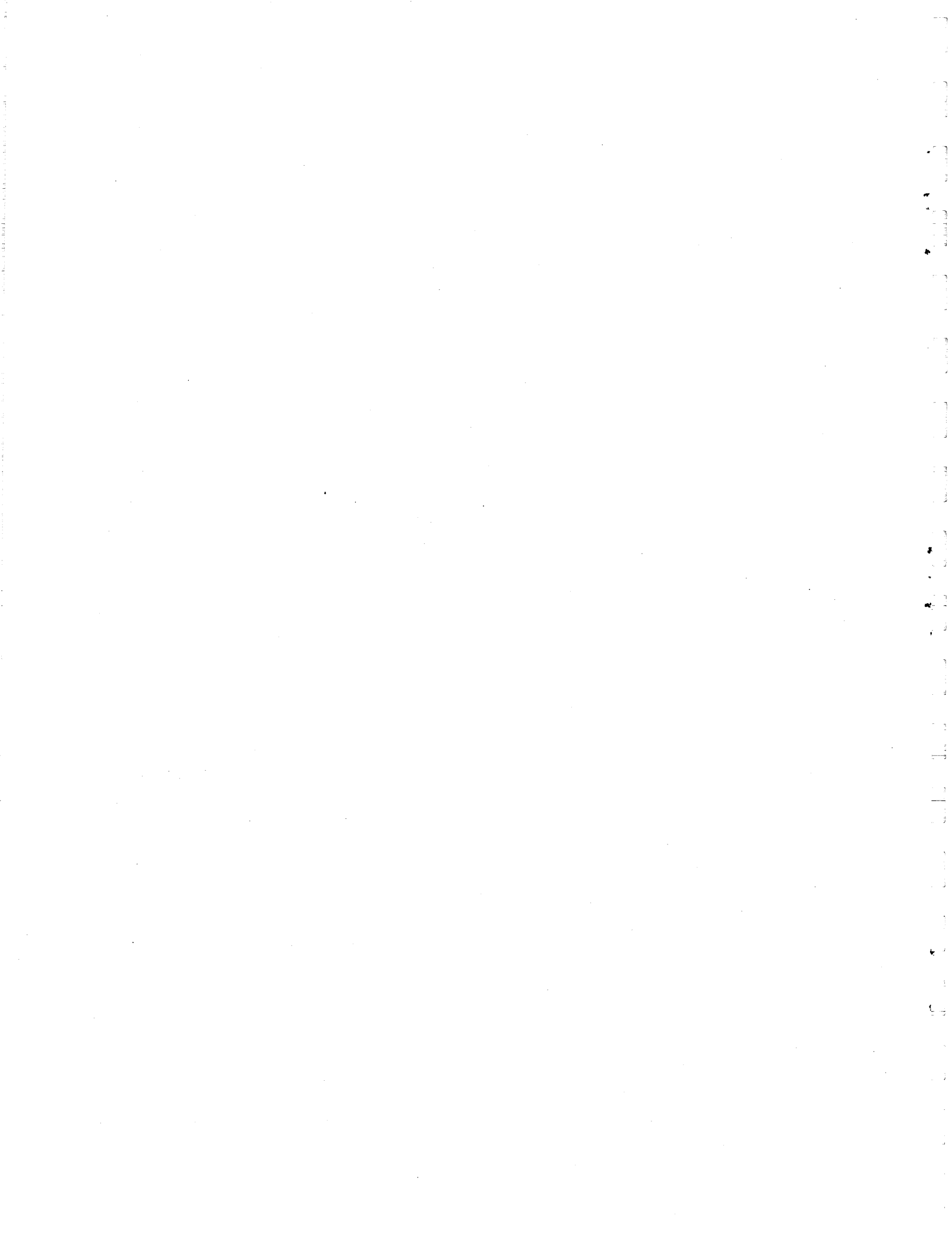


Table 1-4 Con't

Name of Spring or Well (W)	Lab No.	Fe	Mn	Zn	Cu	Co	Ni	Pb	Hg	Sb
Brogdams D	547	699	221	32	2	4	17	< 5	<0.1	<0.2
Brogdams E	548	661	296	13	2	6	16	10	<0.1	<0.2
Burrows	899	10	5	92	9	<3	8	11	0.72	<0.5
Lead Mine Road	900	33	5	137	4	2	4	<7	<0.1	<0.5
Lead Mine Road	901	2236	701	27	2	15	17	30	<0.1	<0.5
(No Name)	902	39	19	21	2	7	12	11	<0.1	<0.5
(No Name)	903	41	10	12	1	10	15	15	<0.1	<0.5
(No Name)	904	25	89	48	14	20	25	34	0.59	<0.5
Womack (W)	905	7	6	<5	1	10	15	23	0.51	<0.5
Womack	906	23	25	6	2	18	30	33	<0.1	<0.5
Lolla Bell (W)	907	2062	344	21	2	15	32	37	0.63	<0.5
(No Name)	908	10082	6	7	2	39	80	42	2.27	<0.5
(No Name)	909	146	23	6	4	4	7	17	1.48	<0.5
(No Name)	910	149	83	20	4	18	21	21	<0.1	<0.5
(No Name)	911	195	31	144	1	10	19	13	0.68	<0.5
(No Name)	912	5	1	<5	2	5	8	17	0.75	<0.5
Sulfur	913	31	7	<5	1	5	10	10	0.29	<0.2
Elliott	914	1347	32	5	4	13	17	14	0.76	<0.2
Caddo River	915	80	7	<5	4	8	12	18	0.28	<0.2
Artesian* Well (W)	916	5222	631	65	2	23	39	37	0.78	0.2
Artesian* Well (W)	917	1642	351	21	2	18	28	25	0.26	<0.2
(No Name)	918	49	10	<5	2	6	8	7	0.32	<0.2
(No Name)	919	49	42	<5	13	10	17	17	0.20	<0.5
(No Name)	920	15	22	<5	9	8	12	12	0.32	0.5
(No Name)	921	10	4	<5	3	5	10	12	1.02	<0.5
Chalybeate	922	43	77	19	10	7	20	21	0.65	-
Caddo Gap (W)	923	580	436	38	<1	11	17	21	0.75	-
Redland Mt.	924	226	345	9	<1	22	24	30	<0.2	-
Lolla Bell A	925	15	5	4	3	4	2	30	0.22	-
Lolla Bell B	926	43	10	16	5	15	32	42	0.44	-
Artesian Well (W)	927	85	86	29	3	15	24	46	0.87	-
Mineral	928	4218	345	16	3	11	21	30	0.65	-
Barite Pit	929	129	13	10	5	7	9	30	0.65	-
Murfreesboro Area	930	15	16	9	5	7	13	17	0.90	-
Murfreesboro Area	931	29	5	5	<1	7	9	34	0.65	-
Lake Greeson Area	932	1375	70	5	<1	15	28	59	<0.2	-
Dierks Area	933	175	14	16	5	11	17	34	<0.2	-
Cox Residence	934	43	16	24	<1	11	24	25	0.22	-
Dierks Area	935	29	10	10	3	11	24	25	0.65	-
L. Greeson, Sulfur	936	9930	370	11	<1	38	54	34	0.65	-
L. Greeson, Possum	937	57	8	8	5	15	20	25	1.10	-
Salem Area	938	287	24	16	3	22	36	30	0.65	-
Bethesda	939	301	743	11	2	30	28	24	0.22	-
Mine Creek	940	60	22	12	2	6	7	9	0.22	-
Rock	941	2	2	4	5	2	3	<5	<0.2	-
West	942	10	3	6	2	2	3	20	<0.2	-
Pigeon Roost	943	3306	800	28	3	31	39	29	0.30	-
Barite Pit	944	100	12	10	2	6	7	13	0.45	-
"Warm"	945	5	8	6	2	6	7	9	<0.2	-
Athens Area	946	13	6	4	2	2	7	18	<0.2	-

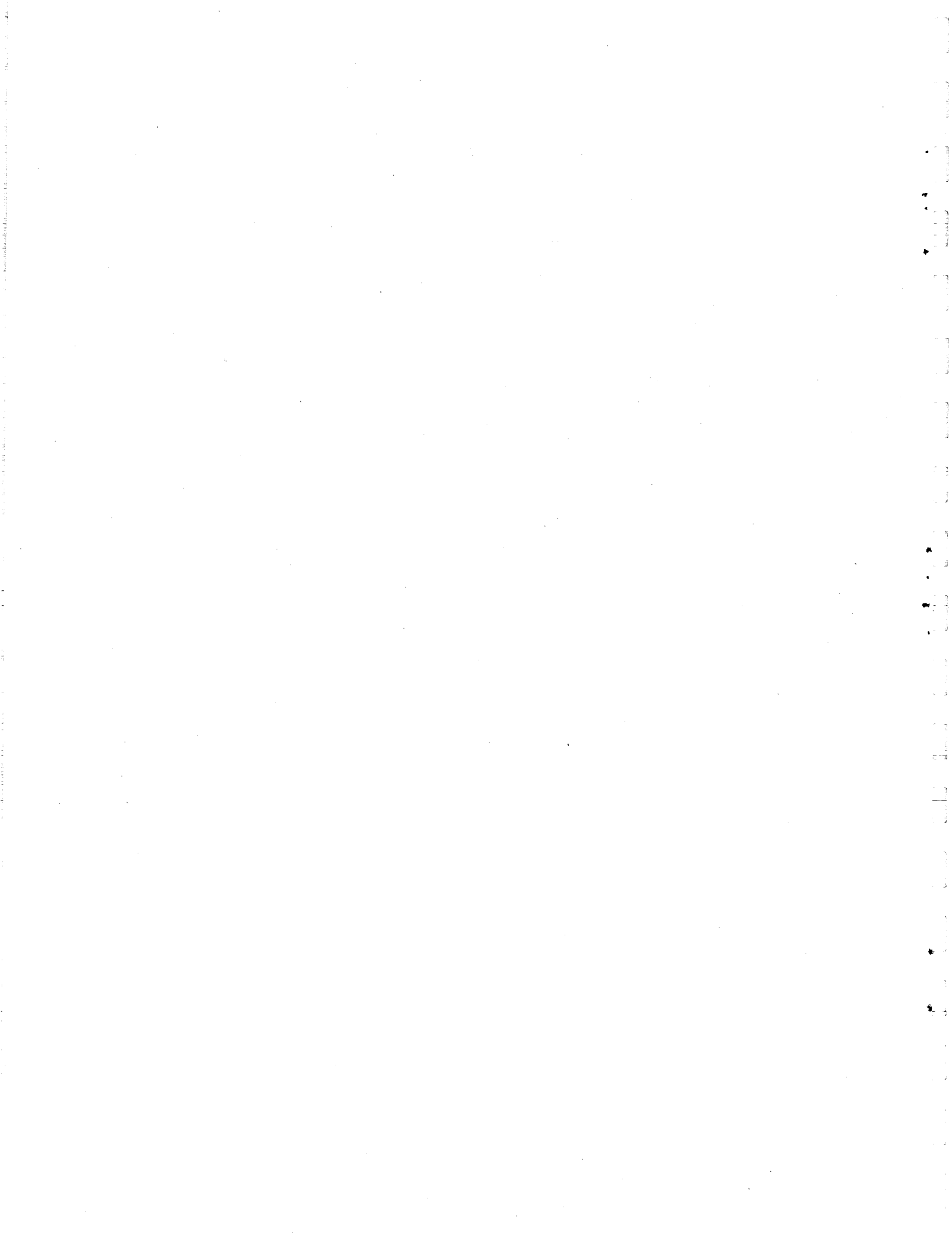


Table 1-4 Con't

Name of Spring or Well(W)	Lab No.	Fe	Mn	Zn	Cu	Co	Ni	Pb	Hg	Sb
Umpire Area	947	27	5	6	4	8	7	11	0.2	-
S. Cox	948	20	110	45	7	25	33	31	<0.2	-
Cossatot	949	17	4	6	2	11	7	13	<0.2	<0.5
Church	950	43	11	7	4	11	2	31	<0.2	<0.5
Defore (W)	951	48	283	111	3	34	33	33	<0.2	<0.5
(No Name)	952	18	14	13	2	11	5	22	<0.2	<0.5
Mena Park	953	10	11	8	2	9	7	13	<0.2	-

* Cored, plastic piping

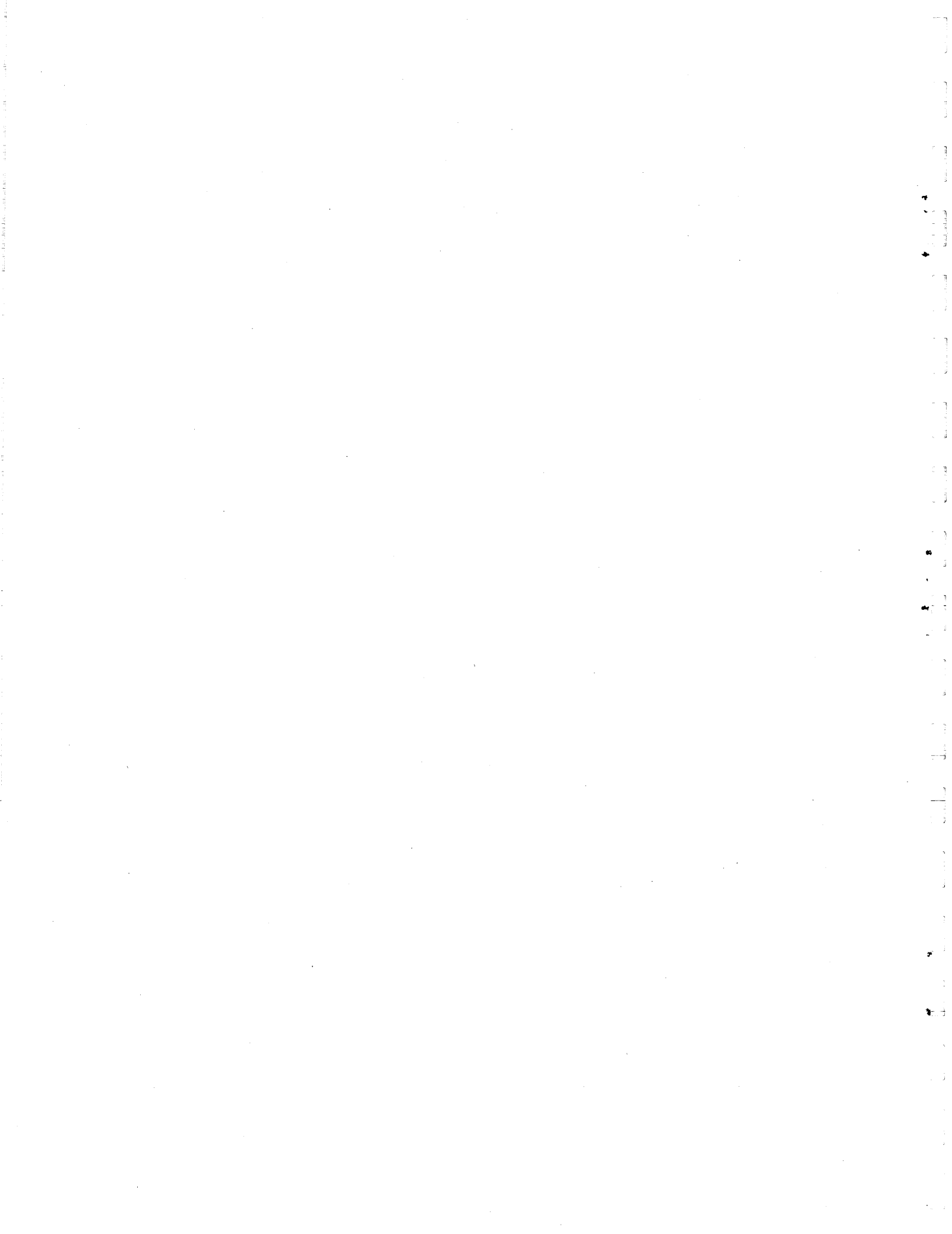


Table 1-5

Median Values (and Ranges) of Various Measurements
And Analyses for Areas I and II

Miscellaneous Measurements

	Water T ₁ ^o C Surface	Water T ^o C Subsurface*	pH Units	Specific Cond. µmhos per cm 25°C	Total Alkalinity mg/l as CaCO ₃
Area I	17(5-35)	40.0(7.1-91.2)	6.4(3.8-7.9)	107(14-385)	60(0-291)
Area II	16(8-25)	45.9(11.2-89.2)	5.8(4.1-8.7)	61(<5-563)	15(0-265)
(II/I)100 (%)	94	115	91	57	25

Anion Concentrations (ppm)

	NO ₃ as N	NH ₃ as N	PO ₄ as P	SO ₄	Cl	SiO ₂
Area I	0.66(<0.01-5.5)	0.1(<0.01-1.52)	0.07(<0.01-0.54)	2.3(<0.3-32)	2.3(0.25-87)	10.3(3.2-34.9)
Area II	1.0(0.06-14.3)	0.20(0.02-152)	0.10(0.02-0.41)	3.2(0.3-55)	3.5(0.5-9.5)	12.1(3.8-37.1)
(II/I)100 (%)	152	200	143	139	152	117

*calculated from SiO₂ ranges

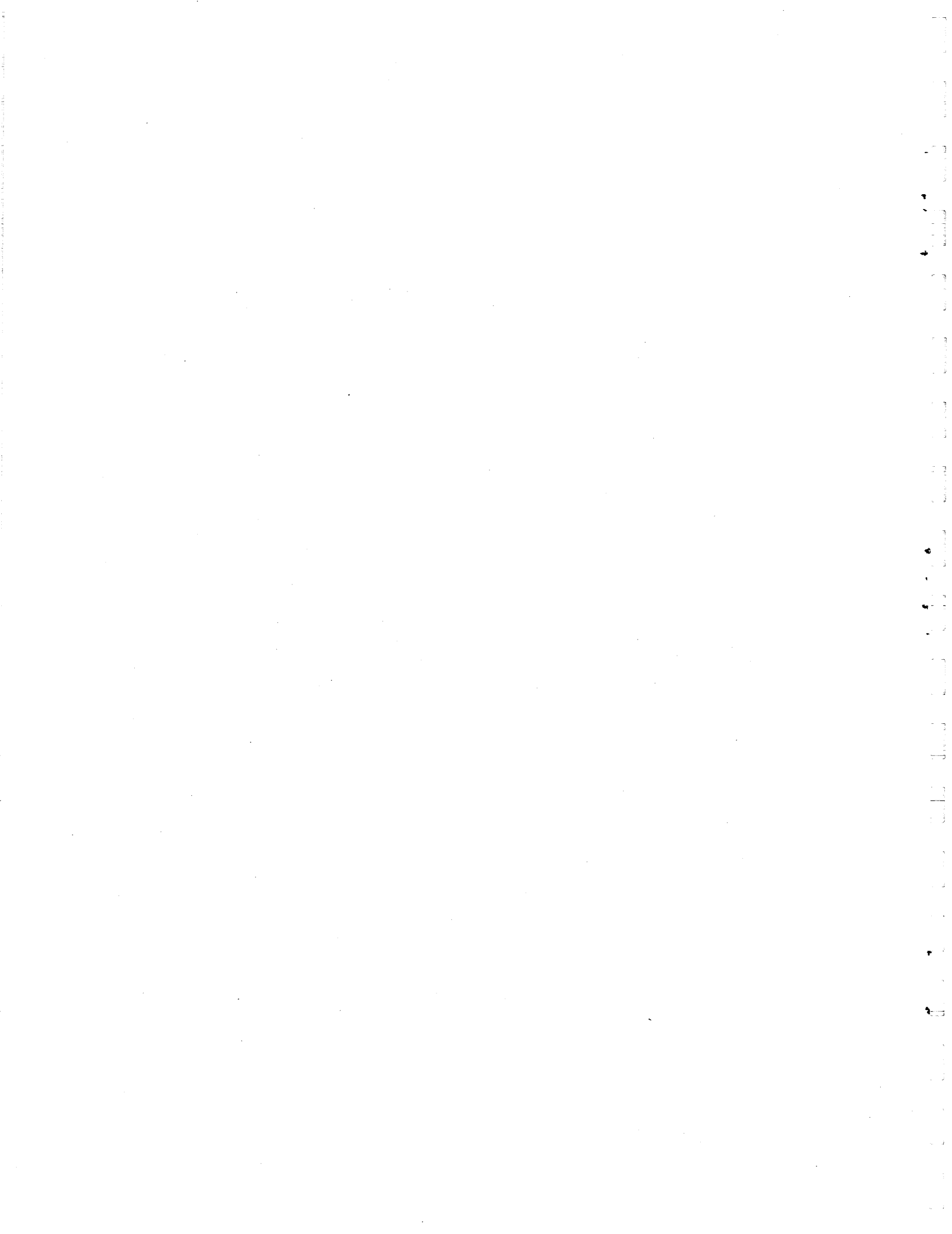


Table 1-5 (con't)

Alkali Metals Concentrations

	<u>Na (ppm)</u>	<u>K (ppm)</u>	<u>Li (ppb)</u>
Area I	1.8(0.95-152)	0.56(0.04-2.9)	2.5(<1-70)
Area II	3.9(0.85-162)	1.10(0.21-8.5)	2.3(<1-86)
(II/I)100 (%)	217	196	92

Alkaline Earth Metals Concentrations

	<u>Ca (ppm)</u>	<u>Mg (ppm)</u>	<u>Sr (ppm)</u>	<u>Ba (ppm)</u>
Area I	13.5(0.03-70)	1.9(0.15-14.9)	0.055(0.002-2.2)	<0.015(0.001-0.42)
Area II	2.7(0.6-78)	1.40(0.19-23.7)	0.03(<0.01-1.4)	0.04(0.01-0.92)
(II/I)100 (%)	20	74	55	>267

Base Metal Concentrations (ppb)

	<u>Fe</u>	<u>Mn</u>	<u>Zn</u>	<u>Cu</u>	<u>Co</u>	<u>Ni</u>	<u>Pb</u>
Area I	56(2-5220)	19(1-1220)	18(2-144)	4.0(1-85)	5(<2-31)	9(2-39)	10(1-37)
Area II	49(7-10082)	14(4-370)	9.5(2-48)	3.0(<1-14)	11(2-39)	20(2-54)	30(<10-59)
(II/I)100 (%)	87	74	53	75	220	222	300

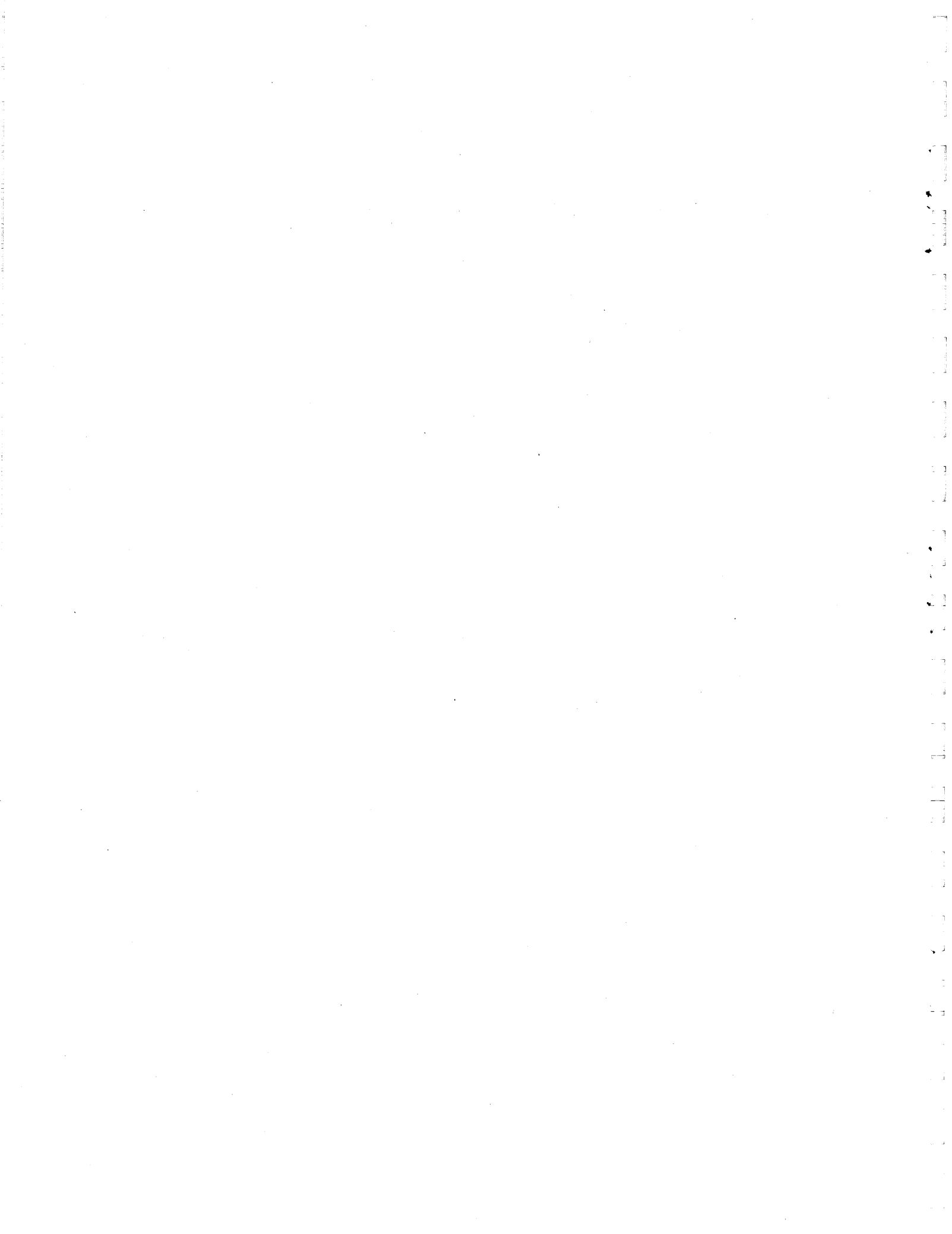


Table 1-5 (con't)

Mercury and Antimony Concentrations (ppb)

	<u>Hg</u>	<u>Sb*</u>
Area I	<0.1(<0.1-2.1)	<0.2(<0.2-0.3)
Area II	0.2(<0.1-1.2)	<0.5(<0.2-0.5)
(II/I)100 (%)	>200	-

*limited number of samples

	<u>% Saturation</u>			
	<u>% BaSO₄ Saturation</u>		<u>% SrSO₄ Saturation</u>	
	<u>Surface Temp.</u>	<u>Subsurface Temp.</u>	<u>Surface Temp.</u>	<u>Subsurface Temp.</u>
Area I	2.0(0.16-96)	1.1(0.09-39)	0.0065(<0.0001-0.504)	0.01(0.0001-0.867)
Area II	8.0(0.46-97)	6.5(0.24-38)	0.0048(0.0004-1.97)	0.007(0.0005-3.38)
(II/I)100 (%)	400	591	44	70

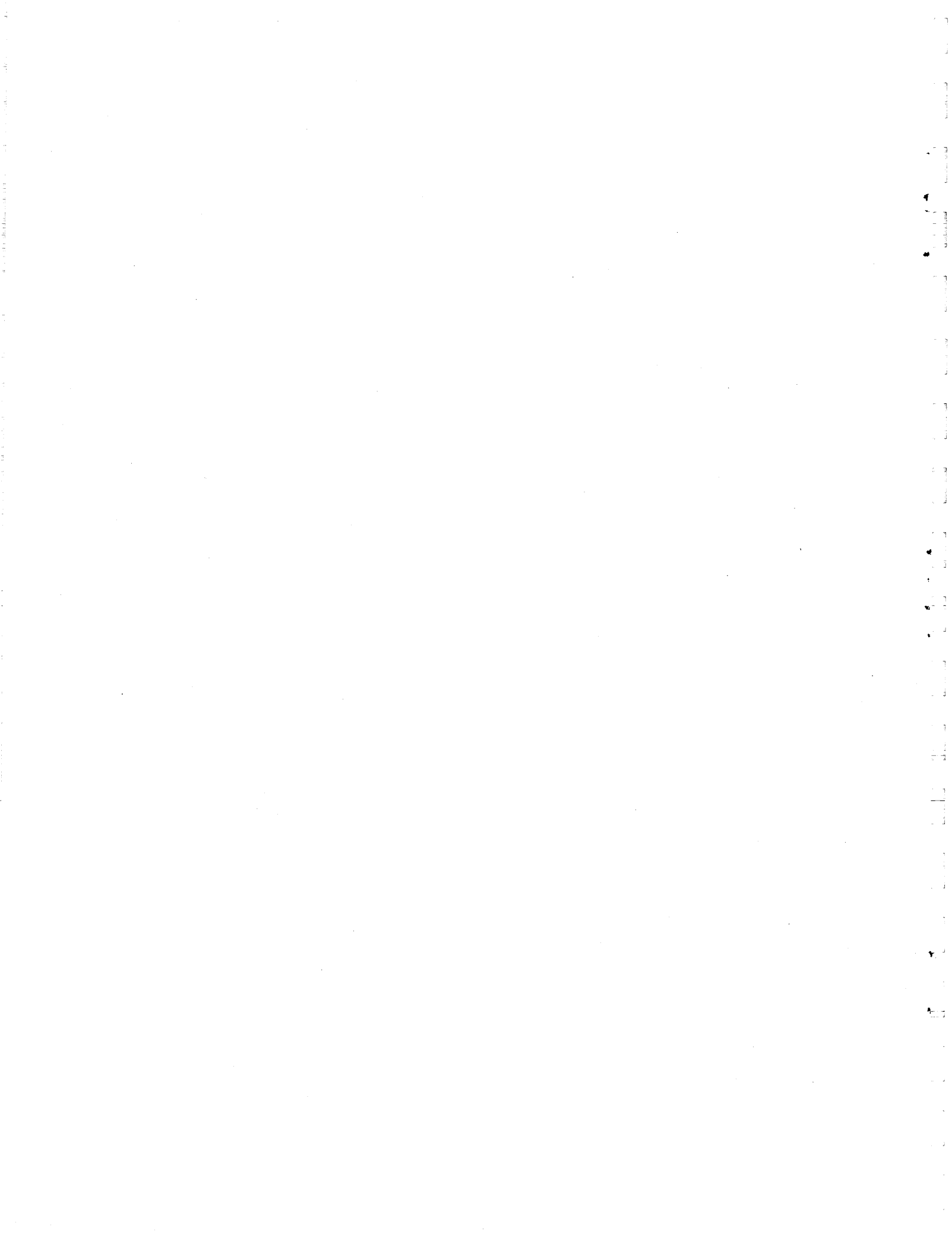


Table I-1

Comparison of Ouachita Mountain Spring Water Ranges
 And Median Values with Limits for Drinking Water
 And with Ranges and Median Values for Wells from the Ouachita Mountains
 (Well Data from Albin and Stephens (1963))

	<u>Limits</u>	<u>Springs Range</u>	<u>Springs Median</u>	<u>Wells Range</u>	<u>Wells Median</u>
Temperature, °C	-	5-35	16.5	13.8-22.2	18.3
pH	6.5-9.2 ^{b*}	3.8-7.9	6.2	5.0-8.6	7.5
Specific cond. μ mhos/cm 25°C	-	<5-563	85	34-1080	214
Total Alkalinity mg/l as CaCO ₃	-	0-291	37	2-378	63
NO ₃ as N ppm	10 ^{**}	<0.01-14.3	0.9	0-68	1.6
NH ₃ as N ppm	0.5 ^a	<0.01-1.52	0.15	-	-
PO ₄ as P ppm	-	<0.01-0.54	0.08	-	-
SO ₄ ppm	400 ^b	<0.3-32	3	0-221	10
Cl ppm	600 ^{b*}	0.25-87	2.9	2-288	11
SiO ₂ ppm	-	3.2-31.9	11	5.1-25.0	8.0
Na ppm	-	0.85-162	2.85	1.6-167.0	13.0
K ppm	-	0.04-8.5	0.83	0.3-22.0	1.2
Li ppb	-	<1-86	2.4	-	-
Ca ppm	200 ^{b*}	0.03-78	16.2	1-131	11
Mg ppm	150 ^b	0.15-23.7	1.65	0.5-40.0	5.1
Sr ppm	-	<0.01-2.2	0.04	-	-
Ba ppm	1 ^a	0.001-0.92	0.024	-	-
Fe ppb	300 ^{a*}	2-10,082	53	-	-
Mn ppb	50 ^{a*}	1-1220	17	-	-
Zn ppb	15,000 ^{b*}	2-144	14	-	-
Cu ppb	1,000 ^{a*}	<1-85	4	-	-
Co ppb	-	<2-39	8	-	-
Ni ppb	-	2-54	15	-	-
Pb ppb	50 ^a	1-59	20	-	-
Hg ppb	2 ^a	<0.1-2.1	0.1	-	-
Sb ppb	-	<0.2-0.5	<0.3	-	-

* limit based solely on welfare
 ** EPA, 1976

^a PHS, 1962
^b WHO, 1971

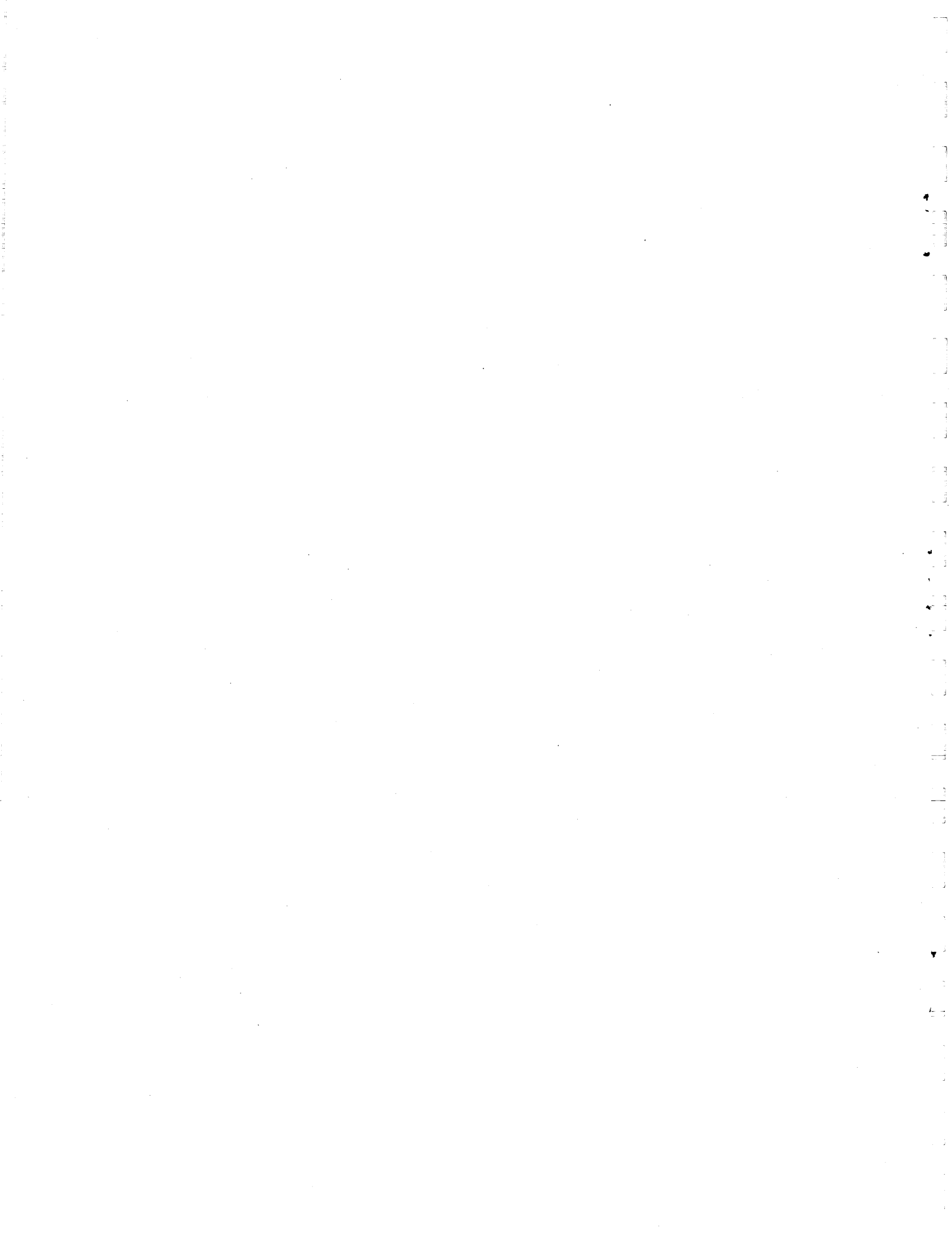


Table I-2
Data on Sulfate and Carbonate Minerals of Ca, Sr and Ba

Mineral	Chemical Formula	Solubility Product (mo6) ² /l ² at 25°C	Reference
gypsum	CaSO ₄ .2H ₂ O	1.49x10 ⁻⁴ =K ₁	Posnjak (1938)
celestite	SrSO ₄	1.37x10 ⁻⁷ =K ₂	Figure III-11
barite	BaSO ₄	1.25x10 ⁻¹⁰ =K ₃	Figure III-11
calcite	CaCO ₃	4.57x10 ⁻⁹ =K ₄	Garrels & Christ (1965)
strontianite	SrCO ₃	4.8x10 ⁻¹⁰ =K ₅	Helz & Holland (1965)
witherite	BaCO ₃	2.0x10 ⁻⁹ =K ₆	Garrels et al. (1960)

Equilibrium Equation for Two Solids

Equilibrium Activity Ratio (Moles/Mole)

SrSO ₄ (C) + Ca ⁺⁺ + 2H ₂ O = CaSO ₄ .2H ₂ O(c) + Sr ⁺⁺	a _{Sr⁺⁺} /a _{Ca⁺⁺} = K ₂ /K ₁ = 0.92 x 10 ⁻³
BaSO ₄ (C) + Ca ⁺⁺ + 2H ₂ O = CaSO ₄ .2H ₂ O(c) + Ba ⁺⁺	a _{Ba⁺⁺} /a _{Ca⁺⁺} = K ₃ /K ₁ = 0.84 x 10 ⁻⁶
BaSO ₄ (C) + Sr ⁺⁺ = SrSO ₄ (c) + Ba ⁺⁺	a _{Ba⁺⁺} /a _{Sr⁺⁺} = K ₃ /K ₂ = 0.91 x 10 ⁻³
SrCO ₃ (C) + Ca ⁺⁺ = CaCO ₃ (c) + Sr ⁺⁺	a _{Sr⁺⁺} /a _{Ca⁺⁺} = K ₅ /K ₄ = 0.10
BaCO ₃ (C) + Ca ⁺⁺ = CaCO ₃ (c) + Ba ⁺⁺	a _{Ba⁺⁺} /a _{Ca⁺⁺} = K ₆ /K ₄ = 0.44
BaCO ₃ (C) + Sr ⁺⁺ = SrCO ₃ (c) + Ba ⁺⁺	a _{Ba⁺⁺} /a _{Sr⁺⁺} = K ₆ /K ₅ = 4.2

Minerals in Equilibrium	Ionic Activity Ratio	Median Values Found For Activity Ratios*		Equilibrium Values* Activity Ratios
		Area I	Area II	
celestite=gypsum	a _{Sr⁺⁺} /a _{Ca⁺⁺}	4.1x10 ⁻³	11x10 ⁻³	2.0x10 ⁻³
barite=gypsum	a _{Ba⁺⁺} /a _{Ca⁺⁺}	1100x10 ⁻⁶	15000x10 ⁻⁶	2.9x10 ⁻⁶
barite=celestite	a _{Ba⁺⁺} /a _{Sr⁺⁺}	<0.27	1.3	1.4x10 ⁻⁶
strontianite=calcite	a _{Sr⁺⁺} /a _{Ca⁺⁺}	0.0041	0.011	0.22
witherite=calcite	a _{Ba⁺⁺} /a _{Ca⁺⁺}	<0.0011	0.015	1.5
witherite=strontianite	a _{Ba⁺⁺} /a _{Sr⁺⁺}	<0.27	1.3	6.6

*units are ppm/ppm



Table II-1

Surface Temperature, Subsurface Temperature,
Heat Flow and Geothermal Gradient

<u>Sample #</u>	<u>Surface Temperature °C</u>	<u>SiO₂ ppm²</u>	<u>Subsurface Temperature °C</u>	<u>Heat Flow w/m²</u>	<u>Geothermal Gradient °C/km</u>
501	15.0	7.7	31.7	27.6	17.5
502	5.0	3.2	7.1	-9.1	-5.8
503	15.0	18.3	60.9	70.6	44.9
504	16.5	12.2	46.1	49.2	31.3
505	15.0	7.7	31.7	27.6	17.5
506	18.0	12.2	46.1	49.2	31.3
507	13.0	13.7	50.2	55.2	35.1
508	13.0	7.7	31.7	27.6	17.5
509	10.0	10.3	40.6	40.8	26.0
510	9.0	6.6	26.7	20.2	12.8
511	16.0	10.2	40.3	40.4	25.7
512	16.0	8.9	36.0	34.1	21.7
513	16.0	19.4	63.2	74.0	47.1
514	18.0	10.2	40.3	40.4	25.7
515	17.0	9.6	38.5	37.8	24.0
516	21.0	12.0	45.6	48.4	30.8
517	17.0	12.0	45.6	48.4	30.8
518	15.0	10.6	41.7	42.5	27.0
519	18.0	9.5	37.9	36.9	23.5
520	17.0	13.5	49.7	54.5	34.7
521	17.0	9.6	38.5	37.8	24.0
522	18.0	9.5	37.9	36.9	23.5
523	22.0	12.0	45.6	48.4	30.8
524	15.5	15.5	54.5	61.6	39.2
525	15.0	9.7	38.8	38.2	24.3
526	17.0	16.8	57.5	66.1	42.1
527	17.5	16.8	57.5	66.1	42.1
528	15.5	16.2	56.1	64.1	40.7
529	17.5	9.2	37.0	35.5	22.6

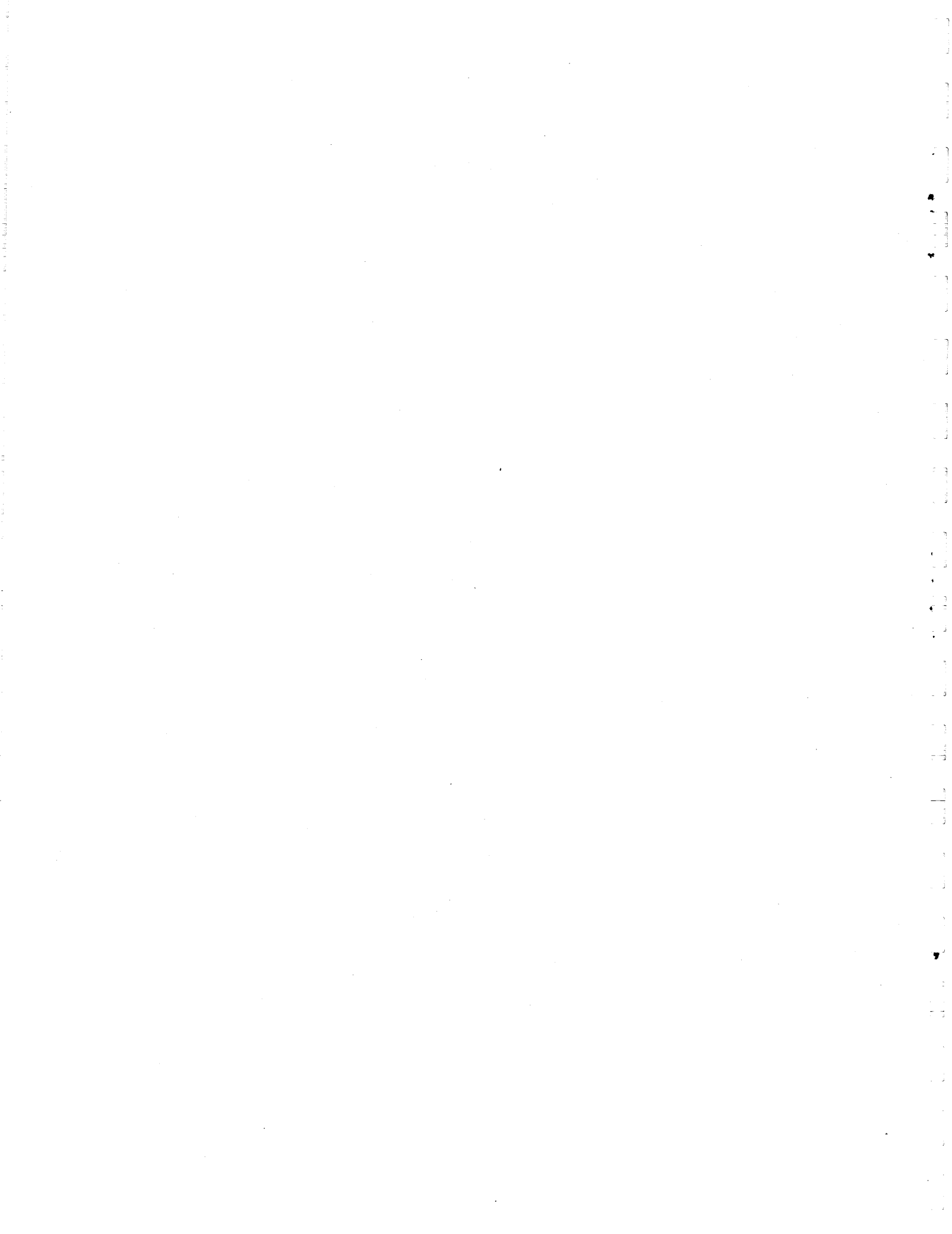


Table II-1 (con't)

Sample #	Surface Temperature °C	SiO ₂ ppm ²	Subsurface Temperature °C	Heat Flow w/m ²	Geothermal Gradient °C/km
530	15.0	9.2	37.0	35.5	22.6
531	17.5	10.1	40.0	40.0	25.4
532	18.0	9.7	38.8	38.2	24.3
533	17.5	10.1	40.0	40.0	25.4
534	15.5	7.0	28.7	23.1	14.7
535	16.0	7.5	30.8	26.3	16.7
536	18.0	9.2	37.0	35.5	22.6
537	15.0	8.3	33.8	30.7	19.5
538	19.0	8.5	34.6	32.0	20.4
539	17.0	7.5	30.8	26.3	16.7
540	18.0	7.2	29.6	24.5	15.6
541	16.0	8.8	35.8	33.7	21.4
542	16.0	8.5	34.6	32.0	20.4
543	20.0	9.7	38.8	38.2	24.3
544	18.0	10.1	40.0	40.0	25.4
545	20.0	9.7	38.8	38.2	24.3
546	21.0	10.1	40.0	40.0	25.4
547	20.0	10.4	41.1	41.7	26.5
548	22.0	10.4	41.1	41.7	26.5
899	17.0	10.8	42.2	43.3	27.6
900	18.0	9.3	37.3	36.0	22.9
901	18.0	9.1	36.7	35.0	22.3
902	20.0	10.5	41.4	42.1	26.8
903	18.0	11.6	44.6	46.9	29.8
904	18.0	8.7	35.4	33.1	21.1
905	20.0	9.9	39.4	39.1	24.9
906	20.0	5.4	21.0	11.7	7.4
907W	19.0	37.1	89.2	112.6	71.7
908	18.0	11.9	45.3	47.9	36.8
909	19.0	13.2	49.0	53.5	34.0
910	18.0	19.3	62.9	73.5	46.7



Table II-1 (con't)

<u>Sample #</u>	<u>Surface Temperature °C</u>	<u>SiO₂ ppm</u>	<u>Subsurface Temperature °C</u>	<u>Heat Flow w/m²</u>	<u>Geothermal Gradient °C/km</u>
943	15.0	11.4	44.1	46.1	29.3
944	18.5	10.8	42.2	43.3	27.6
945	23.0	11.1	43.1	44.7	28.4
946	16.0	13.6	50.0	55.0	35.0
947	18.0	12.5	47.0	50.4	32.1
948	16.0	8.0	32.7	29.1	18.5
949	13.0	16.0	55.6	63.3	40.3
950	18.0	12.5	47.0	50.4	32.1
951W	16.0	29.0	78.3	97.2	61.8
952	14.0	12.8	48.0	51.9	33.0
953	12.8	12.8	48.0	51.9	33.0
Mean	16.8	11.5	44.3	46.4	35.7

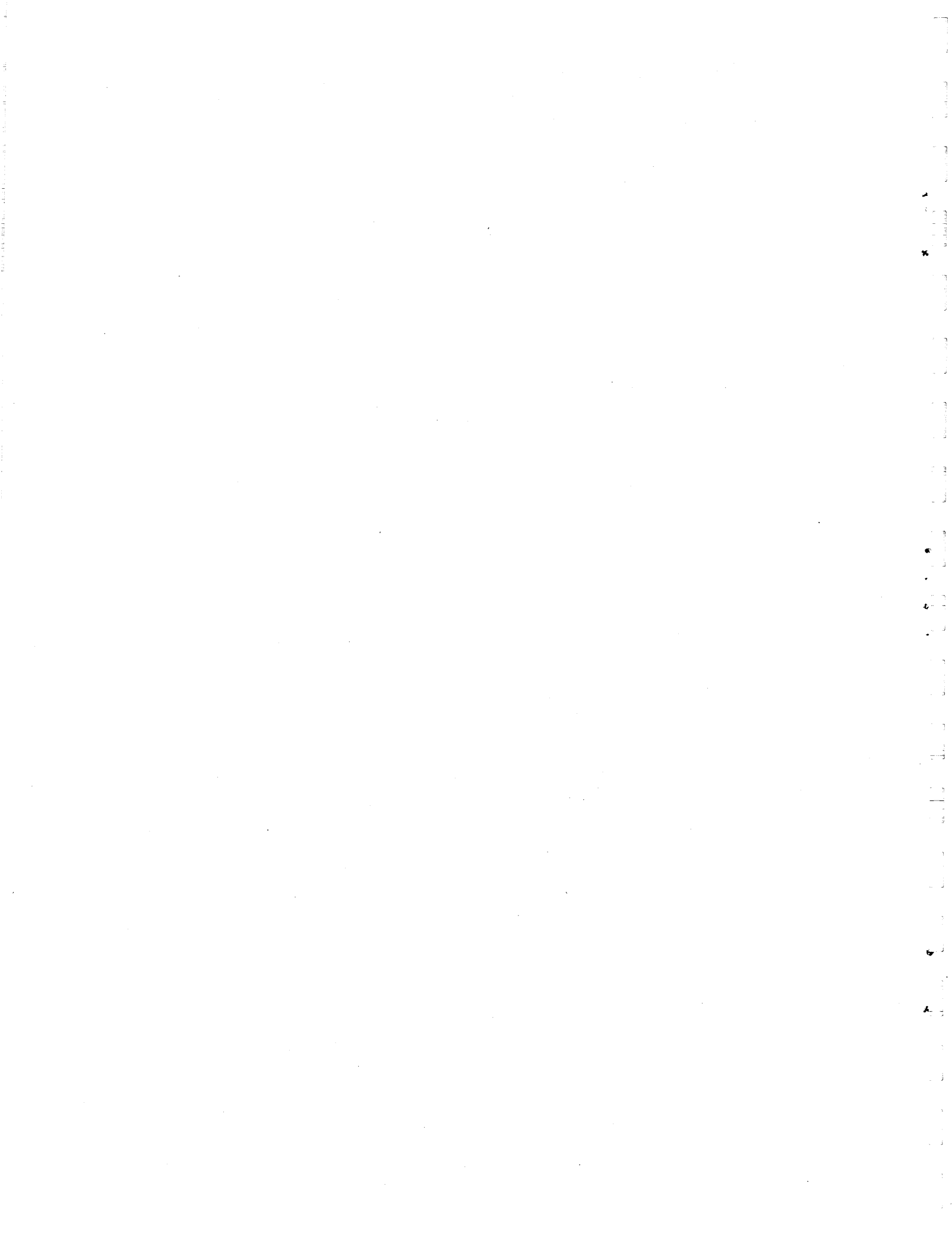
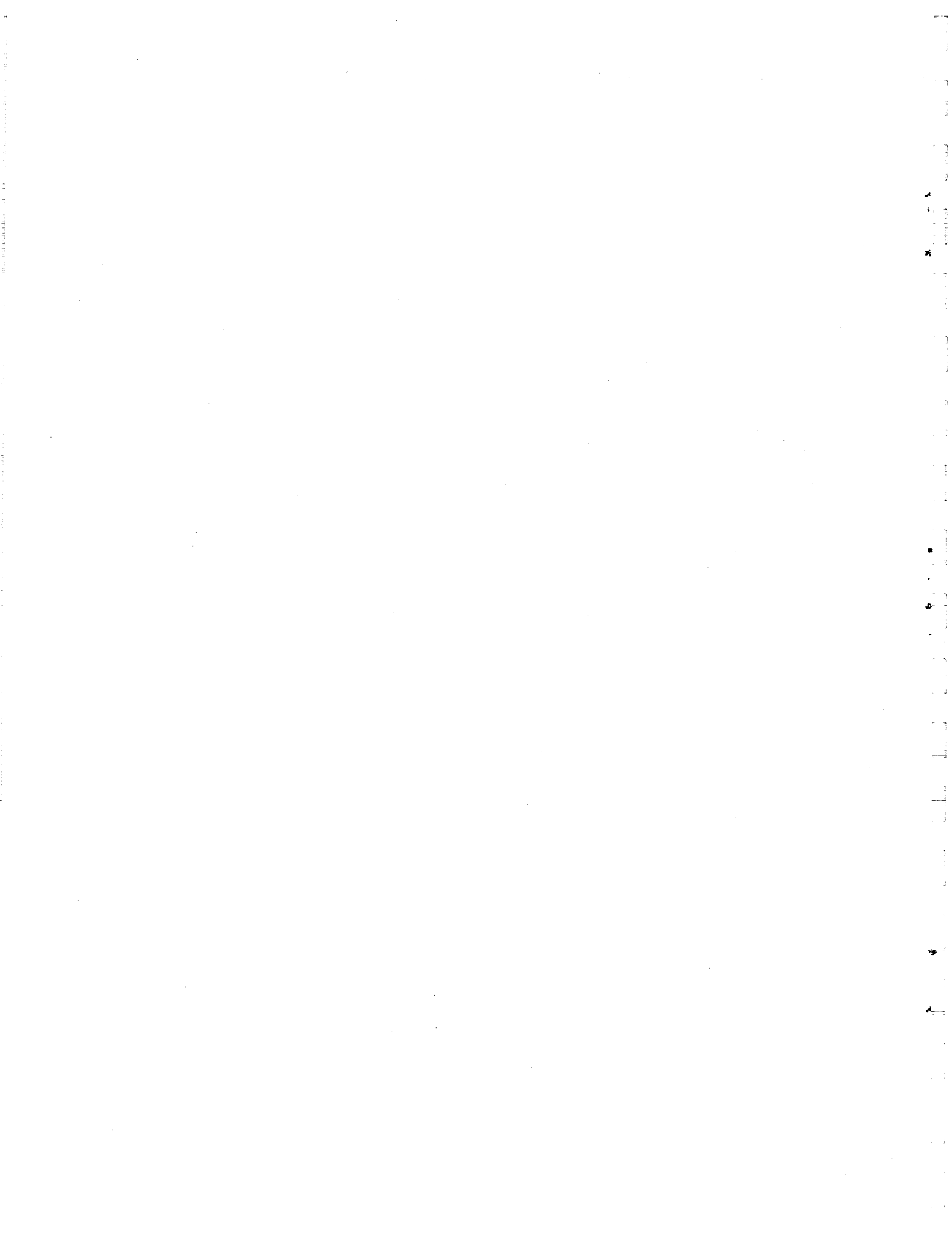


Table II-2

Anomalous Springs Based on Surface and SiO₂ Geotemperatures

Sample No.	Anomalous Surface Temperature °C	Anomalous SiO ₂ Geotemperature °C
503	-	60.4
513	-	63.2
523	22.0	-
548	22.0	-
907W	-	89.2
910	-	62.9
914	-	60.7
915	35.0	-
916W	-	91.2
920	-	68.8
921	30.0	-
923W	22.0	82.9
924	25.0	60.4
927W	-	65.7
928	-	75.8
932	-	73.4
936	-	70.0
945	23.0	-
951W	-	78.3



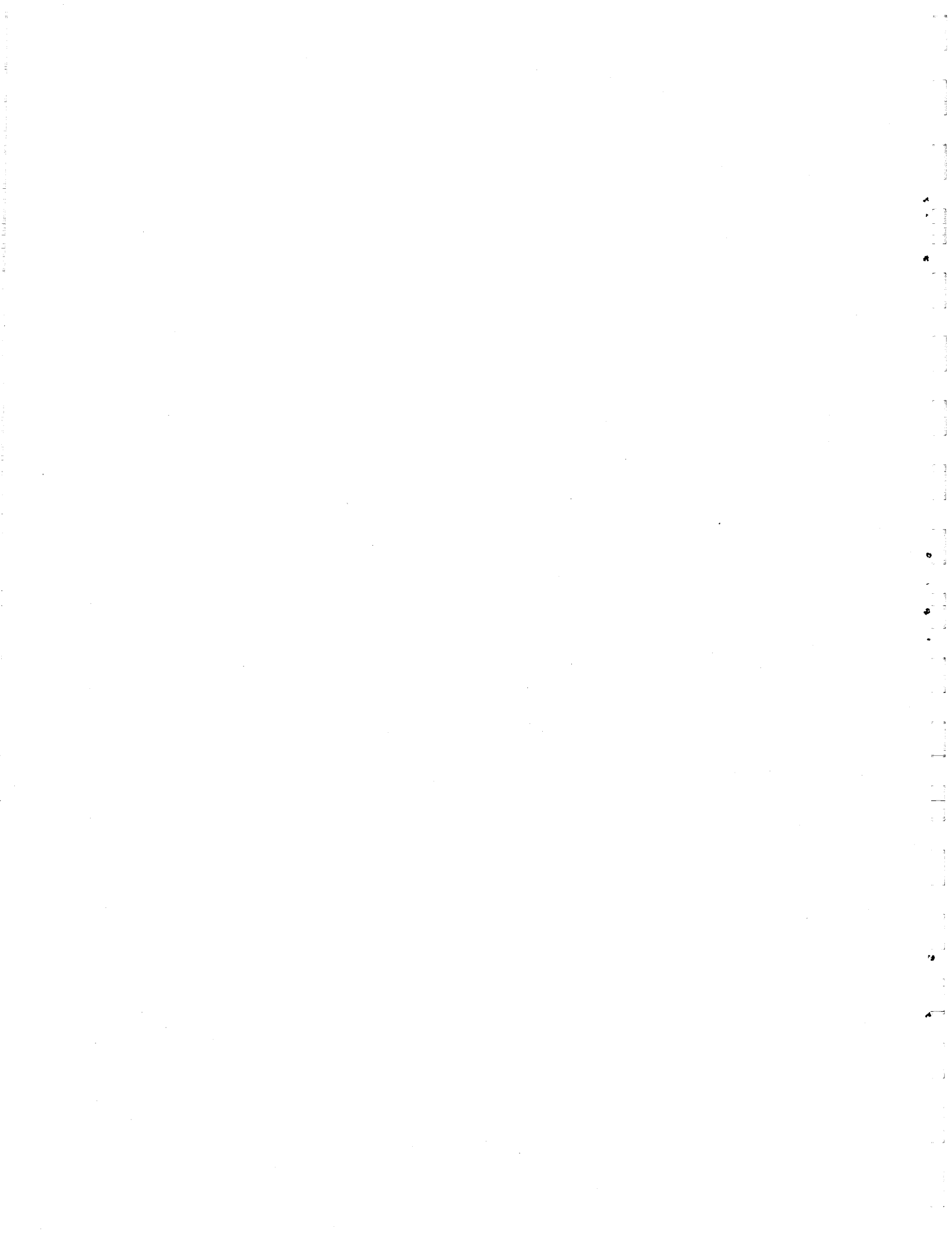
Summary of Anomalous Measurements

Table III-1

Measurement	Area	Units	Median	Threshold*	Average Anomalous Value**	Sample Numbers of Anomalous Springs
pH	I	pH units	6.4	7.5	7.67	506,527,545,899
	I	pH units	-	4.4	4.10	913
	II	pH units	5.8	7.5	8.15	927,937
	II	pH units	-	4.4	4.16	904,906,948
Specific conduc- tance at 25°C	I	µmhos/cm	107	320	364	503,526,527
Specific conduc- tance at 25°C	II	µmhos/cm	61	500	563	937
Total alkalinity	I	Mg CaCO ₃ /L	60	290	291	501
	II	Mg CaCO ₃ /L	15	200	265	937
NO ₃ as N	I	ppm	0.66	3.1	7.4	922,939,945
	II	ppm	1.00	12.0	14.3	904
NH ₃ as N	I	ppm	0.10	0.51	0.95	520,526,913,930
	II	ppm	0.20	1.40	1.52	925,926
PO ₄ as P	I	ppm	0.07	0.24	0.36	517,911,913
	II	ppm	0.10	0.35	0.41	951
Sulfate	I	ppm	2.3	9.0	16.8	503,517,916,917
	II	ppm	3.2	20.0	42.2	927,951
Chloride	I	ppm	2.3	4.4	48	526,527,922,953
Chloride	II	ppm	3.5	6.8	8.7	904,948
Fe	I	ppb	56	5000	5222	916
	II	ppb	49	580	6399	908,928,932,937
Mn	I	ppb	19	900	1220	517
	II	ppb	14	220	353	924,928,936,
Zn	I	ppb	18	100	134	536,537,900,911
	II	ppb	9.5	33	46	904,948
Cu	I	ppb	4.0	11	34	511,513,517,523
	II	ppb	3.0	7	12	904,919,920
Co	I	ppb	5	20	27	916,938,939,943
	II	ppb	11	32	38	908,936
Ni	I	ppb	9	28	36	517,916,938,943
	II	ppb	20	37	54	936
Hg	I	ppb	<0.1	<0.3	2.0	528,544
	II	ppb	0.2	0.6	1.1	909,910,937
Pb	I	ppb	10	31	35	520,916
	II	ppb	30	42	59	932
Ba	I	ppb	<15	320	420	520
	II	ppb	40	320	915	929,944
Sr	I	ppb	55	380	987	501W,526,527,911,913
	II	ppb	30	150	857	927W,951W
Li	I	ppb	2.5	18	46	501,520,526,527
	II	ppb	2.3	54	72	927,937
K	I	ppm	0.56	1.35	2.0	506,520W,922
	II	ppm	1.10	6.0	8.5	948

* 95% frequency value from cumulative frequency curve (Lepeltier, 1969), except Hg=1 ppb

** Anomalous values are those exceeding threshold



Anomalous Springs

AREA I

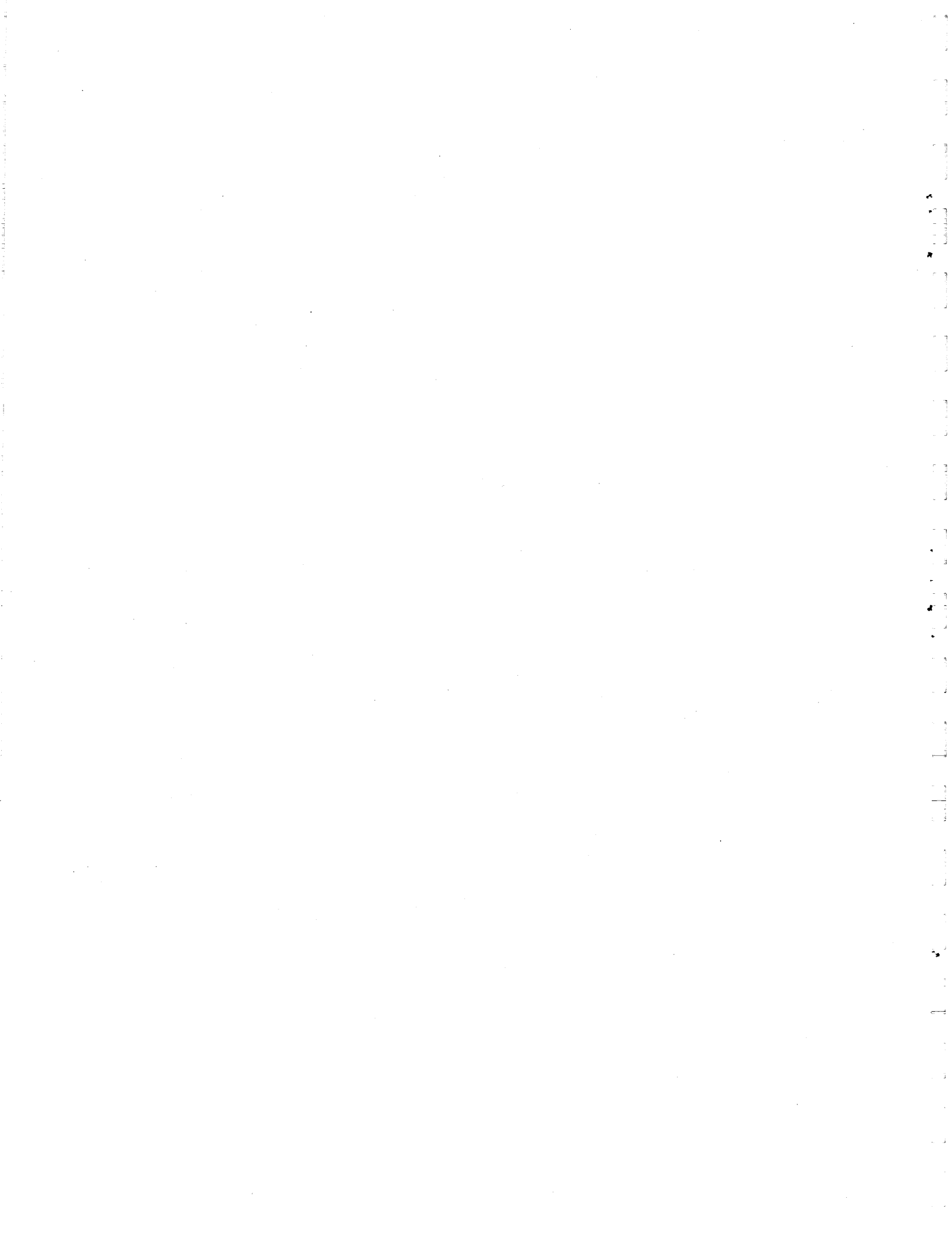
Table III-2

<u>Spring No.</u>	<u>Measurement (Value) for Which Spring is Anomalous</u>
501W	Alkalinity(291), Sr(501 ppb), Li(21 ppb)
503	Specific conductance(324), SO ₄ (32 ppm), SiO ₂ , subsurface T(60.4°C)
506	pH(7.6), K(1.7 ppm)
511	Cu(85 ppb)
513	Cu(15 ppb), SiO ₂ , subsurface T(62.2°C)
517	PO ₄ (0.54 ppm), SO ₄ (12 ppm), Mn(1220 ppb), Cu(21 ppb), Ni(29 ppb)
520W	NH ₃ (0.64 ppm), Pb(33 ppb), Li(23 ppb), Ba(420 ppb), K(1.4 ppm)
523	Cu(14 ppb), surface T(22°C)
526	Specific conductance(353), NH ₃ (0.54 ppm), Cl(86 ppm), Sr(838 ppb), Li(70 ppb)
527	pH(7.6), specific conductance(385), Cl(87 ppm), Sr(802 ppb), Li(70 ppb)
528	Hg(2.1 ppb)
536	Zn(143 ppb)
537	Zn(110 ppb)
544	Hg(1.9 ppb)
545	pH(7.6)
899	pH(7.9)
900	Zn (137 ppb)
911	PO ₄ (0.28 ppm), Zn(144 ppb), Sr(500 ppb)
913	NH ₃ (1.08 ppm), PO ₄ (0.25 ppm), Sr(2200 ppb), pH(4.1)
916W	SO ₄ (13.1 ppm), Fe(5222 ppb), Co(23 ppb), Ni(39 ppb), Pb(37 ppb), SiO ₂ , subsurface T(91.2°C)
917	SO ₄ (10.2 ppm)
922	NO ₃ (12 ppm), Cl(13 ppm), K(2.9 ppm)
930	NH ₃ (1.52 ppm)
938	Co(22 ppb), Ni(36 ppb)
939	NO ₃ (4.8 ppm), Co(30 ppb)
943	Co(31 ppb), Ni(39 ppb)
945	NO ₃ (5.5 ppm), surface T(23°C)
953	Cl(6 ppm)

$$\Sigma(\text{anomalous springs}) = 28, \% \text{ anomalous} = \left(\frac{28}{71}\right)100 = 39$$

$$\Sigma(\text{springs}) = 71$$

$$\Sigma(\text{anomalous springs}), \text{ cations only} = 21, \% = \left(\frac{21}{71}\right)100 = 30$$



Anomalous Springs

AREA II

Table III-3

<u>Spring No.</u>	<u>Measurement (Value) for Which Spring is Anomalous</u>
904	pH(4.1 units), NO ₃ (14.3 ppm), Cl(9.5 ppm), Zn(48 ppb), Cu(14 ppb)
906	pH(4.3 units)
908	Fe(10082 ppb), Co(39 ppb)
909	Hg(1.1 ppb)
910	Hg(1.2 ppb), SiO ₂ , subsurface T(62.9°C)
919	Cu(13 ppb),
920	Cu(9 ppb), SiO ₂ , subsurface T(68.8°C)
924	Mn(345 ppb), SiO ₂ , surface T(25°C), subsurface T(60.4°C)
925	NH ₃ (1.52 ppm)
926	NH ₃ (1.52 ppm)
927W	pH(7.6 units), SO ₄ (29.5 ppm), Sr(1440 ppb), Li(58 ppb), SiO ₂ , subsurface T(65.7°C)
928	Fe(4218 ppb), Mn(345 ppb), SiO ₂ , subsurface T(75.8°C)
929	Ba(910 ppb)
932	Fe(1375 ppb), Pb(59 ppb), SiO ₂ , subsurface T(73.4°C)
936	Mn(370 ppb), Co(38 ppb), Ni(54 ppb), SiO ₂ , subsurface T(70.0°C)
937	pH(8.7 units), specific conductance(563 μmhos), alkalinity (265 mg/L CaCO ₃) Fe(9930 ppb), Li(86 ppb), Hg(1.1 ppb)
944	Ba(920 ppb)
948	pH(4.1 units), Cl(8.0 ppm), Zn(45 ppb), K(8.5 ppm)
951W	PO ₄ (0.41 ppm), SO ₄ (55 ppm), Sr(275 ppb), SiO ₂ , subsurface T(78.3°C)

$$\Sigma(\text{anomalous springs}) = 19, \% \text{ anomalous} = \left(\frac{19}{31}\right)100 = 61$$

$$\Sigma(\text{anomalous springs}), \text{ cations only} = 16, \% \text{ anomalous} = \left(\frac{16}{31}\right)100 = 52$$

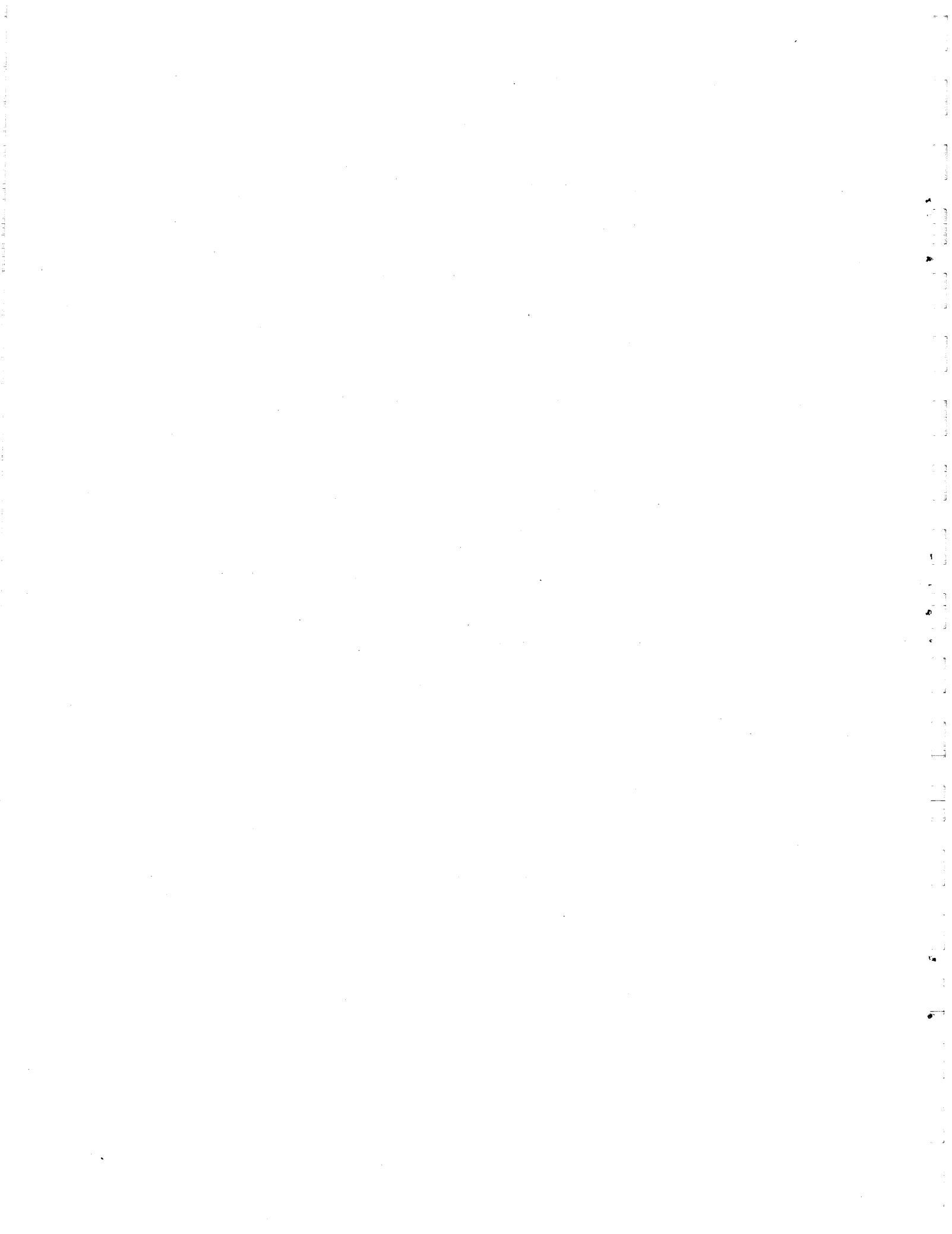
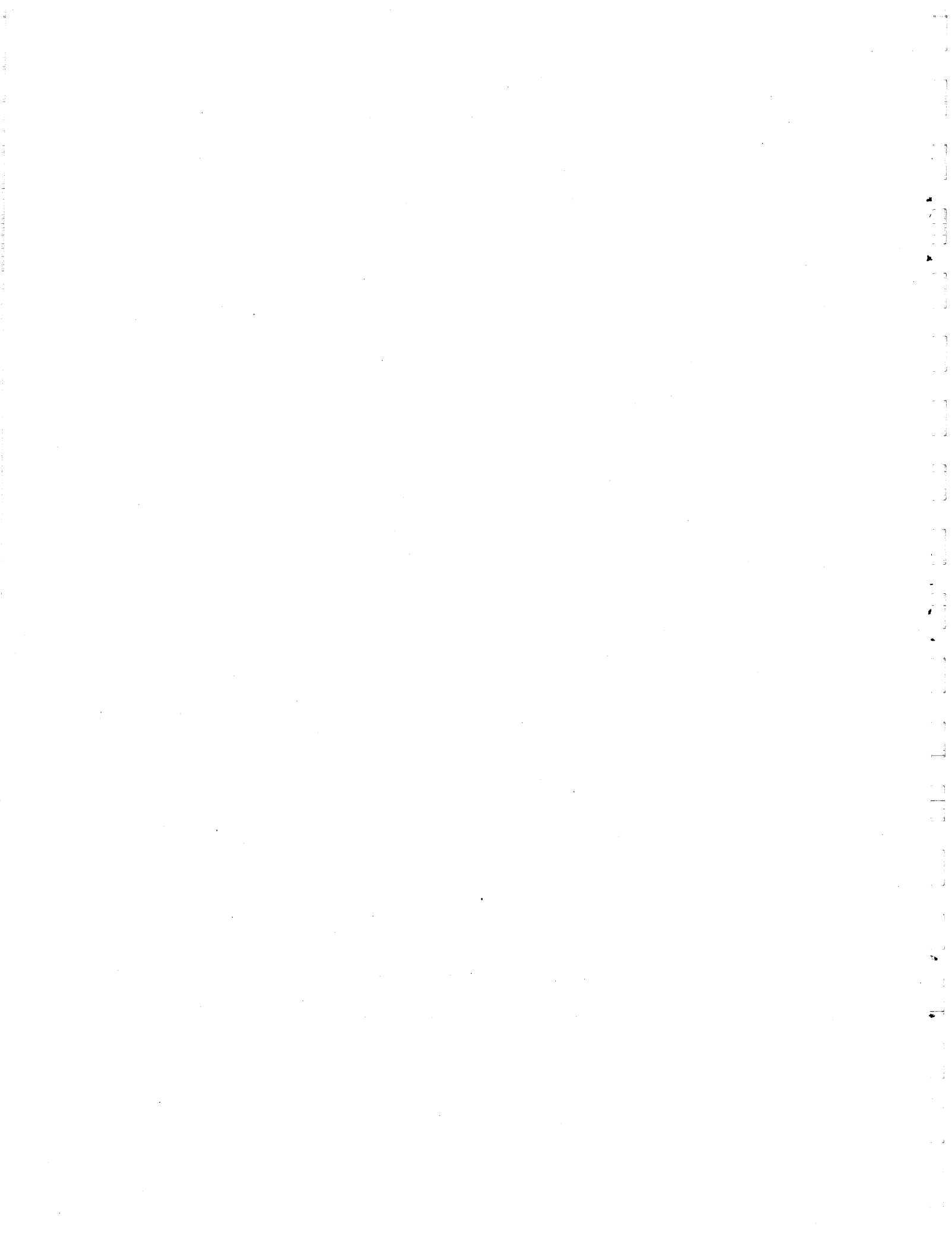


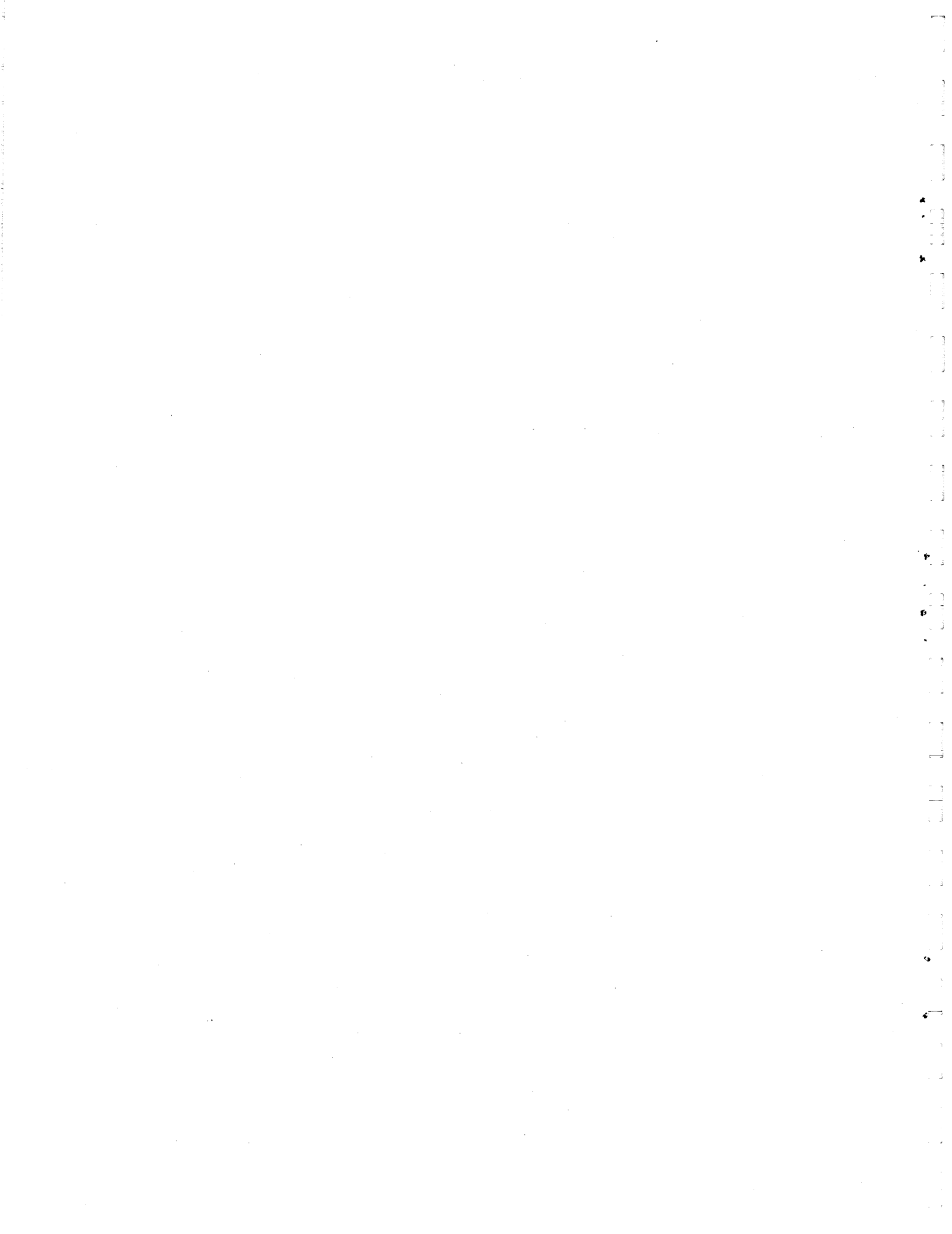
Table III-4
Location and Date of Collection of Spring Samples

Sample No.	Name	Location	Area	County	Topo Quad	Date
501**	Iron	NE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 6, T1N, R9W	I	Garland	Nimrod	3/2/78
502	Iron	NE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 6, T1N, R19W	I	Garland	Nimrod	" "
503	Black	SW $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 19, T3S, R25W	I	Montgomery	Glenwood	3/3/78
504	Collier	SE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 17, T3S, R24W	I	Montgomery	Glenwood*	" "
505	Caddo Valley	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 25, T3S, R27W	I	Montgomery	Athens*	" "
506	Crystal A	SE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 28, T2S, R22W	I	Garland	Crystal Spring	" "
507	Goat Brown	NW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 22, T3S, R31W	I	Polk	Cove	3/6/78
508	Bogg	NW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 16, T5S, R32W	II	Polk	Cove*	" "
509	Queen Wilhelmina	SE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 11, T1S, R32W	I	Polk	Rich Mtn.	3/7/78
510	Queen Wilhelmina	SE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 11, T1S, R32W	I	Polk	Rich Mtn.	" "
511	Silver World	NE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 14, T3S, R29W	I	Polk	Umpire	" "
512	Abernathy	NW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 24, T3S, R28W	I	Polk	Athens*	" "
513	Three Sisters	NW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 29, T1S, R20W	I	Garland	Mtn. Pine	4/8/78
514	(No Name)	NE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 16, T3S, R22W	I	Garland	Percy	" "
515	(No Name)	NE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 16, T3S, R22W	I	Garland	Percy	" "
516	(No Name)	NE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 16, T3S, R22W	I	Garland	Percy	" "
517	Wyatt	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 22, T3S, R23W	I	Montgomery	Bonnerdale*	" "
518	(No Name)	NE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 24, T4S, R24W	I	Montgomery	Glenwood	4/9/78
519	(No Name)	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 20, T4S, R23W	I	Montgomery	Amity	" "
520***, ****	Pigeon Roost (W)	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 20, T4S, R23W	I	Montgomery	Amity	" "
521	Dripping	SE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 11, T5S, R25W	I	Pike	Glenwood*	" "
522	Strawn	SE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood*	" "
523	(No Name)	SE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 16, T1S, R28W	I	Polk	Oden	5/16/78
524	(No Name)	SE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 26, T4S, R32W	I	Polk	Cove	" "



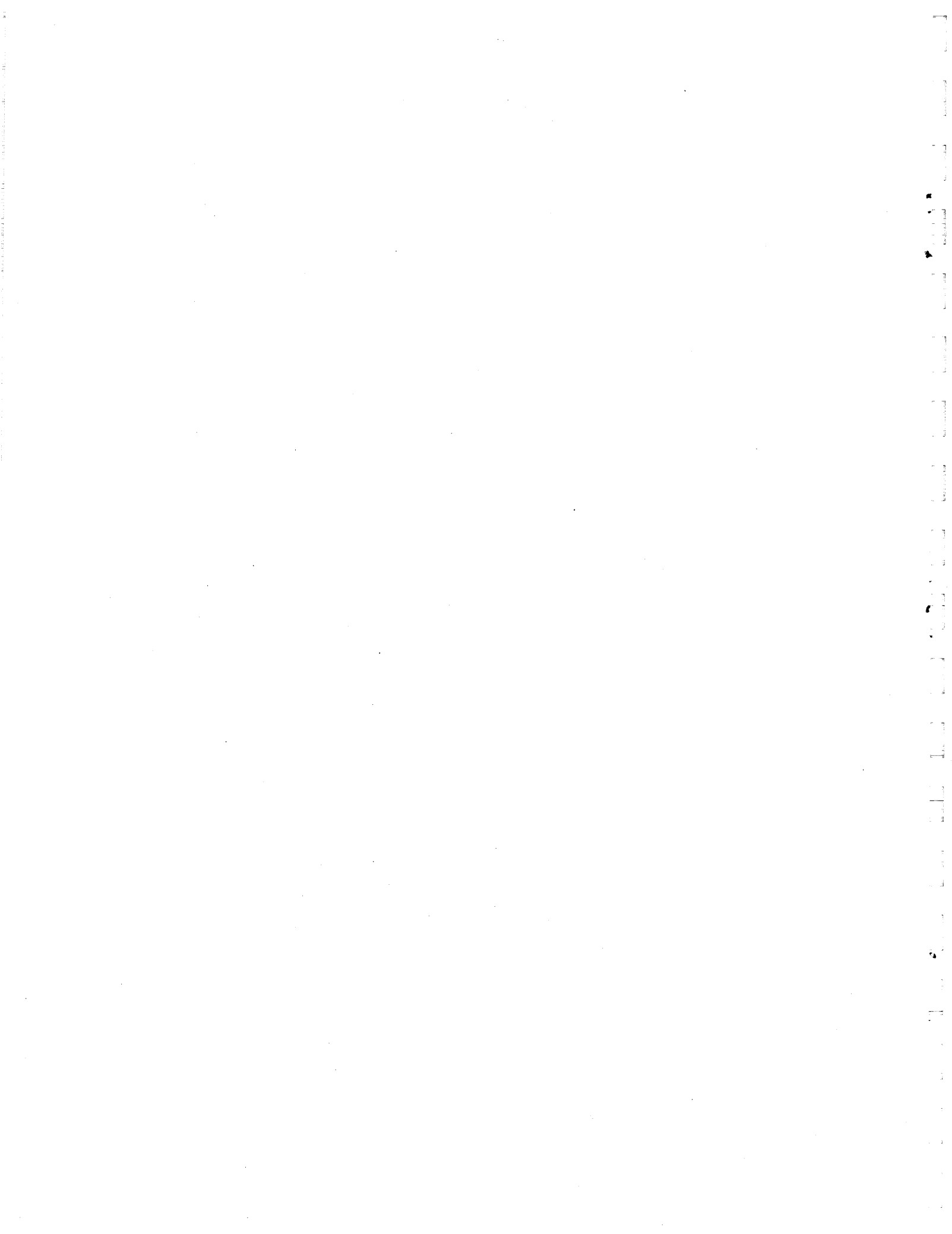
Location and Date of Collection of Spring Samples (con't)

<u>Sample No.</u>	<u>Name</u>	<u>Location</u>	<u>Area</u>	<u>County</u>	<u>Topo Quad</u>	<u>Date</u>
525	(No Name)	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 3, T2S, R32W	I	Polk	Cove*	5/17/78
526	Gilham A	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 22, T4S, R30W	I	Polk	Umpire*	" "
527	Gilham B	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 22, T4S, R30W	I	Polk	Umpire*	" "
528	Bard	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 20, T4S, R28W	I	Polk	Umpire	" "
529	(No Name)	NW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 27, T4S, R27W	I	Montgomery	Athens	" "
530	(No Name)	NW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 22, T4S, R27W	I	Montgomery	Athens	" "
531	McClaine A	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	5/18/78
532	McClaine B	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	" "
533	McClaine C	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	" "
534	McClaine D	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	" "
535	W. Jones	NE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	" "
536	Jones Valley	NE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	" "
537	Buttermilk A	NW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 6, T4S, R24W	I	Montgomery	Glenwood	" "
538	Buttermilk B	NW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 6, T4S, R24W	I	Montgomery	Glenwood	" "
539	(No Name)	NW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 6, T4S, R24W	I	Montgomery	Glenwood	" "
540	McElthanan	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 7, T4S, R24W	I	Montgomery	Glenwood	" "
541	(No Name)	NW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	" "
542	(No Name)	NW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 18, T4S, R24W	I	Montgomery	Glenwood	" "
543	Crystal B	SE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 34, T2S, R22W	I	Garland	Glenwood	" "
544	Brogdams A	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 34, T2S, R22W	I	Garland	Crystal Spring	5/19/78
545	Brogdams B	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 34, T2S, R22W	I	Garland	Crystal Spring	" "
546	Brogdams C	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 34, T2S, R22W	I	Garland	Crystal Spring	" "



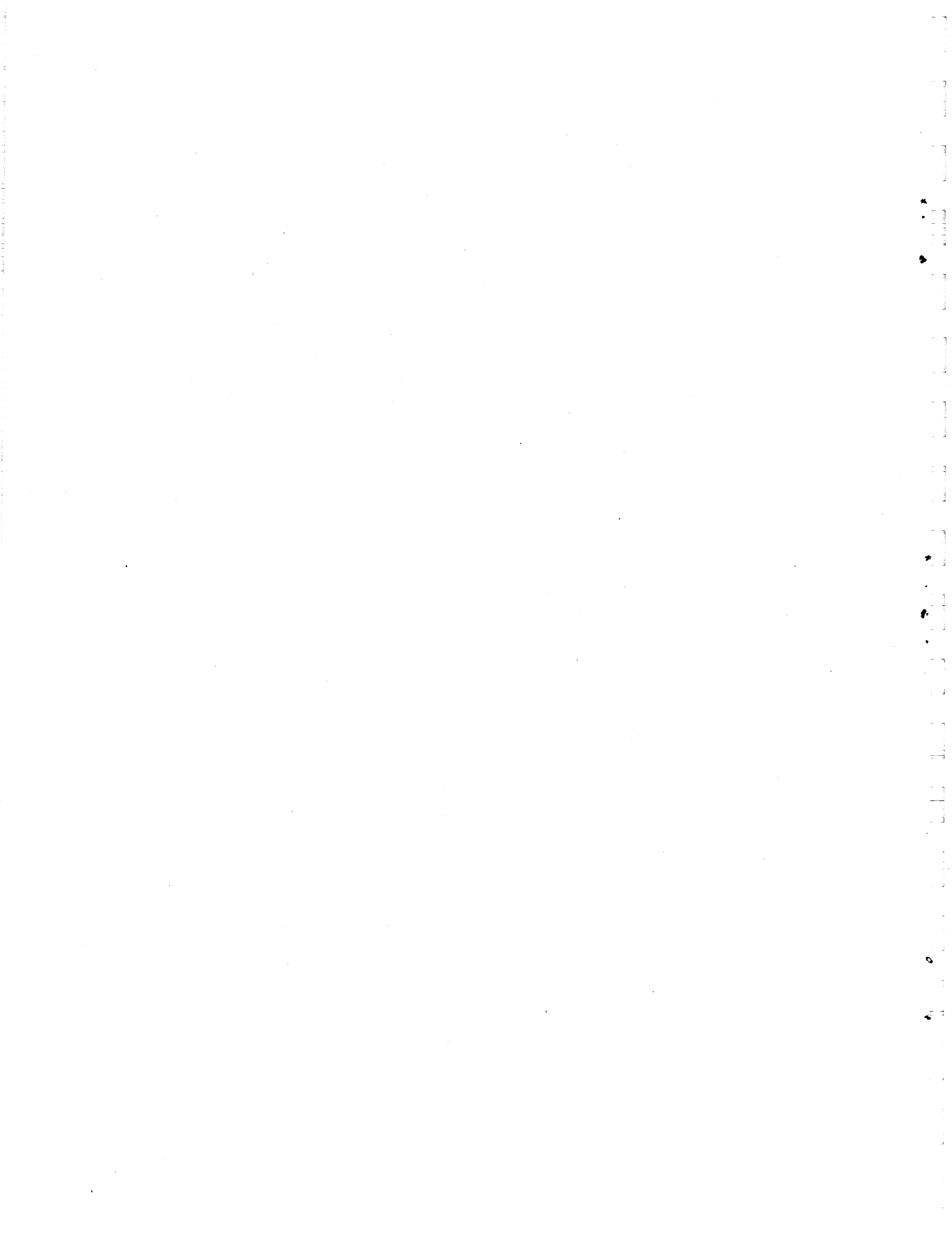
Location and Date of Collection of Spring Samples (con't)

<u>Sample No.</u>	<u>Name</u>	<u>Location</u>	<u>Area</u>	<u>County</u>	<u>Topo Quad</u>	<u>Date</u>
547	Brogdams D	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 34, T2S, R22W	I	Garland	Crystal Spring	5/19/78
548	Brogdams E	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 34, T2S, R22W	I	Garland	Crystal Spring	" "
899	Burrows	SE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 6, T5S, R22W	I	Clark	Amity	7/13/78
900	Lead Mine Road	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 30, T4S, R21W	I	Hot Spring	Point Cedar	" "
901	Lead Mine Road	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 30, T4S, R21W	I	Hot Spring	Point Cedar	" "
902	(No Name)	SE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 31, T4S, R21W	I	Hot Spring	Point Cedar*	" "
903	" "	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 36, T4S, R21W	I	Hot Spring	Point Cedar	" "
904	" "	SE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 35, T7S, R25W	II	Pike	Murfreesboro*7/14/78	" "
905**	Womack	NW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 25, T7S, R25W	II	Pike	Delight	" "
906	Womack	SE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 25, T7S, R25W	II	Pike	Delight	" "
907**	Lolla Bell	NW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 3, T7S, R25W	II	Pike	Narrows Dam	" "
908	(No Name)	NW $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 26, T6S, R24W	II	Pike	Murfreesboro NE	" "
909	" "	SE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 32, T6S, R24W	II	Pike	Murfreesboro NE	" "
910	" "	SE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 31, T7S, R25W	I	Pike	Murfreesboro	7/15/78
911	" "	SE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 20, T3S, R26W	I	Montgomery	Athens	" "
912	" "	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 33, T3S, R26W	I	Montgomery	Athens	" "
913	Sul fur	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 10, T3S, R27W	I	Montgomery	Athens	" "
914	Elliott	NE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 21, T2S, R29W	I	Polk	Board Camp	7/18/78
915	Caddo River	NW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 19, T4S, R24W	I	Montgomery	Glenwood	" "
916***,****	Artesian Well	SE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 19, T4S, R26W	I	Montgomery	Athens	" "
917***,****	Artesian Well	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 19, T4S, R26W	I	Montgomery	Athens	" "
918	(No Name)	NE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 28, T4S, R26W	I	Montgomery	Athens	" "
919	(No Name)	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 16, T6S, R27W	II	Pike	Newhope	" "



Location and Date of Collection of Spring Samples (con't)

<u>Sample No.</u>	<u>Name</u>	<u>Location</u>	<u>Area</u>	<u>County</u>	<u>Topo Quad</u>	<u>Date</u>
920	(No Name)	SE $\frac{1}{4}$, sec. 20, T7S, R27W	II	Howard	Newhope	7/18/78
921	(No Name)	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 26, T4S, R25W	I	Montgomery	Glenwood	" "
922	Chalybeate	NE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 20, T4N, R24W	I	Yell	Gravelly	3/3/79
923	Caddo Gap Well	SE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 26, T4S, R25W	I	Montgomery	Glenwood	" "
924	Redland Mtn.	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 12, T5S, R26W	I	Pike	Glenwood	3/4/79
925	Lolla Bell A	NW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 3, T7S, R25W	II	Pike	Narrows Dam	" "
926	" B	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 3, T7S, R25W	II	Pike	Narrows Dam	" "
927**	Artesian Well	SE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 3, T10S, R27W	II	Howard	Nashville	" "
928	Mineral	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 19, T10S, R27W	II	Howard	Mineral Springs S	" "
929	Barite Pit	NE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 24, T7S, R28W	II	Howard	Newhope	" "
930	Murfreesboro Area	NE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 4, T8S, R25W	II	Pike	Murfreesboro	" "
931	Murfreesboro Area	NE $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 10, T7S, R25W	II	Pike	Narrows Dam	3/5/79
932	Lake Greeson Area	NE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 24, T6S, R27W	II	Pike	Center Pt. NE	" "
933	Dierks Area	NE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 4, T8S, R29W	II	Sevier	Dierks	" "
934	Cox Residence	SE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 33, T7S, R29W	II	Sevier	Dierks	" "
935	Dierks Area	SW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 34, T7S, R29W	II	Howard	Dierks	" "
936	L. Greeson, Sulfur	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 34, T6S, R26W	II	Pike	Center Pt. NE	3/6/79
937	L. Greeson, Possum	NW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 34, T6S, R26W	II	Pike	Center Pt. NE	" "
938	Salem Area	NE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 7, T5S, R24W	I	Pike	Glenwood	" "
939	Bethesda	NW $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 31, T2S, R30W	I	Polk	Mena	5/15/79
940	Mine Creek	SW $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 14, T3S, R28W	I	Polk	Athens	" "
941	Rock	SW $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 25, T2S, R27W	I	Montgomery	Oden*	" "
942	West	SE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 2, T2S, R25W	I	Montgomery	Mt. Ida*	" "
943	Pigeon Roost	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 20, T4S, R23W	I	Montgomery	Amity	5/16/79
944	Barite Pit	NE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 24, T7S, R28W	II	Howard	Newhope	" "



Location and Date of Collection of Spring Samples (con't)

<u>Sample No.</u>	<u>Name</u>	<u>Location</u>	<u>Area</u>	<u>County</u>	<u>Topo Quad</u>	<u>Date</u>
945	"Warin"	SW $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 8, T4S, R27W	I	Montgomery	Athens	5/17/79
946	Athens Area	SW $\frac{1}{4}$, SE $\frac{1}{4}$, sec. 16, T5S, R28W	II	Howard	Athens	" "
947	Umpire Area	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 13, T5S, R30W	II	Howard	Umpire	" "
948	S. Cox	SW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 33, T7S, R29W	II	Sevier	Dierks	" "
949	Cossatot	NW $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 19, T7S, R30W	II	Sevier	Gillham Dam	" "
950	Church	NW $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 16, T7S, R31W	II	Sevier	Gillham*	" "
951	Defore Well	NE $\frac{1}{4}$, SW $\frac{1}{4}$, sec. 20, T7S, R31W	II	Sevier	Gillham	5/18/79
952	(No Name)	SE $\frac{1}{4}$, NE $\frac{1}{4}$, sec. 6, T7S, R31W	II	Sevier	Gillham*	" "
953	Mena Park	NE $\frac{1}{4}$, NW $\frac{1}{4}$, sec. 18, T2S, R30W	I	Polk	Mena	" "

* spring is shown on map

** well with metal pipe

*** well with plastic pipe

**** artesian well

NOTE: 944 and 929 are duplicate collections
of the same site.



Spring Water Surface and Subsurface Temperatures
And Corresponding % Saturations of the Waters with BaSO₄ and SrSO₄

Table III-5

Spring	Area	Water Temperature (°C)		% BaSO ₄ Saturation ^c		% SrSO ₄ Saturation ^d	
		Surface ^a (T ₁)	Subsurface ^b (T ₂)	Surface	Subsurface	Surface	Subsurface
501	I	15.0	31.7	0.51	0.30	0.0077	0.0089
502	I	5.0	7.1	<0.47	<0.38	0.0001	0.0001
503	I	15.0	60.9	<14.4	<5.38	0.504	0.867
504	I	16.5	46.1	1.37	0.67	0.0014	0.0022
505	I	15.0	31.7	<0.80	<0.47	0.0097	0.0113
506	I	18.0	46.1	<0.50	<0.25	0.0035	0.0052
507	I	13.0	50.2	<0.76	<0.28	0.0020	0.0032
508	II	13.0	31.7	<1.4	<0.73	0.0072	0.0086
509	I	10.0	40.6	<2.5	<0.93	0.0024	0.0034
510	I	9.0	26.7	1.33	0.59	0.0004	0.0005
511	I	16.0	40.3	<2.60	<1.31	0.0311	0.0401
512	I	16.0	36.0	<1.10	<0.61	0.0152	0.0185
513	I	16.0	63.2	<3.40	<1.31	0.0303	0.0531
514	I	18.0	40.3	<0.62	<0.34	<0.0001	<0.0002
515	I	17.0	38.5	<0.92	<0.50	<0.0001	<0.0001
516	I	21.0	45.6	<1.95	<1.11	0.0069	0.0097
517	I	17.0	45.6	<8.0	<3.90	<0.0016	<0.0024
518	I	15.0	41.7	<0.46	<0.22	0.0002	0.0003
519	I	18.0	37.9	16.5	9.41	0.0733	0.0210
520	I	17.0	49.7	52.8	24.3	0.0424	0.0652
521	I	17.0	38.5	<1.3	<0.71	0.0011	0.0014
522	I	18.0	37.9	<1.8	<0.97	0.0469	0.0565
523	I	22.0	45.6	<0.32	<0.19	<0.0002	<0.0004
524	I	15.5	54.5	<0.21	<0.08	<0.0002	<0.0004
525	I	15.0	38.8	<2.0	<0.99	0.0063	0.0081
526	I	17.0	57.5	2.71	1.14	0.0111	0.0130
527	I	17.5	57.5	<2.60	<1.12	<0.0102	<0.0166
528	I	15.5	56.1	<3.8	<1.5	0.0517	0.0857
529	I	17.5	37.0	<0.19	<0.11	<0.0002	<0.0004

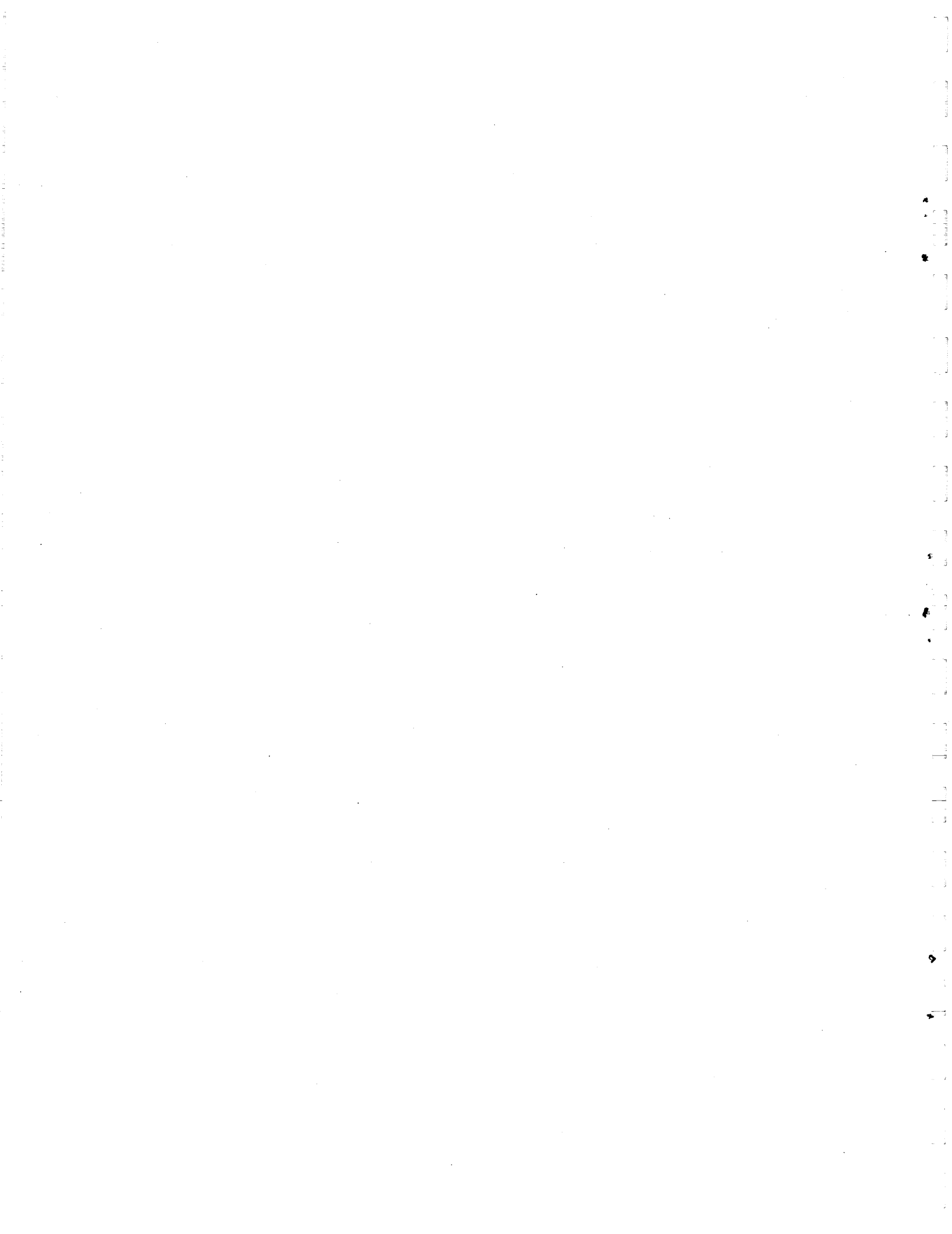


Table III-5 (con't)

Spring	Area	Water Temperature ($^{\circ}\text{C}$)		% BaSO_4 Saturation ^c		% SrSO_4 Saturation ^d	
		Surface ^a (T_1)	Subsurface ^b (T_2)	Surface	Subsurface	Surface	Subsurface
530	I	15.0	37.0	<1.83	<0.95	0.0032	0.0040
531	I	17.5	40.0	2.16	1.18	0.0327	0.0409
532	I	18.0	38.8	1.90	1.07	0.0150	0.0185
533	I	17.5	40.0	<2.20	<1.21	0.0390	0.0487
534	I	15.5	28.7	2.96	1.87	0.0025	0.0028
535	I	16.0	30.8	<0.20	<0.12	0.0007	0.0008
536	I	18.0	37.0	<2.76	<1.61	0.0592	0.0706
537	I	15.0	33.8	1.57	0.87	0.0026	0.0032
538	I	19.0	34.6	<3.17	<2.05	0.0563	0.0644
539	I	17.0	30.8	1.65	1.06	0.0025	0.0028
540	I	18.0	29.6	<0.20	<0.14	0.0003	0.0003
541	I	16.0	35.8	0.86	0.48	0.0003	0.0003
542	I	16.0	34.6	0.16	0.09	0.0001	0.0003
543	I	20.0	38.8	<3.47	<2.16	0.0354	0.0419
544	I	18.0	40.0	<0.99	<0.55	0.0169	0.0209
545	I	20.0	38.8	<0.98	<0.61	0.0107	0.0126
546	I	21.0	40.0	<4.40	<2.75	0.0417	0.0493
547	I	20.0	41.1	<2.65	<1.58	0.0240	0.0293
548	I	22.0	41.1	<2.94	<1.86	0.0325	0.0385
899	I	17.0	42.2	<0.41	<0.21	<0.0015	<0.0022
900	I	18.0	37.3	1.29	0.74	0.0002	0.0003
901	I	18.0	36.7	0.66	0.39	0.0066	0.0078
902	I	20.0	41.4	1.70	1.01	0.0002	0.0002
903	I	18.0	44.6	0.49	0.25	0.0026	0.0039
904	II	18.0	35.4	35.0	21.0	0.0131	0.0155
905	II	20.0	39.4	2.27	1.39	0.0222	0.0266
906	II	20.0	21.0	6.66	6.48	0.0019	0.0019
907W	II	19.0	89.2	68.7	25.2	0.0314	0.0647
908	II	18.0	45.3	8.00	3.80	0.0005	0.0008
909	II	19.0	49.0	3.67	1.85	0.0004	0.0006
910	II	18.0	62.9	63.8	26.2	0.0094	0.0164
911	I	16.0	44.6	13.4	6.35	0.1170	0.1730

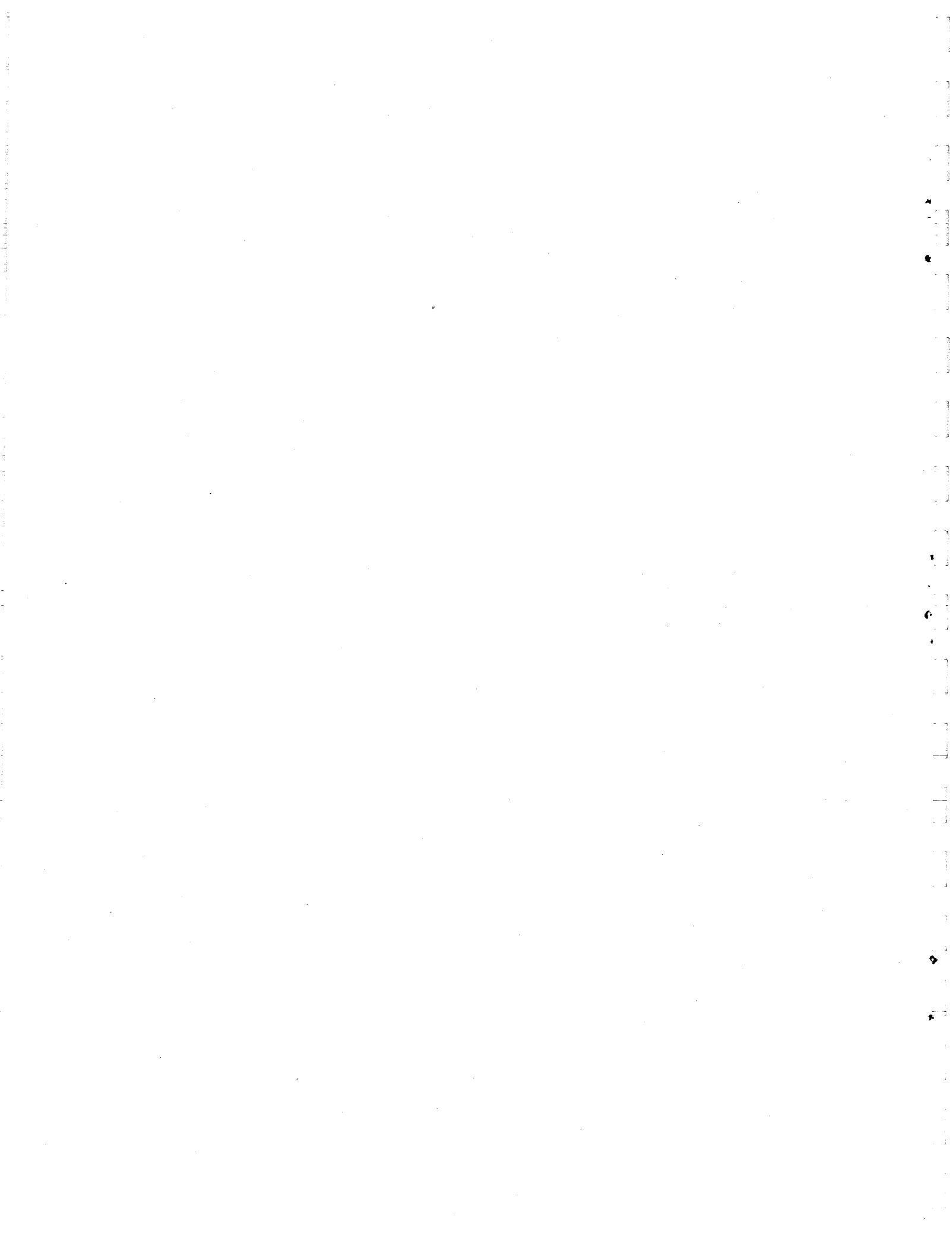


Table III-5 (con't)

Spring	Area	Water Temperature ($^{\circ}\text{C}$)		% BaSO_4 Saturation ^c		% SrSO_4 Saturation ^d	
		Surface ^a (T_1)	Subsurface ^b (T_2)	Surface	Subsurface	Surface	Subsurface
912	I	16.5	37.3	4.09	2.23	0.0367	0.0451
913	I	17.0	50.9	4.52	2.05	0.1520	0.2360
914	I	20.0	60.7	0.36	0.17	0.0003	0.0005
915	I	35.0	58.3	0.60	0.42	0.0021	0.0031
916W	I	18.0	91.2	96.0	34.7	0.0207	0.0480
917	I	18.0	44.6	51.6	26.3	0.0237	0.0349
918	I	20.0	35.4	11.3	7.46	0.0055	0.0063
919	II	21.0	58.4	1.10	0.54	0.0022	0.0036
920	II	20.0	68.8	0.37	0.20	0.0016	0.0024
921	I	30.0	41.4	<0.96	<0.77	0.0171	0.0205
922	I	10.0	31.8	54.2	23.8	0.0196	0.0245
923W	I	22.0	82.9	22.4	9.46	0.0264	0.0529
924	II	25.0	60.4	5.90	3.17	0.0343	0.0511
925	II	12.0	41.3	9.29	3.77	0.0057	0.0080
926	II	12.0	39.5	17.8	7.34	0.0049	0.0067
927W	II	18.5	65.7	92.0	37.8	1.97	3.38
928	II	16.0	75.8	3.93	1.36	0.0186	0.0456
929	II	10.0	41.3	45.1	16.5	0.0011	0.0015
930	I	11.0	29.3	23.3	11.3	0.0036	0.0043
931	II	11.0	17.4	16.4	11.6	0.0036	0.0043
932	II	13.0	73.4	96.7	28.7	0.0128	0.0261
933	II	10.0	31.0	16.8	7.54	0.0023	0.0028
934	II	12.0	30.1	13.1	6.64	0.0079	0.0094
935	II	10.0	26.8	8.41	4.20	0.0025	0.0029
936	II	11.0	70.0	50.8	13.8	0.0092	0.0187
937	II	8.0	46.0	7.50	2.09	0.0045	0.0075
938	I	12.0	39.5	23.6	9.73	0.0081	0.0110
939	I	19.0	48.8	<1.96	<1.00	0.0096	0.0141
940	I	16.0	36.9	<2.73	<1.48	0.0444	0.0532
941	I	17.0	49.0	<3.60	<1.69	0.0211	0.0318
942	I	15.0	31.8	1.91	1.10	0.0017	0.0020
943	I	15.0	44.1	86.0	39.3	0.0280	0.0424



Table III-5 Con't

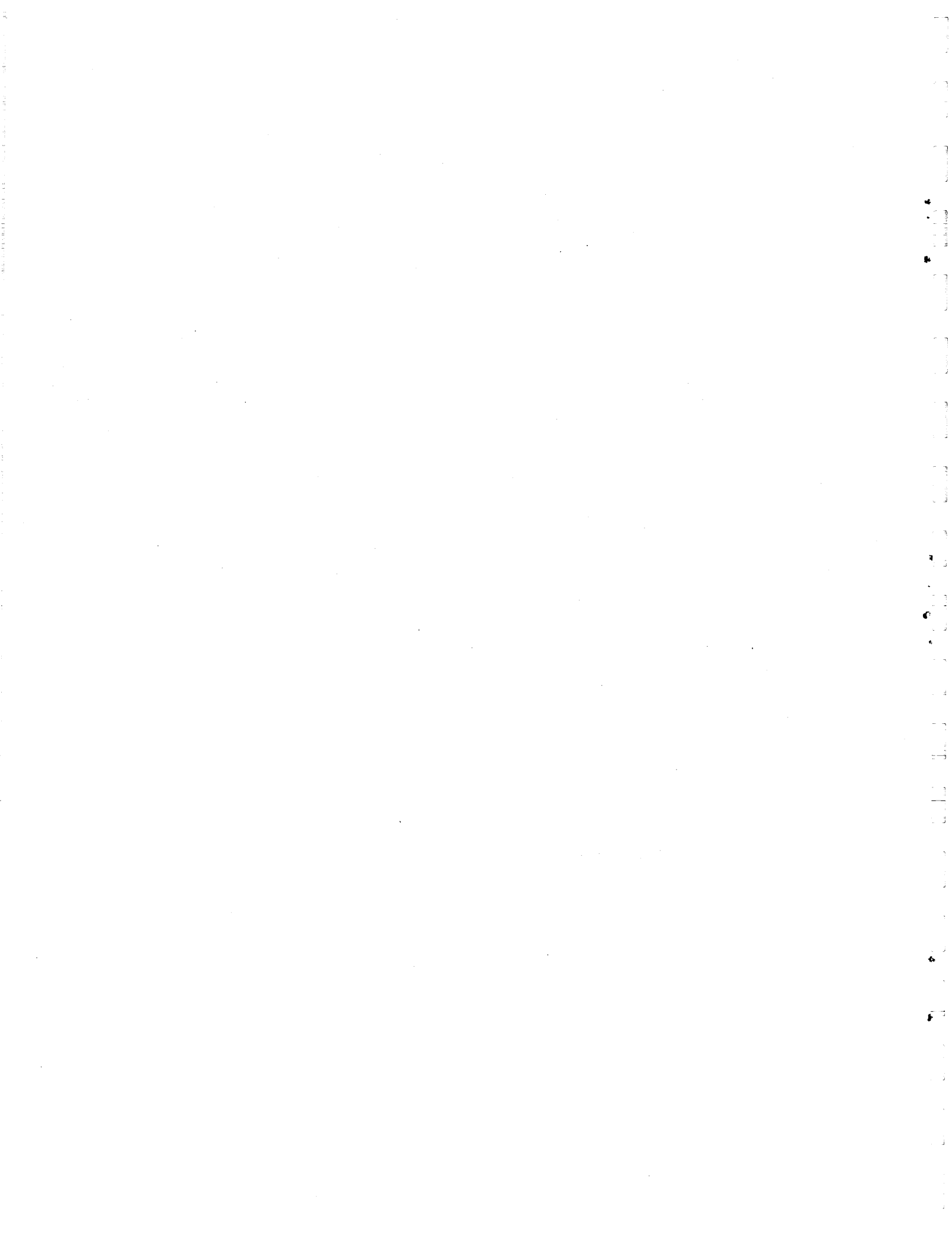
Spring	Area	Water Temperature (°C)		% BaSO ₄ Saturation ^c		% SrSO ₄ Saturation ^d	
		Surface ^a (T ₁)	Subsurface ^b (T ₂)	Surface	Subsurface	Surface	Subsurface
944	II	18.5	42.2	63.10	34.3	0.0015	0.0022
945	I	23.0	43.1	<2.30	<1.46	0.0056	0.0073
946	II	16.0	50.0	3.71	1.63	<0.0018	<0.0030
947	II	18.0	47.0	<0.68	<0.34	0.0004	0.0006
948	II	16.0	32.7	15.20	9.00	0.0057	0.0065
949	II	13.0	55.6	<2.50	<0.86	0.0033	0.0059
950	II	18.0	47.0	<3.20	<1.59	0.0047	0.0070
951W	II	16.0	78.3	43.5	14.8	0.606	1.22
952	II	14.0	48.0	2.22	0.90	0.0016	0.0018
953	I	16.0	48.0	<11.1	<5.00	0.0110	0.0171

^a Measured in flowing spring at surface

^b Calculated from $T\text{ }^{\circ}\text{C} = 1315 / [5.205 - \log_{10}(\text{ppm SiO}_2)] - 273.15$

^c $a_{\text{Ba}} \cdot a_{\text{SO}_4} / (\text{solubility product})_{T_1} \times 100$

^d $a_{\text{Sr}} \cdot a_{\text{SO}_4} / (\text{solubility product})_{T_2} \times 100$ where a = activity = activity coefficient x concentrations (molarity)



FIGURES

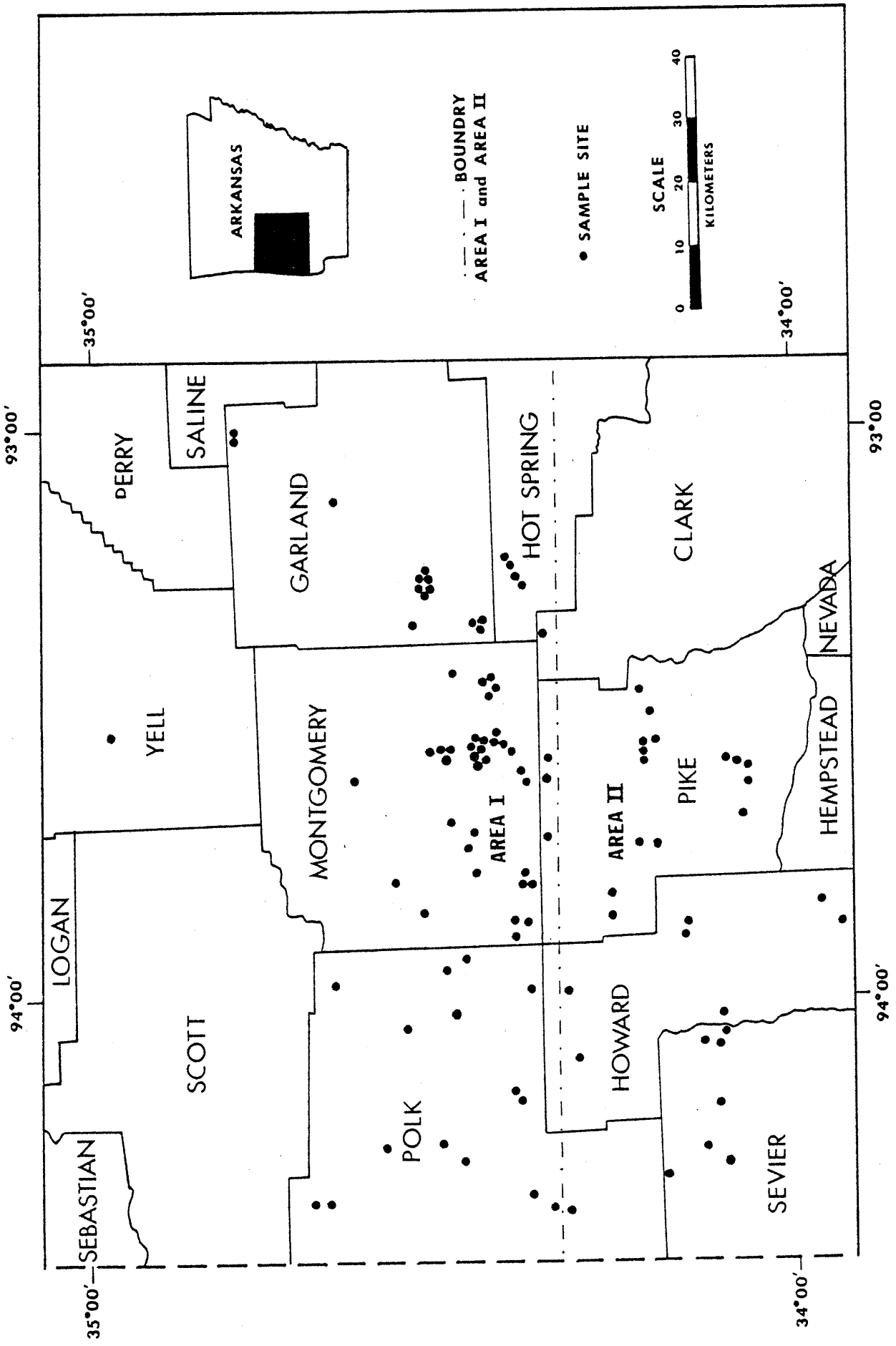


Figure 1-1
Location Map of Study Area and Spring Sites

SYSTEM	SERIES	FORMATION	MEMBER	THICKNESS (METERS)	LITHOLOGY	
QUATERNARY		ALLUVIUM			SILT, SAND, AND GRAVEL ALONG STREAM CHANNELS.	
		TERRACE DEPOSITS			GRAVEL, SAND, AND SILT OCCURRING ALONG MAJOR STREAMS.	
CRETACEOUS		TRINITY FORMATION			GRAVEL, SILT, CLAY, SILTSTONE, AND SANDSTONE WITH SOME BARITE IN WESTERN PART OF ITS OUTCROP.	
PENNSYLVANIAN	ATOKAN	ATOKA	UPPER	6000+	SHALE, LIGHT GRAY, SILTY, MICACEOUS, AND FLAKY WITH INTERBEDDED FINE TO COARSE-GRAINED, MICACEOUS SANDSTONE WITH VERY ABUNDANT SOLE MARKINGS. THIN SILICEOUS SHALES NEAR BASE AND IN LOWER PART OF FORMATION.	
			MIDDLE			
			LOWER			
MORROWAN	JOHNS VALLEY SHALE	JACKPORN SANDSTONE		500	SHALE, LIGHT GRAY TO TAN, DARK GRAY NEAR BASE, AND THIN BEDS OF SANDSTONE AND LIMESTONE. LARGE ERRATIC MASSES OF LIMESTONE OR SHALE ARE FOUND NEAR THE BASE OF THE FORMATION, AND EXOTIC BOULDERS, PEBBLES, AND GRANULES OCCUR AT NUMEROUS HORIZONS.	
				2300	SANDSTONE, MEDIUM TO COARSE GRAINED, HARD, WITH INTERBEDDED SHALE. SOLE MARKINGS ARE ABUNDANT IN THE SANDSTONES. FOUR BEDS OF SILICEOUS SHALE AND ONE BED OF MAROON TO GREEN SHALE ARE IDENTIFIABLE OVER LONG DISTANCES AND FORM MARKER BEDS.	
MISSISSIPPIAN	CHESTERIAN	STANLEY GROUP	CHICKASAW CREEK	4000	SHALE, DARK COLORED, MOSTLY GRAY, INTERBEDDED WITH DARK GRAY ARGILLACEOUS SILTSTONE AND VERY POORLY SORTED FINE- TO VERY FINE-GRAINED ARGILLACEOUS CHLORITIC SANDSTONE. BEDS OF SILICEOUS SHALE ARE IDENTIFIABLE OVER LONG DISTANCES IN SEVERAL HORIZONS. CONE-IN-CONE CONCRETIONS ARE ABUNDANT AT PLACES.	
	MERAMECIAN		NATTON TUFF		0-50	FELSIC VITRIC TUFF
			HOT SPRINGS SANDSTONE		0-20	SANDSTONE, HARD, QUARTZOSE, FINE TO VERY FINE GRAINED, SMALL AMOUNTS OF INTERBEDDED SHALE AND LOCALLY CONGLOMERATIC NEAR BASE. CROPS OUT ONLY IN RELATIVELY SMALL AREA NEAR HOT SPRINGS.
	OSAGEAN	ARKANSAS NOVACULITE	UPPER	310	GREEN, BROWN, AND GRAY RADIOLARIAN CHERT AND RADIOLARIAN SHALE.	
KINDERHOOKIAN	UPPER MIDDLE		RED AND GREEN RADIOLARIAN SHALE SILICEOUS SHALE RADIOLARIAN CHERT AND BITUMINOUS CHERT.			
DEVONIAN	UPPER AND MIDDLE	LOWER MIDDLE	LIGHT GRAY TO BLACK BITUMINOUS SPORE-BEARING CHERT AND BLACK PAPERY BITUMINOUS SHALE.			
		LOWER	WHITE TO GREEN MASSIVE SPICULITIC CHERT AND GREEN LAMINATED SILICEOUS SHALE.			
SILURIAN	NIAGARAN	MISSOURI MOUNTAIN SHALE		100	SHALE, HARD, GREEN SILICEOUS, SANDY IN PART. THIN BEDS OF FINELY LAMINATED CHERT AND QUARTZOSE SANDSTONE AND LOCAL LENSES OF SANDY CHERT CONGLOMERATE.	
	ALEXANDRIAN	BLAYLOCK SANDSTONE		500	SANDSTONE, GRAY TO GREEN, THIN BEDDED, FINE GRAINED, WITH INTERBEDDED SHALEY MICACEOUS SILTSTONE AND DARK FISSILE SHALES. VEINS OF QUARTZ AND SMOKY QUARTZ ARE ABUNDANT. FORMATION PRESENT ONLY IN PART OF BROKEN BOW-BENTON UPLIFT.	
MIDDLE AND UPPER ORDOVICIAN	CINCINNATIAN	POLK CREEK SHALE		36	SHALE, SOFT, BROWN, PLATY IN MOST OF FORMATION; HARD, BLACK, BITUMINOUS, AND SILICEOUS NEAR BASE. ABUNDANT GRAPTOLITES. THIN STREAKS OF QUARTZITIC SANDSTONE AND OOLITIC LIMESTONE.	
	TRENTONIAN	BIG FORK CHERT	UPPER	250	BLACK, NONCALCAREOUS, BITUMINOUS CHERT AND BLACK BITUMINOUS PAPERY SHALE.	
			LOWER		GRAY TO BROWN CALCAREOUS CHERT, SILICEOUS LIMESTONE, CLASTIC LIMESTONE AND CHERTY SHALE.	
LOWER ORDOVICIAN	BLACKRIVERIAN	WOMBLE SHALE		1100	SHALE, BLACK TO GREEN, WITH THIN INTERBEDS OF QUARTZOSE SANDSTONE AND LIMESTONE. SOME SILICEOUS BITUMINOUS SHALE NEAR CONTACT WITH BIGFORK CHERT.	
	CHAZYAN	BLAKELY SANDSTONE		150	SHALE, BLACK TO GREEN, INTERBEDDED WITH FINE TO MEDIUM GRAINED QUARTZOSE SANDSTONE. SOME VEINS OF SMOKY QUARTZ.	
	CANADIAN	BAZARN SHALE		1000	SHALE, BLACK TO GREEN, BANDED, CLAYEY, FISSILE, WITH THIN LAYERS OF GREEN SANDSTONE AND BLuish-BLACK LIMESTONE. VEINS OF QUARTZ AND CALCITE.	
		CRYSTAL MOUNTAIN SANDSTONE		280	SANDSTONE, MASSIVE, LIGHT GRAY, CALCAREOUS TO QUARTZITIC. MANY QUARTZ VEINS AND CRYSTALS.	
		COLLIER SHALE		330	SHALE, BLACK, GRAPHITIC, AND DARK-COLORED SILICEOUS LIMESTONE. SOME DENSE BLACK CHERT.	

Figure 1-3. Stratigraphic column for study Area I. After Vogelpohl (1977).

Figure 1-2
Stratigraphic Column for Study Area

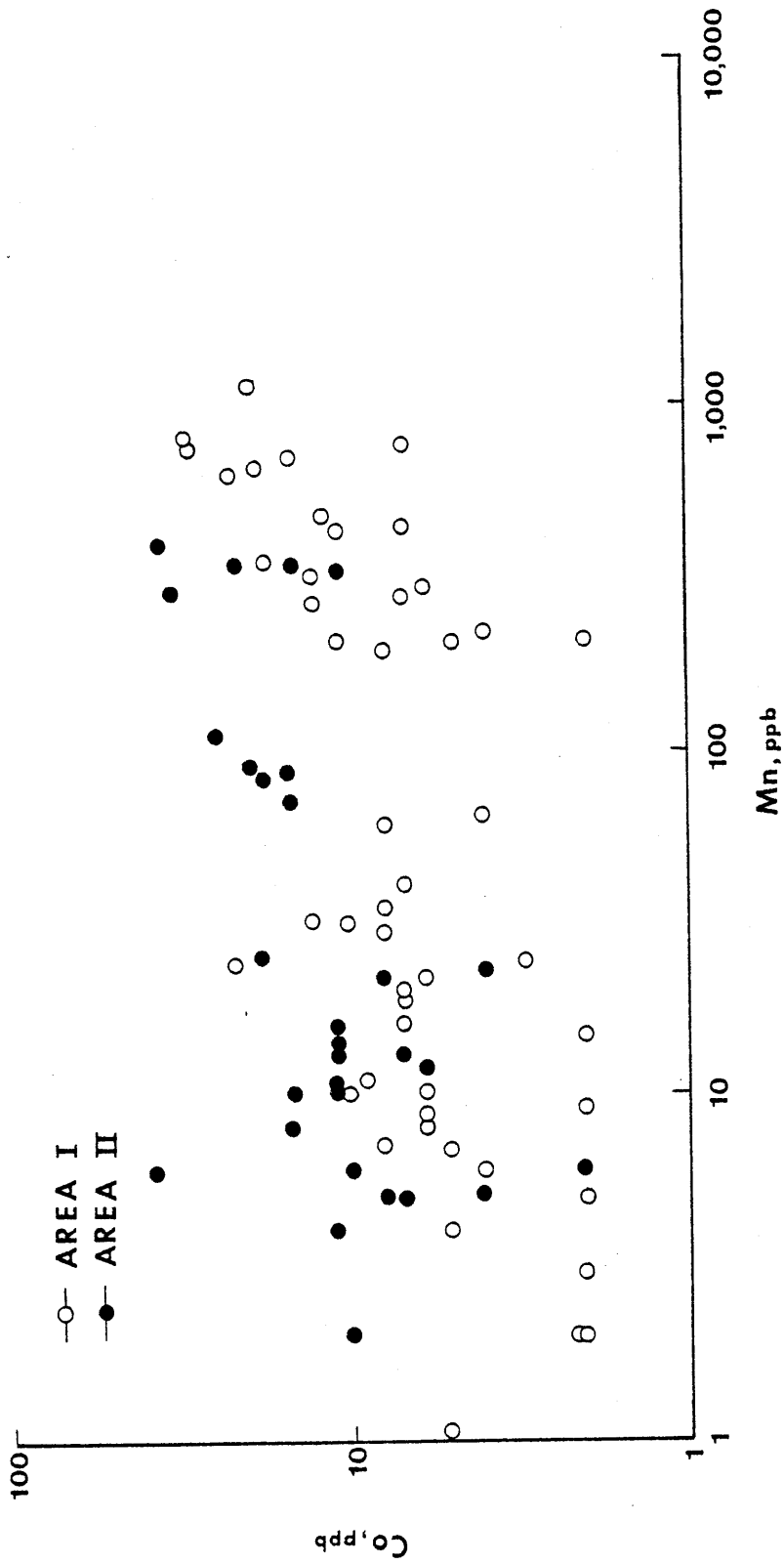


Figure I-1
Correlation Plot for Co vs Mn, Areas I and II

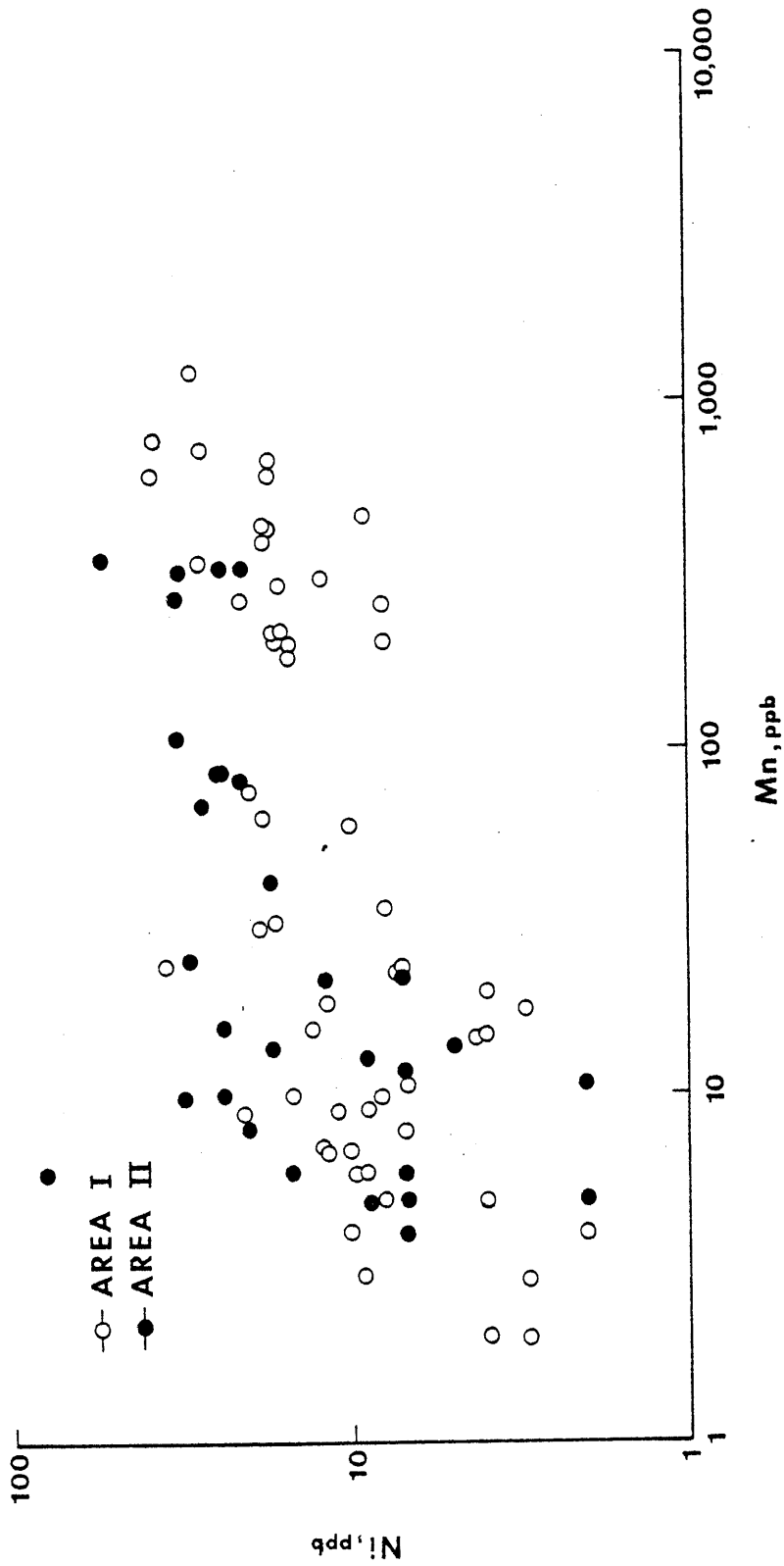


Figure I-2
Correlation Plot for Ni vs Mn, Areas I and II

59

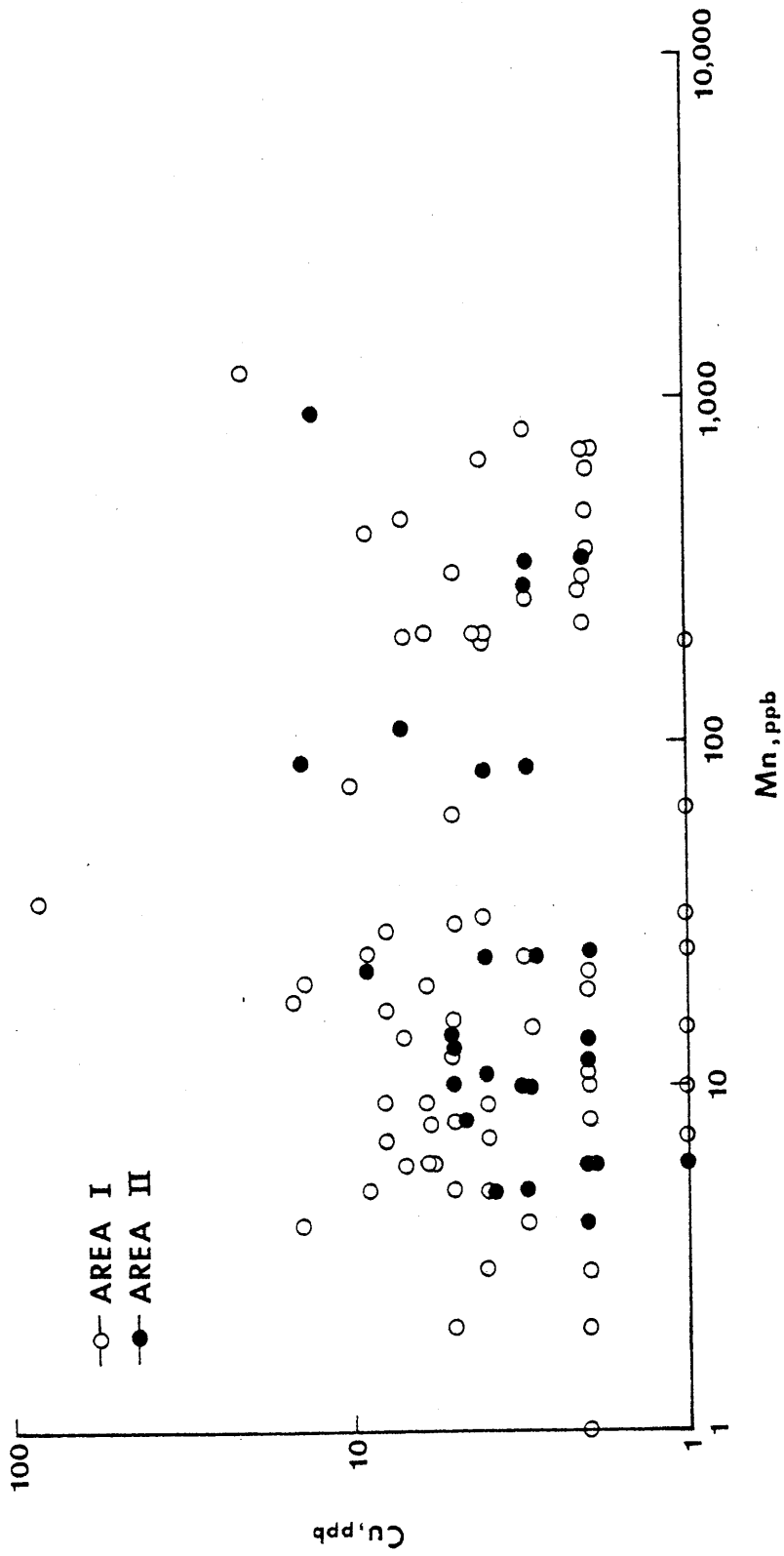
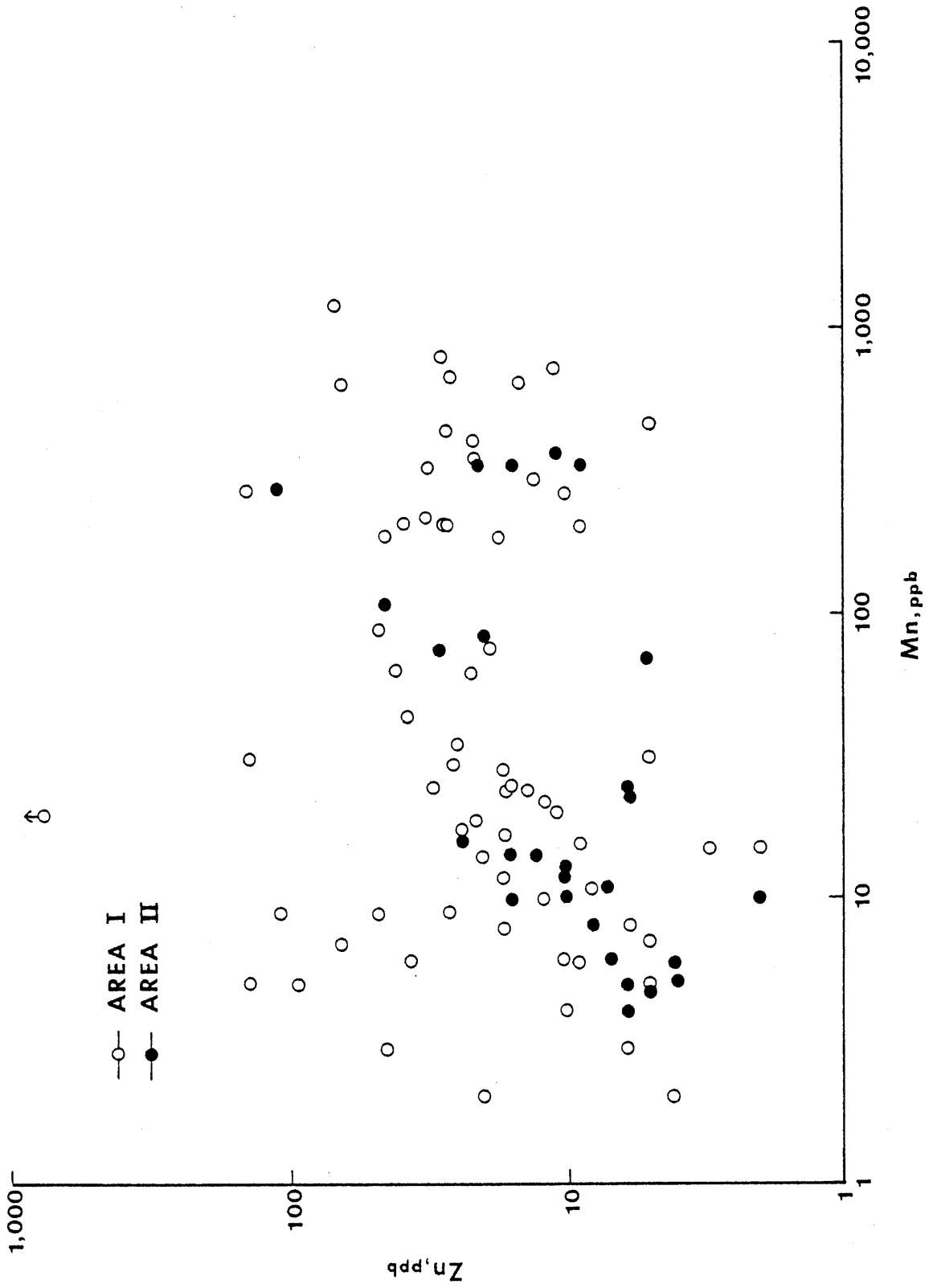


Figure I-3
Correlation Plot for Cu vs Mn, Areas I and II



♂

○ AREA I
 ● AREA II

Zn, ppb

Min, ppb

Figure I-4
Correlation Plot for Zn vs Mn, Areas I and II

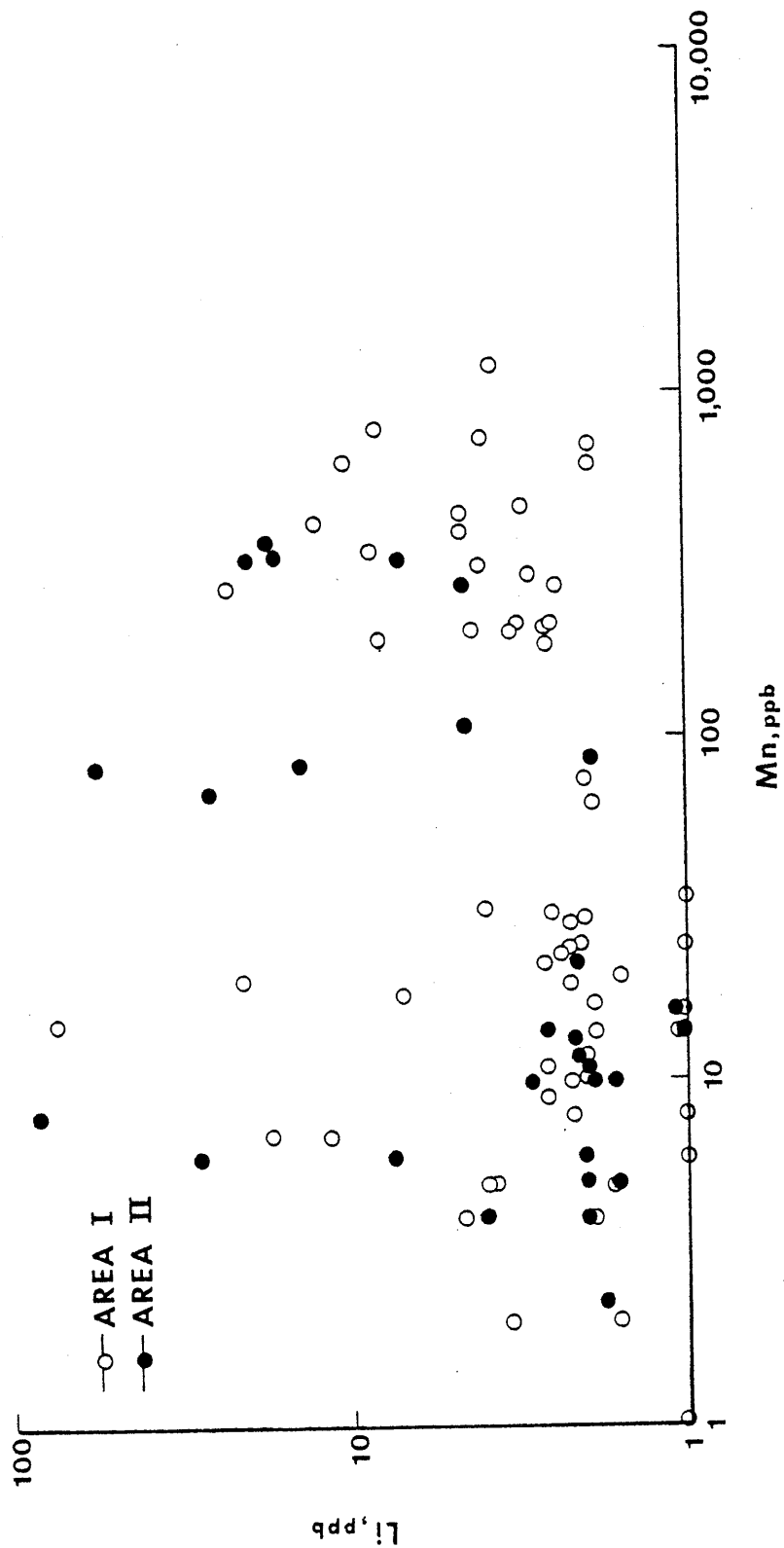


Figure I-5
Correlation Plot for Li vs Mn, Areas I and II

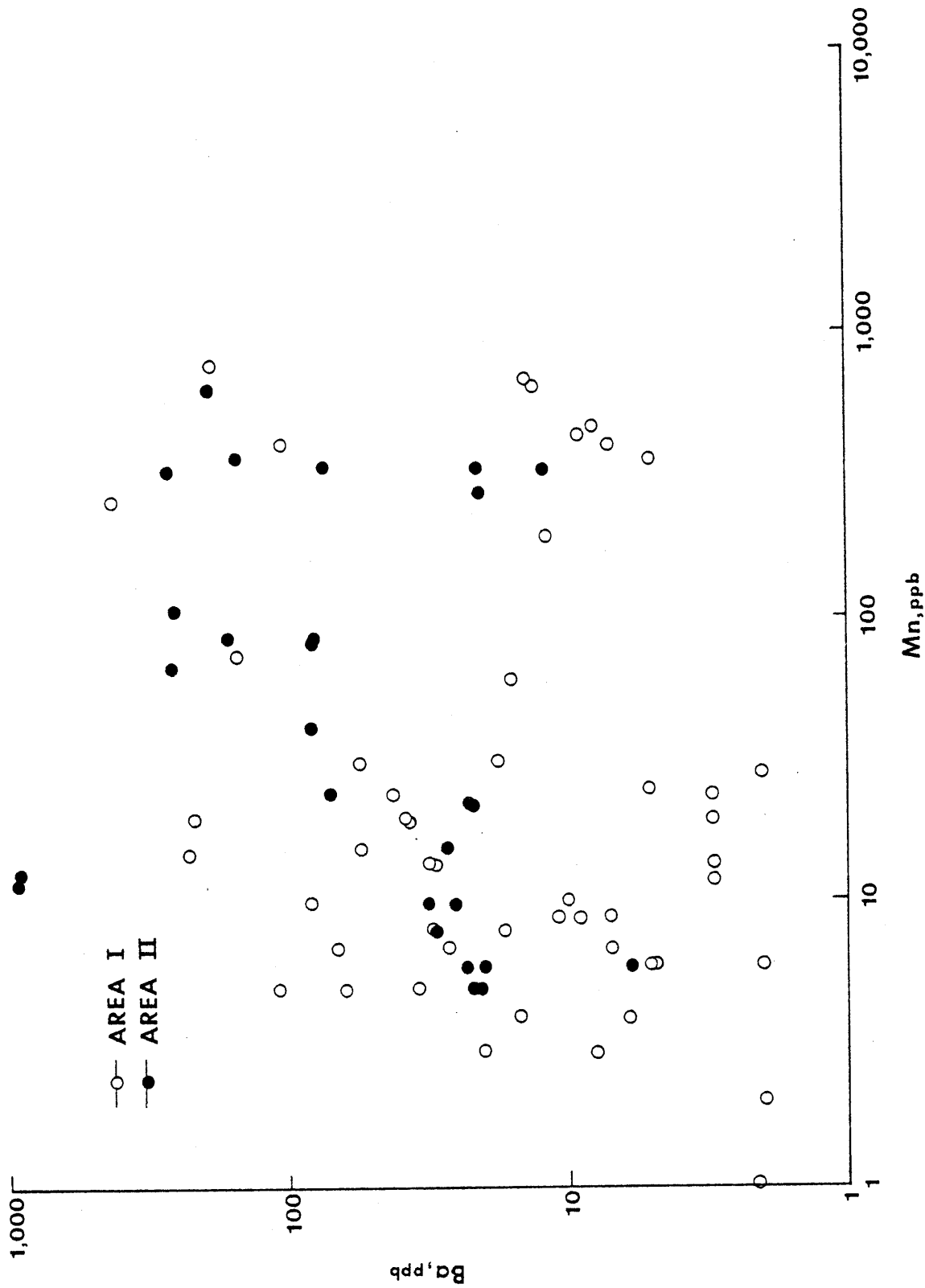
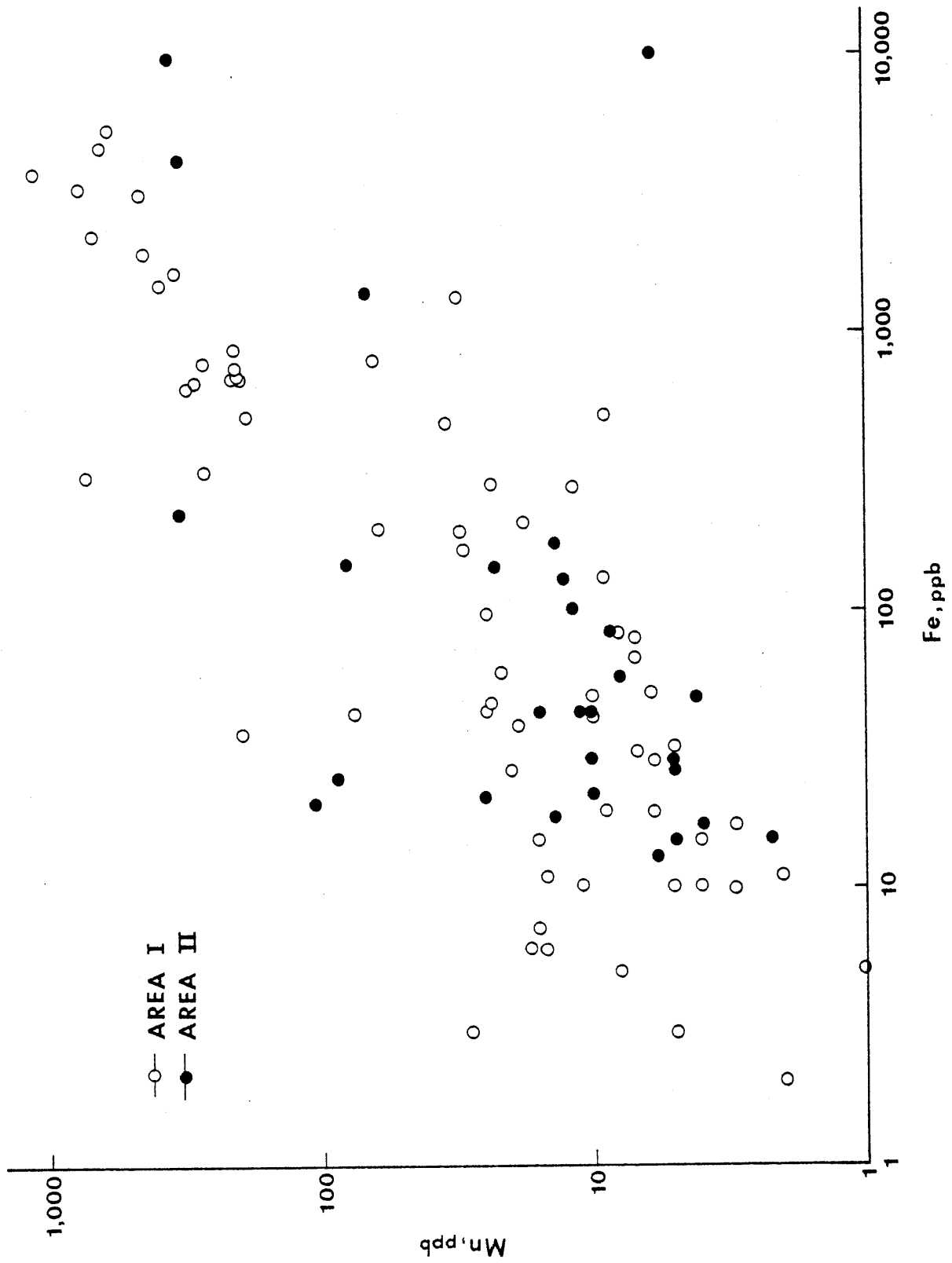
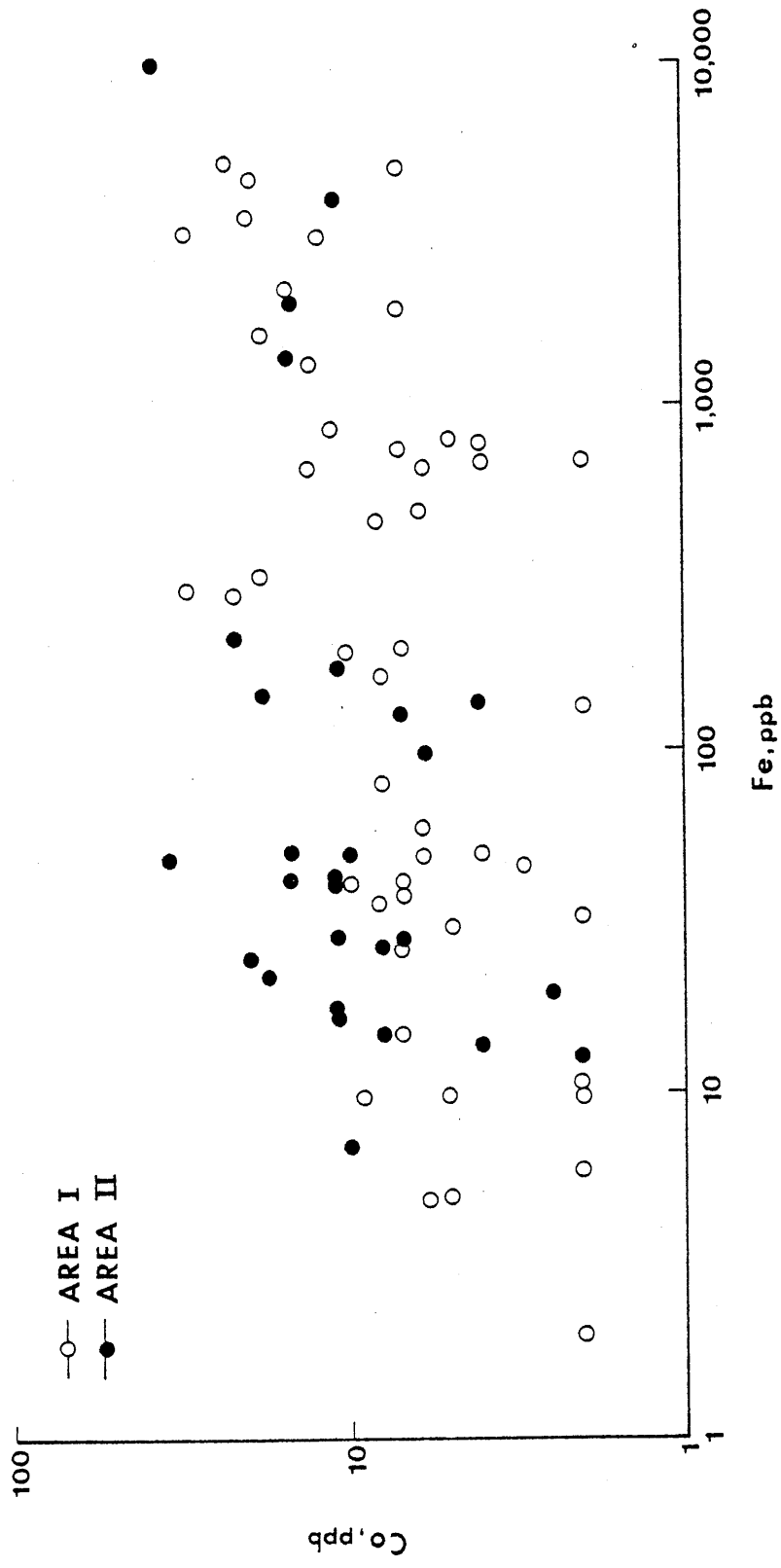


Figure I-6
Correlation Plot for Ba vs Mn, Areas I and II



99

Figure I-7
Correlation Plot for Mn vs Fe, Areas I and II



100

Figure I-8

Correlation Plot for Co vs Fe, Areas I and II

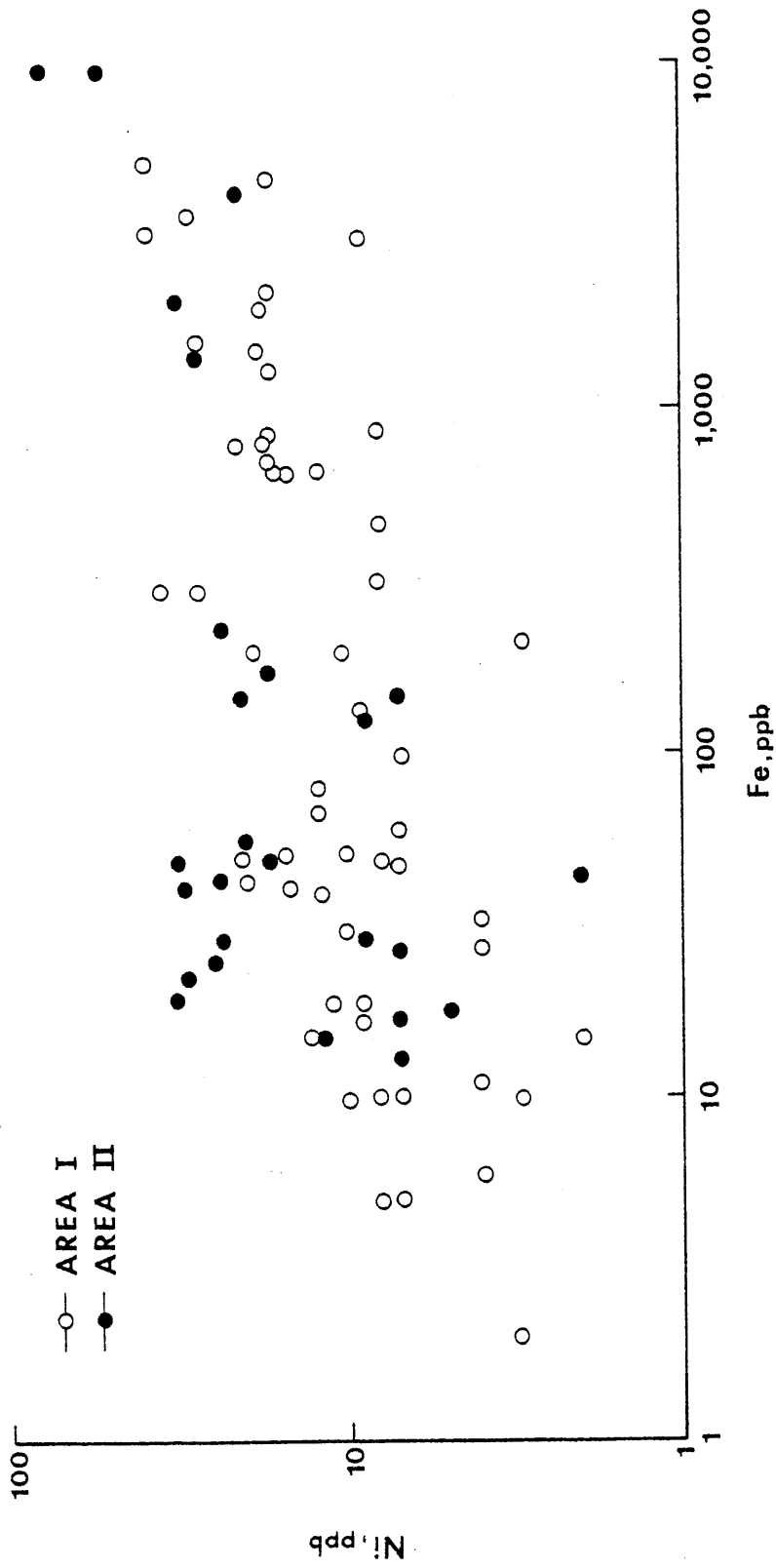
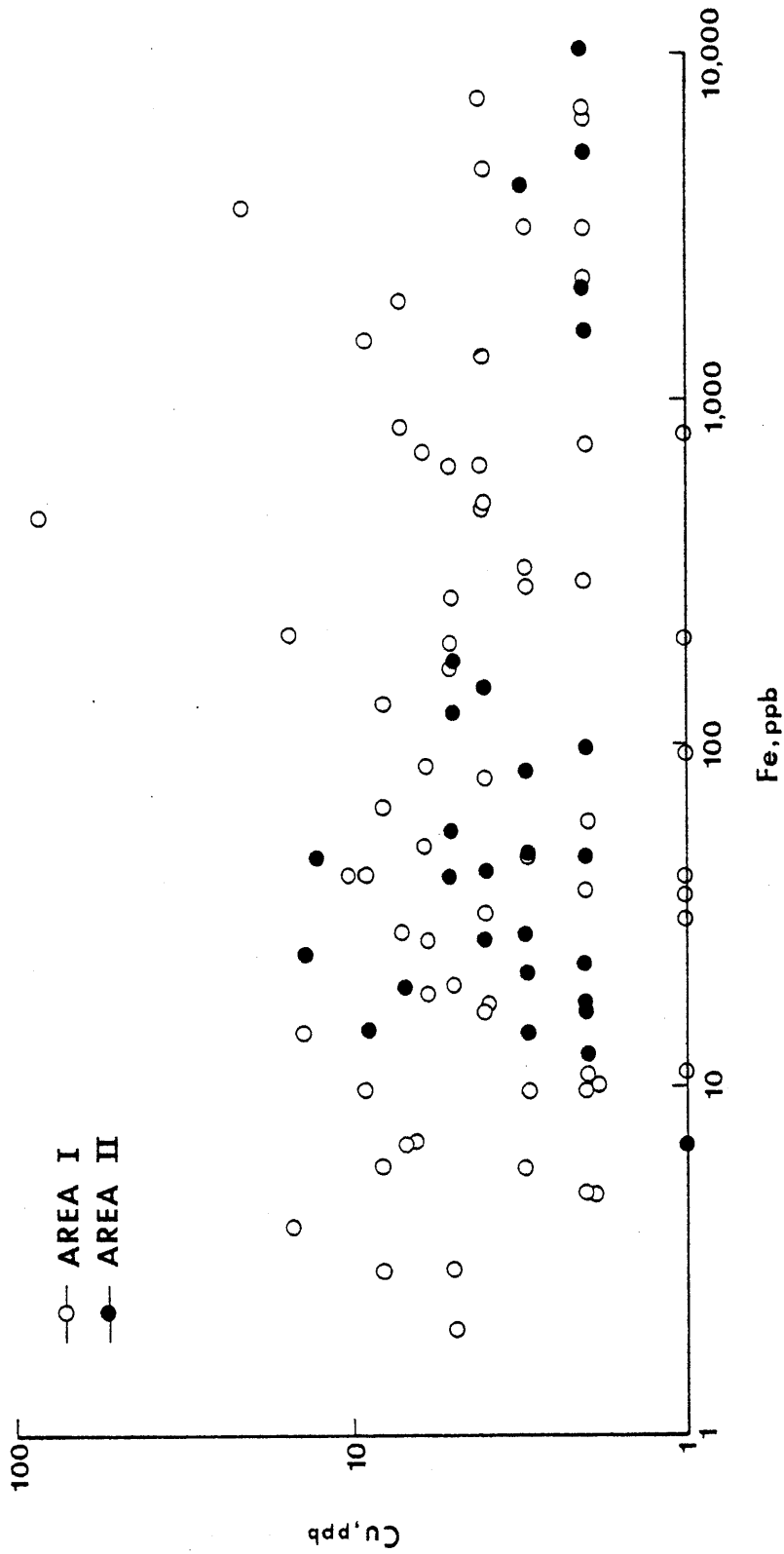


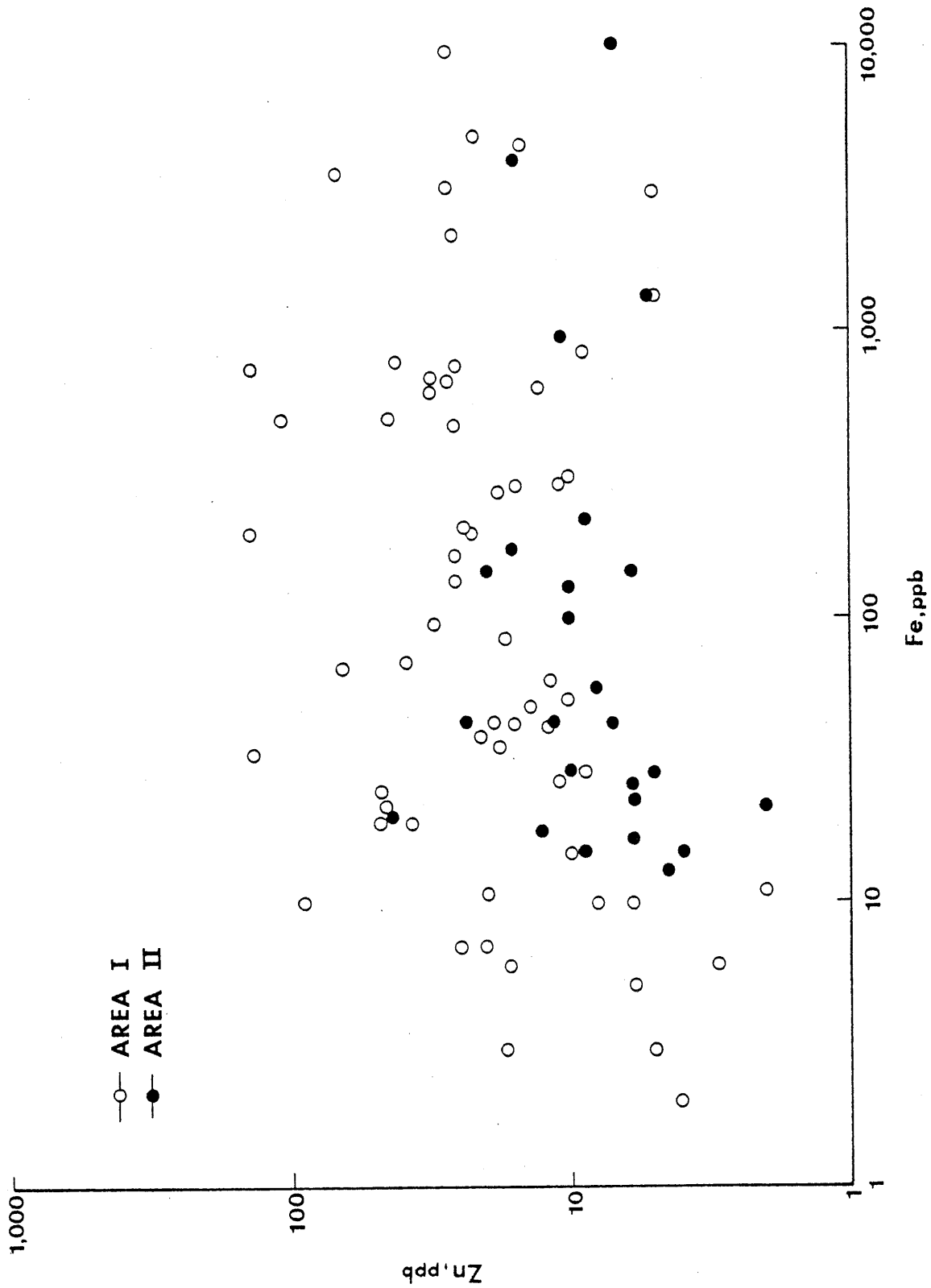
Figure I-9

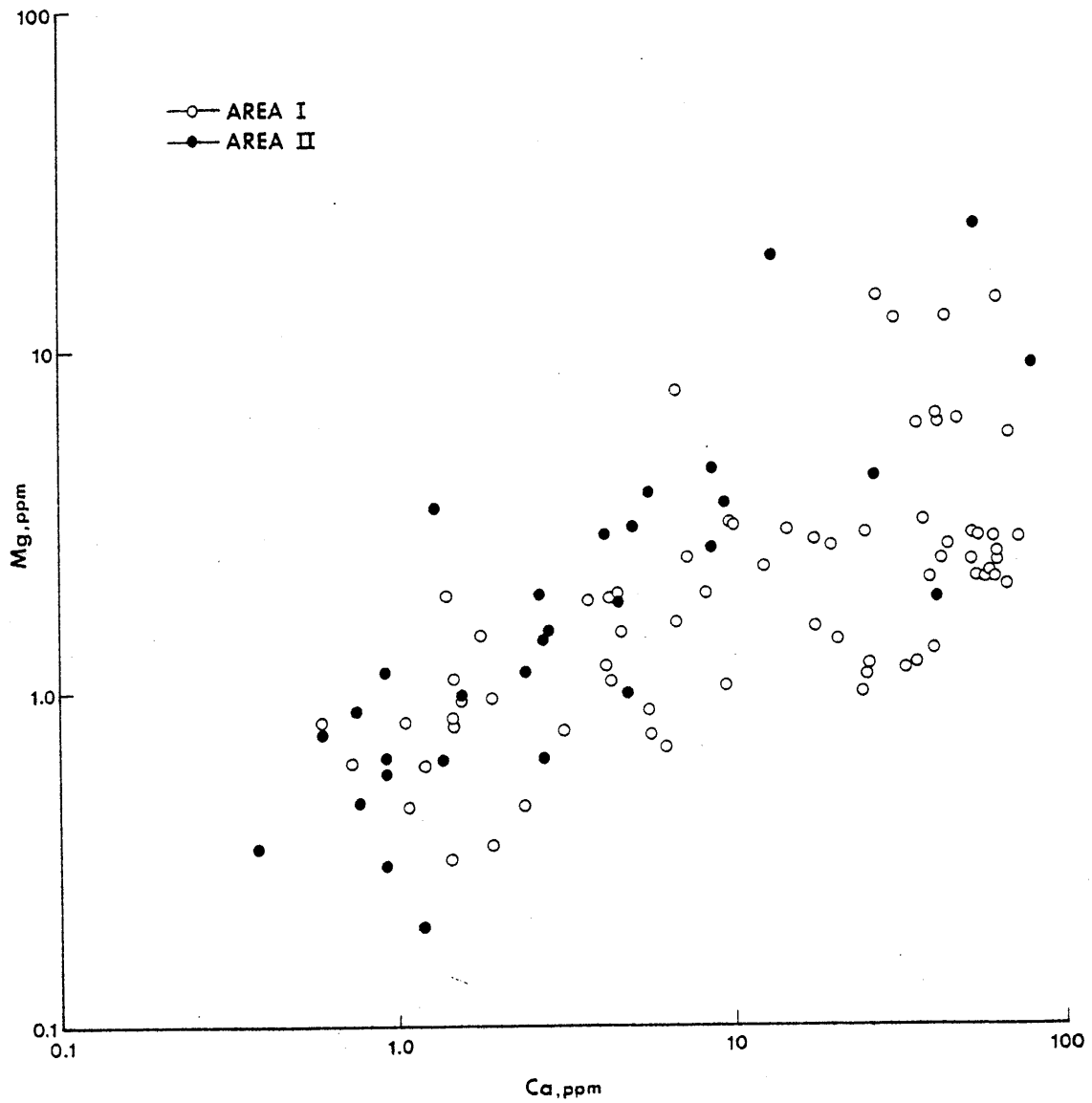
Correlation Plot for Ni vs Fe, Areas I and II



102

Figure I-10
Correlation Plot for Cu vs Fe, Areas I and II





104

Figure I-11
Correlation Plot for Zn vs Fe, Areas I and II

Figure I-12
Correlation Plot for Mg vs Ca, Areas I and II

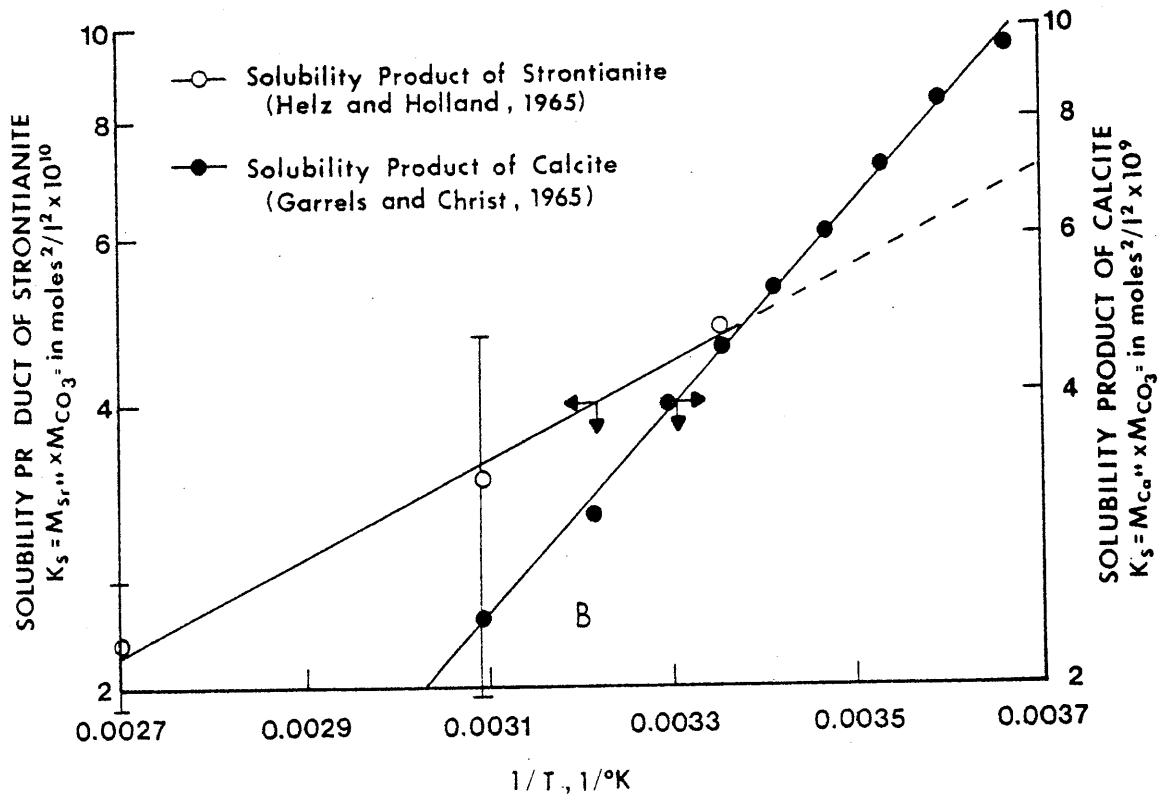
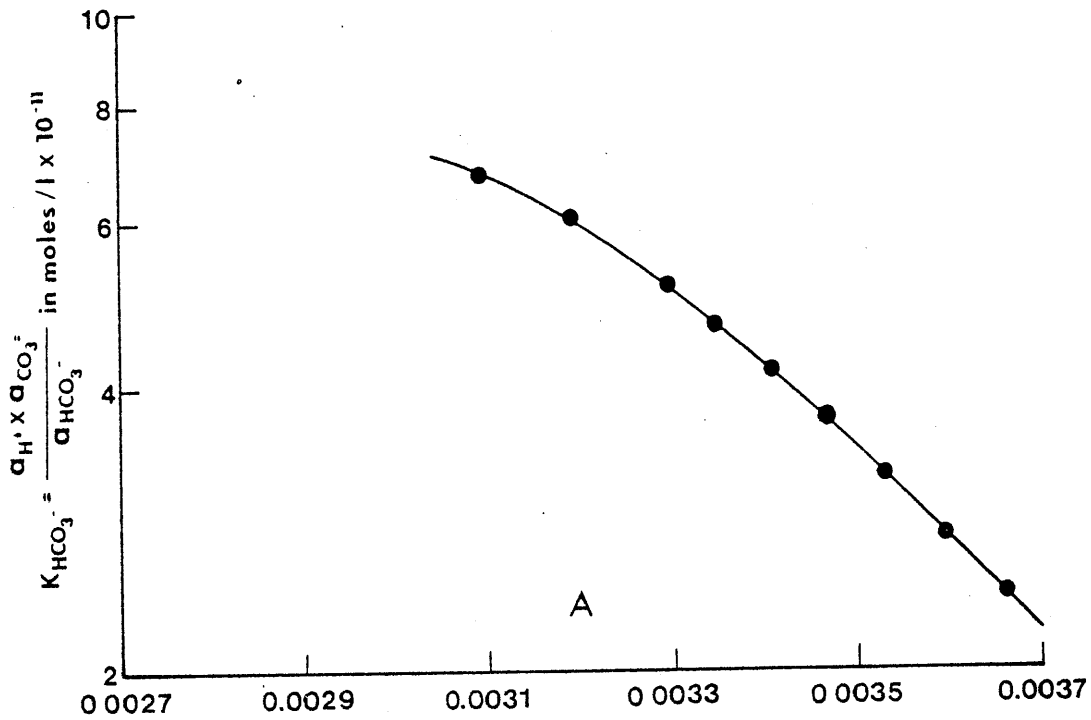


Figure I-13A

Equilibrium Constant for HCO_3^- vs T

Figure I-13B

Solubility Product of Calcite (CaCO_3) and Strontianite (SrCO_3) vs T

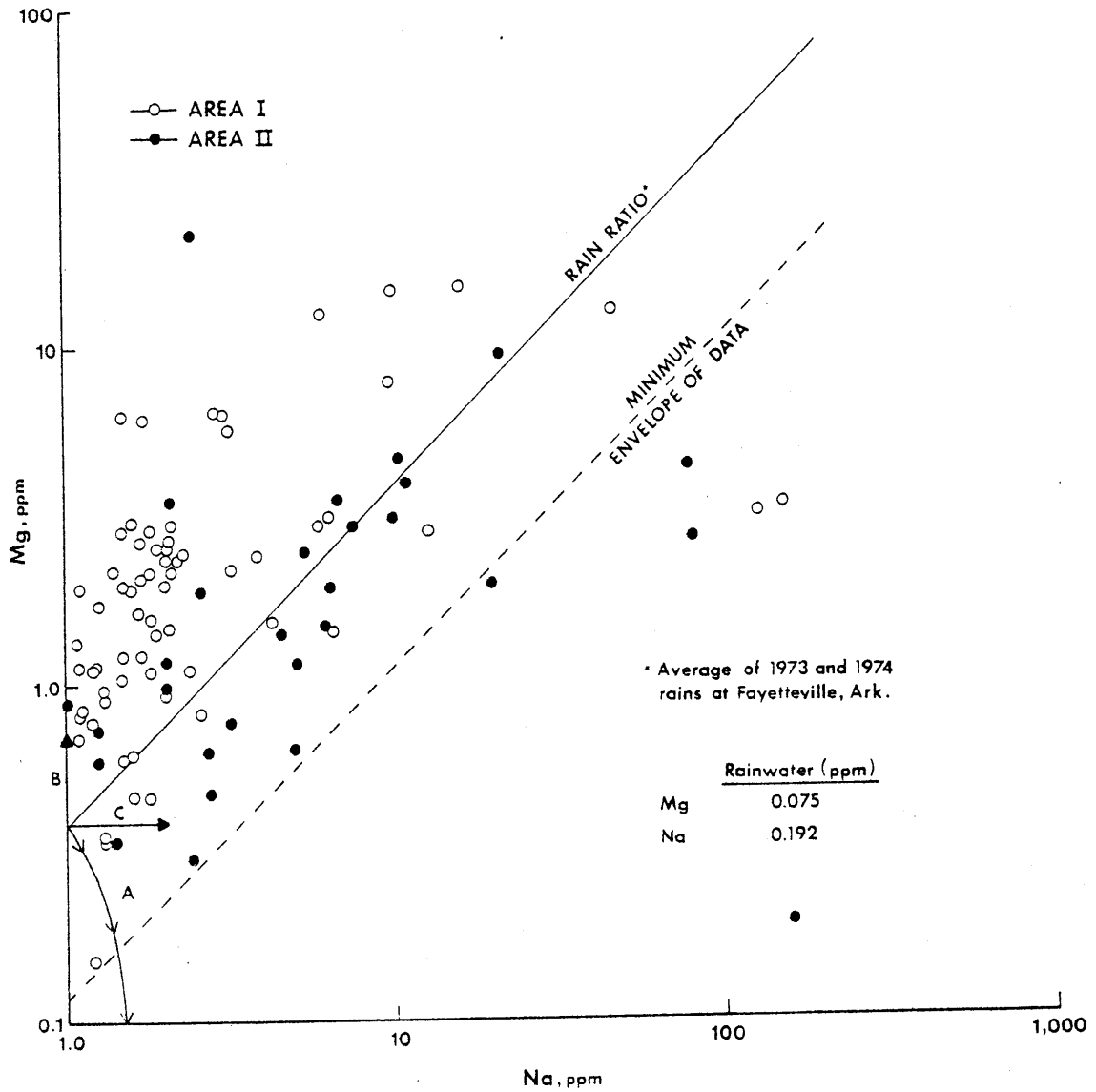


Figure I-14

Concentration of Mg vs Concentration of Na for Areas I and II

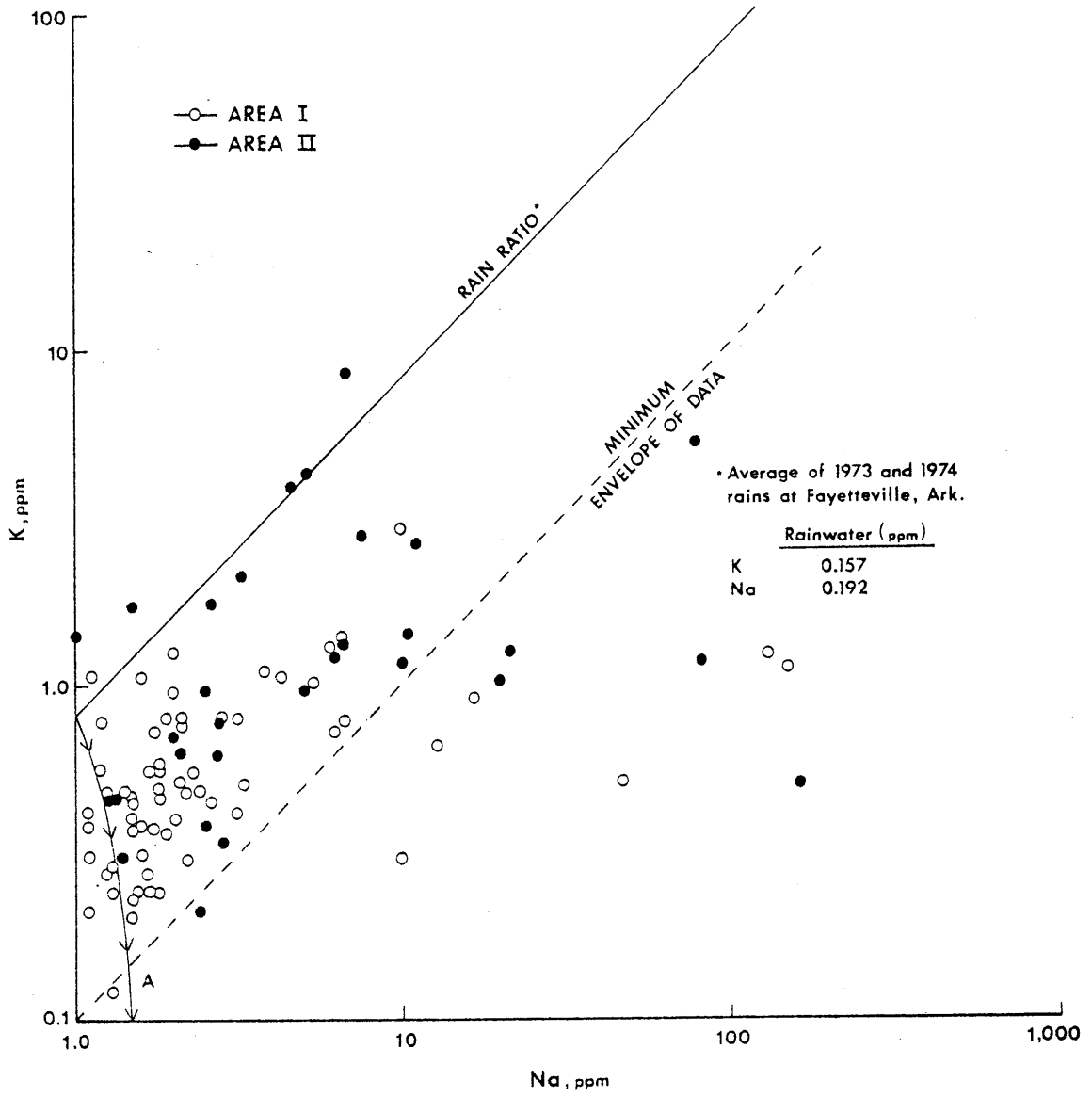


Figure I-15

Concentration of K vs Concentration of Na for Areas I and II

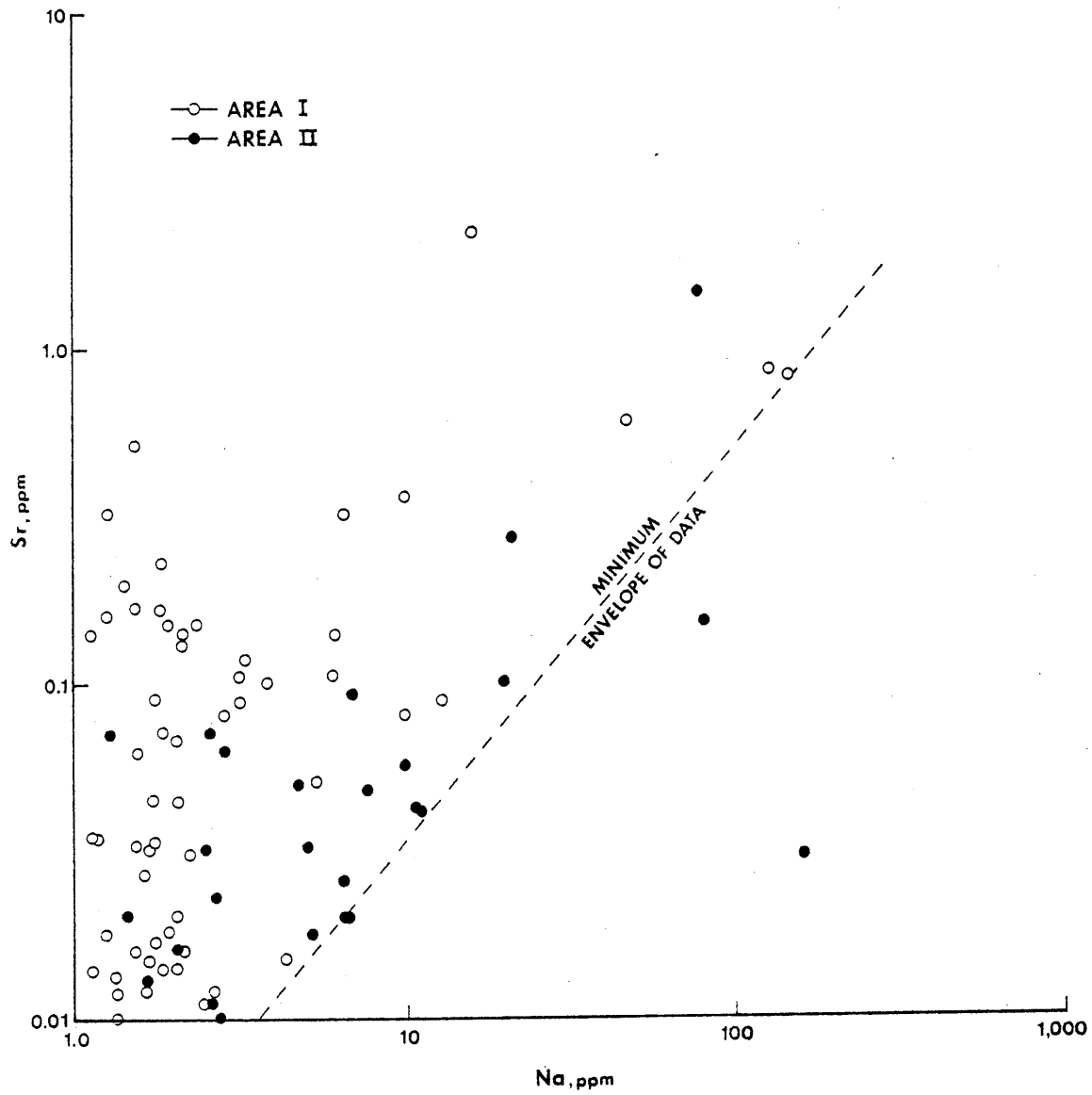


Figure I-16

Concentration of Sr vs Concentration of Na for Areas I and II

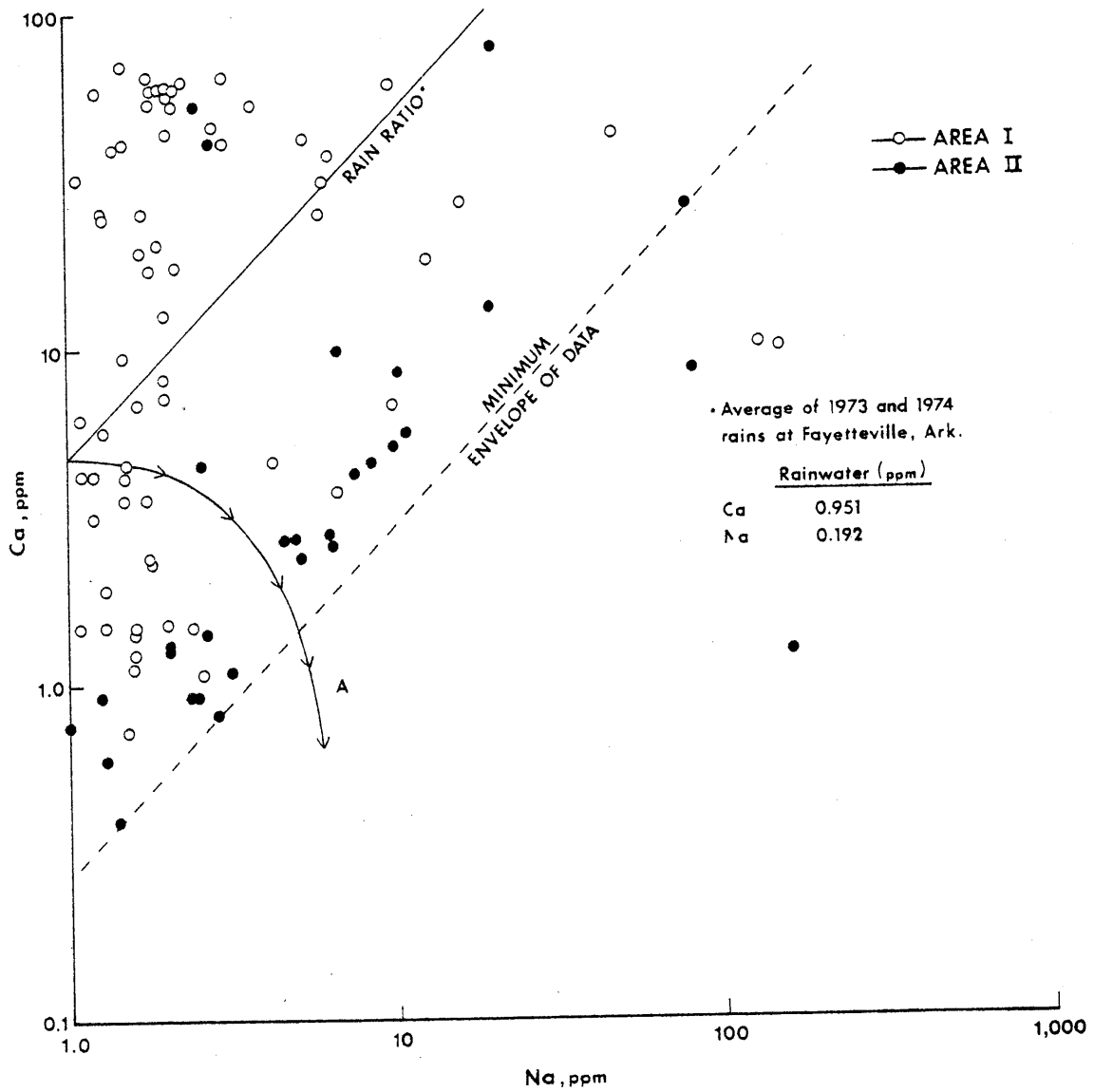


Figure I-17

Concentration of Ca vs Concentration of Na for Areas I and II

110

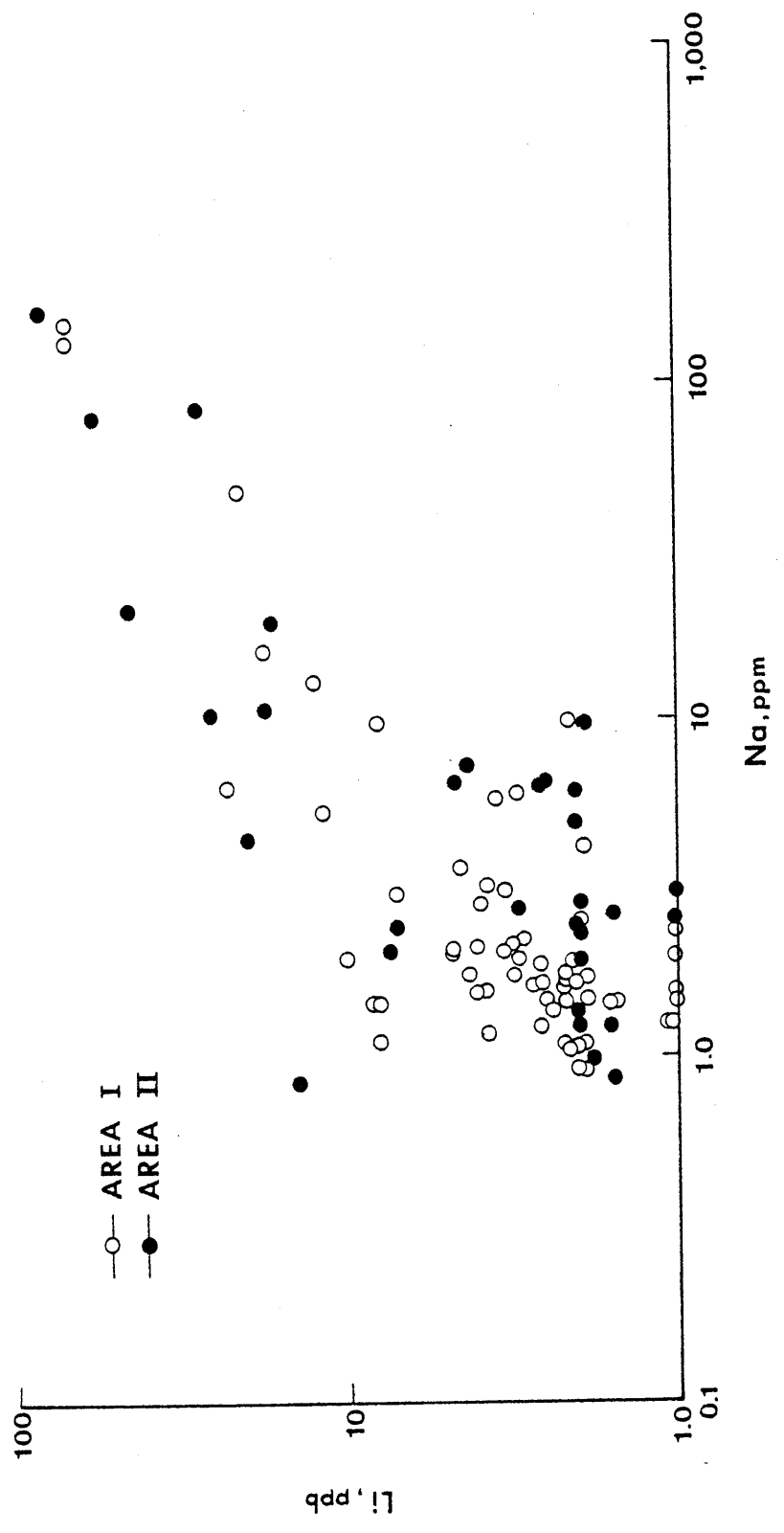


Figure I-18
Concentration of Li vs Concentration of Na for Areas I and II

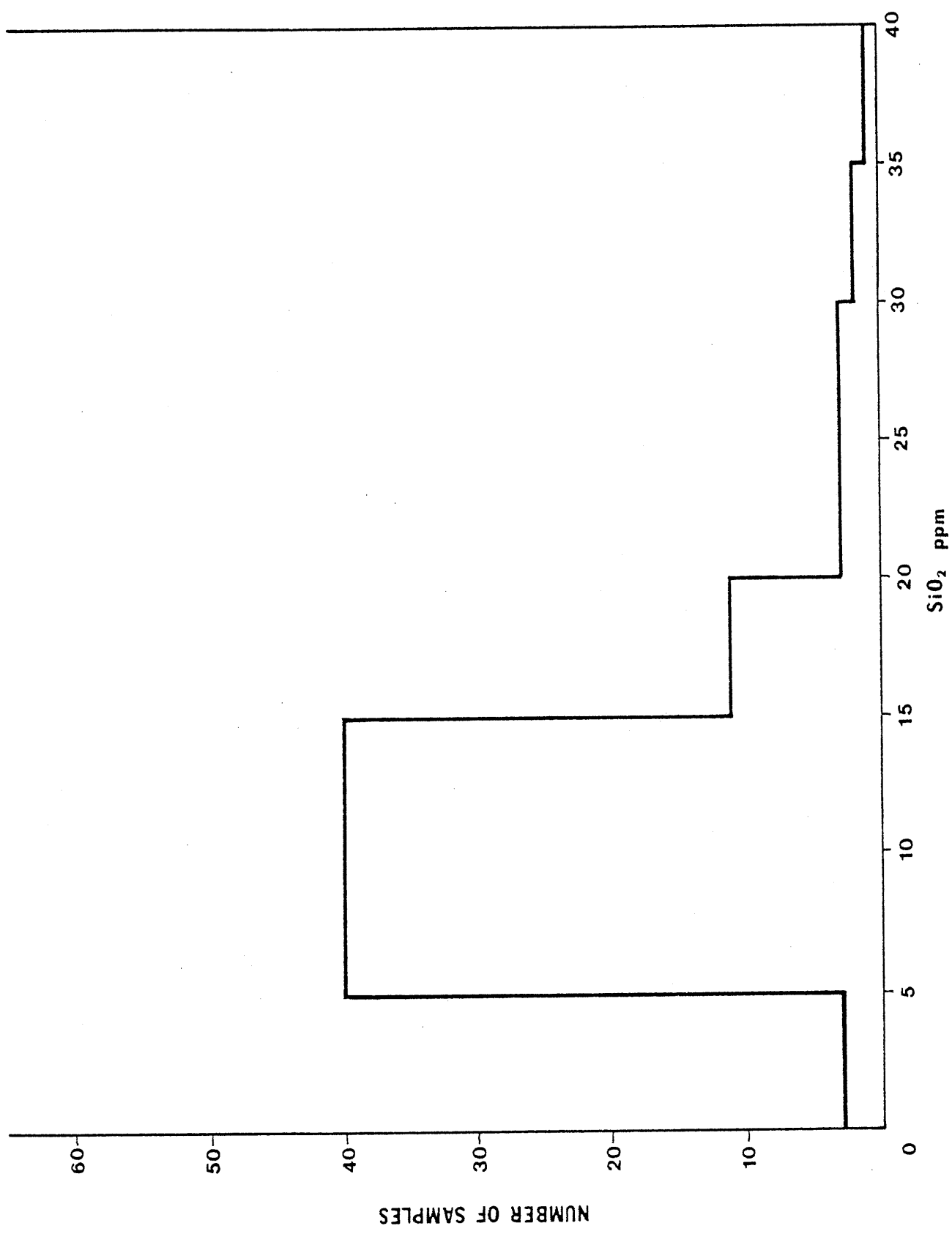


Figure II-1
Histogram of Silica Concentration in Spring Waters
of Areas I and II Combined

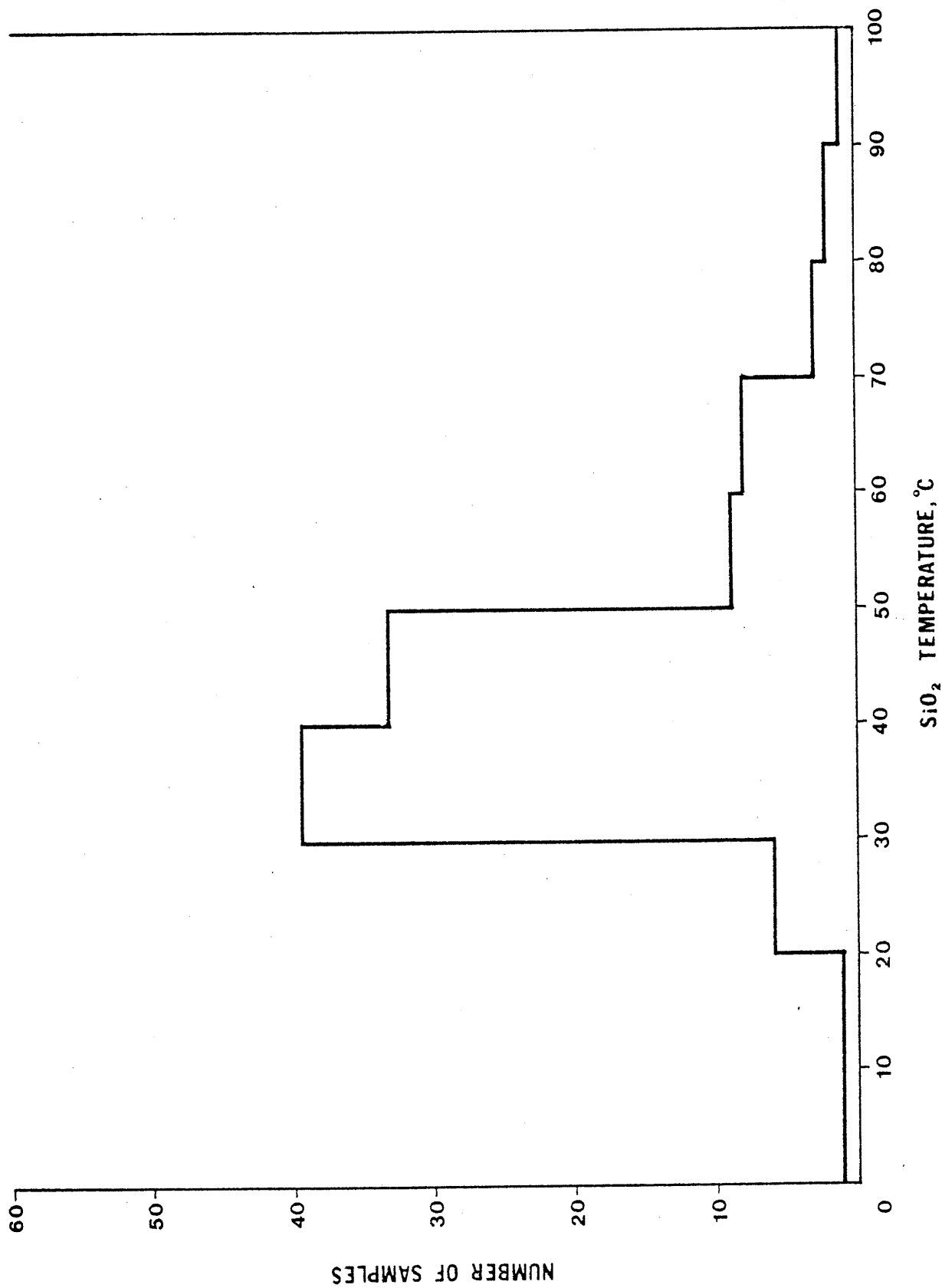
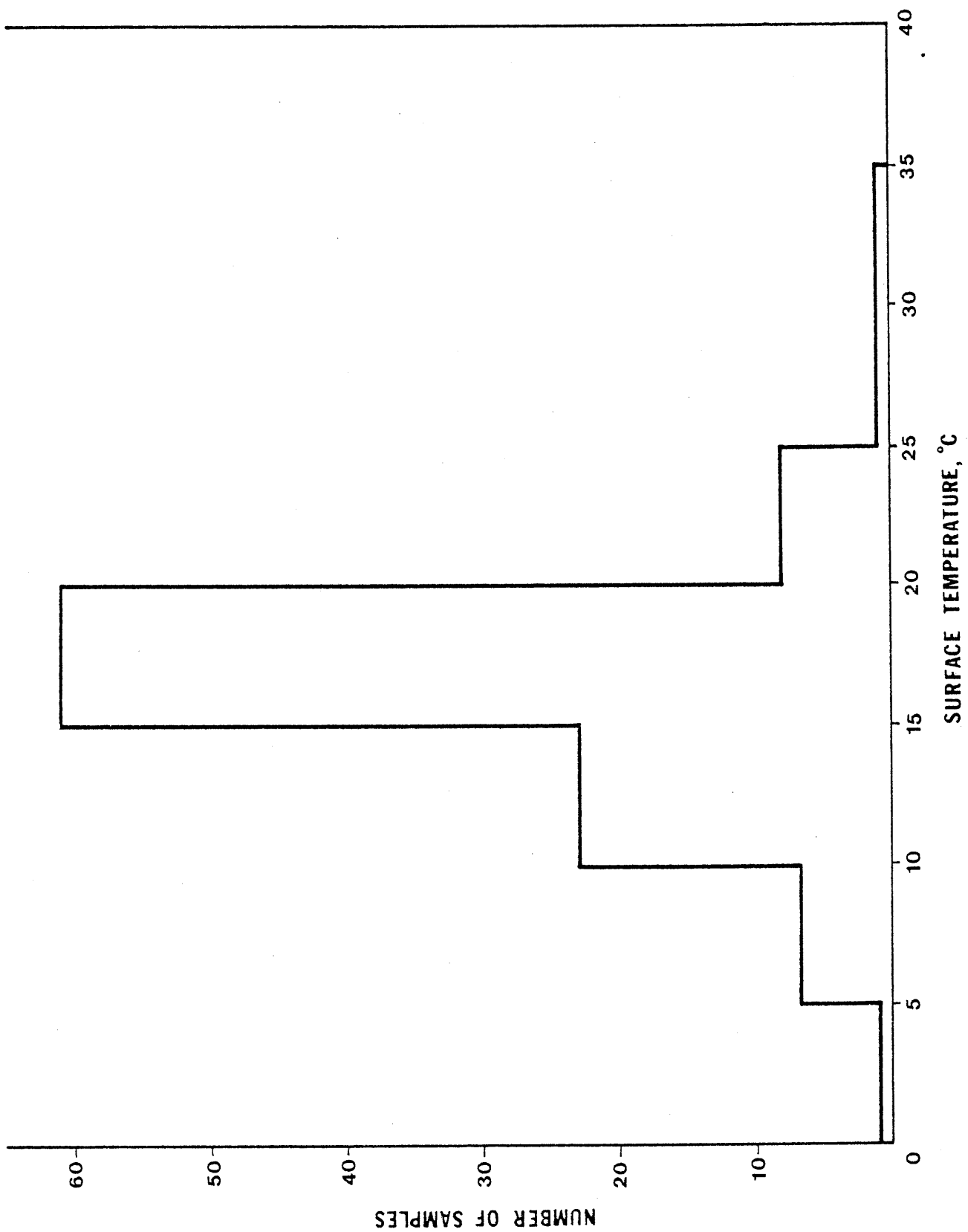


Figure II-2

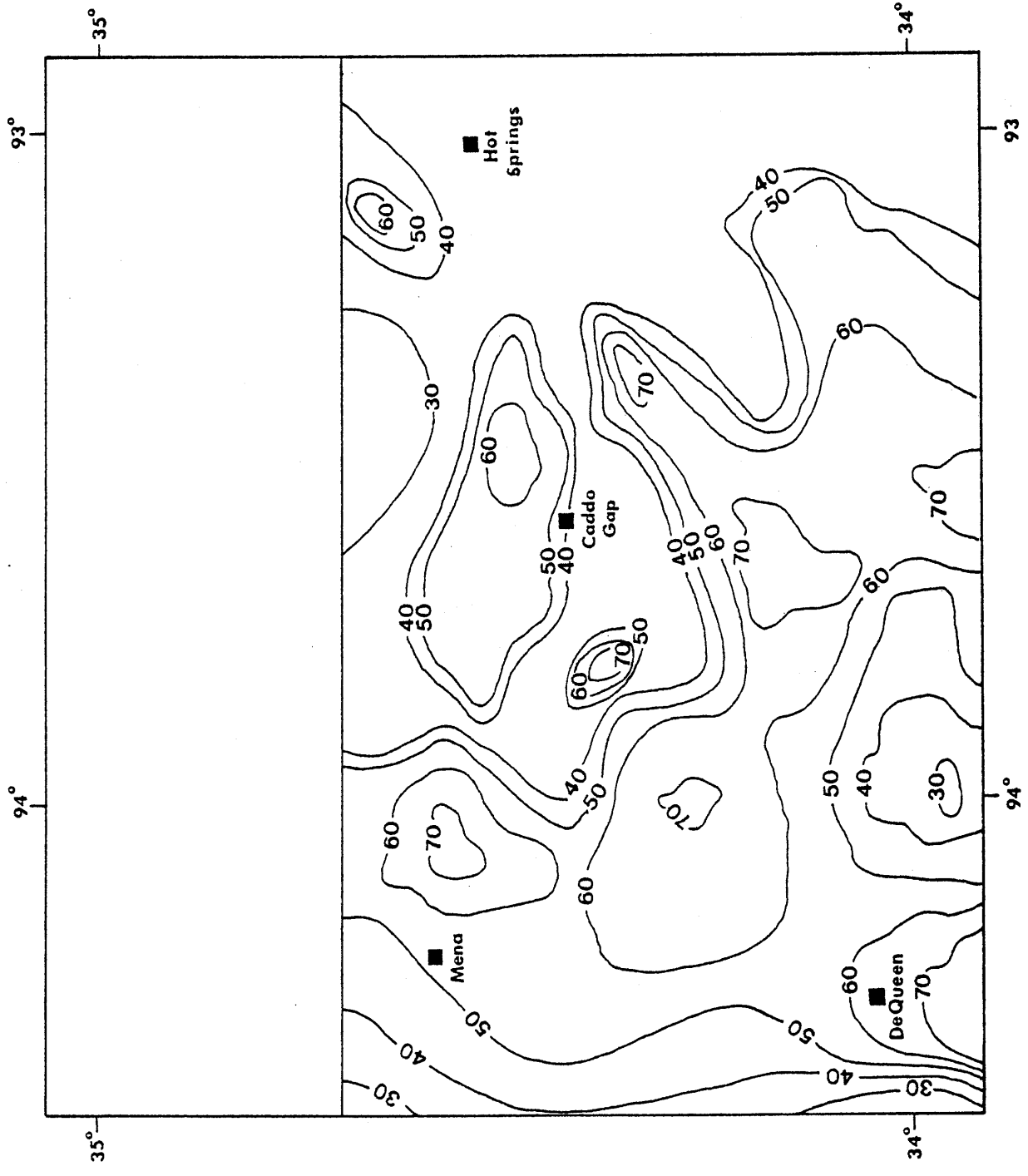
Histogram of Silica Geotemperatures of Spring Waters
for Areas I and II Combined



211

Figure II-3

Histogram of Surface Temperatures of Spring
Waters for Areas I and II Combined



CONTOUR
 INTERVAL
 10 w/m²

SCALE

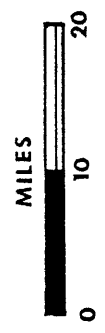
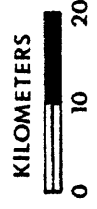
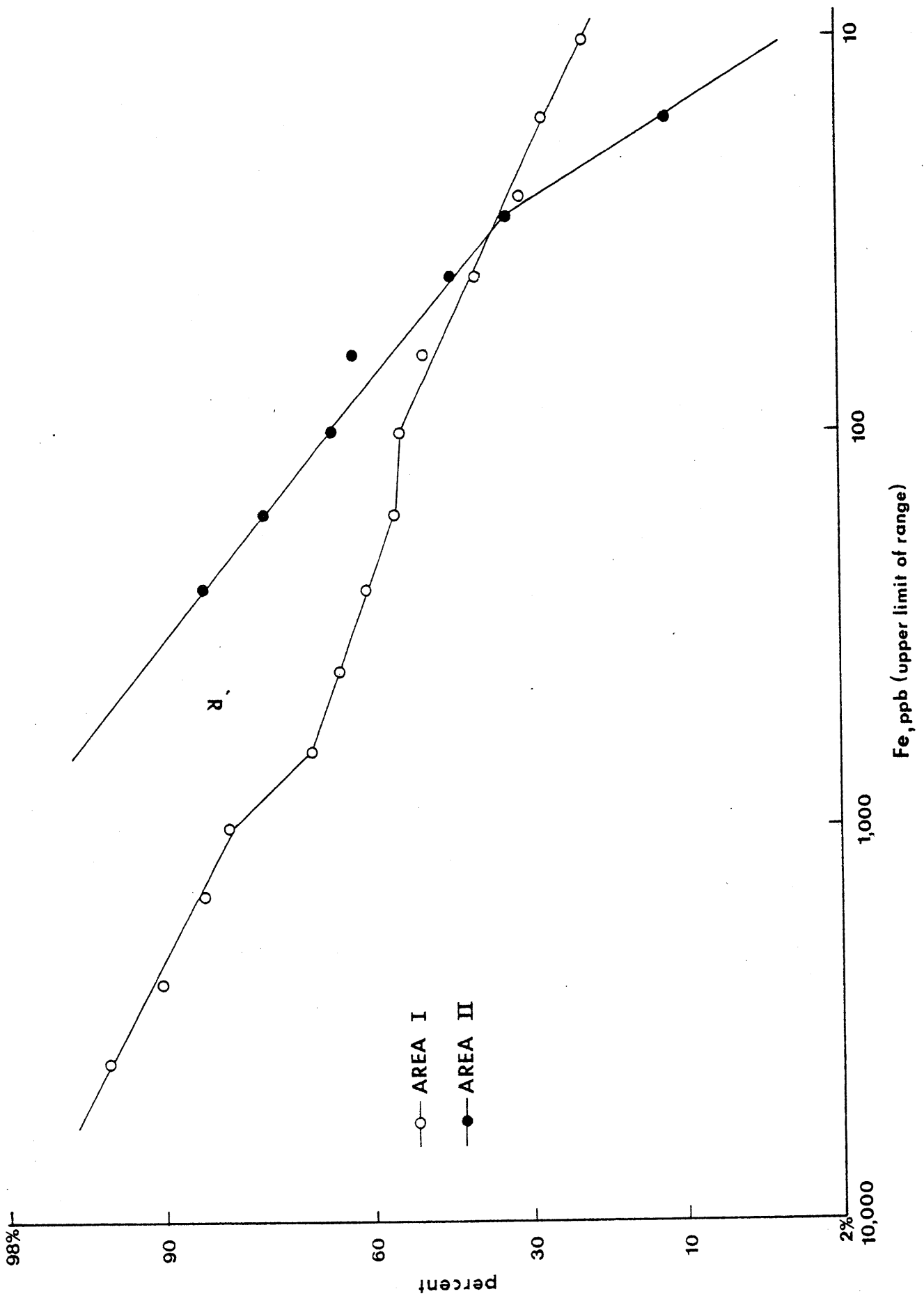


Figure II-4
Contour Map of Heat Flow



AREA I
AREA II

R

Percent

Fe, ppb (upper limit of range)

115

Figure III-1
Cumulative Frequency Curve for Fe, Areas I and II

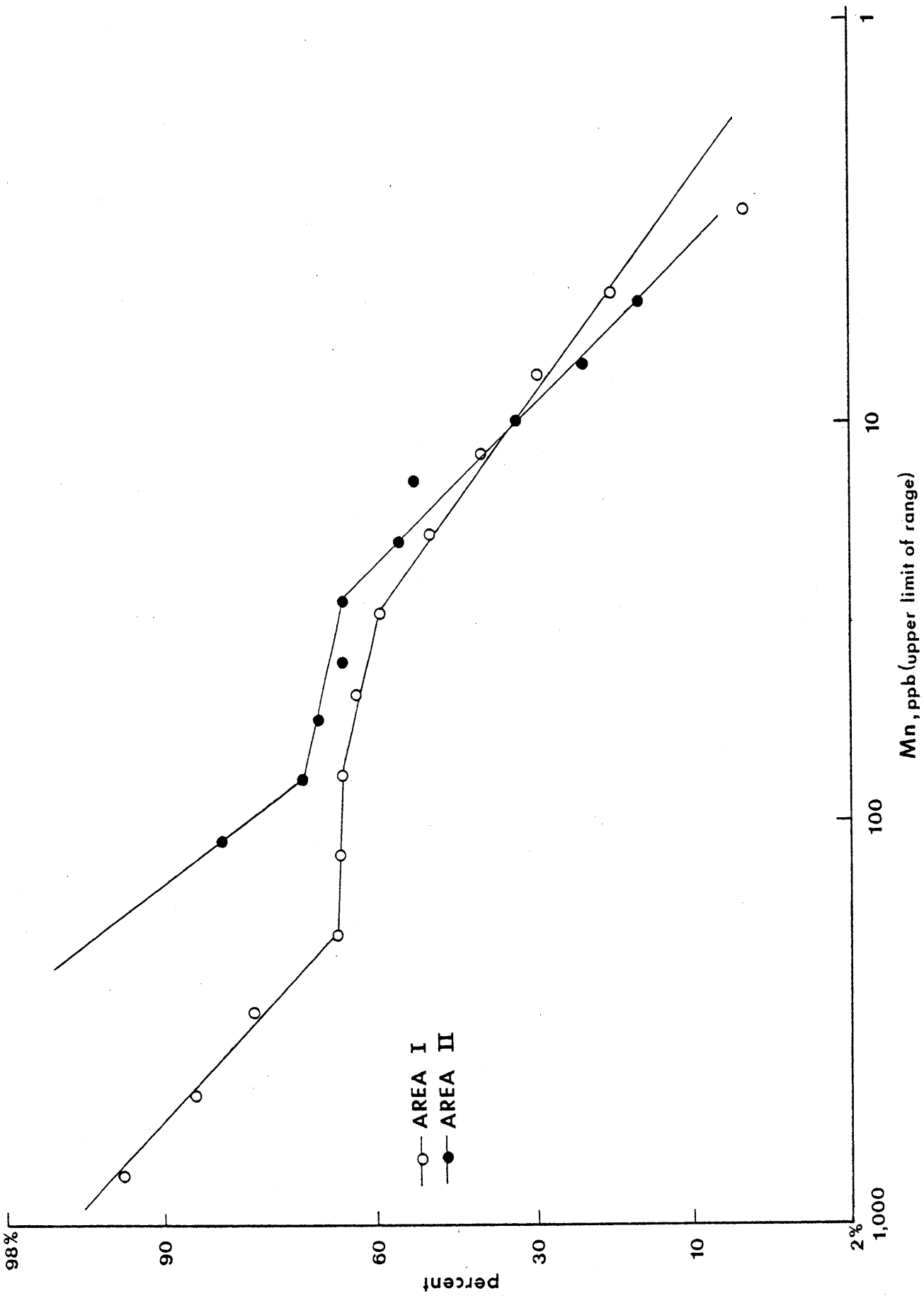
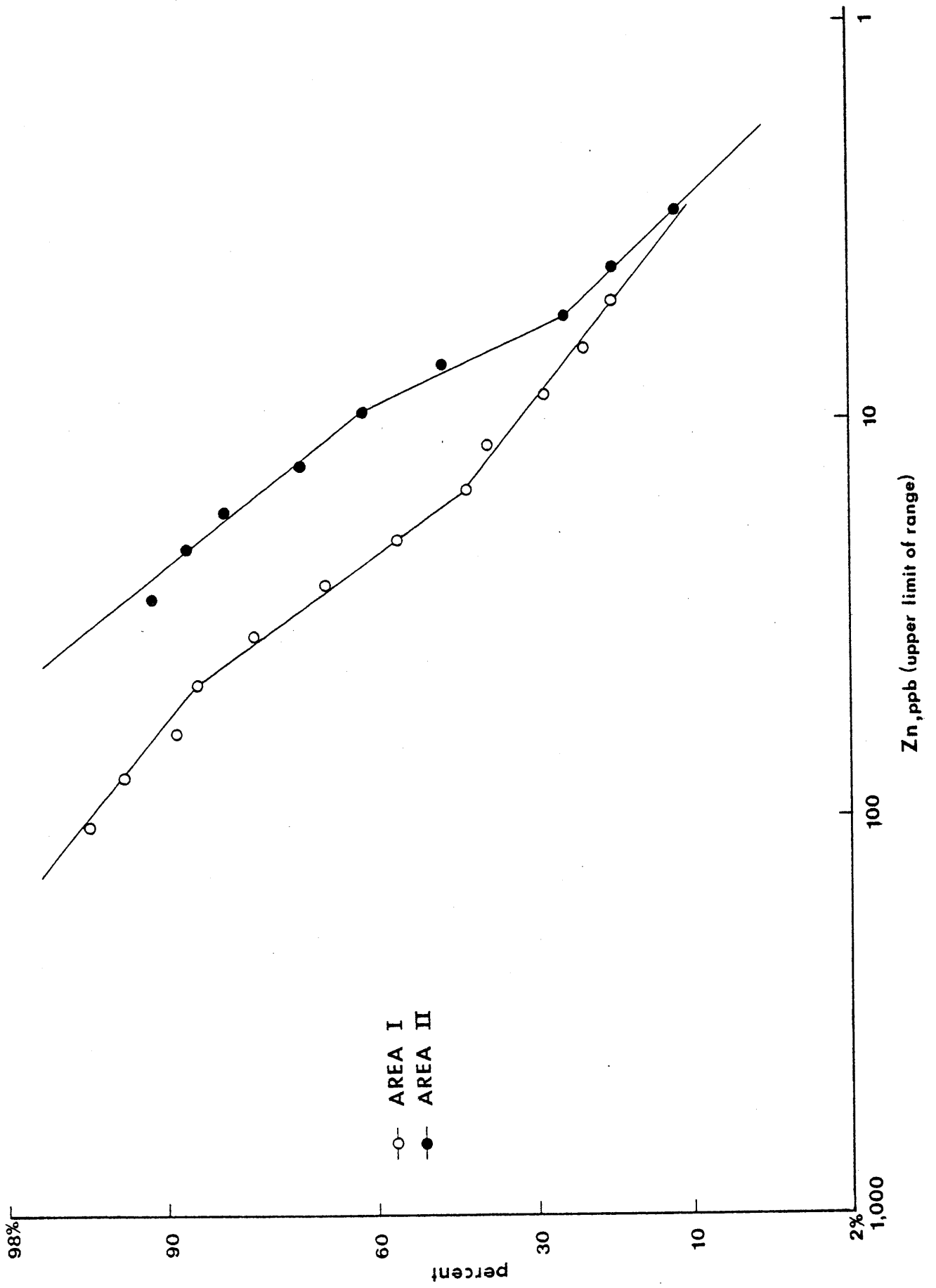


Figure III-2
Cumulative Frequency Curve for Mn, Areas I and II



L11

Figure III-3
Cumulative Frequency Curve for Zn, Areas I and II

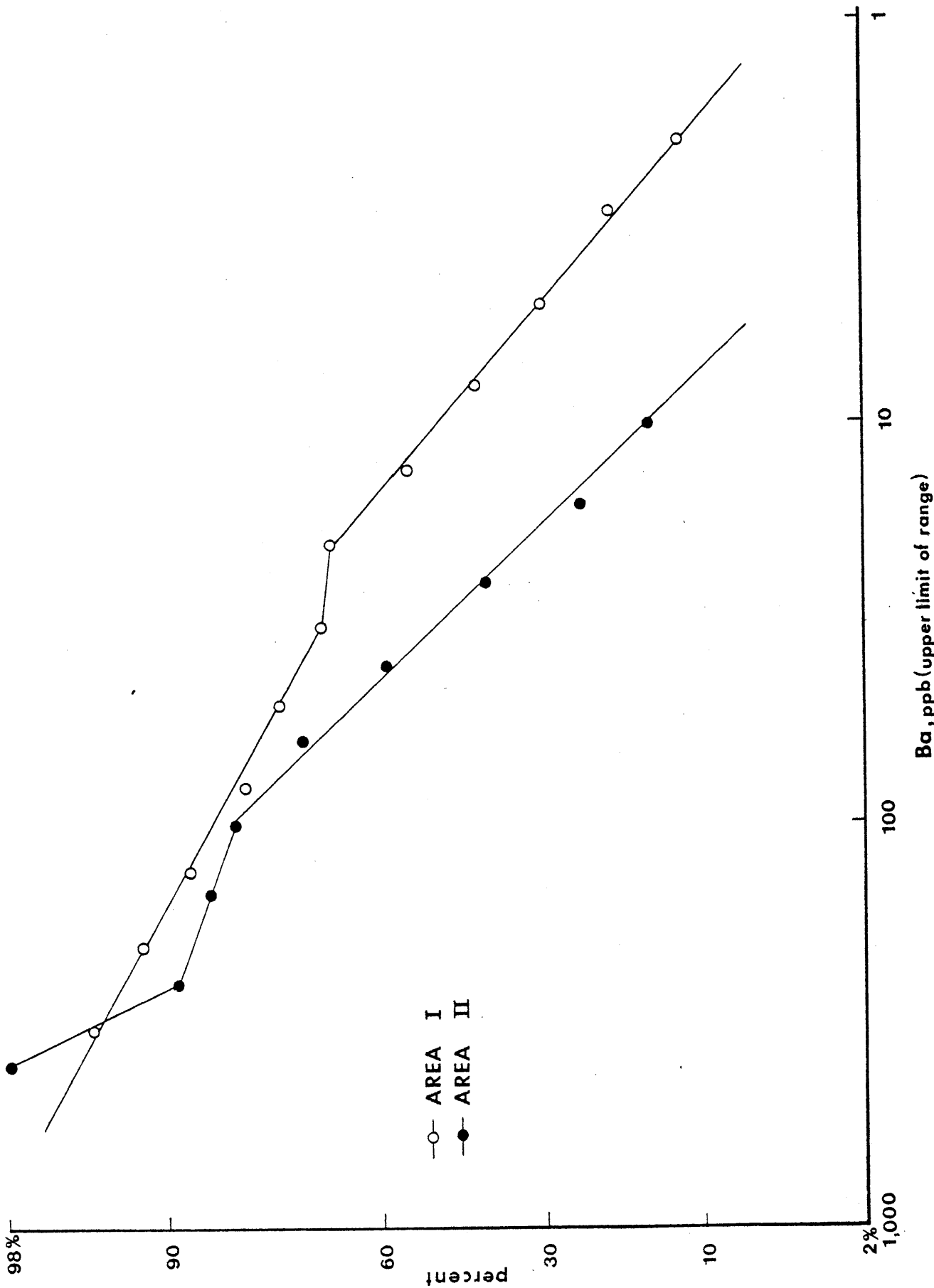


Figure III-4
Cumulative Frequency Curve for Ba, Areas I and II

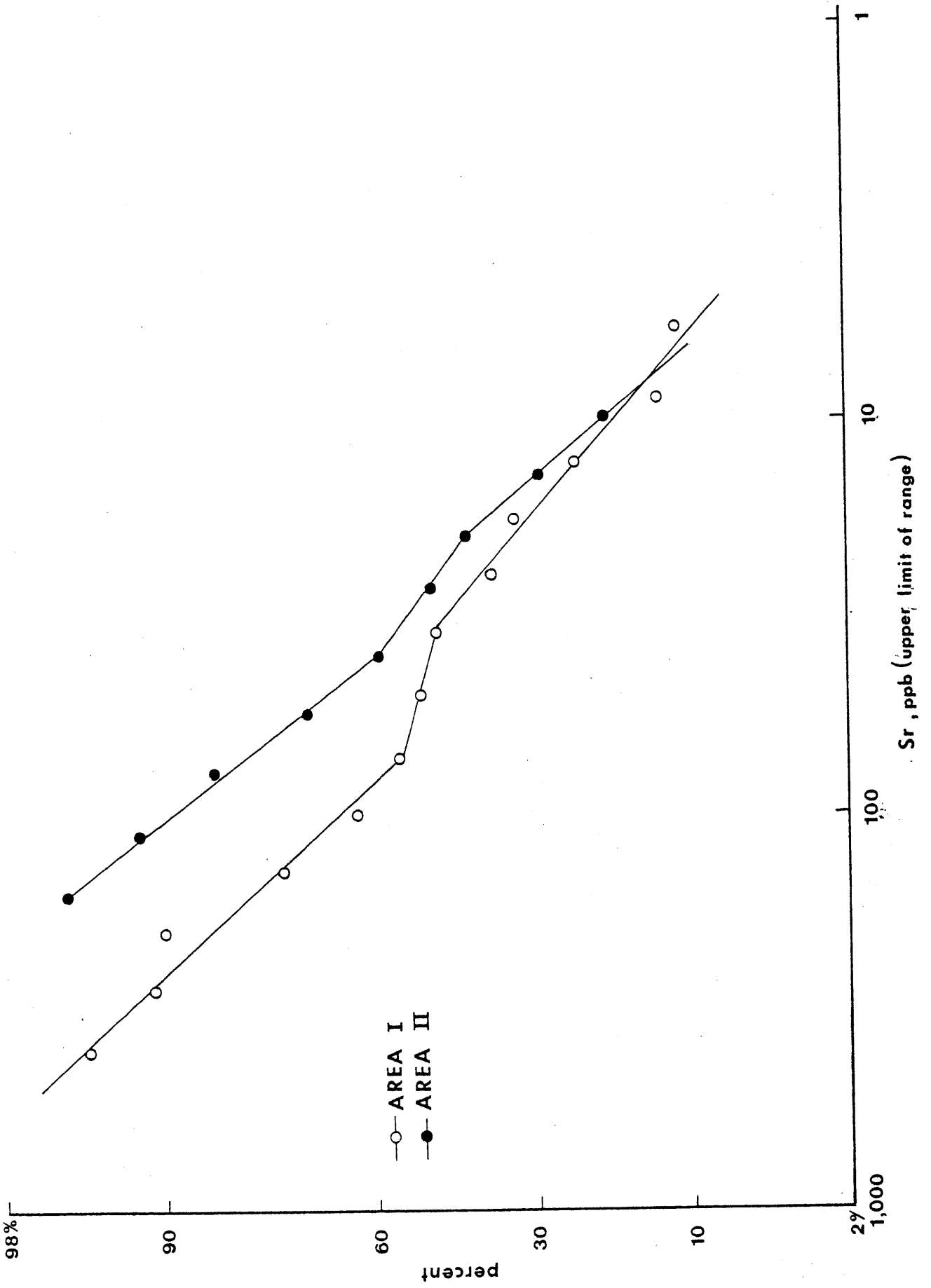
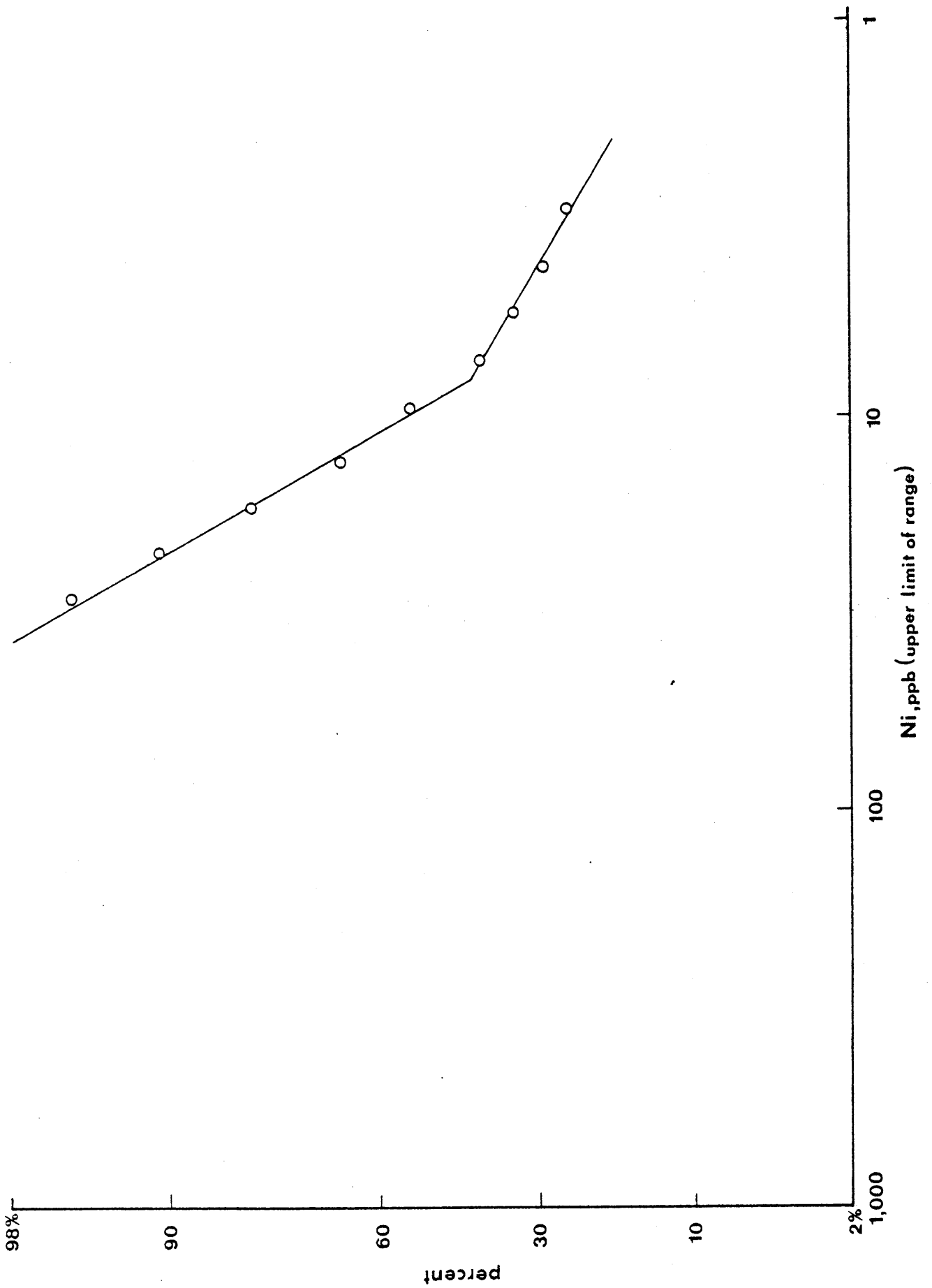
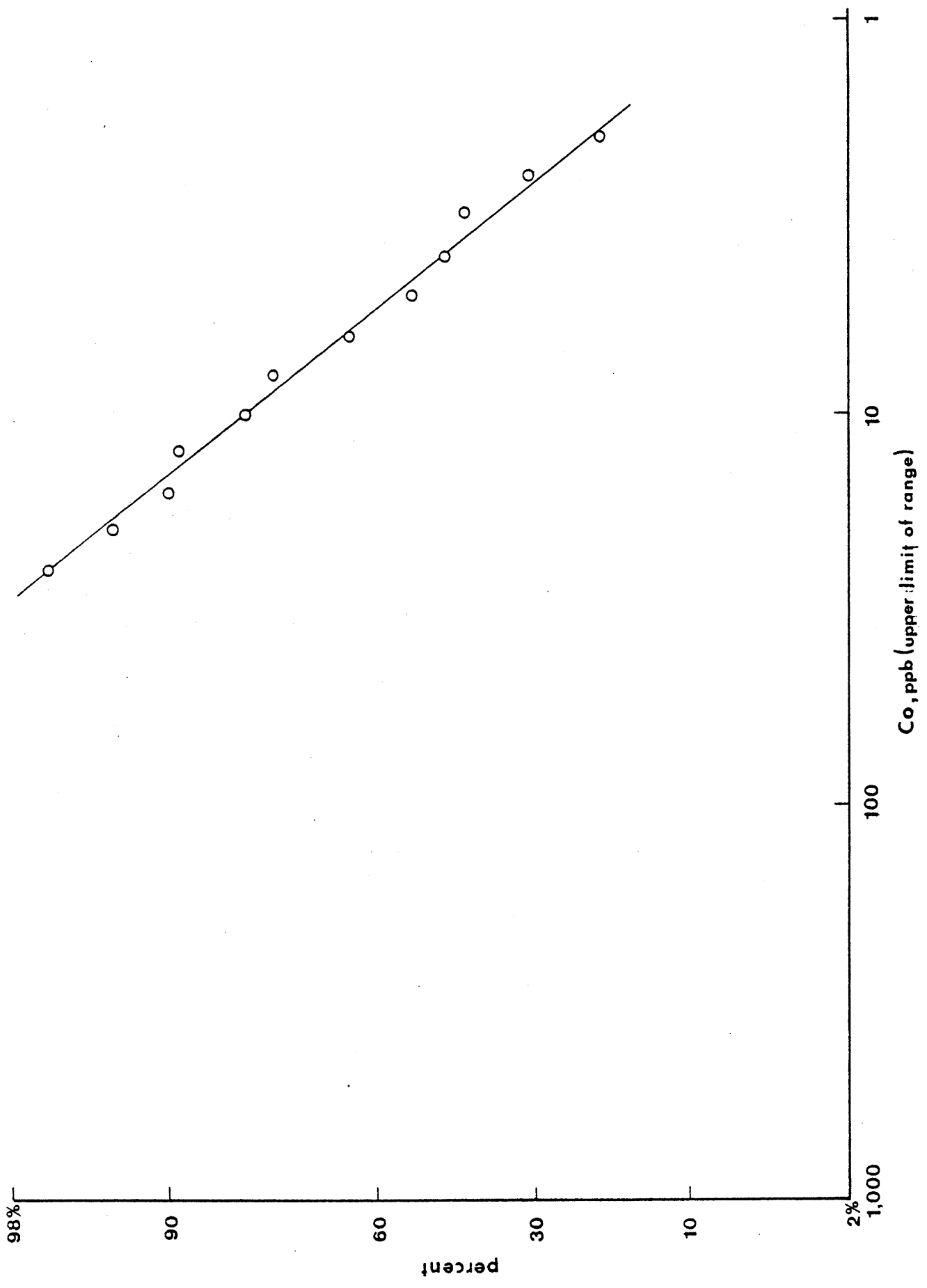


Figure III-5
Cumulative Frequency Curve for Sr, Areas I and II



120

Figure III-6
Cumulative Frequency Curve for Ni, Area I



121

Figure III-7
Cumulative Frequency Curve for Co, Area I

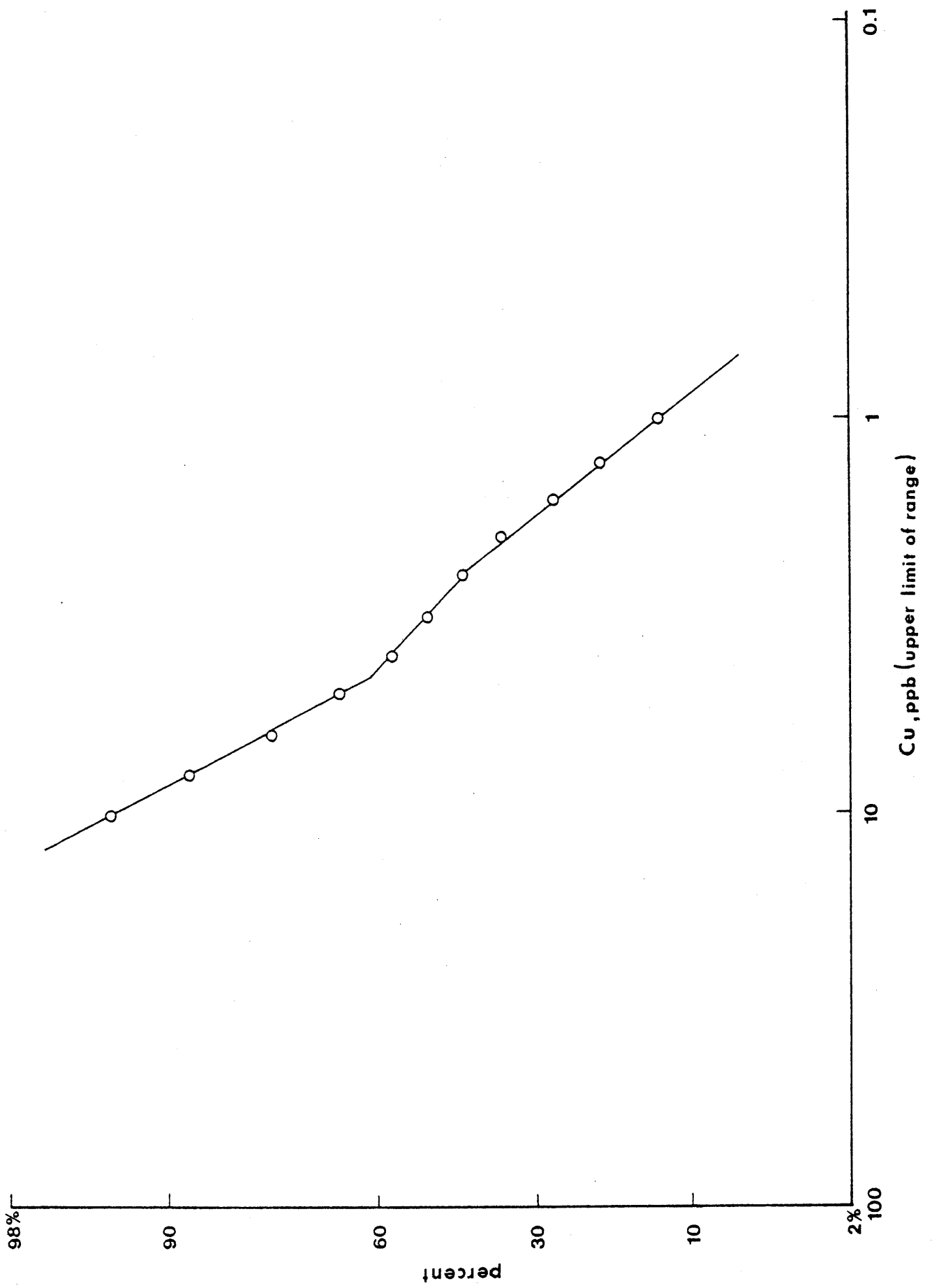
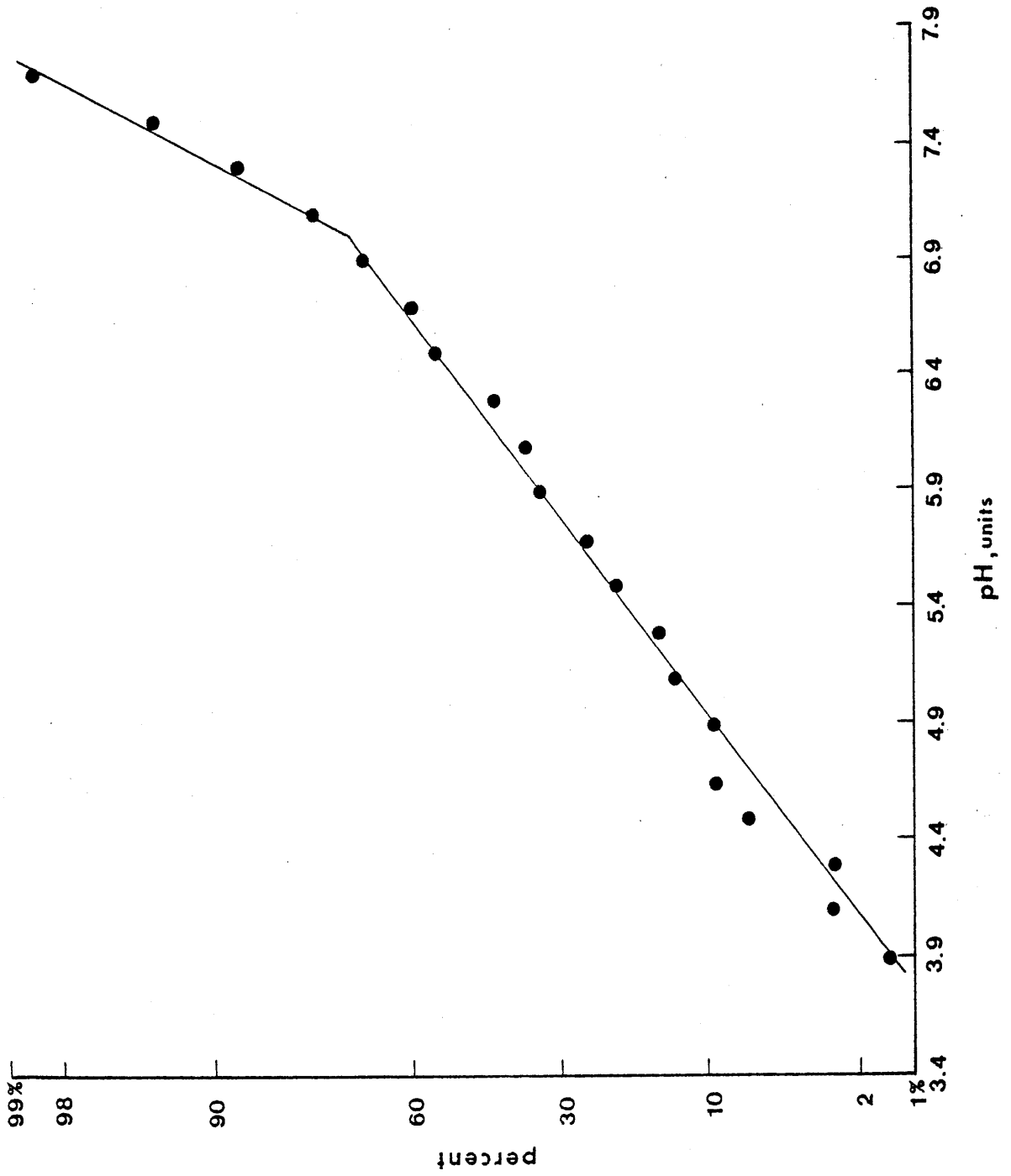
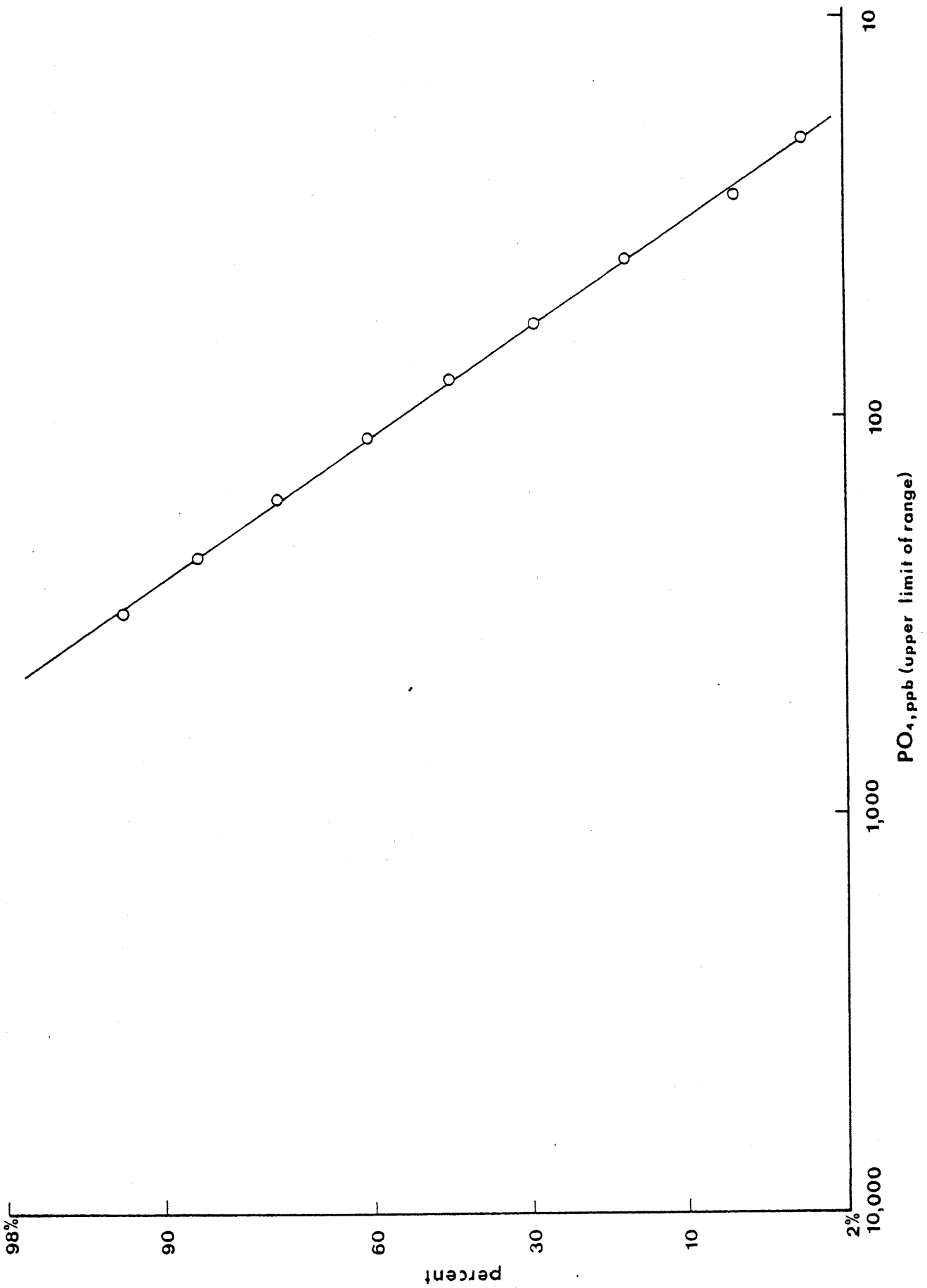


Figure III-8
Cumulative Frequency Curve for Cu, Area I



123

Figure III-9
Cumulative Frequency Curve for pH, Area I



124

Figure III-10
Cumulative Frequency Curve for PO_4 , Area II

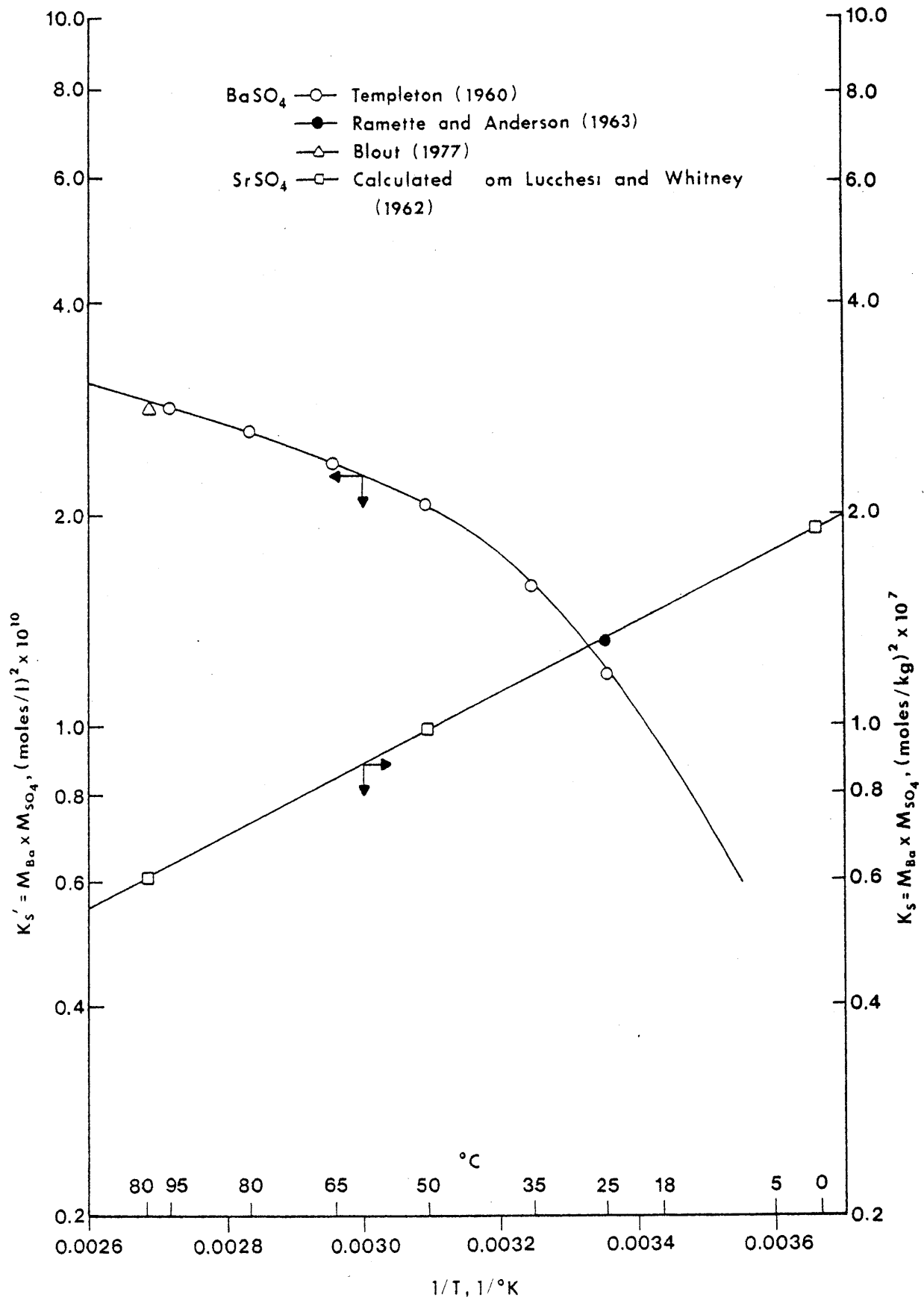


Figure III-11
Solubility Product of Barite (BaSO_4) and
Celestite (SrSO_4) vs Temperature

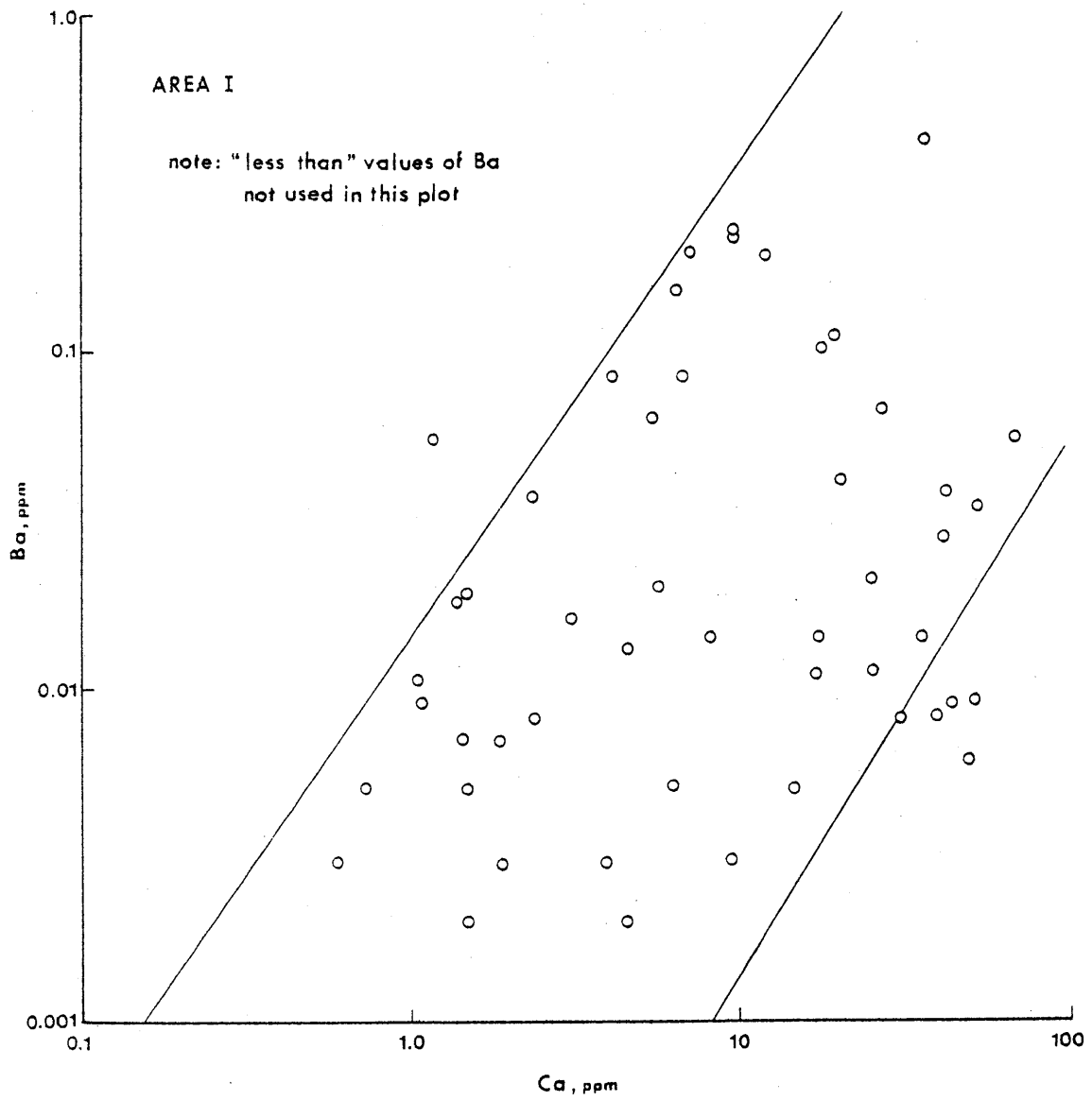


Figure III-12

Ba⁺⁺ Concentration vs Ca⁺⁺ Concentration
For Ouachita Spring Waters of Area I

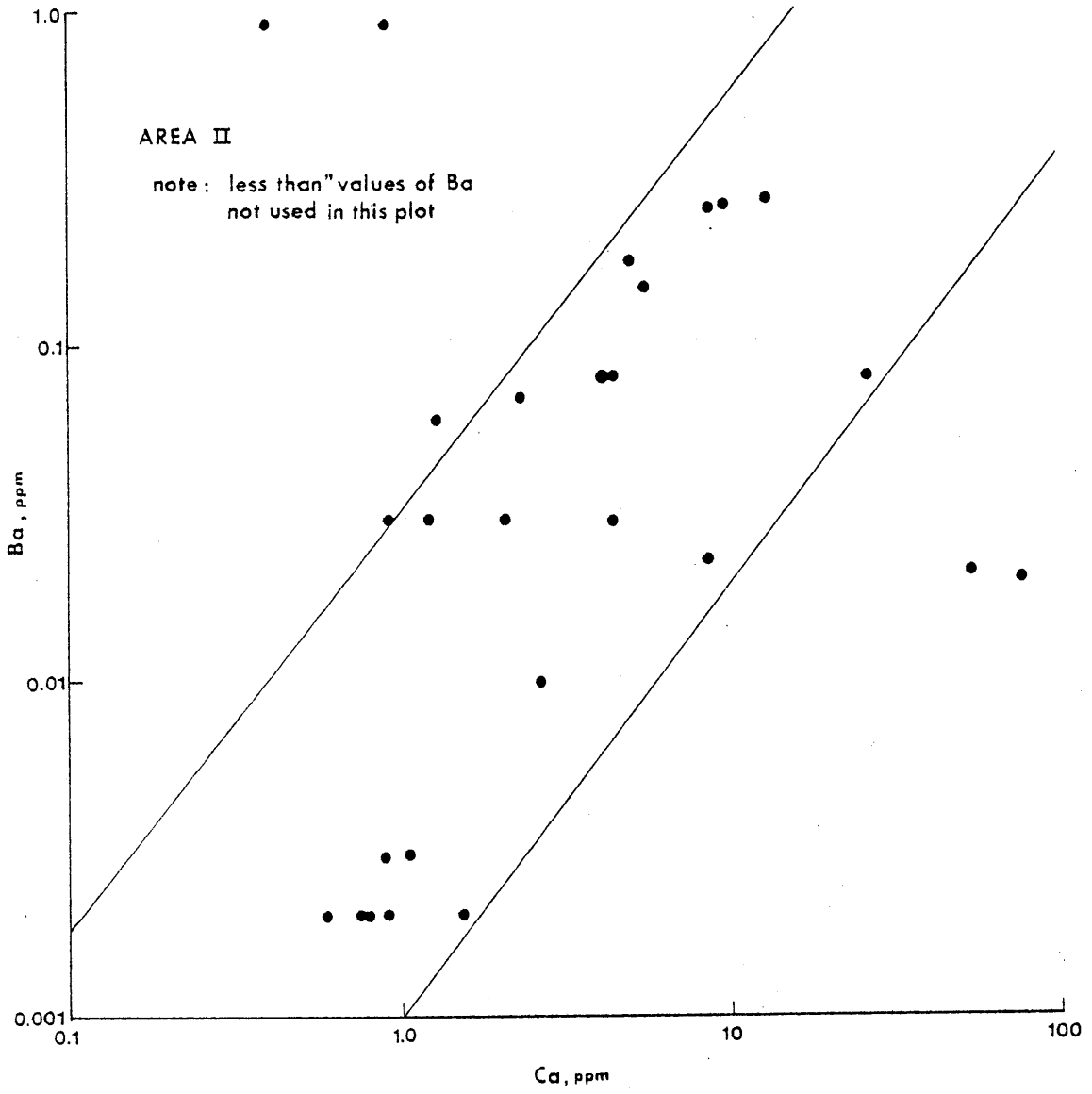


Figure III-13
Ba⁺⁺ Concentration vs Ca⁺⁺ Concentration
For Ouachita Spring Waters of Area II

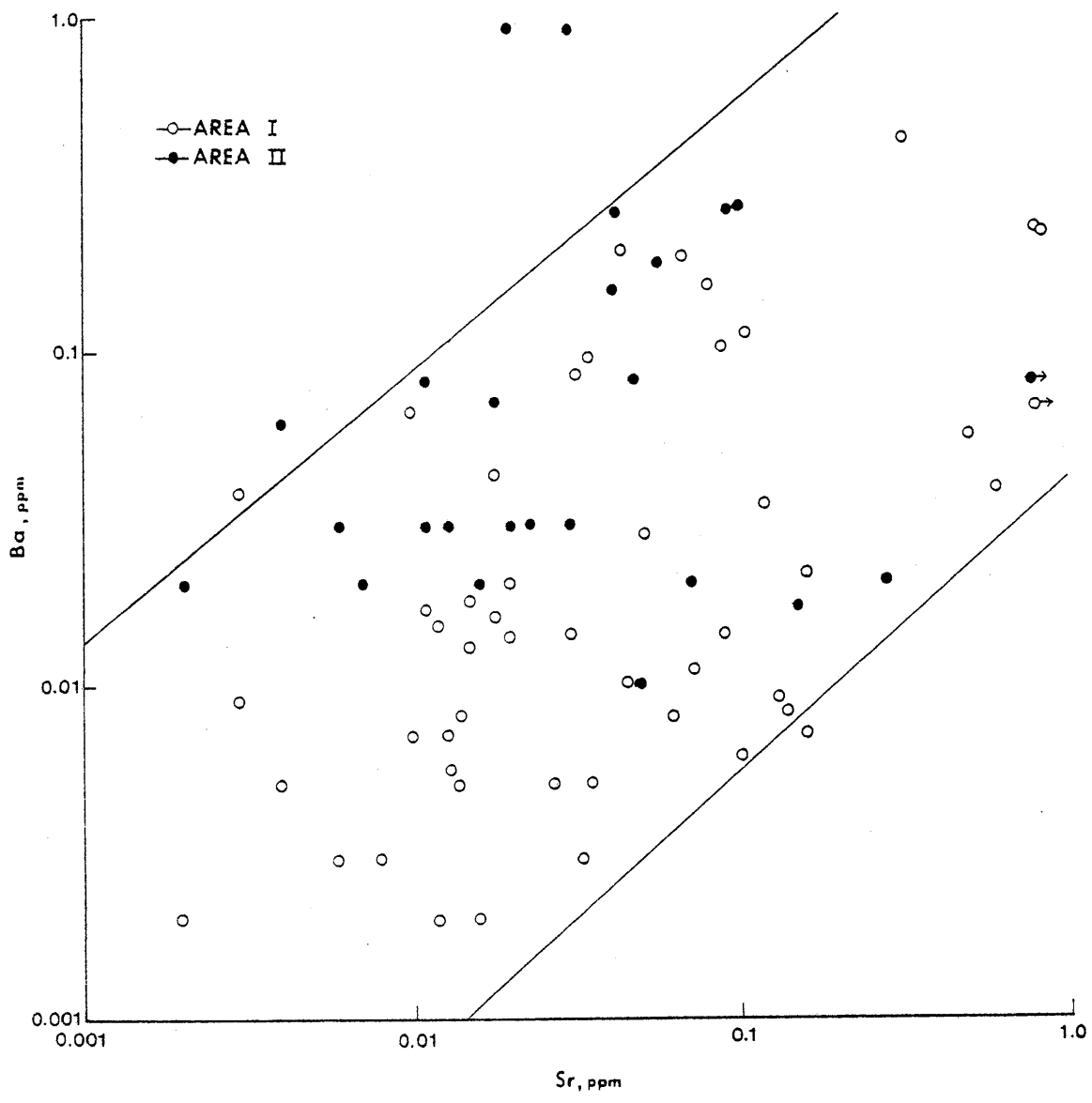


Figure III-14

Ba⁺⁺ Concentration vs Sr⁺⁺ Concentration
For Ouachita Spring Waters of Areas I and II

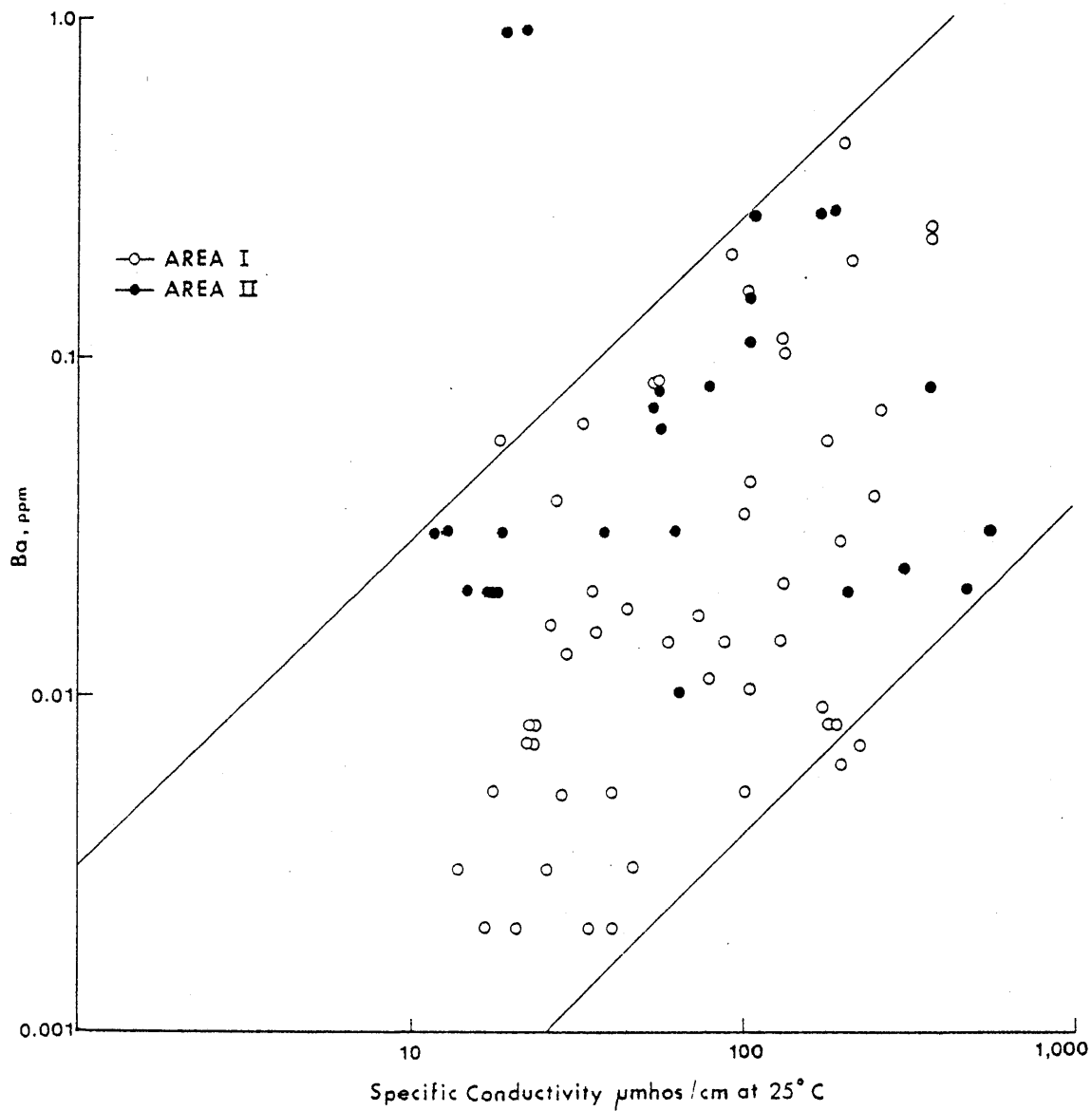
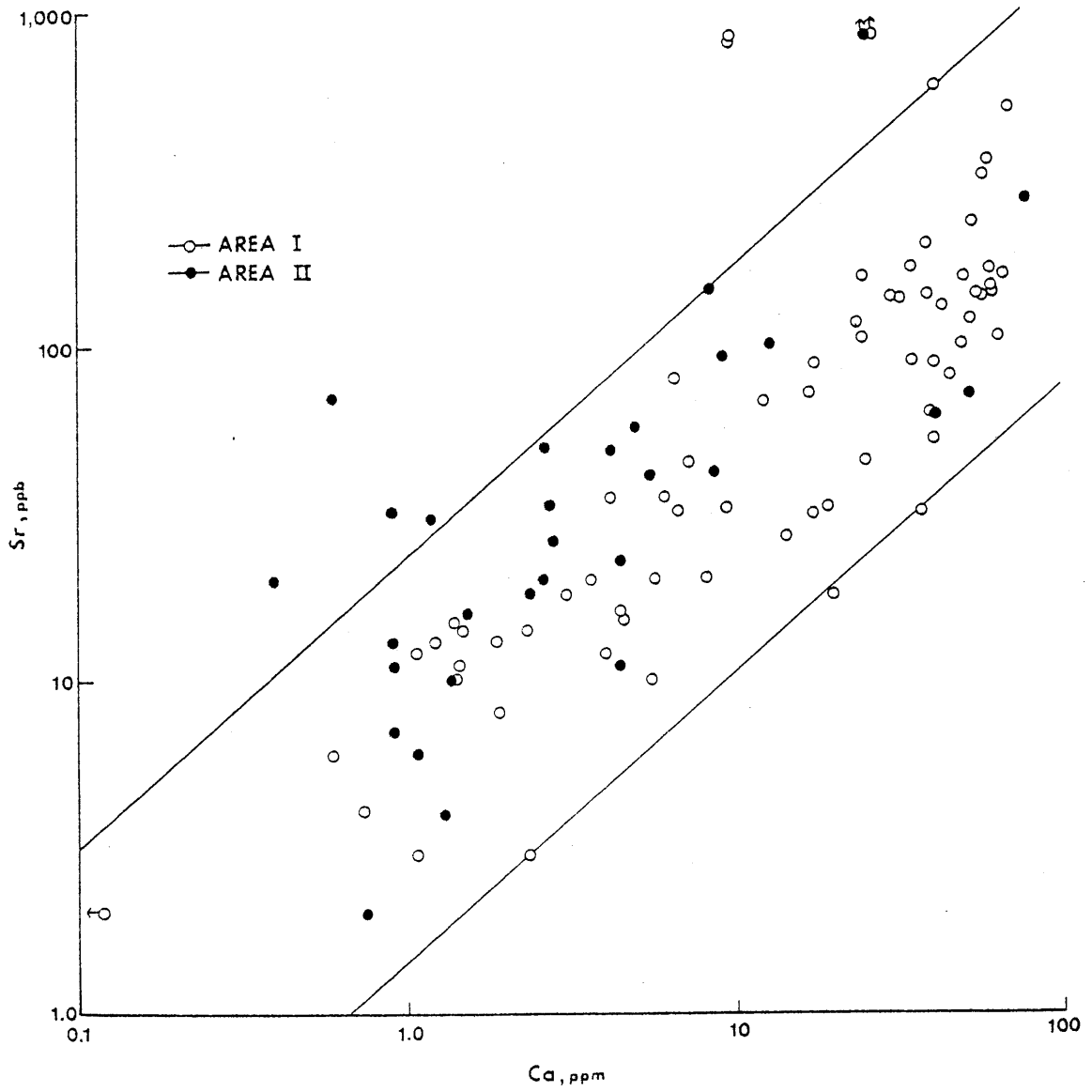


Figure III-15

Ba⁺⁺ Concentration vs Specific Conductivity
For Ouachita Spring Waters in Areas I and II



130

Figure III-16
Sr⁺⁺ Concentration vs Ca⁺⁺ Concentration
For Ouachita Spring Waters in Areas I and II

